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THE ELEMENTS OF EXPERIMENTAL CHEMISTRY,

BY

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THE SIXTH EDITION,

GREATLY ENLARGED; AND ILLUSTRATED WITH NINE PLATES
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TO

MR. JOHN DALTON,

OF MANCHESTER,

AS A

TESTIMONY OF RESPECT

FOR THE

ZEAL, DISINTERESTEDNESS, AND SUCCESS,

WITH WHICH

HE HAS DEVOTED HIMSELF TO THE ADVANCEMENT OF

CHEMICAL PHILOSOPHY,

THIS WORK IS INSCRIBED,

BY HIS FRIEND

THE AUTHOR.

Manchester.
October, 1810.

PREFACE

TO THE

SIXTH EDITION.

THE following work, of which a new impression has for some time past been demanded by the public, though progressively enlarged in bulk, has undergone little change as to its distinguishing plan and objects. The increase of matter, which, in the present edition, is so considerable as to render the former title no longer eligible, bears only a due proportion to the progress of chemical philosophy during the two last years; a progress unexampled in the history of the science, not only for the number and novelty of discoveries, but for the importance of the generalizations to which they have led.

In the FIRST PART of the book, I have

still limited myself to an assemblage of those facts, which are best adapted to convey a knowledge of the elements of chemistry, connected however by a more ample discussion of its general doctrines, and especially by a fuller explanation of the laws of affinity; of the principles of the reformed nomenclature; and of that new department of science, which has been created by the genius and industry of Mr. Davy. A more complete and systematic view of the subject, besides being incompatible with the design of an elementary treatise, appears to me to be rendered unnecessary by several excellent works, of which our language may already boast. The plan of arrangement (the same which I have followed from the first) though perhaps not suited to the maturity of chemical science, is sufficiently adapted to its present state; because it admits of being readily accommodated to newly discovered facts, or to future modifications of theory.

In the SECOND PART, I have detailed the most simple and easy methods of analyzing mineral waters, and mineral bodies in general. With few exceptions, I have had occasion to verify, by my own experience, the practical directions which are given throughout the work; and in those instances, where I have not described facts or processes from my own knowledge, I trust that long established habits of observing chemical phenomena have qualified me to select, from different authorities, those which are most entitled to confidence.

The APPENDIX contains an account of the discoveries, which have been announced while the work was passing through the press; and in a Postscript, which will be found at the close of the second volume, I have been enabled, by the obliging communication of Mr. Davy, to give an abstract of the last paper which he has read before the Royal Society. The TABLES, to which I have added a few new ones, will be found useful both by the speculative and practical chemist. A ninth plate is annexed to this edition; and the series, I believe, now comprehends every article of apparatus, that is essential to the pursuit of experimental chemistry. They have been engraved by Mr. Lowry with his usual accuracy and elegance, from sketches* made for the purpose in the orthographic manner, which appears to me to exhibit chemical instruments much better than the common mode of drawing in perspective. Being executed, also, to a scale, which is placed at the bottom of each plate, a workman is enabled to construct what is represented, much more easily and accurately, than from any verbal description or perspective figure.

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* For these sketches I am indebted to my friends Mr. Ewart of Manchester, and Messrs. Creighton, of Soho, near Birmingham.

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ERRATA.

VOL. I. Page	18,	l. 6,	*for* "refined" *read* "defined."
	64,	4,	"effected" *read* "effective."
	67,	19,	"probably" *read* "probable."
	104,	7,	"loosing" *read* "losing."
	123,	16,	"thin" *read* "tin."
	173,	28,	"fig. 20" *read* "fig. 26."
	189,	11,	"a circular piece," *read* "two circular pieces."
	191,	16,	"fig. 28" *read* "fig. 20."
	215,	last,	"deliquidated" *read* "deliquiated."
	268,	30,	*after* "those" *insert* "metals."
	281,	3,	*for* "increased" *read* "decreased."
VOL. II. Page	482,	18,	"carbon" *read* "hydrogen."
		19,	"hydrogen" *read* "carbon."

INTRODUCTION*.

IT has so long been a custom to preface a course of lectures with the history of the science which is their subject, that it may be necessary to state, briefly, the reasons that have induced me to depart from this established usage.

The history of chemistry may either be merely a history of the science, that is, a view of the progressive development of the facts and doctrines of which the science is composed; or it may comprehend, also, the biography of chemists. The detail of the progress of discovery, however, concerning particular objects of chemical research, would certainly be premature, at a period, when the student may be supposed to be ignorant of the external forms, and even of the existence, of no inconsiderable part of them. Respecting chemists themselves, little can be said that can contribute to information or amusement; for their lives, devoted to the abstract pursuits of science, have seldom been productive of events, that are suited to awaken or gratify general curiosity. Our interest, indeed, respecting philosophers, is seldom excited, unless by a knowledge of the additions which they have made to the facts or theories of a science; and with these a lecturer may fairly presume, however the fact may really be, that his hearers, at the commencement of a course, are wholly unacquainted. On these grounds, therefore, I hope to be excused for devoting to other purposes the time, that would have been allotted to the history of the science. For this, will be substituted a brief view of the nature and objects of chemistry; of its connexion with the arts and with other sciences; and an outline of the plan on which the following lectures will be conducted.

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* The following discourse formed, originally, the introduction to a series of lectures delivered in Manchester, and was afterwards published under the title of "A General View of the Nature and Objects of Chemistry, and of its Application to Arts and Manufactures." As the readers of an elementary book may be presumed to require a similar plan of instruction, with the hearers of a popular course of lectures, I have thought it unnecessary to alter the form under which the essay first appeared, though a few passages are applicable chiefly to the persons to whom it was originally addressed.

Natural philosophy, in its most extensive sense, is a term comprehending every science, that has for its objects the properties and affections of matter. But it has attained, by the sanction of common language, a more limited signification; and chemistry, though strictly a branch of natural philosophy, is generally regarded as a distinct science. Between the two it may, perhaps, be difficult to mark out precisely the line of separation: but, an obvious character of the facts of natural philosophy is, that they are always attended with sensible motion; and the determination of the laws of motion is peculiarly the office of its cultivators. Chemical changes, on the other hand, of the most important kind, often take place without any apparent motion, either of the mass, or of its minute parts; and where the eye is unable to perceive that any change has occurred. The laws of gravitation, of central forces, and all the other powers that fall under the cognizance of the natural philosopher, produce, at most, only a change of place in the bodies that obey their influence. But, in chemical changes, we may always observe an important difference in the properties of things: their appearances and qualities are completely altered, and their individuality destroyed. Thus, two highly corrosive substances, by uniting chemically together, may become mild and harmless; the combination of two colourless substances may present us with a compound of brilliant complexion; and the union of two fluids, with a compact and solid mass.

Chemistry, therefore, may be defined, that science, the object of which is to discover and explain the changes of composition that occur among the integrant and constituent parts of different bodies*.

From this definition, it may readily be conceived, how wide is the range of chemical inquiry; and, by applying it to the various events that daily occur in the order of nature, we shall be enabled to separate them with accuracy, and to allot, to the sciences of natural philosophy and chemistry, the proper objects of the cultivation of each. Whenever a change of place is a necessary part of any event, we shall call in the aid of the former. When this condition may be dispensed with, we shall resort to chemistry for the light of its principles. But it will be often found, that the concurrence of the two sciences is essential to the full explanation of phenomena. The water of the ocean, for example, is raised into the atmosphere by its chemical combination with the matter of heat; but the clouds, that are thus formed, maintain their elevated situation by virtue of a specific gravity inferior to that of the lower regions of the air,—a law, the discovery and application of which are due to the natural philosopher, strictly so called.

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* The reader, who wishes to examine other definitions of chemistry, will find a variety of them, collected by Dr. Black, in the first volume of his "Lectures."

It has not been unusual to consider chemistry, under the twofold view of a science and of an art. This arrangement, however, appears to have had its origin in an imperfect discrimination between two objects, that are essentially distinct. Science consists of assemblages of facts, associated together in classes, according to circumstances of resemblance or analogy. The business of its cultivators is, first, to investigate and establish individual truths, either by the careful observation of natural appearances, or of new and artificial combinations of phenomena produced by the instruments of experiment. The next step is the induction, from well ascertained facts, of general principles or laws, more or less comprehensive in their extent, and serving, like the classes and orders of natural history, the purposes of an artificial arrangement. Of such a body of facts and doctrines, the science of chemistry is composed. But the employment of the artist consists merely in producing a given effect, for the most part by the sole guidance of practice or experience. In the repetition of processes, he has only to follow an established rule; and, in the improvement of his art, he is benefited generally by fortuitous combinations, to which he has not been directed by any general axiom. An artist, indeed, of enlarged and enlightened mind, may avail himself of general principles, and may employ them as an useful instrument in perfecting established operations: but the art and the science are still marked by a distinct boundary. In such hands, they are auxiliaries to each other; the one contributing a valuable accession of facts; and the other, in return, imparting fixed and comprehensive principles, which simplify the processes of art, and direct to new and important practices.

The possession of the general principles of chemistry enables us to comprehend the mutual relation of a great variety of events, that form a part of the established course of nature. It unfolds the most sublime views of the beauty and harmony of the universe; and developes a plan of vast extent, and of uninterrupted order, which could have been conceived only by perfect wisdom, and executed by unbounded power. By withdrawing the mind, also, from pursuits and amusements that excite the imagination, its investigations may tend, in common with the rest of the physical sciences, to the improvement of our intellectual and moral habits; to strengthen the faculty of patient and accurate thinking; and to substitute placid trains of feeling, for those which are too apt to be awakened by the contending interests of men in society, or the imperfect government of our own passions.

The class of natural events, that call for the explanation of chemical science, is of very considerable extent; and the natural philosopher (using this term in its common acceptation) is wholly incompetent to unfold their connexion. He may explain, for example, on the principles of his own science, the annual and diurnal revolutions of the Earth, and part of the train of consequences depending on these rotations. But here he must stop; and the chemist must trace the effects, on the Earth's surface, of the caloric and light derived from the sun; the absorption of caloric by the various bodies on which it falls; the consequent fluidity of some, and volatilization of others; the production of clouds, and their condensation in the form of rain; and the effects of this rain, as well as of the sun's heat, on the animal, vegetable, and mineral kingdoms. In these minuter changes, we shall find, there is not less excellence of contrivance, than in the stupendous movements of the planetary system. And they interest us even more nearly; because, though not more connected with our existence or comfort, yet they are more within our sphere of observation; and an acquaintance with their laws admits of a more direct application to human affairs.

There is another branch of knowledge (that of natural history), which is materially advanced by the application of chemical science. The classifications of the naturalist are derived from an examination and comparison of the external forms, both of animate and inanimate bodies. He distributes the whole range of nature into three great and comprehensive kingdoms,—the animal, the vegetable, and the mineral. Each of these, again, is subdivided into several less extensive classes; and individual objects are referred to their place in the system, by the agreement of their characters, with those assigned to the class, order, and genus. In the different departments of natural history, these resemblances vary in distinctness, in facility of observation, and in certainty of description. Thus, the number and disposition of the parts of fructification in vegetables afford marks of discrimination, which are well defined, and easily ascertained. But minerals, that are not possessed of a regularly crystallized form, are distinguished by outward qualities that scarcely admit of being accurately conveyed by language; such as minute shades of colour; or trifling differences of hardness, transparency, &c. To the evidence of these loose and varying characters, that of the chemical composition of minerals has within the few past years been added; and mineralogy has been advanced, from a confused assemblage of its objects, to the dignity of a well methodized and scientific system. In the example of crystallized bodies, the correspondence between external form and chemical composition, has been most successfully traced by the genius of Haüy; whose method of investigation has enabled him, in numerous instances, to anticipate, from physical characters, the results of the most skilful and laborious analysis.

It is unnecessary to pursue this part of the subject to a greater extent; because, to all who have been in the habit of philosophical investigation, the connexion between the sciences must be sufficiently apparent; and because there is another ground, on which chemistry is more likely to claim, with success, the respect and attention of the great mass of mankind. This is, its capacity of ministering to our wants and our luxuries, and of instructing us to convert to the ordinary purposes of life, many substances which nature presents in a rude and useless form. The extraction of metals from their ores; the conversion of the rudest materials into the beautiful fabrics of glass and porcelain; the production of wine, ardent spirits, and vinegar; and the dyeing of linen and woollen manufactures,—are only a few of the arts that are dependent on chemistry for their improvement, and even for their successful practice.

It cannot, however, be denied, that all the arts which have been mentioned, were practised in times when the rank of chemistry, as a science, was extremely degraded; and that they are the daily employment of unlettered and ignorant men. But to what does this confession amount? and how far does it prove the independence of the above arts on the science of chemistry?

The skill of an artist is compounded of knowledge and of manual dexterity. The latter, it is obvious, no science can teach. But the acquirement of experience, in other words, a talent for accurate observation of facts, and the habit of arranging facts in the best manner, may be greatly facilitated by the possession of scientific principles. Indeed, it is hardly possible for any one to frame rules for the practice of a chemical art, or to profit by the rules of others, who is unacquainted with the general doctrines of the science. For, in all rules, it is implied, that the promised effect will only take place, when circumstances are precisely the same as in the case under which the rule was formed. To ensure an unerring uniformity of result, the substances, employed in chemical processes, must be of uniform composition and excellence; or, when it is not possible to obtain them thus unvaried, the artist should be able to judge precisely of the defect, that he may proportion his agents according to their qualities. Were chemical knowledge more generally possessed, we should hear less of failures and disappointments in chemical operations; and the artist would commence his proceedings, not, as at present, with distrust and uncertainty, but with a well grounded expectation of success.

It will scarcely be contended, that any one of the arts has hitherto attained the extent of its possible perfection. In all, there is yet a wide scope for improvement, and an extensive range for ingenuity and invention. But from what class of men are we to expect useful discoveries? Are we to trust, as hitherto, to the favour of chance and accident; to the fortuitous success of those who are not guided in their experiments by any general principles? Or shall we not rather endeavour to inform the artist, and to induce him to substitute, for vague and random conjecture, the torch of induction and of rational analogy? In the present imperfect state of his knowledge, the artist is even unable fully to avail himself of those fortunate accidents, by which improvements sometimes occur in his processes; because, to the eye of common observation, he may have acted agreeably to established rules, and have varied in circumstances which he can neither perceive nor appreciate. The man of science, in these instances, sees more deeply, and, by availing himself of a minute and accidental difference, contributes at once to the promotion of his own interest, and to the advancement of his art.

But it is the union of theory with practice that is now recommended. And, "when theoretical knowledge and practical skill are happily combined in the same person, the intellectual power of man appears in its full perfection, and fits him equally to conduct, with a masterly hand, the details of ordinary business, and to contend successfully with the untried difficulties of new and perplexing situations. In conducting the former, mere experience may frequently be a sufficient guide; but experience and speculation must be combined to prepare us for the latter*." "Expert men," says Lord Bacon, "can execute and judge of particulars one by one; but the general counsels, and the plots, and the marshalling of affairs, come best from those that are learned."

This recommendation to artists, of the acquirement of scientific knowledge, is happily sanctioned by the illustrious success, in our own days, of the application of theory to the practice of certain arts. Few persons are ignorant of the benefits, that have resulted to the manufactures of this country, from the inventions of Mr. Watt and Mr. Wedgwood; both of whom have been not less benefactors of philosophy than eminent for practical skill. The former, by a clear insight into the doctrine of latent heat, resulting, in a great measure, from his own acuteness and patience of investigation, and seconded by an unusual share of mechanical skill, has perhaps brought the steam-engine to its acme of perfection. Mr. Wedgwood, aided by the possession of extensive chemical knowledge, made rapid advances in the improvement of the art of manufacturing porcelain; and, besides raising himself to great opulence and distinction, has created for his country a source of most profitable and extensive industry. In an art, also, which is nearly connected with the manufactures of our own town, and the improvement of which must, therefore, "come home to our business and bosoms," we owe unspeakable obligations to two speculative chemists, — to Scheele, who first discovered the oxygenized muriatic acid; and to Berthollet, who first instructed us in its application to the art of bleaching.

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* Stewart's Elements of the Philosophy of the Human Mind, chap. iv. sect. 7.

Examples, however, may be urged against indulgence in theory; and instances are not wanting, in which the love of speculative refinement has withdrawn men entirely from the straight path of useful industry, and led them on gradually to the ruin of their fortunes. But from such instances, it would be unfair to deduce a general condemnation of theoretical knowledge. It would be the common error of arguing against things that are useful, from their occasional abuse.—In truth, projects which have, for their foundation, a dependence on chemical principles, may be undertaken with a more rational confidence, than such as have in view the accomplishment of mechanical purposes; because, in chemistry, we are better able, than in mechanics, to predict, from an experiment on a small scale, the probable issue of more extensive attempts. No one, from the successful trial of a small machine can affirm, with unerring certainty, that the same success will attend one on a greatly enlarged plan; because the amount of the resistances, that are opposed to motion, increases often in a *ratio* greater than, from theory, could ever have been foreseen: but the same law, by which the mineral alkali is extracted from a pound of common salt, must equally operate on a thousand times the quantity; and, even when we augment our proportions in this immense degree, the chemical affinities, by which so large a mass is decomposed, are exerted only between very small particles. The failures of the mechanic, therefore, arise from the nature of things; they occur, because he has not in his power the means of foreseeing and calculating the causes that produce them. But, if the chemist fail in perfecting an economical scheme on a large scale, it is either because he has not sufficiently ascertained his facts on a small one, or has rashly embarked in extensive speculations, without having previously ensured the accuracy of his estimates.

The benefits which we are entitled to expect, from the efforts of the artist and the man of science, united in one person, and at the same time tempered and directed by prudential wisdom, affect not only individual but national prosperity. To the support of its distinction, as a commercial nation, this country is to look for the permanency of its riches, its power, and, perhaps, even of its liberties; and this preeminence is to be maintained, not only by local advantages, but on the more certain ground of superiority in the productions of its arts. Impressed with a full conviction of this influence of the sciences, a neighbouring and rival people offer the most public and respectful incitements to the application of theory in the improvement of the chemical arts; and, with the view of promoting this object, national institutions have been formed among them, which have been already, in several instances, attended with the most encouraging success. It may be sufficient, at present, to mention, as an example, that France, during a long war, supplied, from her own native resources, her enormous, and, perhaps, unequalled consumption of nitre.

The general uses of chemistry have been thus fully enlarged upon, because it is a conviction of the utility of the science, that can alone recommend it to attentive and persevering study. It may now be proper to point out, in detail, a few of its more striking applications.

I. The art which is, of all others, the most interesting, from its subserviency to wants that are interwoven with our nature, is AGRICULTURE, or the art of obtaining, from the earth, the largest crops of useful vegetables at the smallest expense.

The vegetable kingdom agrees with the animal one, in the possession of a living principle. Every individual of this kingdom is regularly organized, and requires for its support an unceasing supply of food, which is converted, as in the animal body, into substances of various forms and qualities. Each plant has its periods of growth, health, disease, decay, and death; and is affected, in most of these particulars, by the varying condition of external circumstances. A perfect state of agricultural knowledge would require, therefore, not only a minute acquaintance with the structure and economy of vegetables, but with the nature and effects of the great variety of external agents, that contribute to their nutriment, or influence their state of health and vigour. It can hardly be expected, that the former attainment will ever be generally made by practical farmers; and it is in bringing the agriculturist acquainted with the precise composition of soils and manures, that chemistry promises the most solid advantages. Indeed, any knowledge that can be acquired on this subject, without the aid of chemistry, must be vague and indistinct, and can neither enable its possessor to produce an intended effect with certainty, nor be communicated to others in language sufficiently intelligible. Thus we are told, by Mr. Arthur Young, that, in some parts of England, any loose clay is called marl, in others marl is called chalk, and in others clay is called loam. From so confused an application of terms, all general benefits of experience in agriculture must be greatly limited.

Chemistry may, to agriculturists, become a universal language, in which the facts, that are observed in this art, may be so clothed, as to be intelligible to all ages and nations. It would be desirable, for example, when a writer speaks of clay, loam, or marl, that he should explain his conception of these terms, by stating the chemical composition of each substance expressed by them. For, all the variety of soils and manures, and all the diversified productions of the vegetable kingdom, are capable of being resolved, by chemical analysis, into a small number of elementary ingredients. The formation of a well refined language, expressing the proportion of these elements in the various soils and manures, now so vaguely characterized, would give an accuracy and precision, hitherto unknown, to the experience of the tillers of the earth.

It has been said, by those who contend for pure empiricism in the art of agriculture, that it has remained stationary, notwithstanding all improvements in the sciences, for more than two thousand years. "To refute this assertion," says Mr. Kirwan, "we need only compare the writings of Cato, Columella, or Pliny, with many modern tracts, or, still better, with the modern practice of our best farmers."—"If the exact connexion of effects with their causes," he adds, "has not been so fully and extensively traced in this as in other subjects, we must attribute it to the peculiar difficulty of the investigation. In other subjects, exposed to the joint operation of many causes, the effect of each, singly and exclusively taken, may be particularly examined, and the experimenter may work in his laboratory, with the object always in his view. But the secret processes of vegetation take place in the dark, exposed to the various and indeterminable influences of the atmosphere, and require, at least, half a year for their completion. Hence the difficulty of determining on what peculiar circumstance success or failure depends; for, the diversified experience of many years alone can afford a rational foundation for solid, specific conclusions*."

II. To those who study MEDICINE as a branch of general science, or with the more important view of practical utility, chemistry may be recommended with peculiar force and propriety.—The animal body may be regarded as a living machine, obeying the same laws of motion as are daily exemplified in the productions of human art. The arteries are long, flexible, and elastic canals, admitting, in some measure, the application of the doctrine of hydraulics; and the muscles are so many levers, of precisely the same effect with those which are employed to gain power in mechanical contrivances. But there is another view, in which, with equal justice, the living body may be contemplated. It is a laboratory, in which are constantly going forward processes of various kinds, dependent on the operation of chemical affinities. The conversion of the various kinds of food into blood, a fluid of comparatively uniform composition and qualities; the production of animal heat by the action of the air on that fluid, as it passes through the lungs; and the changes, which the blood afterwards undergoes in its course through the body,—are all, exclusively, subjects of chemical inquiry. To these, and many other questions of physiology, chemistry has of late years been applied with the most encouraging success; and it is to a long continued prosecution of the same plan, that we are to look for a system of physiological science, which shall derive new vigour and lustre from the passing series of years.

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* See Kirwan on Manures.

It must be acknowledged, however, as has been observed by Mr. Davy*, that "the connexion of chemistry with physiology has given rise to some visionary and seductive theories; yet even this circumstance has been useful to the public mind, in exciting it by doubt, and in leading it to new investigations. A reproach, to a certain degree just, has been thrown upon those doctrines known by the name of the chemical physiology; for, in the applications of them, speculative philosophers have been guided rather by the analogies of words than of facts. Instead of endeavouring slowly to lift up the veil, which conceals the wonderful phenomena of living nature; full of ardent imaginations, they have vainly and presumptuously attempted to tear it asunder."

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* In his excellent "Discourse, Introductory to a Course of Lectures," &c. London. Johnson. 1802.

III. There is an extensive class of arts, forming, when viewed collectively, a great part of the objects of human industry, which do not, on a loose and hasty observation, present any general principle of dependency or connexion. But they appear thus disunited, because we have been accustomed to attend only to the productions of these arts, which are, in truth, subservient to widely different purposes. Who would conceive, for instance, that iron and common salt; the one a metal, the use of which results from its hardness, ductility, and malleability; the other a substance, chiefly valuable from its acting as a preservative and seasoner of food,—are furnished by arts alike dependent on the general principles of chemistry? The application of science, in discovering the principles of these arts, constitutes what has been termed ECONOMICAL CHEMISTRY; amongst the numerous objects of which, the following stand most distinguished:

1st. *Metallurgy*, or the art of extracting metals from their ores, comprehending that of *Assaying*, by which we are enabled to judge, from the composition of a small portion, of the propriety of working large and extensive strata. To the metallurgist, also, belong the various modifications of the metals when obtained, and the union of them together, in different proportions, so as to afford compounds adapted to particular uses.—Throughout the whole of this art, much practical knowledge may be suggested by attention to the general doctrines of chemistry. The artist may receive useful hints respecting the construction of furnaces for the fusion of ores and metals; the employment of the proper fluxes; the utility of the admission or exclusion of air; and the conversion of the refuse of his several operations to useful purposes. When the metals have been separated from their ores, they are to be again subjected to various chemical processes. Cast or pig iron is to be changed into the forms of wrought or malleable iron and of steel. Copper, by combination with zinc or tin, affords the various compounds of brass, pinchbeck, bell-metal, gun-metal, &c. Even the art of printing owes something of its present unexampled perfection to the improvement of the metal of types.

2d. Chemistry is the foundation of those arts that furnish us with *saline substances*, an order of bodies highly useful in the businesses of common life. Among these, the most conspicuous are, sugar in all its various forms; the vegetable and mineral alkalis, known in commerce by the names of potash, pearlash, and barilla; common salt; green and blue vitriol, and alum; nitre or saltpetre; sugar of lead; borax; and a long catalogue, which it is needless to extend farther.

3d. The manufacturer of *glass*, and of various kinds of *pottery* and *porcelain*, should be thoroughly acquainted with the nature of the substances he employs: with their fusibility, as affected by difference of proportion, or by the admixture of foreign ingredients; with the means of regulating and measuring high degrees of heat; with the principles on which depend the hardness of his products, and their fitness for bearing the vicissitudes of heat and cold; and with the chemical properties of the best adapted colours and glazings. —Even the humble art of making bricks and tiles has received, from the chemical knowledge of Bergman, the addition of several interesting facts.

4th. The preparation of various kinds of *fermented liquors*, of wine, and ardent spirits, is intimately connected with chemical principles. Malting, the first step in the production of some of these liquors, consists in the conversion of part of the grain into saccharine matter, essential in most instances to the success of the fermentative change. To acquire a precise acquaintance with the circumstances, that favour or retard the process of fermentation, no small share of chemical knowledge is required. The brewer should be able to ascertain, and to regulate exactly, the strength of his infusions, which will vary greatly when he has seemingly followed the same routine. He should be aware of the influence of minute changes of temperature in retarding or advancing fermentation; of the means of promoting it by proper ferments; and of the influence of the presence or exclusion of atmospherical air. A complete acquaintance with the chemical principles of his art, can hardly fail to afford him essential aid in its practice.

The production of ardent spirits is only a sequel of the vinous fermentation, and is, therefore, alike dependent on the doctrines of chemistry.

5th. The arts of *bleaching*, *dyeing*, and *printing*, are, throughout, a tissue of chemical operations. It is not unusual to hear the new mode of bleaching distinguished by the appellation of the chemical method; but it is, in truth, not more dependent on the principles of this science, than the one which it has superseded, nor than the kindred arts of dyeing and printing. In the instance of bleaching, the obligation due to the speculative chemist is universally felt and acknowledged. But the dyer and the printer have yet to receive from the philosopher some splendid invention, which shall command their respect, and excite their attention to chemical science. From purely speculative men, however, much less is to be expected, than from men of enlightened experience, who endeavour to discover the design and reason of each step in the processes of their arts, and fit themselves for more effectual observation of particular facts, by diligently possessing themselves of general truths.

The objects of inquiry that present themselves to the dyer and printer, are of considerable number and importance. The preparation of goods for the reception of colouring matter; the application of the best bases, or means of fixing fugitive colours; the improvement of colouring ingredients themselves; and the means of rendering them permanent, so that they shall not be affected by soap, or by the accidental contact of acids or other corrosive bodies; are among the subjects of chemical investigation. It is the business of the dyer, therefore, to become a chemist; and he may be assured that, even if no brilliant discovery should be the reward of the acquisition, he will yet be better fitted by it for conducting common operations, with certain and unvaried success.

6th. The *tanning and preparation of leather* are processes strictly chemical, which were involved in mystery, till they were reduced to well established principles, by the researches of Seguin, and by the subsequent experiments of Davy. In this, as in most other examples, the application of science to the practical improvement of an art, has to encounter the obstacles of ignorance and prejudice. But the interests of men are sure finally to prevail; and the most bigotted attachment to established forms must give place to the clearly demonstrated utility of new practices. Such a demonstration is generally furnished by some artist of more enlightened views than his neighbours, who has the spirit to deviate from ordinary rules; and thus becomes (not unfrequently with some personal sacrifice) a model for the imitation of others, and an important benefactor of mankind.

Many other chemical arts might be enumerated; but enough, I trust, has been said, to evince the connexion between practical skill and the possession of scientific knowledge. I shall now proceed to develope the plan, on which the following course of instruction will be conducted.

There are two methods of delivering the general doctrines of chemistry, and the facts connected with them. The one consists in a historical detail of the gradual progress of the science; and, in pursuing this plan, we follow the natural progress of the human mind, ascending from particular facts to the establishment of general truths.

But a strong objection to its adoption is, that we are thus led into a minuteness of detail, which is ill suited to the plan of elementary instruction. In the other mode of arrangement, we neglect wholly the order of time in which facts were discovered, and class them under general divisions so framed as to assist the mind in apprehending and retaining the almost infinite variety of particular truths.

In a classification of the objects of chemistry, we may either begin with those substances, which are deemed to be simple, and proceed gradually to the more complicated;—or we may take bodies, as they are usually presented to us, and arrange them according to the resemblances of their external characters; making the development of their composition a subordinate part of the plan. To the former, or synthetic method, there is this strong objection,—that as we are probably still very remote from a knowledge of the true elements of matter, it must be liable, in the progress of science, to frequent and fundamental changes. It has been found necessary, for example, in consequence of Mr. Davy's discoveries, to remove the earths from the class of simple to that of compound bodies. Besides, it may be urged, where are we to place those substances, which have hitherto resisted all attempts at their analysis, and yet have a close resemblance, in natural characters, to the bodies with which they are now associated? Thus the muriatic acid can, with no propriety, be removed from the class of acids to that of simple bodies. For these reasons it appears to me, that one arrangement is preferable to another, on no other ground, than as it is better adapted for communicating a knowledge of the subject; for all must be equally remote from that perfection, which cannot be attained, till the science of chemistry is no longer capable of improvement.

The order, which I have adopted as most eligible, is to commence with those facts, which lead most directly to the establishment of general principles. Attraction or affinity, as the great cause of all chemical changes, and as admitting of illustration by phenomena that are sufficiently familiar, has a primary claim to consideration. Next to that of attraction, the influence of Heat, over the forms and properties of bodies, is the most generally observed fact; and as heat is a power, which is constantly opposed to that of affinity, there is the more propriety in contrasting their operation. With heat, Light also, as a repulsive agent, is frequently associated; and electricity belongs to the same class of powers. But as the action of electricity consists, chiefly, in effecting the disunion of chemical compounds, I have removed it from that place in the system, which seems naturally to belong to it. For before we can understand the general laws of electro-chemical agency, it is necessary to know something of oxygen and a few of the inflammable bodies; nor can the theory of the excitation of galvanic electricity be made at all intelligible, without this previous knowledge.

The phenomena of heat, and the laws deduced from them, conduct us naturally to the great source of that fluid, which will be traced to a class of bodies agreeing, in mechanical properties, with the air of our atmosphere, and called airs or *gases*. These gases, we shall find, consist partly of gravitating matter, and partly of an extremely subtile fluid, which impresses on our organs the sensation of heat, and is called *caloric*[*]. When the ponderable ingredients, usually called the *bases*, of these gases, combine together, or with other bodies, caloric is given out and new compounds are generated. It is on the possession or absence of the property of decomposing one of them, oxygen gas, that a comprehensive division has been made of bodies into *combustible* and *incombustible*. The latter, however, it appears probable, are such, only in consequence of being already combined with oxygen. Hence we might divide all the variety of substances into oxygen and inflammable matter; and of these, the last mentioned, it appears not improbable from Mr. Davy's discoveries, will turn out to be universally metallic.

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[*] Light and electricity are probably, also, constituents of the gases.

The next class of bodies, that claim our attention, includes those compounds, which are formed either by the union of the simple gases or of their bases. Thus oxygen and nitrogen gases compose atmospheric air; and hydrogen and oxygen water. Nitrogen and hydrogen, by their union, afford ammonia; and with this fluid the fixed alkalis are naturally associated. The detail of properties belonging to the alkalis and earths, is, indeed, a necessary preliminary to that of the acids, the most important quality of which is, that they constitute, with the alkalis and earths, the extensive class of neutral salts. The consideration of the bases of the alkalis and earths has been made to follow that of the bodies themselves, because these bases are the products of refined and complicated operations, which could scarcely have been otherwise understood. The fixed alkalis, also, precede the volatile one, on account of the singular effects of potassium on ammonia.

The next class of compounds is, that of acids. With each of these I have connected the history of its base, when known; for as several of these bodies (sulphur and phosphorus for instance) have lost, in consequence of Mr. Davy's discoveries, their title to be considered as elementary, it becomes merely a question of convenience where they should be placed. In treating of the acids, their relation will be traced to those bodies only which have already been described; for it would be unreasonable to detail their action on metals, till that class of substances has been specifically discussed.

Having dismissed the consideration of such elementary bodies as are distinguished by affording acids when combined with oxygen, of the properties of acids thus generated, and of the compounds afforded by the union of acids with alkalis; an important division of elementary substances will next claim our attention, *viz.* the metals.

This class of bodies, it is usual to introduce at a much earlier period: but I have adopted a different order, from the consideration, that, with the previous knowledge of the constitution and qualities of acids, the history of the metals may be made much more complete; and, especially, that all the various modes and phenomena of their combinations with oxygen may be more distinctly explained. The more complex productions of the vegetable and animal kingdoms will be the last step in our progress through the science.

For the exclusive adoption of the new doctrines of chemistry, and of the nomenclature connected with them, no apology is necessary. Every one, who will be at the pains of attentively comparing the new with the old theory, I can venture to predict, will prefer the lucid arrangements and precision of the former, to the confused order and illogical inferences of the advocates of phlogiston. From those who have been in the habit of teaching chemistry, both before and since the revolution in this science, we have the strongest testimony, that the labour of acquiring a knowledge of it is diminished beyond all comparison.—"I have adopted the new nomenclature," says Mr. Chaptal, "in my lectures and writings, and I have not failed to perceive how very advantageous it is to the teacher; how much it relieves the memory; how greatly it tends to produce a taste for chemistry; and with what facility and precision the ideas and principles concerning the nature of bodies fix themselves in the minds of the auditors." We have the approbation, also, of the most distinguished metaphysicians of the age, of the connexion of new doctrines with a new and more accurate language. "The new nomenclature of chemistry," it is observed by Mr. Dugald Stewart, in his Elements of the Philosophy of the Human Mind, "seems to furnish a striking illustration of the effect of appropriate and well-defined expressions, in aiding the intellectual powers; and the period is probably not far distant, when similar innovations will be attempted in other sciences*."

These doctrines, and the nomenclature dependent on them, have not, it must be acknowledged, been long established; and though the progress of chemistry, during and since their development, has advanced with unexampled rapidity, we should still so limit our approbation, as not to esteem them beyond the reach of improvement.—For my own part, I adopt them, not from a belief that they are perfect; but because they are better adapted than any hitherto offered, for explaining and classing phenomena; and with this qualification, I strongly recommend them to general acceptance.

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* See also Condorcet's "*Esquisse d'un Tableau Historique des Progrès de l'Esprit Humain.*"

ELEMENTS

OF

EXPERIMENTAL CHEMISTRY.

PART I.

AN ARRANGED SERIES OF EXPERIMENTS AND PROCESSES TO BE PERFORMED BY THE STUDENT OF CHEMISTRY.

CHAPTER I.

OF A CHEMICAL LABORATORY AND APPARATUS.

A CHEMICAL laboratory, though extremely useful, and even essential, to all who embark extensively in the practice of chemistry, either as an art, or as a branch of liberal knowledge, is by no means required for the performance of those simple experiments, which furnish the evidence of the fundamental truths of the science. A room that is well lighted, easily ventilated, and destitute of any valuable furniture, is all that is absolutely necessary for the purpose. It is even advisable, that the construction of a regular laboratory should be deferred, till the student has made some progress in the science; for he will then be better qualified to accommodate its plan to his own peculiar views and convenience.

It is scarcely possible to offer the plan of a laboratory, which will be suitable to every person, and to all situations; or to suggest any thing more than a few rules that should be generally observed. Different apartments are required for the various classes of chemical operations. The principal one may be on the ground-floor; twenty-five feet long, fourteen or sixteen wide, and open to the roof, in which there should be contrivances for allowing the occasional escape of suffocating vapours. This will be destined chiefly for containing furnaces, both fixed and portable. It should be amply furnished with shelves and drawers, and with a large table in the centre, the best form of which is that of a double cross. Another apartment may be appropriated to the minuter operations of chemistry; such as those of precipitation on a small scale, the processes that require merely the heat of a lamp, and experiments on the gases. In a third, of smaller size, may be deposited accurate balances, and other instruments of considerable nicety, which would be injured by the acid fumes that are constantly spread through a laboratory.

The following are the principal instruments that are required in chemical investigations; but it is impossible, without entering into very tedious details, to enumerate all the apparatus that should be in the possession of a practical chemist.

I. Furnaces. These may be formed either of solid brick-work, or of such materials as admit of their removal from place to place.

The directions generally laid down in elementary books of chemistry, for the construction of FIXED FURNACES, appear to me deficient in precision, and such as a workman would find it difficult to put in practice. I have, therefore, given plans and sections, in the last two plates, of the various kinds of furnaces: and, in the Appendix, minute instructions will be found for erecting them *.

The furnaces of most general utility are, 1st, the *Wind Furnace*, in which an intense heat is capable of being excited for the fusion of metals, &c. In this furnace, the body submitted to the action of heat, or the vessel containing it, is placed in contact with the burning fuel. Fig. 60 exhibits one of the most common construction. Fig. 61 is the section of a wind furnace; the plan of which was obligingly communicated to me by Mr. Knight, of Foster-lane, London, to whom, also, I am indebted for that represented, fig. 62. The wind furnace of Mr. Chenevix is shown by fig. 74. 2dly, The *Evaporating Furnace* is formed of iron plates, joined together by rabbiting, and placed over horizontal returning flues. Figs. 64 and 65, are two views of this furnace, as recommended by Mr. Knight. When evaporation is performed by the naked fire, the vessel may be placed on the top of the furnace, fig. 60 or 61; and when effected through the intervention of a water bath, a shallow kettle of water, in which is placed the evaporating dish and its contents, may be set in the same situation. For the purposes of evaporating liquids, and drying precipitates on a small scale, at a temperature not exceeding 212° Fah^{t}, a convenient apparatus is represented in plate 10th to the 6th vol. of Nicholson's Journal, 8vo. 3dly, The *Reverberatory Furnace* is represented by figs. 66, 67, 68. 4thly, *The Furnace for Distilling by a Sand Heat*, is constructed, by setting upon the top of the brick-work, fig. 60, the iron pot, fig. 71; a door being made in the side of the furnace for introducing fuel. Distillation by the naked fire, is performed with the wind furnace, figs. 62, 63. 5thly, The *Cupelling*, or *Enamelling Furnace*, is shown by figs. 69, 70.

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* See the Description of the 7th and 8th plates in the Appendix.

Portable furnaces, however, are amply sufficient for all the purposes of the chemical student, at the outset of his pursuit. The one which I prefer, is that shown by figs. 58 and 59. It was originally contrived, I believe, by Mr. Schmeisser*; and is made, with considerable improvements, and sold by Mr. Knight. Its size is so small, that it may be set on a table, and the smoke may be conveyed by an iron pipe, into the chimney of the apartment. In the furnace, as it is usually sold, the chimney, adapted for distillation with a sand heat, passes directly through the sand-bath, the form of which is necessarily altered, from the common to a very inconvenient one. I have found it a great improvement to make the aperture for the chimney at *k*. This allows us to have a sand-bath of the usual shape, as shown by fig. 59; or even to place evaporating dishes, or a small boiler, on the top of the furnace. The aperture may be closed by a stopper, when we dispose the furnace as shown by fig. 58. Dr. Black's furnace is generally made of a larger size; and is adapted to operations on a more considerable scale. (See figs. 72 and 73.) Both these furnaces are constructed of thin iron plates, and are lined with fire-clay. They will be minutely described in the references to the plates.

For the purpose of exciting a sudden heat, and of raising it to great intensity, nothing can be better adapted than a very simple, cheap, and ingenious furnace, contrived by Mr. Charles Aikin, fig. 55. It is formed out of pieces of black-lead melting pots, in a manner to be afterwards described; and is supplied with air by a pair of double bellows, *d*. By a slight alteration, this furnace may occasionally be employed for the operation of cupelling. (See fig. 57.)

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* See his Mineralogy, Tab. iii. and iv.

II. For containing the materials, which are to be submitted to the action of heat in a wind furnace, vessels called CRUCIBLES are employed. They are most commonly made of a mixture of fire-clay and sand, occasionally with the addition of plumbago, or black lead. The Hessian crucibles are best adapted for supporting an intense heat without melting; but they are liable to crack when suddenly heated or cooled. The porcelain ones, made by Messrs. Wedgwood, are of much purer materials, but are still more apt to crack on sudden changes of temperature; and, when used, they should, therefore, be placed in a common crucible of larger size, the interval being filled with sand. The black-lead crucibles resist very sudden changes of temperature, and may be repeatedly used; but they are destroyed when some saline substances (such as nitre) are melted in them, and are consumed by a current of air. For certain purposes, crucibles are formed of pure silver, or platina. Their form varies considerably, as will appear from inspecting plate vi. figs. 49, 50, 51, and 54. It is necessary, in all cases, to raise them from the bars of the grate, by a stand, fig. 53, *a* or *b*. For the purpose of submitting substances to the continued action of a red heat, and with a considerable surface exposed to the air, the hollow arched vessel, with a flat bottom, fig. 52, termed a *muffle*, is commonly used.

In fig. 69, *d*, *e*, the muffle is shown, placed in a furnace for use.

III. Evaporating vessels should always be of a flat shape, so as to expose them extensively to the action of heat. (See a section of one, fig. 12.) They are formed of glass, of earthen ware, and of various metals. Those of glass are with difficulty made sufficiently thin, and are often broken by changes of temperature; but they have a great advantage in the smoothness of their surface, and in resisting the action of most acid and corrosive substances. Evaporating vessels of porcelain, or Wedgwood's ware, are next in utility, are less costly, and less liable to be cracked. They are made both of glazed and unglazed ware. For ordinary purposes the former are to be preferred; but the unglazed should be employed when great accuracy is required, since the glazing is acted on by several chemical substances. Evaporating vessels of glass, or porcelain, are generally bedded, up to their edge, in sand (see fig. 65); but those of various metals are placed immediately over the naked fire. When the glass or porcelain vessel is very thin, and of small size, it may be safely placed on the ring of the brass stand, plate i. fig. 18, and the flame of an Argand's lamp, cautiously regulated, may be applied beneath it. A lamp thus supported, so as to be raised or lowered, at pleasure, on an upright pillar, to which rings, of various diameters, are adapted, will be found extremely useful; and, when a strong heat is required, it is advisable to employ a lamp, provided with double concentric wicks.

IV. In the process of evaporation, the vapour for the most part is allowed to escape; but of certain chemical processes, the collection of the volatile portion is the principal object. This process is termed, DISTILLATION. It is performed in vessels of various forms and materials. The common still is so generally known, that a representation of it in the plates was deemed unnecessary. (See Aikin's Chem. Dict. pl. ii. fig. 31.) It consists of a vessel, generally of copper, shaped like a tea-kettle, but without its spout and handle. Into the opening of this vessel, instead of a common lid, a hollow moveable head is affixed, which ends in a narrow, open pipe. This pipe is received into another tube of lead, which is twisted spirally, and fixed in a wooden tub, so that it may be surrounded by cold water. (Fig. 40, *dd.*) When the apparatus is to be used, the liquid intended to be distilled is poured into the body of the still, and the head is fixed in its place, the pipe, which terminates it, being received into the leaden worm. The liquid is raised into vapour, which passes into the worm, is there condensed by the surrounding cold water, and flows out at the lower extremity.

The common still, however, can only be employed for volatilizing substances that do not act on copper, or other metals, and is, therefore, limited to very few operations. The vessel, fig. 2, is of glass, or earthen ware, and is also intended for distillation. It is termed an *alembic*, and consists of two parts; the body *a* for containing the materials, and the head *b* by which the vapour is condensed; the pipe *c* conveying it to a receiver. Vessels, termed *retorts*, however, are more generally used. Fig. 1, *a* shows the common form, and fig. 13, *a* represents a stoppered, or tubulated retort. Retorts are made of glass, of earthen ware, or of metal. When a liquid is to be added at distant intervals during the process, the best contrivance is that shown fig. 26, *a*, consisting of a bent tube, with a funnel at the upper end. When the whole is introduced at first, it is done either through the tubulure, or, if into a plain retort, through the funnel, fig. 10.

To the retort, a *receiver* is a necessary appendage; and this may either be plain, fig. 1, *b*, or tubulated, as shown by the dotted lines at *c*. To some receivers a pipe is added (fig. 13, *b*), which may enter partly into a bottle beneath. This vessel, which is principally useful for enabling us to remove the distilled liquid, at different periods of the process, is termed a *quilled receiver*. For some purposes, it is expedient to have the quilled part accurately ground to the neck of the bottle, *c*, which should then be furnished with a tubulure, or second neck, having a ground stopper, and should be provided, also, with a bent tube, to be occasionally applied, for conveying away any gases that may be produced. The condensation of the vapour is much facilitated, by lengthening the neck of the retort with an *adopter* (fig. 11), the wider end of which slips over the retort neck, while its narrow extremity is admitted into the mouth of the receiver. (See fig. 63.)

Heat may be applied to the retort in several modes. When the vessel is of earthen ware, and when the distilled substance requires a strong heat to raise it into vapour, the naked fire is applied, as shown fig. 63. Glass retorts are generally placed in heated sand (fig. 59); and, when of a small size, the flame of an Argand's lamp, cautiously regulated, may be conveniently used (fig. 13).

In several instances, the substance raised by distillation is partly a condensable liquid, and partly a gas, which is not condensed till it is brought into contact with water. To effect this double purpose, a series of receivers, termed *Woulfe's Apparatus*, is employed. The first receiver (*b*, fig. 30) has a right-angled glass tube, open at both ends, fixed into its tubulure; and the other extremity of the tube is made to terminate beneath the surface of distilled water, contained, as high as the horizontal dotted line, in the three-necked bottle *c*. From another neck of this bottle, a second pipe proceeds, which ends, like the first, under water, contained in a second bottle *d*. To the central neck a straight tube, open at both ends, is fixed, so that its lower end may be a little beneath the surface of the liquid. Of these bottles any number may be employed that is thought necessary.

The materials being introduced into the retort, the arrangement completed, and the joints secured in the manner to be presently described, the distillation is begun. The condensable vapour collects in a liquid form in the balloon *b*, while the evolved gas passes through the bent pipe, beneath the surface of the water in *c*, which continues to absorb it till saturated. When the water of the first bottle can absorb no more, the gas passes, uncondensed, through the second right-angled tube, into the water of the second bottle, which, in its turn, becomes saturated. Any gas that may be produced, which is not absorbable by water, escapes through the bent tube *e*, and may be collected, if necessary.

Supposing the bottles to be destitute of the middle necks, and, consequently, without the perpendicular tubes, the process would be liable to be interrupted by an accident: for if, in consequence of a diminished temperature, an absorption or condensation of gas should take place, in the retort *a*, and, of course, in the balloon *b*, it must necessarily ensue that the water of the bottles *c* and *d* would be forced, by the pressure of the atmosphere, into the balloon, and possibly into the retort: but, with the addition of the central tubes, a sufficient quantity of air rushes through them to supply any accidental vacuum.

This inconvenience, however, is still more conveniently obviated by Welther's tube of safety (fig. 31. *b*), which supersedes the expediency of three-necked bottles. The apparatus being adjusted, as shown by the figure, a small quantity of water is poured into the funnel, so as to about half fill the ball *b*. When any absorption happens, the fluid rises in the ball, till none remains in the tube, when a quantity of air immediately rushes in. On the other hand, no gas can escape, because any pressure from within is instantly followed by the formation of a high column of liquid in the perpendicular part, which resists the egress of gas. This ingenious invention I can recommend, from ample experience of its utility.

Very useful alterations in the construction of Woulfe's apparatus have been contrived also by Mr. Pepys and Mr. Knight. That of the former is shown (fig. 32), where the balloon *b* is surmounted by a vessel accurately ground to it, and furnished with a glass valve, resembling that affixed to Nooth's apparatus. This valve allows gas to pass freely into the vessel *c*, but prevents the water which it contains from falling into the balloon. Mr. Knight's improvement is described, and represented in a plate, in the Philosophical Magazine, vol. xx*.

When a volatile substance is submitted to distillation, it is necessary to prevent the escape of the vapour through the junctures of the vessels; and this is accomplished by the application of LUTES. The most simple method of confining the vapour, it is obvious, would be to connect the places of juncture accurately together by grinding; and accordingly the neck of the retort is sometimes ground to the mouth of the receiver. This, however, adds too much to the expense of apparatus to be generally practised.

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* Another modification of this apparatus, by Mr. Murray, is represented in Nich. Journ. 8vo. vol. iii. or in Murray's System of Chemistry, vol. i. plate v. fig. 40. Fig. 41 of the same plate exhibits a cheap and simple form of this apparatus, contrived by the late Dr. Hamilton, and depicted originally in his translation of Berthollet on Dyeing. Mr. Burkitt's improvement of this apparatus may be seen in Nicholson's Journal, 4to. vol. v. 349.

When the distilled liquid has no corrosive property (such as water, alcohol, ether, &c.), slips of moistened bladder, or of paper, or linen, spread with flour paste, white of egg, or mucilage of gum arabic, sufficiently answer the purpose. The substance which remains, after expressing the oil from bitter almonds, and which is sold under the name of almond-meal, or flour, forms a useful lute, when mixed, to the consistency of glaziers' putty, with water or mucilage. For confining the vapour of acid, or highly corrosive substances, the fat lute is well adapted. It is formed by beating perfectly dry and finely sifted tobacco pipe-clay, with painters' drying oil, to such a consistence that it may be moulded by the hand. The same clay, beat up with as much sand as it will bear, without losing its tenacity, with the addition of cut tow, or of horse-dung, and a proper quantity of water, furnishes a good lute, which has the advantage of resisting a considerable heat, and is applicable in cases where the fat lute would be melted or destroyed. Various other lutes are recommended by chemical writers; but the few that have been enumerated I find to be amply sufficient for every purpose.

On some occasions, it is necessary to protect the retort from too sudden changes of temperature, by a proper coating. For glass retorts, a mixture of common clay, or loam, with sand, and cut shreds of flax, may be employed.

If the distillation be performed by a sand heat, the coating needs not to be applied higher than that part of the retort which is bedded in sand; but if the process be performed in a wind furnace (fig. 63), the whole body of the retort, and that part of the neck also which is exposed to heat, must be carefully coated. To this kind of distillation, however, earthen retorts are better adapted; and they may be covered with a composition originally recommended by Mr. Willis. Two ounces of borax are to be dissolved in a pint of boiling water, and a sufficient quantity of slaked lime added, to give it the thickness of cream. This is to be applied by a painter's brush, and allowed to dry. Over this a thin paste is afterwards to be applied, formed of slaked lime and common linseed-oil, well mixed and perfectly plastic. In a day or two, the coating will be sufficiently dry to allow the use of the retort.

For joining together the parts of iron vessels, used in distillation, a mixture of the finest China clay, with solution of borax, is well adapted. In all cases, the different parts of any apparatus made of iron should be accurately fitted by boring and grinding, and the above lute is to be applied to the part which is received into an aperture. This will generally be sufficient without any exterior luting; otherwise the lute of clay, sand, and flax, already described, may be used.

In every instance, where a lute or coating is applied, it is advisable to allow it to dry before the distillation is begun; and even the fat lute, by exposure to the air during one or two days after it's application, is much improved in its quality. The clay and sand lute is perfectly useless, except it be previously quite dry. In applying a lute, the part immediately over the juncture should swell outwards, and its diameter should be gradually diminished on each side. (See fig. 13, where the luting is shown, applied to the joining of the retort and receiver.)

Besides the apparatus already described, a variety of vessels and instruments are necessary, having little resemblance to each other, in the purposes to which they are adapted. Glass vessels are required for effecting *solution*, which often requires the application of heat, and sometimes for a considerable duration. In the latter case, it is termed digestion, and the vessel, fig. 4, called a *mattras*, is the most proper for performing it. When solution is quickly effected, the bottle, fig. 5, with a rounded bottom, may be used; or a common Florence oil flask serves the same purpose extremely well, and bears, without cracking, sudden changes of temperature. For *precipitations*, and separating liquids from precipitates, the decanting-jar (fig. 16) will be found useful; or, if preferred, it may be shaped as in fig. 26, *f*. Liquids, of different specific gravities, are separated by the vessel fig. 3; the heavier fluid being drawn off through the cock *b*, and air being admitted by the removal of the stopper *a*, to supply its place. Glass rods, of various lengths, and spoons of the same material, or of porcelain, are useful for stirring acid and corrosive liquids; and a stock of cylindrical tubes, of various sizes, is required for occasional purposes. It is necessary also to be provided with a series of glass measures, graduated into drachms, ounces, and pints.

For the drying of precipitates, and other substances, by a heat not exceeding 212°, a very useful apparatus is sold in London. It is represented, supported by the ring of a lamp-stand, by fig. 27. The vessel *a* is of sheet-iron or copper japanned and hard-soldered; *c* is a conical vessel of very thin glass, having a rim, which prevents it, when in its place, from entirely slipping into *a;* and *d* is a moveable ring, which keeps the vessel *c* in its place. When the apparatus is in use, water is poured into *a* about as high as the dotted line; the vessel *c*, containing the substance to be dried, is immersed in the water, and secured by the ring *d;* and the whole apparatus set over an Argand's lamp. The steam escapes by means of the chimney *b*, through which a little hot water may be occasionally poured, to supply the waste by evaporation. By changing the shape of *c* to the segment of a sphere, still retaining the rim, I have found it a most convenient vessel for evaporating fluids.

Accurate beams and scales, of various sizes, with corresponding weights, some of which are capable of weighing several pounds, while the smaller size ascertains a minute fraction of a grain, are essential instruments in the chemical laboratory. So also are mortars of different materials, such as of glass, porcelain, agate, and metal. Wooden stands, of various kinds, for supporting receivers, should be provided. (See Aikin's Chem. Dict. pl. iv. fig. 59, *e*.) For purposes of this sort, and for occasionally raising to a proper height any article of apparatus, a series of blocks, made of well seasoned wood, eight inches (or any other number) square, and respectively eight, four, two, one, and half an inch in thickness, will be found extremely useful; since, by combining them in different ways, thirty-one different heights may be obtained.

The blow-pipe is an instrument of much utility in chemical researches. A small one, invented by Mr. Pepys, with a flat cylindrical box for condensing the vapour of the breath, and for containing caps, to be occasionally applied with apertures of various sizes, is perhaps the most commodious form. (See Aikin's Chem. Dict. pl. vii. fig. 71, 72, 73.) One of a much smaller size, for carrying in the pocket, has been contrived by Dr. Wollaston, and may be had from Mr. Knight of Foster-Lane*. A blow-pipe, which is supplied with air from a pair of double bellows worked by the foot, may be applied to purposes that require both hands to be left at liberty, and will be found useful in blowing glass, and in bending tubes. The latter purpose, however, may be accomplished by holding them over an Argand's lamp with double wicks. Occasionally, when an intense heat is required, the flame of the blow-pipe, instead of being supported by the mouth, may be kept up by a stream of oxygen gas, expelled from a bladder or from a gas-holder. (See a representation of the apparatus in the Chemical Conversations, pl. ix.)

In the course of this work, various other articles of apparatus will be enumerated, in detailing the purposes to which they are adapted, and the principles on which they are constructed. It must be remembered, however, that it is no part of my object to describe every ingenious and complicated invention, which has been employed in the investigation of chemical science: but merely to assist the student in attaining apparatus for general and ordinary purposes. For such purposes, and even for the prosecution of new and important inquiries, very simple means are sufficient; and some of the most interesting chemical facts may be exhibited with the aid merely of Florence flasks, of common vials, and of wine glasses. In converting these to the purposes of apparatus, a considerable saving of expense will accrue to the experimentalist; and he will avoid the encumbrance of various instruments, the value of which consists in show, rather than in real utility.

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* It is described in Nich. Journ. xv. 284.

In the selection of experiments, I shall generally choose such as may be undertaken by persons not possessed of an extensive chemical apparatus. On some occasions, however, it may be necessary, in order to complete the series, that others should be included, requiring, for their performance, instruments of considerable nicety. The same experiment may, perhaps, in a few instances, be repeatedly introduced in illustration of different principles; but this repetition will be avoided as much as possible. Each experiment will be preceded by a brief enunciation of the general truth which it is intended to illustrate.

CHAPTER II.

OF CHEMICAL AFFINITY.

ALL bodies, composing the material system of the universe, have a mutual tendency to approach each other, whatsoever may be the distances at which they are placed. The operation of this force extends to the remotest parts of the planetary system, and is one of the causes that preserve the regularity of their orbits. The smaller bodies, also, that are under our more immediate observation, are influenced by the same power, and fall to the Earth's surface, when not prevented by the interference of other forces. From these facts, the existence of a property has been inferred, which has been called *attraction*, or more specifically the *attraction of gravitation.* Its nature is entirely unknown to us; but some of its laws have been investigated, and successfully applied to the explanation of phenomena. Of these, the most important are, that the force of gravity acts on bodies directly in proportion to the quantity of matter in each; and that it decreases in the reciprocal proportion of the squares of the distances.

From viewing bodies in the aggregate, we may next proceed to contemplate them as composed of minute particles. These, also, are influenced by the force of attraction, but not unless when placed in apparent contact. Hence a distinction has been made between gravitation, and that kind of attraction, which is effective only at insensible distances. The latter has been called *contiguous attraction* or *affinity;* and it has been distinguished, as it is exerted between particles of matter, of the *same* kind, or between particles of a *different* kind.

By the *affinity of aggregation*, the *cohesive affinity*, or more simply *cohesion*, is to be understood that force or power, by which particles of matter *of the same kind* attract each other, the only effect of this affinity being an aggregate or mass. Thus a lump of copper may be considered as composed of an infinite number of minute particles or integrant parts, each of which has precisely the same properties, as those that belong to the whole mass. These are united by the cohesive affinity. But if the copper be combined with another metal (such as zinc), we obtain a compound, the constituent parts of which, copper and zinc, are combined by the power of chemical affinity. In simple bodies, therefore, the cohesive affinity alone is exerted; but compounds are influenced by both affinities, their *constituent* or dissimilar parts being united by chemical affinity, and their *integrant* or similar parts by the affinity of aggregation.

SECTION I.

Of Cohesion, Solution, and Crystallization.

The cohesive affinity is a property, which is common to a great variety of bodies. It is most strongly exerted in solids; and in these it is directly proportionate to the mechanical force required for effecting their disunion. In liquids, it acts with considerably less energy; and in aëriform bodies we have no evidence that it exists at all; for their particles, as will afterwards be shown, if not held together by pressure, would probably separate to immeasurable distances. Its force is not only different among different bodies, but in various states of the same body. Thus in the cohesion of certain metals, important changes are produced by the rate of cooling, by hammering, and by other mechanical operations. Water, also, in a solid state, has considerable cohesion, which is much diminished when it becomes liquid, and entirely destroyed when it is changed into vapour.

The most important view, in which the chemist has to consider cohesion, is that of a force either counteracting or modifying chemical affinity; for the more strongly the particles of any body are united by this power, the less are they disposed to enter into combination with other bodies. In many cases, a very powerful affinity between two substances may be rendered wholly inefficient, by the strong cohesion of one or both of them. Hence it has been received as an axiom, that *the affinity of composition is inversely proportionate to the cohesive affinity.*

The cohesion of bodies may be overcome either by mechanical disintegration; by the action of caloric; or by the predominance of chemical affinity. The operations of pulverizing, rasping, grinding, &c., which form part of a variety of ordinary processes, are examples of the first method of destroying cohesion. The effect of caloric is perhaps best exemplified in the case of two metals (suppose tin and copper) which do not unite, however minutely divided, until brought by heat into a state of fluidity. One of the most simple cases of the predominance of chemical over cohesive affinity, is presented by the solution of a solid in a liquid, a process which affords a clear illustration of the agency of the opposite powers of chemical affinity and cohesion.

The term SOLUTION comprehends an extensive class of phenomena; for it is applied not only to the disappearance of solids in liquids, but to such combinations of liquids with each other, and of solid or liquid with aëriform bodies, as are attended with perfect transparency. It appears to be entirely an effect of the operation of chemical affinity, exerted between the more fluid body termed the *solvent*, and the body which is to be dissolved. Thus when water is poured upon any salt, the solid may be considered as acted upon by two forces. Its cohesive attraction, on the one hand, tends to preserve it in a solid state, and its affinity for water, on the other, to bring it into a state of solution. In this case either the cohesion of the salt may be so strong, as to resist entirely any sensible action of the water; or it may be inferior, in various degrees, to the force of chemical affinity. If the latter force prevail, solution goes on to a certain point, beyond which it cannot be carried. This term, at which water ceases to act upon a salt, is called the point of *saturation*, and the resulting liquid a *saturated solution.*

In the instance of a soluble salt, the force of affinity prevails then over that of cohesion. But conformably to a general law, which will afterwards be explained, water acts with gradually diminishing energy on all solids, as the solution approaches the term of saturation. When this is attained, the affinity of the fluid, and the cohesion of the solid, may be supposed to have arrived at an equilibrium. No farther action can be exerted by the fluid upon the solid, unless we can either increase the energy of affinity, or diminish that of cohesion, by some extraneous force. In these cases we have generally recourse to the agency of caloric, which acts chiefly by diminishing cohesion. But even the operation of heat has limits; for it is incapable of impairing cohesion beyond a certain degree, when the two opposing forces are again balanced, and we obtain a new point of saturation. In this and all similar instances, the term of saturation must vary at almost every different temperature. It does not, however, appear that the increase of the solvent power of water bears any regular proportion to the increase of temperature; and in several cases, it remains nearly the same at all temperatures.

When water has ceased to act upon a solid, in consequence of its having attained the term of saturation, the solution may take up another solid of a different kind. In this instance the solid first dissolved exerts an affinity for the second, sufficient to overcome the cohesion of the latter; and the solution goes on, till these two forces attain their equilibrium.

In these examples of the most simple effect of chemical affinity, and its predominance over cohesion, the properties of the bodies, which are combined, undergo little or no alteration. A solid, when thus dissolved, retains its most remarkable sensible qualities of taste, smell, &c., as well as its chemical relations. The affinities, indeed, of most solids, are rendered much more efficient by solution; because one of the principal obstacles to chemical action (the cohesive affinity) is thus in a great measure overcome. There are even many bodies that do not combine at all, without previous solution. Hence it was formerly received as an universal axiom, "*corpora non agunt nisi sint soluta:*" a proposition, to which the advancement of chemical science has since discovered many exceptions.

In solids, which are thus retained in solution, the cohesive attraction, though overcome by a counteracting force, must still be considered as existing, and as tending to reunite the integrant particles, which are suspended in the fluid. The smallest addition to the cohesive attraction renders this force efficient, and a part of the solid resumes its original form. Thus, every solution, which, after being saturated at the common temperature, has received a second quantity of the solid by the application of heat, deposits a portion of the solid on cooling; because when the increase of temperature, which overcame its cohesion, ceases to act, that force again becomes efficient. The same thing happens, if we expel, by heat, a portion of the water which holds the solid in solution. The effect of this, is to approximate the particles of the solid, and thus to bring them within the sphere of their mutual attraction.

When the cohesive affinity of a dissolved solid once more becomes efficient, it may produce, according to circumstances, either shapeless masses, or solids of a regular figure, bounded by plane surfaces and determinate angles. A heated solution of any salt, which is greatly more soluble in hot than in cold water, if suddenly cooled, deposits the salt in the form of a powder, and affords one of the simplest cases of what is termed in chemistry *precipitation*. Again, a salt with a less difference as to solubility at different temperatures, such as sulphate of soda, forms, by the rapid cooling of a heated and saturated solution, an irregular but transparent mass, not unlike ice in its appearance. From the same solution, when very slowly cooled, the salt is deposited in regularly shaped figures, termed *crystals*. Those saline solutions also, which retain the same quantity of salt at all temperatures, and from which therefore crystals cannot be obtained by cooling, afford either an irregular mass or perfect crystals, accordingly as the solvent is removed by rapid or slow evaporation. In order, then, to produce regular arrangements of the particles of a dissolved solid, the cause, overcoming their cohesive affinity, must be very slowly and gradually withdrawn.

There are other circumstances, which influence the cohesive affinity of dissolved solids, and determine it to become effective. Of this kind is the introduction of another solid, either of the same or of a different kind into the fluid mass. On this principle, pieces of wood are often introduced into a solution, to promote crystallization; and, for the same reason, the first crystals are formed upon the sides of the containing vessel. But a still more effectual method of inducing crystallization, is to immerse, in the solution, a crystal *of the same kind* with that which we expect to be formed. The crystal, thus exposed, receives successive additions to its several surfaces, and preserves its form, with a considerable addition to its magnitude. This curious fact was originally noticed by Le Blanc, who has founded on it a method of obtaining large and perfect crystals.

Another circumstance, promoting the cohesion of dissolved solids, is that of pressure on the surface of the solution. A saturated solution of a salt may be prepared under a diminished pressure, which will immediately become solid on restoring the weight of the atmosphere. This can only be explained on the supposition, that by the force of extraneous pressure, the particles of the solid are brought nearer to each other, and more within the sphere of cohesive attraction. On the same principle, liquids must necessarily have their solvent power increased by a diminished pressure of the air, and perhaps in an appreciable degree, even by common barometrical changes.

The cohesion of the particles of a solid, which has separated from its solvent, is never so powerful as to unite them to the entire exclusion of the fluid, a portion of which enters into the composition of a crystal, and is essential to the regularity of its form. In the instance of salts, a quantity of water is almost invariably thus combined, and is termed their *water of crystallization.* Its proportion is variable in different salts. In some it is extremely small; in others it constitutes the principal part of the salt, and is even so abundant as to liquefy them on the application of heat, producing what is called the *watery fusion.* The water of crystallization is retained, also, by different salts, with very different degrees of force. Some crystals abandon their water by mere exposure to the atmosphere, and thus lose their transparency, and in time, their form. Such salts are said to *effloresce.* Others, on the contrary, not only hold their water of crystallization very strongly combined, but eagerly attract more, and on exposure to the atmosphere become liquid, or *deliquiate.* The property itself is called *deliquescence.*

It is not, however, every solid that is capable of crystallizing from its solution in water; for the affinity of some saline bodies for water is so strong, as to overcome, under all circumstances, the cohesive attraction. Even these salts, however, may be brought to crystallize, by dissolving them in a fluid which they attract less strongly than water. Thus several salts of this kind are separated in a crystalline form, from their solution in alcohol.

Since the cohesive affinity is a property which belongs to solids in very different degrees, and since solids crystallize merely in consequence of the predominance of this affinity, it must be obvious that when two solids are dissolved in one fluid, they may be separated from each other by their different tendencies to cohesion. That solid, whose particles have the greatest cohesive attraction, will separate first; and the other, which is less disposed to crystallize, may be afterwards obtained by reducing the quantity or temperature of the solvent. The separation, however, of the two solids from each other is seldom, if ever, perfect, on account of their mutual affinity for each other. When nitre and common salt, for example, exist together in the same solution, after separating most of the nitre by its greater disposition to crystallize, there still remains a portion in the saline solution. The crystals, also, which we obtain at first, are not pure nitre, but consist of that salt combined with a portion of common salt.

Every solid, that is susceptible of crystallization, has a tendency to assume a peculiar shape. Thus common salt, when most perfectly crystallized, forms regular cubes; nitre has the shape of a six-sided prism; and alum that of an octaedron. The same solid, however, admits of variations of its crystalline figure. Calcareous spar, for example, forms six-sided prisms, and three or six-sided pyramids. These varieties are occasioned by accidental circumstances, which modify the operation of the force of cohesion. The diversities are, on first view, extremely numerous; and yet, upon a careful examination and comparison, they are found to be reducible to a small number of simple figures.

The attempt to trace all the observed forms of crystals to a few simple or primary ones, seems to have originated with Bergman.* In the instance of calcareous spar, this distinguished chemist demonstrated that its numerous modifications may possibly result from one simple figure, the rhomb, by the accumulation of which, in various ways, crystals of the most opposite forms may be generated. This theory he extended to crystals of every kind; and he accounted for the differences of their external figures, by varieties of their mechanical elements or minute molecules.

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* Bergman's Essays, vol. ii.

About the same period with Bergman, or immediately afterwards, M. Romé de l'Isle pursued still farther the theory of the structure of crystals. He reduced the study of crystallography to principles more exact, and more consistent with observation. He classed together, as much as he was able, crystals of the same nature. From among the different forms belonging to the same species, he selected, for the primitive form, one which appeared to him to be the most proper, on account of its simplicity. Supposing this to be truncated in different manners, he deduced the other forms, and established a certain gradation, or series of passages, from the primitive form to complicated figures, which on first view would scarcely appear to have any connexion with it. To the descriptions and figures of the primitive forms, he added the mechanical measurement of the principal angles, and showed that these angles are constantly the same in each variety. It must be acknowledged, however, that the primitive forms, assumed by this philosopher, were entirely imaginary, and not the result of any experimental analysis. His method was to frame an hypothesis; and then to examine its coincidence with actual appearances. On his principles any form might have been the primitive one, and any other have been deduced from it.

It was reserved for the sagacity of the Abbe Haüy to unfold the true theory of the structure of crystals, and to support it both by experimental and mathematical evidence. By the mechanical division of a complicated crystal, he first obtains the simple form, and afterwards constructs, by the varied accumulation of the primitive figure, according to mathematical synthesis, all the observed varieties of that species.

Every crystal may be divided by means of proper instruments; and, if split in certain directions, presents plane and smooth surfaces. If split in other directions, the fracture is rugged, is the mere effect of violence, and is not guided by the natural joining of the crystal. This fact had been long known to jewellers and lapidaries; and an accidental observation of it proved, to the Abbe Haüy, the key of the whole theory of crystallization. By the skilful division of a six-sided prism of calcareous spar, he reduced it to a rhomb, precisely resembling that which is known under the name of Iceland crystal. Other forms of calcareous spar were subjected to the same operation; and, however different at the outset, finally agreed in yielding, as the last product, a rhomboidal solid. It was discovered also by Haüy, that if we take a crystal of another kind (the cubic fluor spar for instance), the nucleus, obtained by its mechanical division, will have a different figure, *viz.* an octaedron. Other crystallized bodies produce still different forms; which are not, however, very numerous. Those, which have hitherto been discovered, are reducible to six; the parallelopipedon, which includes the cube, the rhomb, and all the solids which are terminated by six faces, parallel two and two; the tetraedron; the octaedron; the regular hexaedral prism; the dodecaedron with equal and similar rhomboidal planes; and the dodecaedron with triangular planes.

The solid of the primitive form or nucleus of a crystal, obtained by mechanical division, may be subdivided in a direction parallel to its different faces. All the sections thus produced being similar, the resulting solids are precisely similar in shape to the nucleus, and differ from it only in size, which continues to decrease as the division is carried farther. To this division, however, there must be a limit, beyond which we should come to particles so small, that they could no longer be divided. At this term, therefore, we must stop; and to these last particles, the result of an analysis of the primitive nucleus, and similar to it in shape, Haüy has given the name of the *integrant molecule.* If the division of the nucleus can be carried on in other directions than parallel to its faces, the integral molecule may then have a figure different from that of the nucleus. The forms, however, of the integrant molecule, which have hitherto been discovered, are only three; the tetraedron, the simplest of pyramids; the triangular prism, the simplest of prisms; and the parallelopipedon, including the cube and rhomboid, the simplest of solids which have their faces parallel two and two.

The primitive form, and that of the integral molecule having been experimentally determined by the dissection of a crystal, the next step is to discover the law, according to which these molecules are arranged, in order to produce, by their accumulation around the primitive figure, the great variety of secondary forms. What is most important in the discoveries of Haüy, and what constitutes in fact the essence of his theory, is the determination of these laws, and the precise measurement of their action. He has shown that all the parts of a secondary crystal, superadded to the primitive nucleus, consist of laminæ, which decrease gradually by the subtraction of one or more layers of integrant molecules; so that theory is capable of determining the number of these ranges, and, by a necessary consequence, the exact form of the secondary crystal.

By the development of these laws of decrement, Haüy has admirably shown how, from variations of the arrangement of the integrant molecules, a great variety of secondary figures may be produced. Their explanation, however, would involve a minuteness of detail, altogether unsuitable to the purpose of this work; and I refer, therefore, for a very perspicuous statement of them to the first and ninth volumes of the Philosophical Magazine.

Such are the principal views which interest the chemical philosopher in the consideration of the force of cohesion. He has sometimes to regard it as a cause impeding or modifying the action of chemical affinity; and sometimes as producing the re-union of a solid, either in a regular or irregular form, whose particles have been separated by the predominance of chemical affinity. The cohesive attraction is, therefore, to be considered as the sole cause of the crystallization of solids. It appears to be exerted between the integrant molecules; for these are most probably the bodies that are suspended in a fluid, from which crystallization is about to happen. To these molecules, the property of polarity has been ascribed, or a tendency to arrange themselves by certain sides in preference to others. The explanation, however, is too hypothetical to be received as a satisfactory account of the process of crystallization. It assigns, moreover, a cause, the relative force of which, in different cases, we have no means of measuring or even of estimating.

The influence of cohesion over chemical affinity, it will appear however from the sequel, is much more extensive than can be inferred from the foregoing observations. To the full development of this interference, it is essential that the general laws of chemical affinity should be understood, and to the explanation of these I shall proceed in the following section.

SECTION II.

Of Chemical Affinity, and the General Phenomena of Chemical Action.

CHEMICAL affinity, like the cohesive attraction, is effected only at insensible distances; but it is distinguished, from the latter force, in being exerted between the particles of bodies of *different kinds.* The result of its action is not a mere aggregate, having the same properties as the separate parts, and differing only by its greater quantity or mass, but a new compound, in which the properties of the components have either entirely or partly disappeared, and in which new qualities are also apparent. The combinations effected by chemical affinity are permanent, and are destroyed only by the interference of a more powerful force, either of the same or of a different kind.

As a general exemplification of chemical action, we may assume that which takes place between potash and sulphuric acid. In their separate state, each of these bodies is distinguished by peculiarities of taste and other qualities. The alkali, on being added to blue vegetable infusions, changes their colour to green; and the acid turns them red. But if we add the one substance to the other, very cautiously and in small quantities, examining the effect of each addition, we shall at length attain a certain point, at which the liquid will possess neither acid nor alkaline qualities; the taste will be converted into a bitter one; and the mixture will produce no effect on the blue vegetable colours. Here then, the qualities of the constituent parts, or at least some of their most important ones, are destroyed by combination. When opposing properties thus disappear, the bodies combined have been said to *saturate* each other; and the precise term, at which this takes place, has been called the *point of saturation.* It is adviseable, however, to restrict this expression to weaker combinations, where there is no remarkable alteration of qualities, as in cases of solution; and to apply, to those results of more energetic affinities, which are attended with loss of properties, the term *neutralization.*

At the same time that the properties of bodies disappear on combination, other new qualities, both sensible and chemical, are acquired; and the affinities of the components for other substances become in some cases increased, in others, diminished in energy. Sulphur, for example, is destitute of taste, smell, or action on vegetable colours; and oxygen gas is, in these respects, equally inefficient. But the compound of sulphur and oxygen is intensely acid; the minutest portion instantly reddens blue vegetable infusions, and the acid is disposed to enter into combination with a variety of bodies, for which its components evinced no affinity. Facts of this kind sufficiently confute the opinion of the older chemists, that the properties of compounds are *intermediate* between those of their component parts; for in instances like the foregoing the compound has qualities, not a vestige of which can be traced to either of its elements.

Besides the alteration of properties, which usually accompanies chemical action, there are certain other phenomena, which are generally observed to attend it. Of these the principal are change of form, change of density, and change of temperature.

1. *Change of form* may take place under various modifications. Thus two or more solid bodies, by combination, may become liquid or aëriform, or the reverse. A

solid and an aëriform body, by chemical union, may produce a liquid, as in the instance of sulphur and oxygen; two gases may produce a solid or liquid; or two liquids a solid.

2. *Change of density* is an almost invariable concomitant of chemical union. In most instances, the density of bodies is increased by combination, or is greater than the mean density of the components. If equal weights of sulphuric acid and of water be mixed, the resulting density is not the mean, but considerably greater. In some cases, a change of form and of density happen together. Thus a salt, when dissolved in water, has exchanged the solid for the fluid form, and its density is diminished. The solvent, on the contrary, has experienced an increase of density; but the former effect more than counterbalances the latter, and, on the whole, an enlargement of volume, or diminution of density, is the result. The same thing happens with certain combinations of the metals, the compound of which has a specific gravity less than that of the components.

3. *Change of temperature* is observed in almost every instance of chemical action; and this, like density, may either be increased or diminished. During the solution of certain solids in water, the thermometer indicates generally a production of cold, or in other words a diminished temperature. When sulphuric acid and water are suddenly mixed, a contrary change happens, and a considerable increase of temperature results. The cause of these changes cannot be explained, until the subject of the various quantities of caloric in bodies has been considered. It is sufficient at present to observe that the alteration of temperature is invariably accompanied either by a change of form, of density, or of both.

Chemical affinity may be exerted, 1stly, between two *simple* substances, or among three, four, or more, giving rise to what may be termed *primary* compounds, or compounds of elementary substances with each other. 2dly, It may be effective between simple bodies and compounds; or, 3dly, between compounds and other compounds. By the combinations of elementary bodies with each other, either new affinities are acquired, or at least existing ones are for the first time rendered apparent. Oxygen, for example, has no affinity for lime; but the compound of oxygen and sulphur is powerfully attracted by that earth, and we obtain a compound of sulphur, oxygen, and lime. In this example, and in all similar ones, it becomes an interesting question, whether the resulting combination is to be ascribed to affinities exerted by the compound (sulphuric acid) as such; or, to the affinities of the constituent principles rendered active by combination. If the latter be the true conclusion (as will appear extremely probably in the sequel), it will follow that we cannot, in any instance, infer from the want of action between two substances, that they have no affinity; and that, on the contrary, affinity may exist when not manifested by any effect. In the latter case, its activity is suppressed by counteracting forces, the nature of which will presently be considered.

Chemical affinity, when it actually exists between two bodies, is not exerted with the same energy in all proportions. 1stly, Some bodies unite in one proportion only. Thus hydrogen and oxygen combine only in the proportions that constitute water, which are nearly 15 of the former to 85 of the latter. 2dly, Other bodies combine in several proportions; but these are definite; and, in the intermediate ones, no combination ensues. Oxygen and nitrogen, for instance, unite in the proportions of 37 to 63, 56 to 44, and 70½ to 29½, but cannot be combined in any other. 3dly, In some cases, combination takes place in all proportions to a certain extent, beyond which it cannot be carried. This is best exemplified by the solution of a salt in water, which may take place in all the intermediate degrees to the point of saturation. 4thly, In other instances, bodies unite without limitation in all proportions; but the affinities, exerted in these cases, are far from being energetic, and the bodies combined undergo little alteration of properties. Thus sulphuric acid unites with any quantity of water; and the characteristic properties of the acid (its taste and action on vegetable colours) continue to be apparent in the compound.

SECTION III.

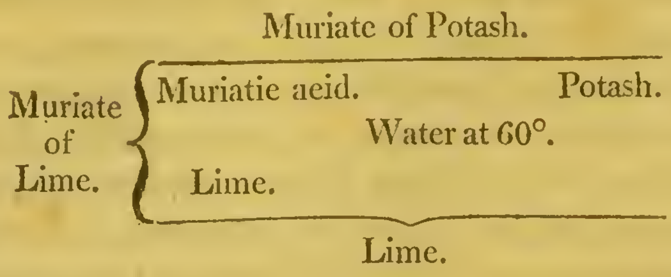
Of Elective Affinity.

An important law of affinity, which is the basis of almost all chemical theory, is, that one body has not the same force of affinity towards a number of others, but attracts them unequally. Thus A will combine with B in preference to C, even when these two bodies are presented to it under equally favourable circumstances. Or, when A is united with C, the application of B will detach A from C, and we shall have a new compound consisting of A and B, C being set at liberty. Such cases are examples of what is termed in chemistry *simple decomposition*, by which it is to be understood that a body acts upon a compound, and unites with one of its constituents; leaving the other at liberty. And as the forces of affinity of one body to a number of others vary, this body has been considered as making an election; and the affinity has been called *single elective affinity.* Thus if to the muriate of lime, consisting of lime and muriatic acid, we add potash, the muriatic acid exerts a stronger elective affinity for the potash than for the lime; and the lime falls down in the state of a powder, or is *precipitated.* Of facts of this kind a great variety have been comprehended in the form of tables, the first idea of which seems to have occurred, nearly a century ago, to Geoffrey, a celebrated French chemist. The substance, whose affinities are to be expressed, is placed at the head of a column, and is separated from the rest by a horizontal line. Beneath this line are arranged the bodies, with which it is capable of combining, in the order of their respective forces of affinity; the substance, which it attracts most strongly, being placed nearest to it, and that, for which it has the least affinity, at the bottom of the column. The affinities of muriatic acid, for example, are exhibited by the following plan.

MURIATIC ACID.

Barytes,
Potash,
Soda,
Lime,
Ammonia,
Magnesia,
&c. &c.

Simple decompositions may be expressed, also, by another form, contrived by Bergman. Thus the following scheme illustrates the decomposition of muriate of lime by potash.



The original compound (muriate of lime) is placed on the outside and to the left of the vertical bracket. The included space contains the original principles of the compound, and also the body which is added to produce decomposition. Above and below the horizontal lines, are placed the results of their action. The point of the lower horizontal line being turned downwards, denotes that the lime falls down or is precipitated; and the upper line, being perfectly straight, shows, that the muriate of potash remains in solution. If both the bodies had remained in solution, they would both have been placed above the upper line; or, if both had been precipitated, beneath the lower one. If either one or both had escaped in a volatile form, this would have been expressed by placing the volatilized substance above the diagram, and turning upwards the middle of the upper horizontal line. But since decompositions vary under different circumstances, it is necessary to denote, by the proper addition to the scheme, that the bodies are dissolved in water of the temperature of 60°.

No chemical facts can appear, on first view, more simple or intelligible, than those which are explained by the operation of single elective affinity. It will be found, however, on a more minute examination, that this force, abstractedly considered, is only one of several causes which are concerned in chemical decompositions, and that its action is modified, and sometimes even subverted, by counteracting forces.

SECTION IV.

Of the Causes, which modify the action of Chemical Affinity.

Chemical affinity is by no means an uniform force, accompanied, when exerted between the same substances, with invariably the same effects; but these effects are modified by other circumstances. The apparent exceptions were first attentively examined and explained by Bergman; but it is to the recent labours of Berthollet, that we are indebted for their full elucidation. The views of this distinguished philosopher have effected, indeed, a complete revolution in our notions of affinity; and have shown that it is a force much less exclusively concerned, than was formerly imagined, in the production of chemical phenomena.

The circumstances, which modify the exertion of chemical affinity, are the following:

1. *Cohesion.* The influence of this force over chemical affinity has already been in part explained in a preceding section; and other illustrations of its interference will be given, when we consider the subject of the limitations to chemical combination.

2. *Quantity of matter.* If affinity were an invariable

force, a body A, united by a strong affinity to another B, could never be separated from it by a third body C, for which A has a weaker affinity. Or if the three bodies A, B, and C, were mixed together, under certain circumstances, A should combine with B to the entire exclusion of C. In neither of these instances, however, does such an exclusive action take place; for if the bodies A B and C be brought into contact, we have a combination of A with B, and another of A with C, in proportions regulated by the quantities of B and C. Again, the compound A B may be partly decomposed by C. Muriatic acid, for instance, has a stronger affinity for barytes than for potash; and yet, in a mixture of the three bodies, the acid unites not only with barytes but with potash, even though more barytes may be present than is necessary to neutralize the whole of the acid. The combination of muriatic acid and barytes is, also, partly decomposed by potash.

It is found, moreover, by experiment, that when two bodies, having different affinities for a third, are brought into contact with it, under equal circumstances, this third body is divided between them, not only in proportion to the energy of their affinities, but also in proportion to their quantities. A larger quantity, therefore, may compensate a weaker affinity; and the reverse. To obtain, in such cases, a measure of the action of two bodies on a third, if their respective affinities were precisely determined, it would only be necessary to multiply the number indicating the affinity by the quantity. To the product of this multiplication Berthollet has given the appellation of *mass*. Thus let us suppose, for the sake of illustration (what is not accurate in fact), that the affinity of barytes for muriatic acid is twice as strong as that of potash, or that these affinities are respectively denoted by the numbers 4 and 2. In this case the same *mass* * will result from 4 parts of barytes as from 8 of potash; because the same product (16) is obtained, in each instance, by multiplying the number indicating the affinity into that denoting the quantity; for 4 (the affinity of barytes) multiplied by 4 (the quantity assumed in this example) is equal to 16; and 2 (the affinity of potash) multiplied by 8 (its quantity) is also equal to 16. In this case, therefore, to divide equally a portion of muriatic acid between barytes and potash, these two bodies should be employed in the proportion of 2 of the former to 4 of the latter.

Since the affinity of any substance for another is influenced by the circumstance of quantity, it might be expected that affinity must decrease, as we approach the term of saturation or neutralization; and this is perfectly conformable to fact. Hence was deduced the law, that the *force of affinity is in the inverse proportion of neutralization.* As an exemplification of the law, we may suppose the quantity of potash, required to neutralize a given weight of sulphuric acid, to be divided into 100 equal parts, and each ten of these to be added in succession. If the initial affinity of the acid for the alkali be denoted by 100, the first ten parts will be combined with a force indicated by that number; the second ten with the force only of 90; the third ten with the force of 80; and so on, in a diminishing progression, as we approach the point of neutralization. It is scarcely necessary to add that these numbers are not intended to express accurately the force of affinity; but are arbitrarily assumed for the purpose of illustration.

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* This term is objectionable, on account of the different meaning, which is affixed to it in mechanical philosophy.

The influence of quantity explains, also, the difficulty which is observed in effecting, in any instance, the total decomposition of a compound of two principles by means of a third. The immediate effect of a third body C, when added to a compound A B, is to abstract from B a portion of the substance A; and consequently a portion of B is set at liberty, the attraction of which for A, is opposed to that of the uncombined part of C. The farther this decomposition is carried, the greater will be the proportion of B which is brought into an uncombined state; and the more powerfully will it oppose any farther tendency of C to detach the substance A. At a certain point, the affinities of B and C for A will be exactly balanced, and the decomposition will proceed no farther. In a few cases, it must be confessed, a third body separates the whole of one of the principles of a compound; but this happens in consequence of the operation of extraneous forces, the nature of which will presently be explained.

3. *Insolubility* is another force, which essentially modifies the exertion of affinity. It is to be considered, indeed, merely as the result of cohesion, with respect to the liquid in which the effect takes place.

Insolubility may be regarded as opposing an obstacle to chemical action, both by the cohesion, of which it is an effect, and by its preventing that contiguity, which is essential to the exertion of affinity. The latter impediment we endeavour to overcome by the mechanical diffusion of an insoluble substance through the fluid, which is intended to act upon it; and when the new compound has considerable solubility, this expedient is sufficiently effectual. Thus lime, by mechanical diffusion in muriatic acid, may be easily brought into permanent union. But if in attempting to combine two bodies, the new compound be also insoluble, the combination is effected with considerable difficulty, and we can scarcely bring it to the point of saturation; as in the case of sulphuric acid and lime.

When a soluble substance and an insoluble one are presented, at the same time, to a third, for which they have nearly an equal affinity, the soluble body is brought into the sphere of action with great advantages over its antagonist. Its cohesion at the outset is but little, and by solution is reduced almost to nothing; while that of the insoluble body remains the same. The whole of the soluble substance also exerts its affinity at once; while a part only of the insoluble one can oppose its force. Hence the soluble substance may prevail, and may attach to itself the greatest proportion of the third body, even though it has a weaker affinity than the insoluble one to the subject of combination.

Insolubility, however, under certain circumstances, is a force which turns the balance in favour of the affinity of one body when opposed to the affinity of another. For example, if to the soluble compound, sulphate of soda, we add barytes, the new compound, sulphate of barytes, is precipitated the instant it is formed; and being removed from the sphere of action, the soda can exert no effect upon it by its greater quantity or mass. For the same reason, when soda is added to sulphate of barytes, the latter is protected from decomposition both by its insolubility and by its cohesion.

These facts sufficiently prove that the order of precipitation, which was formerly assumed as the basis of tables of elective affinity, can no longer be considered as an accurate measure of that force; and that the body, which is precipitated, may, in some cases, be superior in affinity to the one which has caused precipitation. In these cases, a trifling superiority in affinity may be more than counterbalanced by the cohesive force, which causes insolubility.

4. *Great specific gravity* is a force, which must concur with insolubility or cohesion in originally impeding combination; and when chemical union has taken place, it must come in aid of affinity, by removing the new compound from the sphere of action. It is scarcely necessary to enlarge on the operation of a force, the nature of which must be so obvious.

5. *Elasticity*. Cohesion, it has already been stated, may prove an impediment to combination; and, on the other hand, it is possible that the particles of bodies may be separated so widely, as to be removed out of the sphere of their mutual attraction. Such appears to be the fact with regard to a class of bodies called airs or gases. The bases of several of these have powerful attractions for the bases of others, and for various liquids, and yet they do not combine on simple admixture. Thus elastic aëriform bodies, again, may unite with each other or with liquids to a certain point only, at which affinity is balanced by elasticity; for as we approach saturation, affinity is gradually diminished, while elasticity remains the same. The example, which Berthollet has given in illustration of this principle, is the affinity of water for carbonic acid, which occasions it to combine with a certain quantity. But as we approach saturation, this affinity is gradually diminished, while the elasticity of the gas remains the same, and the absorption can go on no longer. If the elasticity be counteracted by mechanical pressure, the affinity of the water for the acid gas again becomes efficient. Such, at least, is the explanation given by Berthollet; but the fact, which

is indubitable, may be explained by causes altogether mechanical.

Again if two bodies, one of which has an elastic and the other a liquid form, be presented at the same time to a solid, for which they have both an affinity, the solid will unite with the liquid in preference to the gas. Or if we add to the compound of an elastic substance with an inelastic one, a third body also inelastic, the two latter combine to the exclusion of the former. For example, if to the compound of potash and carbonic acid we add sulphuric acid, the latter acid, acting both by its affinity and its quantity, disengages a portion of carbonic acid. This, by its elasticity, is removed from the sphere of action, and presents no obstacle to the farther operation of the sulphuric acid. Hence elastic bodies act only by their affinity; whereas liquids act both by their affinity and quantity conjoined. And though the affinity of the liquid, abstractedly considered, may be inferior to the affinity of the elastic body, yet, united with quantity, it prevails. In the above instances, the whole of the elastic acid may be expelled by the fixed acid; whereas, as it has already been observed, decomposition is incomplete, if the substance which is liberated remain within the sphere of action.

6. *Efflorescence* is a circumstance which occasionally influences the exertion of affinity; but this is only of very rare occurrence. The simplest example of it is that of lime, and muriate of soda. When a paste composed of these two substances with a great excess of lime, is exposed, in a moist state, to the air, the lime, acting by its quantity, disengages soda from the common salt, which appears in a dry form, on the outer surface of the paste, united with carbonic acid absorbed from the atmosphere. In this case the soda, which is separated, being removed

from contiguity with the interior part of the mass, presents no obstacle to the farther action of the lime, and the decomposition is carried farther than it would have been, had no such removal happened.

6. The influence of *temperature* over chemical affinity is extremely extensive and important; but at present a very general statement only of its effects is required. In some cases an increased temperature acts in promoting, and at others in impeding, chemical combination: and it materially affects also the order of decompositions.

An increased temperature promotes chemical union by diminishing or overcoming cohesion. Thus metals unite by fusion, and salts are rendered more soluble in water. Whenever heat is an obstacle to combination, it produces its effect by increasing elasticity. Hence water absorbs a less proportion of gas at a high than at a low temperature. A reduction of the temperature of elastic bodies, by lessening their elasticity, facilitates their union with other substances. In certain cases, an increased temperature has the combined effects of diminishing cohesion and increasing elasticity. When sulphur is exposed to oxygen gas, no combination ensues, until the sulphur is heated; and though the elasticity of the gas is thus increased, yet the diminution of cohesion of the solid is more than proportionate, and chemical union ensues between the two bodies.

Such are the most important circumstances, that modify the exertion of chemical affinity. Of their influence, sufficient illustrations have been given to prove, that in every case of combination and decomposition, we are not to consider the force of affinity abstractedly; but are to take into account the agency of other powers, as cohesion, quantity, insolubility, elasticity, efflorescence, and temperature. By the action of these extraneous powers, Berthollet has endeavoured to explain certain facts which are not easily understood on any other principle. Of these the most important are, 1stly, the establishment of proportions in chemical compounds; and 2dly, the modification produced in the affinities of bodies by chemical union.

1. Independently of these extraneous forces, Berthollet imagines that there are no limits to combination, or that two bodies, which are now susceptible of union only in one or in few proportions, might, if these forces were annihilated, be united in every proportion. The causes which he has assigned, as chiefly regulating proportion, are cohesion and elasticity. To take one of the simplest cases, the proportion, in which a salt can be combined with water, depends on the balance between the chemical affinity of the bodies for each other, and the cohesive attraction of the salt. In this case, then, cohesion is the limiting power. As an example of the influence of this force when more energetic affinities are exerted, if we add to diluted sulphuric acid a solution of barytes, a compound is formed, consisting of sulphuric acid and barytes, which, in consequence of its great insolubility or cohesion, is instantly removed from contact with the redundant acid, and with established proportions.

The agency of elasticity in limiting proportion, may be exemplified by the combination of hydrogen and oxygen. If a mixture of the two gases be inflamed, the new compound, water, is immediately separated, from what is superfluous of both ingredients, by its superior density. In other instances, the bases of aëriform substances are combined in various proportions, and in such examples, there are several terms of greatest condensation, as in the case of oxygen and nitrogen.

2. Another important part of the theory of Berthollet is, that the affinities of a compound are not newly acquired; but are merely the modified affinities of its constituents, the action of which, in their separate state, was counteracted by the prevalence of opposing forces. By combination, these forces are so far overcome, that the affinities of the constituents are enabled to exert themselves.

The action of different affinities existing in one compound, Berthollet terms *resulting affinities*, while the individual affinities of the constituents he calls *elementary affinities*. Thus nitric acid acts on potash by an affinity, which results from those of oxygen and azote for potash. And as all affinity is mutual, the term *resulting affinity* is applied, also to that force, with which a simple body acts on a compound; to the affinity for example, which any simple body may exert on nitric acid. A simple body, indeed, may exert towards a compound both an elementary and resulting affinity. If the elementary affinity prevails, it will unite only with one of the principles of the compound, as when a simple body, by its affinity for oxygen, decomposes nitric acid, and liberates its nitrogen in a separate form. If the resulting affinity be predominant, the simple body will unite with the whole compound without effecting any disunion of its elements.

From these views it may be inferred, that we are not, in any case, to deny the existence of an affinity between two bodies, merely because they do not combine when presented to each other; for an affinity may exist, but may be suppressed by the prevalence of opposing forces. According to the doctrine of Berthollet, affinity is a force exerted by every body towards every other; even though not made apparent by any effect. On this principle, we are able to explain certain phenomena, which are wholly unintelligible on any other, and especially those which have been referred to *disposing affinity*. The action of sulphuret of potash, for example, on oxygen gas, has been ascribed to the disposing affinity of potash for sulphuric acid. This, however, is ascribing an affinity to a compound before that compound has existence. It is much more probable, that besides the diminished cohesion of the sulphur, the affinity of potash for oxygen has some share in producing the combination. On this principle the united affinities of the potash and sulphur for oxygen (in other words the *resulting affinities* of the sulphuret of potash) are the efficient causes of chemical union. This explanation, however, at least, does not, like the theory of disposing affinities, involve an absurdity.

In opposition to the theory that chemical affinity has a tendency to unite bodies in unlimited proportions, an hypothesis has lately been proposed by Mr. Dalton*, which appears more consonant to the general simplicity of nature. When two elements unite to form a new compound, he takes it for granted, unless some sufficient reason can be given to the contrary, that one ultimate particle or atom of the one unites with one ultimate particle or atom of the other. In this way, an atom of oxygen united with an atom of hydrogen forms water; and an atom of hydrogen, combined with an atom of nitrogen, constitutes ammonia. When only one compound of any two elements can be obtained, he presumes, unless the contrary can be proved, that the combination is of the kind which has been already described, and which he calls *binary*. But if several compounds can be obtained from the same elements, they combine, he supposes, in proportions, which are expressed by some simple multiple of the number of atoms. The following table exhibits a view of some of these combinations.

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* "New System of Chemical Philosophy." 8vo. London. 1808.

1 atom of A + 1 atom of B = 1 atom of C, binary.
1 atom of A + 2 atoms of B = 1 atom of D, ternary.
2 atoms of A + 1 atom of B = 1 atom of E, ternary.
1 atom of A + 3 atoms of B = 1 atom of F, quaternary.
3 atoms of A + 1 atom of B = 1 atom of G, quaternary,
&c. &c.

This theory Mr. Dalton has not confined to the combinations of bodies generally deemed simple, but has extended to those which take place between simple bodies and compounds, or between compounds and other compounds. Thus the neutral salts, which result from the combination of the same base with different proportions of acid, he has included under the same general law. If, for example, any neutral salt consist of an atom of base united to one atom of acid, the superacid salt will consist of an atom of base united with two atoms of acid. This doctrine, it must be confessed, cannot at present be regarded in any other light than that of an hypothesis. But it is an hypothesis which has been developed with great ingenuity and patience of investigation, and which is supported by many striking and daily increasing analogies.

SECTION V.

Of the Estimation of the Forces of Affinity.

The affinities of one body for a number of others are not all of the same degree of force. This is all that the present state of our knowledge authorizes us to affirm; for we are ignorant *how much* the affinity of one body for another is superior to that of a third. The determination of the precise forces of affinity would be an important step in chemical philosophy; for its phenomena would then be reduced to calculation; and we should be enabled to anticipate the result of experiment.

The observed order of decomposition, it has already been stated, does not enable us to assign the order of the forces of affinity; because, in all decompositions, other forces are concerned. We are, therefore, obliged to seek some other method of determining the problem. Of these several have been proposed.

When the surface of one body is brought into contact with another surface of the same kind, as when the smooth surfaces of a divided leaden bullet are pressed together, they adhere by the force of cohesion, their particles being all of *the same kind.* But when the surfaces of different bodies are thus brought into apparent contact, it is reasonable to suppose that their adhesion arises from chemical affinity, because their particles are of *different kinds.* Guyton proposed, therefore, the comparative force, with which different surfaces adhere, as a competent measure of chemical affinity. His experiments were made on plates of different metals, of precisely the same size and form, suspended by their centres from the arm of a sensible balance. The lower surfaces of these plates were successively brought into contact with mercury, which was changed for each experiment, and the weight was observed, which it was necessary to add to the opposite scale, in order to detach the several metals. Those which required the largest weight were inferred to have the greatest affinity; and it is remarkable that the order of affinities, as determined in this way, correspond with the affinities as ascertained by other methods. The following were the results:

Gold adhered to mercury with a force of .	446 grains.
Silver	429
Tin	418
Lead	397
Bismuth	372
Zinc	204
Copper	142
Antimony	126
Iron	115
Cobalt	8

This method, it must be obvious, is of too limited application to be of much utility; for few bodies have the mechanical conditions, which can enable us to subject them to such a test. How, for example, could the affinities of acids for alkalis be examined on this principle? It may be doubted, also, whether in the cases to which it may be applied, it does not measure the facility of combination, rather than the actual force of affinity.

To determine the absolute forces, which one body exerts towards a number of others, Mr. Kirwan has proposed the quantity of each, which is required to produce neutralization. This he has ascertained by experiment in a great variety of instances, a few of which are contained in the following tables.

100 Parts of SULPHURIC ACID require for Neutralization	
200	parts of barytes.
138	- - of strontites.
121	- - of potash.
78	- - of soda.
70	- - of lime.
57	- - of magnesia.
26	- - of ammonia.

100 Parts of POTASH require	
105	of carbonic acid.
85	of nitric acid.
82½	of sulphuric.
56	of muriatic.

In judging of the affinities of the same acid for different bases, Mr. Kirwan assumes that they are represented by the numbers indicating the quantities of each base required for neutralization. Thus, because 100 parts of sulphuric acid neutralize 200 of barytes, and 121 of potash, the affinity of the former is superior to that of the latter in the proportion of 200 to 121. So far the inference corresponds with the order of decomposition; for barytes takes sulphuric acid from potash. But if we examine the affinities of potash, as represented in the second table, we shall find that, on this principle, they are directly contradictory to fact. Thus the affinity of sulphuric acid should be greatly inferior to that of the carbonic; whereas it is well known that the former displaces the latter from all its combinations. Mr. Kirwan was, therefore, driven to the necessity of establishing a precisely opposite rule in determining the affinities of different acids for the same base, and of assuming that they are *inversely* proportionate to the affinity of the saturating acid. Thus the affinity of carbonic acid for potash would be 82½, and that of sulphuric acid 105. This, however, involves a contradiction; since it is implied that a stronger affinity, in one instance, requires a greater quantity of the saturating principle, as in the relation of barytes and potash to sulphuric acid; and that, in the other, it requires a less quantity, as in the instance of the sulphuric and carbonic acids with respect to potash.

Neutralization is plainly an effect of chemical affinity, and must in all cases bear a proportion to its cause. It has been assumed, therefore, with much greater probability, by Berthollet, that the substance which, in the *smallest quantity*, neutralizes another, is the one possessing the strongest affinity. On this principle, the affinities of sulphuric acid for different bases, will be exactly the reverse of the order established by Mr. Kirwan; and to that order, which would have been assigned from observed decompositions. Thus ammonia will have a stronger affinity for sulphuric acid, than any of the substances which are placed above it in the table; though it is separated, by each of these, from its union with that acid.

It is in the extraneous forces, which have been enumerated as influencing chemical affinity, that we are to seek for the explanation of this apparent anomaly, and especially in those of cohesion and elasticity. The elasticity of ammonia, for example, turns the balance in favour of magnesia, lime, &c. There is an obvious difficulty, however, in the application of the theory. For as the elasticity of ammonia is suppressed by its combination with sulphuric acid, what, it may be asked, but a superior affinity can occasion the first commencement of decomposition?

This difficulty, perhaps, will be removed, if we admit that the affinity of the last portions of ammonia, which produce saturation, is inferior to the affinity of lime, which, being in an insulated state, acts on the sulphuric acid with all the force of its initial affinity. The decomposition then goes on, because the ammonia, being removed from the sphere of action, presents no obstacle by its quantity.

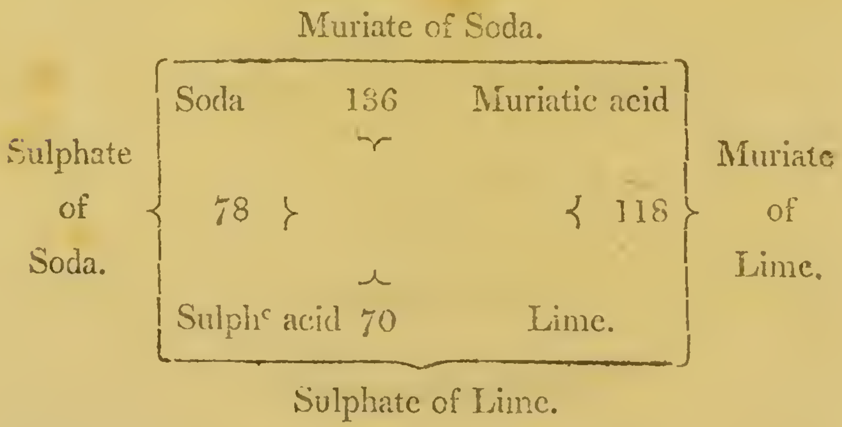
The problem, therefore, of determining the absolute forces of affinity can scarcely be admitted to be solved. Even if it were, we should not be able to predict the order of decomposition, unless the modifying forces of cohesion, elasticity, &c., could be at the same time subjected to precise admeasurement. Until both these objects are accomplished, the results of chemistry can in no case be obtained by calculation, but the science must remain a collection of general principles, derived from experiment and induction.

SECTION VI.

Of Complex Affinity.

UNDER the more general name of *complex affinity,* Berthollet includes that which has hitherto been considered as produced by the action of four affinities, and which has commonly been denominated *double elective affinity.* It frequently happens that the compound of two principles cannot be destroyed either by a third or a fourth separately applied; but if the third and fourth be combined, and placed in contact with the former compound, a decomposition, or a change of principles will ensue. Thus when lime water is added to a solution of the sulphate of soda, no decomposition happens, because the sulphuric acid attracts soda more strongly than it attracts lime. If the muriatic acid be applied to the same compound, still its principles remain undisturbed, because the sulphuric acid attracts soda more strongly than the muriatic. But if the lime and muriatic acid, previously combined, be mixed with the sulphate of soda, a double decomposition is effected.

The lime, quitting the muriatic acid, unites with the sulphuric, and the soda, being separated from the sulphuric acid, combines with the muriatic. These decompositions are rendered more intelligible by the following diagram, contrived by Bergman.



On the outside of the vertical brackets are placed the original compounds; and above and below the diagram, the new compounds. The upper line, being straight, indicates that the muriate of soda remains in solution; and the middle of the lower line, being directed downwards, that the sulphate of lime is precipitated.

In all cases similar to the foregoing, Mr. Kirwan conceives that we may trace the operation of two distinct series of affinities. The affinities tending to preserve the original compounds (which in the above example are those between sulphuric acid and soda, and between muriatic acid and lime), he terms the *quiescent affinities*; because they resist any change of composition. On the other hand the affinities, which tend to disunite the original compounds and to produce new ones (such as those between muriatic acid and soda, and between sulphuric acid and lime), he terms *divellent affinities.* In order that an effect may be produced, the divellent affinities must necessarily be superior to the quiescent. Now assuming the numbers in Mr. Kirwan's tables to express accurately the forces of affinities, the double exchange of principles, which happens in the preceding instance, is readily explained. Thus the quiescent affinities are

Those of lime to muriatic acid =	118
of soda to sulphuric acid =	78
	196

The divellent affinities, opposed to these, consist of

The affinity of soda to muriatic acid =	136
lime to sulphuric acid =	70
	206

The original compound, therefore, is preserved by a force equivalent to 196, and the tendencies to produce new compounds are represented by the number 206. The divellent affinities are, therefore, predominant.

The theory of quiescent and divellent affinities, however, though highly attractive from its simplicity, and from the facility with which it solves certain phenomena, is completely defective in the explanation of others. For example, sulphate of potash is decomposed by muriate of barytes. Yet, estimating in the above manner the quiescent and divellent affinities, an exchange of principles ought not to ensue. The affinities, tending to preserve the original compound, are those of sulphuric acid for potash = 121, and of muriatic acid for barytes = 314. The divellent affinities are that of muriatic acid for potash = 177 + that of sulphuric acid for barytes = 200. The quiescent affinities then are 121 + 314 = 435, and the divellent 177 + 200 = 377. This leaves a balance of 58 in favour of the quiescent affinities; and yet decomposition ensues, when the two compounds are brought into contact.

It must be acknowledged that the numbers, assumed by Mr. Kirwan, do not correspond with the actual forces of affinity. But even if they are taken according to the principle assumed by Berthollet, they will not be found universally applicable. The reason of this is, that the phenomena produced by complex affinity, like those occasioned by simple affinity, are materially influenced by the extraneous forces of cohesion, quantity, elasticity, temperature, &c. The effect of quantity is shown by the fact, that if two salts be mixed together in certain proportions, decomposition will ensue, but not if mixed in other proportions. Thus from the mingled solutions of two parts of muriate of lime and one of nitrate of potash, we obtain muriate of potash; but not from equal weights of the two salts. Insolubility, or precipitation, has also a considerable influence on the result. When this occurs, the influence of quantity is destroyed, as in the case of sulphate of potash and muriate of barytes. Elasticity, and an increased temperature (which operates by increasing elasticity), have also a powerful influence in promoting the action of complex affinities. Thus of four principles, two of which are volatile and two fixed, the two which are volatile will be disposed to unite, in preference to combining with either of those which are fixed. Hence the phenomena of complex decomposition concur with those of a more simple kind, in proving that affinity is not an uniform force, but is materially influenced by various modifying circumstances; and that we cannot confidently anticipate results, from comparing the numerical expressions of quiescent and divellent affinities.

One great obstacle to the construction of tables, capable of representing the forces of affinity, is the difficulty of ascertaining, with precision, the quantities of bodies required for neutralization. Notwithstanding all the care employed by Mr. Kirwan, considerable errors appear to have crept into the results of his experiments. This will sufficiently appear, when they are examined by a test, which has been ingeniously proposed by Guyton. It must be obvious that if between two salts, which are mixed together in solution, decomposition should ensue, and the mixture should afterwards be found neutral, the quantity of acid, which has quitted one of the bases, must have been exactly equivalent to the saturation of the other base, also deserted by its acid. If, for example, we mingle the muriate of magnesia and sulphate of soda, the mixture continues neutral; and hence it follows that the muriatic acid which quits the magnesia must have been exactly equal to the neutralization of the soda deserted by the sulphuric acid. But from a calculation, founded on the proportion of the ingredients of these salts, as established by Mr. Kirwan, it appears that the soda, detached from the sulphuric acid, is not adequate to the saturation of the muriatic acid. The mixture, therefore, ought to be acid;

and since this is contrary to fact, we may safely infer that there is an error in the estimation of the ingredients composing these salts. No tables, indeed, can be correct, unless they stand the test of this mode of verification. Such a table has been calculated by Fischer from the experiments of Richter; but even this table seems in several respects to be of questionable accuracy. I have thought it, however, entitled to a place among the tables in the Appendix.

SECTION VII.

Experimental Illustrations of Chemical Affinity, Solution, &c.

For these experiments, a few wine glasses, or, in preference, deep ale glasses, will be required; and a Florence flask for performing the solutions.

I. *Some bodies have no affinity for each other.*—Oil and water, mercury and water, or powdered chalk and water, when shaken together in a vial, do not combine, the oil or water always rising to the surface, and the mercury or chalk sinking to the bottom.

II. *Examples of chemical affinity, and its most simple effect, viz. solution.*—Sugar or common salt disappears or dissolves in water; chalk in dilute muriatic acid*. Sugar and salt are, therefore, said to be *soluble* in water, and chalk in muriatic acid. The liquid in which the solid disappears, is termed a *solvent.* Chalk or sand, on the contrary, when mixed with water by agitation, always subside again. Hence they are said to be *insoluble.*

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* I omit, purposely, the distinction between the solution and dissolution.

III. *Influence of mechanical division in promoting the action of chemical affinity, or in favouring solution.*—Lumps of chalk or marble dissolve much more slowly in dilute muriatic acid, than equal weights of the same bodies in powder. Muriate of lime, or nitrate of ammonia, cast, after liquefaction by heat, into the shape of a solid sphere, is very slowly dissolved; but with great rapidity when in the state of a powder or of crystals. When a lump of the Derbyshire fluate of lime is immersed in concentrated sulphuric acid, scarcely any action of the two substances on each other takes place; but if the stone be finely pulverized, and then mingled with the acid, a violent action is manifested, by the copious escape of vapours of fluoric acid. In the common arts of life, the rasping and grinding of wood and other substances are familiar examples.

IV. *Hot liquids, generally speaking, are more powerful solvents than cold ones.*—To four ounce-measures of water, at the temperature of the atmosphere, add three ounces of sulphate of soda in powder. Only part of the salt will be dissolved, even after being agitated some time. Apply heat, and the whole of the salt will disappear. When the liquor cools, a portion of salt will separate again in a regular form or in *crystals.* This last appearance affords an instance of *crystallization.*

To this law, however, there are several exceptions; for many salts, among which is muriate of soda, or common salt, are equally, or nearly equally, soluble in cold as in hot water. (See the table of solubility of salts in water, in the Appendix.) Hence, a hot and saturated solution of muriate of soda does not, like the sulphate, deposit crystals on cooling. To obtain crystals of the muriate, and of other salts which observe a similar law as to solubility, it is necessary to evaporate a portion of the water; and the salt will then be deposited, even while the liquor remains hot. In general, the more slow the cooling, or evaporation, of saline solutions, the larger and more regular are the crystals.

V. *A very minute division of bodies is effected by solution.*—Dissolve two grains of sulphate of iron in a quart of water, and add a few drops of this solution to a wine-glassful of water, into which a few drops of tincture of galls have been fallen. The dilute infusion of galls will speedily assume a purplish hue. This shows that every drop of the quart of water, in which the sulphate of iron was dissolved, contains a notable portion of the salt.

VI. *Some bodies dissolve much more readily and copiously than others.*—Thus, an ounce measure of distilled water will dissolve half its weight of sulphate of ammonia, one third its weight of sulphate of soda, one sixteenth of sulphate of potash, and only one five-hundredth its weight of sulphate of lime.

VII. *Mechanical agitation facilitates solution.*—Into a wine-glassful of water, tinged blue with the infusion of litmus, let fall a small lump of solid tartaric acid. The acid, if left at rest, even during some hours, will only change to red that portion of the infusion which is in immediate contact with it. Stir the liquor, and the whole will immediately become red.

VIII. *Bodies do not act on each other, unless either one or both be in a state of solution.*—1. Mix some dry tartaric, or, in preference, citric acid with dry carbonate of potash. No combination will ensue till water is added, which, acting the part of a solvent, promotes the union of the acid and alkali, as appears from a violent effervescence.

2. Spread thinly, on a piece of tinfoil, three or four inches square, some dry nitrate of copper*, and wrap it up. No effect will follow. Unfold the tinfoil, and having sprinkled the nitrate of copper with the smallest possible quantity of water, wrap the tinfoil up again as quickly as possible, pressing down the edges closely. Considerable heat, attended with fumes, will now be excited; and, if the experiment has been dexterously managed, even light will be evolved. This shows that nitrate of copper has no action on tin, unless in a state of solution.

IX. *Bodies, even when in a state of solution, do not act on each other at perceptible distances;* in other words, contiguity is essential to the action of chemical affinity.—Thus, when two fluids of different specific gravities, and which have a strong affinity for each other, are separated by a thin stratum of a third, which exerts no remarkable action on either, no combination ensues between the uppermost and lowest stratum. Into a glass jar, or deep ale glass, pour two ounce-measures of a solution of subcarbonate of potash, containing, in that quantity, two drachms of common salt of tartar. Under this introduce, very carefully, half an ounce-measure of water, holding in solution a drachm of common salt; and again, under both these, two ounce-measures of sulphuric acid, which has been diluted with an equal weight of water, and allowed to become cool. The introduction of a second and third liquid beneath the first, is best effected, by filling, with the liquid to be introduced, the dropping tube fig. 15. pl. i., which may be done by the action of the mouth. The finger is then pressed on the upper orifice of the tube; and the lower orifice, being brought to the bottom of the vessel containing the liquid, the finger is withdrawn, and the liquid descends from the tube, without mingling with the upper stratum. When a solution of carbonate of potash is thus separated from diluted sulphuric acid, for which it has a powerful affinity, by the intervention of a thin stratum of brine, the two fluids will remain distinct and inefficient on each other; but, on stirring the mixture, a violent effervescence ensues, in consequence of the action of the sulphuric acid on the potash.

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* To prepare nitrate of copper, dissolve the filings or turnings of that metal in a mixture of one part nitrous acid and three parts water; decant the liquor when it has ceased to emit fumes; and evaporate it to dryness, in a copper or earthen dish. The dry mass must be kept in a bottle.

X. *Two bodies, having no affinity for each other, unite by the intervention of a third.*—Thus, the oil and water which, in Experiment I., could not, by agitation, be brought into union, unite immediately on adding a solution of caustic potash. The alkali, in this case, acts as an intermedium. The fact, indeed, admits of being explained by the supposition, that the oil and alkali form, in the first instance, a compound which is soluble in water.

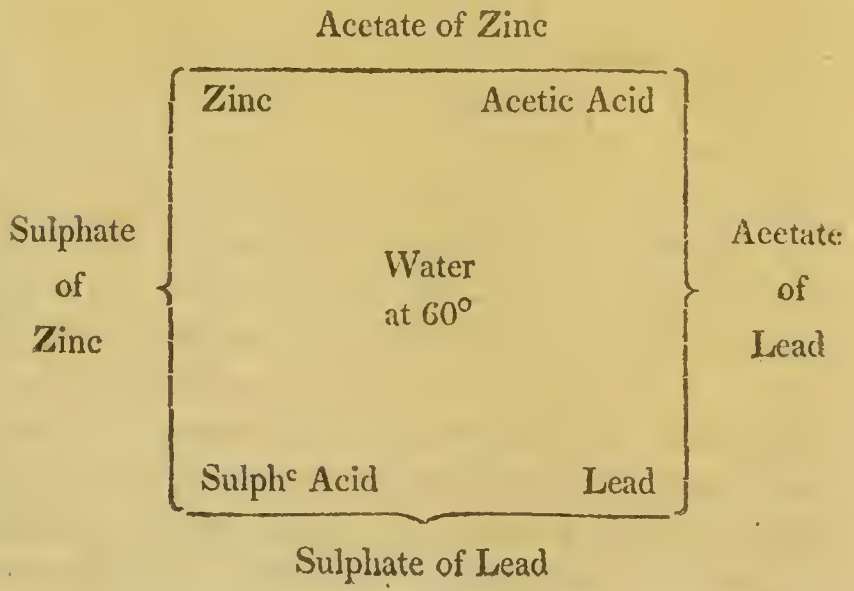
XI. *Saturation and neutralization illustrated.*—Water, after having taken up as much common salt as it can dissolve, is said to be *saturated* with salt. Muriatic acid, when it has ceased to act any longer on lime, is said to be *neutralized.*

XII. *The properties characterizing bodies, when separate, are destroyed by chemical combination, and new properties appear in the compound.*—Thus, muriatic acid and lime, which, in a separate state, have each a most corrosive taste, lose this entirely when mutually saturated; the compound is extremely soluble, though lime itself is very difficult of solution; the acid no longer reddens syrup of violets; nor does the lime change it, as before, to green. The resulting compound, also, muriate of lime, exhibits new properties. It has an intensely bitter taste; is susceptible of a crystallized form; and the crystals, when mixed with snow or ice, generate a degree of cold sufficient to freeze quicksilver.

XIII. *Single elective affinity illustrated.*—1. Add to the combination of oil with alkali, formed in Experiment X., a little diluted sulphuric acid. The acid will seize the alkali, and set the oil at liberty, which will rise to the top. In this instance, the affinity of alkali for acid is greater than that of alkali for oil. 2. To a dilute solution of muriate of lime (prepared in Experiment II.), add a little of the solution of pure potash. The potash will seize the muriatic acid, and the lime will fall down, or be *precipitated.*

XIV. *In every instance, in comparing the affinities of two bodies for a third, a weaker affinity, in one of the two compared, will be found to be compensated by increasing its quantity.*—It is not easy to offer clear and unequivocal examples of this law, and such as the student may submit to the test of experiment. The following, however, may illustrate the proposition sufficiently: Mingle together, in a mortar, one part of muriate of soda (common salt) with half a part of red oxide of lead (litharge, or red lead), and add sufficient water to form a thin paste. The oxide of lead, on examining the mixture after twenty-four hours, will be found not to have detached the muriatic acid from the soda; for the strong taste of that alkali will not be apparent. Increase the weight of the oxide of lead to three or four times that of the salt; and, after the same interval, the mixture will exhibit, by its taste, marks of uncombined soda. This proves, that the larger quantity of the oxide must have detached a considerable portion of muriatic acid from the soda, though the oxide has a weaker affinity for that acid than the soda possesses.

XV. *Double elective affinity exemplified.*—In a watery solution of sulphate of zinc, immerse a thin sheet of lead: the lead will remain unaltered, as also will the sulphate of zinc, because zinc attracts sulphuric acid more strongly than lead. But let a solution of acetate of lead be mixed with one of sulphate of zinc: the lead will then go over to the sulphuric acid, while the zinc passes to the acetic. The sulphate of lead, being insoluble, will fall down in the state of a white powder; but the acetate of zinc will remain in solution. The changes that occur in this experiment will be better understood from the following scheme:



The vertical brackets include the original compounds, *viz.* sulphate of zinc, and acetate of lead; and the horizontal line and bracket point out the new ones, *viz.* acetate of zinc and sulphate of lead. By the upper horizontal line, it is denoted, that the acetate of zinc remains in solution; and, by the point of the lower bracket being directed downwards, it is meant to express, that the sulphate of lead falls down, or is precipitated.

CHAPTER III.

OF HEAT OR CALORIC.

SECTION I.

General Observations on Heat.

WHEN we apply the hand to a body, which is hotter than itself, we are sensible of a peculiar feeling, which we agree to call the *sensation of heat.* At the same time we observe, in almost all bodies that are placed in the same situation with the hand, certain effects, the most remarkable of which is an enlargement of their dimensions. These circumstances, with very few exceptions, so constantly accompany each other, that we can have little or no hesitation in referring them to one and the same cause. Of the nature of this cause we have no satisfactory evidence; and we are unable to demonstrate either that it consists in any general quality of bodies, or that it resides in a distinct and peculiar kind of matter. The opinion, however, which best explains the phenomena, is that which ascribes them to an extremely subtile fluid, of so refined a nature, as to be capable of insinuating itself between the particles of the most dense and solid bodies. To this fluid, as well as to the sensation which it excites, the term *heat* was formerly applied. But there was an obvious impropriety in confounding, under one appellation, two things so distinct as a sensation and its cause; and the term *caloric*, first proposed by Lavoisier, is now, therefore, generally adopted to denote the cause of heat. Occasionally, however, in order to avoid too frequent a repetition of the same word, the term heat is still employed in a more extensive sense, to express not only the sensation which it usually denotes, but also some of the modifications of caloric.

Caloric, so far as its chemical agencies are concerned, may be chiefly considered under two views — as an antagonist to the cohesive attraction of bodies — and as concurring with, and increasing elasticity. By removing the particles of any solid to a greater distance from each other, their cohesive attraction is diminished; and one of the principal impediments to their union with other bodies is overcome. On the other hand, caloric may be infused into bodies in such quantity, as not only to overcome cohesion, but to place their particles beyond the sphere of chemical affinity. Thus, in the class of substances called gases, the ponderable ingredient, whether solid or liquid, is dissolved in so much caloric, that in mechanical properties the gases agree with the air of our atmosphere, and especially in being permanently elastic. Different bodies of this class do not in general unite by simple mixture. But if, of two gases, we employ either one or both in a state of great condensation, or compress their particles nearer to each other by any means, the gravitating matter of each unites, and forms a new compound. Thus hydrogen and oxygen gases remain together in a state of mixture, for any length of time, without combining; but if we force their particles into a state of contiguity by mechanical pressure, they unite and compose water. In many cases, also, when two bodies are combined together, one of which is fixed, and the other becomes elastic by union with caloric, we are able, by its interposition alone, to effect their disunion. Thus carbonate of lime gives up its carbonic acid by the mere application of heat.

We may consider, then, all bodies in nature as subject to the action of two opposite forces, the mutual attraction of their particles on the one hand, and the repulsive power of caloric on the other; and bodies exist in the solid, liquid, or elastic state, as one or the other of these forces prevails. Water, by losing caloric, has its cohesion so much increased, that it assumes the solid form of ice; adding caloric, we diminish again its cohesion, and render it fluid; and, finally, by a still farther addition of caloric, we change it into vapour, and give it so much elasticity, that it may be rendered capable of bursting the strongest vessels. In many liquids, the tendency to elasticity is even so great, that they pass to the gaseous form by the mere removal of the weight of the atmosphere.

Caloric, like all other bodies, may exist in two different states, in a state of freedom, and in a state, either of combination or of something nearly resembling it. In the former state, it is capable of exciting the sensation of heat, and of producing expansion in other bodies. To this modification the terms *free* or *uncombined caloric*, or *caloric of temperature*, have been applied. By the term temperature we are to understand the state of a body relatively to its power of exciting the sensation of heat, and occasioning expansion; effects which, in all probability, bear a proportion to the quantity of free caloric in a given space, or in a given quantity of matter. Thus what we call a high temperature may be ascribed to the presence of a large quantity of free caloric; and a low temperature to that of a small quantity. We are unacquainted, however, with the extremes of temperature; and may compare it to a chain, of which a few middle links only are exposed to our observation.

The degree of expansion produced by caloric, it will afterwards appear, bears a sufficient proportion to its quantity, to afford us a means of ascertaining the latter with tolerable accuracy. In estimating temperature, indeed, our senses are extremely imperfect; for we compare our sensations of heat, not with any fixed or uniform standard, but with those sensations, of which we have had immediately previous experience. The same portion of water will feel warm to a hand removed from contact with snow, and cold to another hand, which has been heated before the fire. To convey, therefore, any precise notion of temperature, we are obliged to describe the degree of expansion produced in some one body, which has been previously agreed upon as a standard of comparison. The standard most commonly employed is a quantity of quicksilver, contained in a glass ball, which terminates in a long narrow tube. This instrument, called a *thermometer*, is of the most important use in acquiring and recording our knowledge of the properties and laws of caloric. The thermometer, however, it must be obvious, is no otherwise a measure of the quantity of caloric, than as it ascertains the amount of one of its principal effects. In this respect, it stands in much the same predicament as the hygrometer, when considered as a mean of determining the moisture of the atmosphere. This last instrument, it may be remembered, is composed of some substance (such as a hair or a piece of whip-chord) which is lengthened by a moist atmosphere and contracted by a dry one; and in a degree proportionate to the moisture or dryness. But all the information, which the hygrometer gives us, is the degree of moisture, between certain points that form the extremities of its scale; and it is quite incompetent to measure the absolute quantity of watery vapour in the air.

In explaining those properties and laws of caloric, which have become known to us by means of the thermometer, it appears a sufficiently natural division of the subject to describe, 1stly, those effects which caloric produces, without loosing its properties of exciting the sensation of heat and occasioning expansion;—and, 2dly, those agencies, in which its characteristic properties are destroyed, and in which it ceases to be cognizable by our senses or by the thermometer.

The EXPANSION OF DILATATION of bodies, it will appear, is almost an universal effect of an increase of temperature. Its amount, however, is not the same in all bodies, but differs very essentially. By the same increase of temperature, liquids expand more than solids, and aëriform bodies more than either. Nor is the same quantity of expansion effected in the *same* solid or liquid, by adding similar quantities of heat; for, generally speaking, bodies expand by equal increments of caloric, more in high than in low temperatures. The explanation of this fact is, that the force opposing expansion (*viz.* cohesion) is diminished by the interposition of caloric between the particles of bodies; and, therefore, when equal quantities of caloric are added in succession, the last portions meet with less resistance to their expansive force than the first. In gases, which are destitute of cohesion, equal increments of heat appear, on the contrary, to be attended with precisely equal augmentations of bulk.

An important property of free caloric, the knowledge of which has been acquired by means of the thermometer, is its *tendency to an equilibrium.* When a heated ball of iron is exposed to the open air, the caloric, which is accumulated in it, flows out; and its temperature is gradually reduced to that of the surrounding medium. This is owing to two distinct causes: the air, immediately surrounding the ball, acquires part of the caloric which escapes; and, having its bulk increased, is rendered specifically lighter and ascends. This is succeeded by a cooler and heavier portion of air from above, which, in its turn, is expanded and carries off a second quantity of caloric. Hence a considerable part of the caloric, which is lost by a heated body, is conveyed away by the ambient air. But the refrigeration cannot be wholly explained on this principle; for it has been long known that heated bodies cool, though with less celerity, under the exhausted receiver of an air pump, and even in a Torricellian vacuum.

When the phenomena accompanying the cooling of bodies are accurately examined, it is found that a part of the caloric, which escapes, moves through the atmosphere with immeasurable velocity. In an experiment of M. Pictet, no perceptible interval took place between the time at which caloric quitted a heated body, and its reception by a thermometer at the distance of sixty-nine feet. It appears, also, to move with equal ease in all directions, and not to be at all impeded by a strong current of air meeting it transversely. Hence it follows that the propagation of caloric, in this state of rapid movement, does not depend on any agency of the medium through which it passes. Like light, it appears to be transmitted in parallel rays; and it has, therefore, under this modification, been called RADIANT CALORIC.

The proportion of caloric, lost by a heated body, in these two different ways, may be approximated by observing what time it takes to cool through the same number of degrees in air and in vacuo. By experiments of this kind, Dr. Franklin thought he had ascertained that a body, which requires five minutes in vacuo, will cool in air, through the same number of degrees, in two minutes. Count Rumford's experiments with a Torricellian vacuum give the proportions of 5 to 3. It will, perhaps, not be very remote from the truth, if it be stated, in general terms, that one half of the caloric, lost by a heated body, escapes by radiation, and that the rest is carried off by the ambient atmosphere.

The radiation of caloric appears to bear a proportion to the elevation of temperature of a body above that of the surrounding medium. Hence in part it is, that a heated body, during refrigeration, loses unequal quantities of caloric in equal times. The series appears to be pretty nearly a geometrical one. Thus, supposing the temperature of a body to be 1000 degrees above the surrounding medium,

In the first minute it will loose $\frac{9}{10}$ of its heat or 900°
In the second . . . . . . $\frac{9}{10}$ of the remainder = 90
In the third . . . . $\frac{9}{10}$ of 10 . . . . = 9

The movement of calorie by radiation occurs only in free space, or through transparent media. But caloric is capable, also, of passing through dense and opake bodies, though with prodigiously impaired velocity. Thus a long bar of iron, heated at one end, soon becomes hot at the other. This property in bodies has been called their CONDUCTING POWER, and it exists, in different bodies, in very different degrees. It is not, however, found to bear a proportion to any other quality of bodies.

All the properties of caloric, which have been hitherto described, belong to it when in a free or uncombined form; for it continues to produce the sensation of heat and to expand the mercury of the thermometer. In the instances of its agency, also, that have been mentioned, no permanent change of form or of properties is effected in the bodies, which have imbibed caloric. A bar of iron, after being expanded by heat, returns on cooling to the same state as before, and exhibits all its former qualities. In certain cases, however, caloric is absorbed by bodies, with the loss of its distinguishing properties. It can then be no longer discovered by our senses or by the thermometer; and it produces important and sometimes permanent changes, in the bodies into which it enters.

Those effects of caloric, in the production of which it loses its distinguishing properties, may be classed under two general heads.

I. *All bodies, in passing from a denser to a rarer state, absorb caloric.*—Thus solids, during liquefaction, imbibe a quantity of caloric, which ceases to be apparent to our senses or to the thermometer; or, as it has been termed, *becomes latent.* In a similar manner, solids and liquids, during their conversion into vapours or gases, render latent a quantity of caloric, which is essential to the elasticity of the new product. In common language cold is, in such cases, said to be produced; but by the production of cold we are to understand, in philosophical language, nothing more than the passage of caloric from a free to a latent form.

II. *All bodies, by an increase of density, evolve or give out caloric, which passes from a latent to a free state.*—The simplest illustration of this law is in the effect of hammering a piece of metal, which may thus be intensely heated, while all that is effected is an augmentation of its density. Liquids by becoming solids, or gases by conversion into liquids, also, evolve caloric, or produce an increase of temperature. A pound of water condensed from steam will render 100 pounds of water at 50° warmer by 11°; whereas a pound of boiling water would produce the same rise of temperature in no more than about 13½ pounds. This is owing to the much greater quantity of caloric, existing in a pound of steam, than in a pound of boiling water, though steam and boiling water affect the thermometer in precisely the same degree.

It is a question which has excited considerable interest among philosophers, whether caloric, when thus absorbed and rendered latent, enters into chemical combination, or is merely united by the same kind of ties as that portion of caloric, that produces the temperature of bodies. Does ice, for example, when changed into water, form a chemical union with caloric, similar to that which exists between potash and sulphuric acid? Such appears to have been the opinion of Dr. Black, who, by the powers of an original and well directed genius, discovered the greater number of those facts, that form the groundwork of the theory of latent heat. The resemblance, however, between chemical union and the disappearance of caloric, which, on first view, appears extremely striking, will be found, it must be confessed, less close on a nearer examination. For caloric may be made to quit those bodies, into which it has entered with the loss of its peculiar properties, merely by reducing their temperature; whereas chemical combinations in general cannot be destroyed, except by the interference of more energetic affinities.

In opposition to the foregoing theory, it has been contended that the absorption of caloric by bodies is a consequence of what has been called a change of their *capacity.*

Thus ice, it is supposed, in becoming water, has its capacity for caloric increased, and the absorption of caloric is a consequence of this increased capacity. The theory, however, is deficient, inasmuch as it fails to explain what is the cause of that change of form, which is assumed to account for the increase of capacity. Notwithstanding this obvious objection, I have retained the term capacity to express, in the abstract, that power, by which bodies absorb and render latent different quantities of caloric; or the property of requiring more or less caloric for raising their temperature an equal number of degrees. The absorption of caloric, then, will always be owing to an increase, and its evolution to a decrease, of capacity. The use of these terms may be exemplified by a slight change of the perspicuous language of Dr. Crawford. "The capacity for containing caloric," he observes (on Heat, page 8), "and the absolute caloric contained, are distinguished as a force from the subject upon which it operates. When we speak of the *capacity*, we mean a power inherent in the heated body; when we speak of the *absolute caloric*, we mean an unknown principle, which is retained in the body by the possession of this power; and when we speak of the *temperature*, we consider the unknown principle as producing certain effects upon the thermometer."

As the capacities of bodies determine their absolute quantities of caloric, it seems reasonable to conclude, that if we can ascertain how much caloric a body absorbs or gives out in changing its form, and in what proportion its capacity is at the same time altered, we may deduce the *absolute quantity of heat* which it contains. Now it will be afterwards shown that the heat, evolved by water in freezing, is equal to 140°; and the capacity of water has been stated to bear to that of ice the proportion of 10 to 9.

Water, then, in becoming ice, must give out $\frac{1}{10}$th of its whole caloric, and as this amounts to 140°, ten times 140 (or 1400°) is the whole quantity of caloric in water at the temperature of 32°; and deducting 140 from 1400, we have 1260° for the caloric contained in the ice itself. This method of determining the problem appears, however, to me, to be liable to several objections, which it would take up too much room to state in this place, and which I have elsewhere urged at considerable length*.

These general observations I have deemed it necessary to make, with the view of connecting together the propositions respecting caloric, and the experiments illustrating them, that form the subject of the following sections. The inquiry respecting heat is one which presents a boundless field for interesting speculation; and it would have been easy to have extended very considerably the discussion of its nature and properties. But in this work, I have no farther object than to lead the student, by easy steps, to a knowledge of what has been actually determined by experiment, or strictly and legitimately deduced from it.

SECTION II.

Illustrations of the Effects of Free Caloric.

I. *Caloric expands all bodies.*—The expansion of *liquids* is shown by that of the mercury of a thermometer, or by immersing in hot water a glass mattras (pl. i. fig. 4), filled, up to a mark in the neck, with spirit of wine, tinged with any colouring substance. The spirit expands immediately when heated, and would overflow if not placed in a cooler situation. The degree of expansion produced in different liquids, by similar elevations of temperature, varies very considerably. Thus, water expands much more than mercury, and alcohol more than water. This difference of expansibility is even sufficiently striking to appear in a remarkable degree, when we immerse, in water heated to 150°, three equal glass vessels of the shape of thermometer tubes, containing the one mercury, the other water, and the third spirit of wine. The spirit will begin to escape from the aperture of the vessel, before the mercury has ascended far in the stem. (See a table of the expansion of liquids, in the Appendix.) The expansion of *aëriform bodies* is shown, by holding, near the fire, a bladder filled with air, the neck of which is closely tied, so as to prevent the enclosed air from escaping. The bladder will soon be fully distended, and may even be burst by continuing and increasing the heat. All aëriform bodies undergo the same expansion by the same additions of heat, or $\frac{1}{483}$ part of their bulk for each degree of Fahrenheit's thermometer, between the freezing and boiling points. The expansion of *solids* is evinced, by heating a rod of iron, of such a length as to be included, when cold, between two points, and the diameter of which is such, as barely to allow it to pass through an iron ring. When heated, it will have become sensibly longer; and it will be found incapable of passing through the ring. This property of metals has been applied to the construction of an instrument for measuring temperature, called a *pyrometer*, a neat and distinct representation of which is given in the first volume of "Chemical Conversations."

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* Manchester Memoirs, vol, v.; or Phil. Mag.

The degree of expansion is not the same for all solids, and even differs materially in substances of the same class. Thus, the metals expand in the following order, the most expansible being placed first; zinc, lead, tin, copper, bismuth, iron, steel, antimony, palladium, platina. (See the table in the Appendix.)

All the above bodies return again, on cooling, to their former dimensions.

II. *Construction of the thermometer founded on the principle of expansion.*—The thermometer is an instrument of so much importance, that it may be expedient to explain the construction of the different kinds which are required in chemical researches.

The instrument employed by Sanctorio, to whom the invention of the thermometer is generally ascribed, was of a very simple kind, and measured variations of temperature by the variable expansion of a confined portion of air. To prepare this instrument, a glass tube (pl. i. fig. 9) is to be provided, eighteen inches long, open at one end, and blown into a ball at the other. On applying a warm hand to the ball, the included air expands, and a portion is expelled through the open end of the tube. In this state, the aperture is quickly immersed in a cup filled with any coloured liquid, which ascends into the tube, as the air in the ball contracts by cooling. The instrument is now prepared. An increase of temperature forces the liquor down the tube; and, on the contrary, the application of cold causes its ascent. These effects may be exhibited, by alternately applying the hand to the ball, and then blowing on it with a pair of bellows. By the application of a graduated scale, the amount of the expansion may be measured.

The ball of the above instrument, it must be obvious, cannot be conveniently applied to measure the temperature of liquids. For adapting it to this purpose, a slight variation may be made in its construction, as represented fig. 8, *a*. To prepare this instrument, a small spherical glass vessel is to be about one 6th or one 4th filled with any coloured liquid. The tube, open at both ends, is then to be cemented into the neck, with its lower aperture beneath the surface of the fluid. The expansion of the included air drives the liquid up the stem, to which we may affix a graduated scale, corresponding with that of a common mercurial thermometer. Other modifications have also been made by different philosophers. One of the most useful and simple forms is represented fig. 8, *b*. It consists merely of a tube of very small bore, from 9 to 12 inches long, at one end of which is blown a ball, from half an inch to an inch in diameter, which is afterwards blackened by paint, or by the smoke of a candle. A small column of coloured liquid, about an inch in length, is then introduced, by a manipulation similar to that already described. To fit the instrument for use, this column ought to be stationary, about the middle of the tube, at the common temperature of the atmosphere. The slightest variation of temperature occasions the movement of the coloured liquid; and a scale of equal parts measures the amount of the effect.

An insuperable objection, however, to the air thermometer, as thus constructed, is, that it is affected, not only by changes of temperature, but by variations of atmospheric pressure. Its utility consists in the great amount of the expansion of air, which, by a given elevation of temperature, is increased in bulk above twenty times more than mercury. Hence it is adapted to detect minute changes of temperature, which the mercurial thermometer would scarcely discover.

An important modification of the air thermometer has been invented by Mr. Leslie, and employed by him, with great advantage, in his interesting researches respecting heat. To this instrument he has given the name of, the Differential Thermometer. Its construction is as follows: "Two glass tubes of unequal length, each terminating in a hollow ball, and having their bores somewhat widened at the other ends (a small portion of sulphuric acid, tinged with carmine, being introduced into the ball of the longer tube), are joined together by the flame of a blow-pipe, and afterwards bent nearly into the shape of the letter U (see fig. 7), the one flexure being made just below the joining, where the small cavity facilitates the adjustment of the instrument. This, by a little dexterity, is performed, by forcing, with the heat of the hand, a few minute globules of air from the one cavity into the other. The balls are blown as equal as the eye can judge, and from four 10ths to seven 10ths of an inch diameter. The tubes are such as are drawn for thermometers, only with wider bores; that of the short one, to which the scale is affixed, must have an exact calibre of a 50th, or a 60th, of an inch. The bore of the long tube need not be so regular, but should be visibly larger, as the coloured liquid will then move quicker under any impression. Each leg of the instrument is from three to six inches in height, and the balls are from two to four inches apart.

"A moment's attention to the construction of this instrument will satisfy us, that it is affected only by the *difference* of heat in the corresponding balls; and is calculated to measure such difference with peculiar nicety. As long as both balls are of the same temperature, whatever this may be, the air contained in both will have the same elasticity, and, consequently, the intercluded coloured liquor, being pressed equally in opposite directions, must remain stationary. But if, for instance, the ball which holds a portion of the liquor be warmer than the other, the superior elasticity of the confined air will drive the liquid forwards, and make it rise, in the opposite branch, above the zero, to an elevation proportional to the excess of elasticity, or of heat." The amount of the effect is ascertained by a graduated scale, the interval between freezing and boiling being distinguished into 100 equal degrees. This instrument, it must be obvious, cannot be applied to measure variations in the temperature of the surrounding atmosphere, for the reason already assigned. It is peculiarly adapted to ascertain the difference of the temperatures of two contiguous spots in the same atmosphere; for example, to determine the heat in the focus of a reflector.

Thermometers, filled with spirit of wine (a liquid which has not been congealed by any degree of cold hitherto produced), are best adapted to the measurement of very low temperatures, at which mercury would freeze. The amount of the expansion of alcohol, also, which exceeds that of mercury above eight times, fits it for ascertaining very slight variations of temperature. But it cannot be applied to measure high degrees of heat; because the conversion of the spirit into vapour would burst the instrument.

The fluid, best adapted for filling thermometers, is mercury, which, though it expands less in amount than air, or alcohol, still undergoes this change to a sufficient degree;

and, in consequence of its difficult conversion into vapour, may be applied to the admeasurement of more elevated temperatures. As a considerable saving of expense will accrue to the experimentalist, who is able to construct mercurial thermometers, I shall offer some rules for this purpose. In general, however, I should deem it preferable merely to superintend their construction, and to be satisfied, by actual inspection, that the necessary accuracy is observed; because much time must be unavoidably lost, in acquiring the manual skill which is essential to construct them neatly.

Thermometer tubes may be had at the glass-house, and of various philosophical instrument makers. In purchasing them, those should be rejected that are not hermetically sealed at both ends; because the smallest condensation of moisture, which must take place when air is freely admitted within the tube, is injurious to the accuracy of the instrument. A small bottle of elastic gum should be provided, in the side of which a brass valve is fixed, or a piece of brass perforated by a small hole, to be occasionally stopped by the hand. A blow-pipe is also an essential part of the apparatus; and, in addition to one of the ordinary kind, it will be found useful to have one which is supplied with air by a pair of double bellows, worked by the foot.

Before proceeding to the construction of the thermometer, it is necessary to ascertain, that the tube is of equal diameter in different parts. This is done, by breaking off both of the sealed ends, immersing one of them an inch or two deep in clean and dry mercury, and then closing the other end with the finger. On withdrawing the tube from the mercury, a small column of that fluid remains in it, the length of which is to be examined, by laying the tube horizontally on a graduated ruler*. By inclining the tube, this column may be gradually moved through its whole length; and if the tube be of uniform bore, it will measure the same in every part. Such a degree of perfection, however, is scarcely ever to be observed throughout tubes of considerable length; but, in general, a portion of the tube will be found perfect, of sufficient length for a thermometer, and this part is to be broken off.

On one end of the tube let the neck of the elastic bottle be firmly tied; and let the other end be heated by the flame of the blow-pipe, till the glass softens. The softened part must then be pressed, by a clean piece of metal, into the form of a rounded button; and to this the flame of the lamp must be steadily applied, till it acquires a white heat, and seems about to enter into fusion. To prevent its falling on one side, the tube, during this time, must be constantly turned round by the hand. When the heated part appears perfectly soft, remove it quickly from the lamp, and, holding the tube vertically, with the elastic bottle uppermost, press this last gently with the hand. The glass will be blown into a small ball, but not into one sufficiently thin for the purpose. To this the flame of the lamp must again be applied, turning it quickly round; and, on a second or third repetition of the process of blowing, the ball will be completely formed. The proportion of the size of the ball to the bore of the tube, can only be learned by some experience.

To fill the ball, which has been thus formed, with mercury, the air must first be expelled by holding it over the flame of an Argand's lamp, and then quickly immersing the open end of the tube in very clean and dry quicksilver. As the ball cools, the mercury will ascend, and will partly fill it. Let a paper funnel be tied firmly over the open end of the tube; into this pour a small portion of quicksilver, and apply the heat of the lamp to the ball. Any remaining portion of air will thus be expelled; and if the heat be raised, so as to boil the mercury, the ball and stem will be filled with mercurial vapour, the condensation of which, on removing the ball from the lamp, will occasion a pretty complete vacuum. Into this vacuum, quicksilver will descend from the paper cone; and the instrument will be completely filled. But for the purpose of a thermometer, it is necessary that the mercury should rise only to a certain height of the stem; and a few drops may, therefore, be expelled by cautiously applying the heat of the lamp. To estimate whether the proper quantity of quicksilver has been left in the instrument, immerse the ball first in ice-cold water, and then in the mouth. The space between these two points will comprise 63 degrees, or pretty nearly one 3d of the whole space between the freezing and boiling points of water. If the empty part of the tube exceeds, in length, about three times the portion thus filled by the expanded quicksilver, we may proceed (when an instrument is wanted with a scale including only from 32° to 212°) to seal it hermetically: which is done as follows: The part to be sealed is first heated with the blow-pipe, and drawn out to a fine capillary tube; the bulb is then heated, till a few particles of quicksilver have fallen from the top of the tube: at this moment, the flame of another candle is directed, by the blow-pipe, on the capillary part of the tube, the candle is withdrawn from the ball, and the tube is sealed, at the instant when the mercury begins to descend. If this operation has been skilfully performed, so as to leave no air in the tube, the whole of the tube should be filled with quicksilver on holding the instrument with the ball uppermost.

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* If the tube be of an extremely small bore, the mercury will not enter, and must be drawn in by the action of the elastic bottle, and not by the mouth.

To have very large degrees, the ball must bear a considerable proportion to the tube; but this extent of scale cannot be obtained without sacrificing, in some measure, the sensibility of the instrument. The whole of the process of constructing thermometers neatly and accurately, is connected with the possession of manual skill, which practice only can confer; and it is scarcely possible, without the most tedious minuteness, to describe all the necessary precautions and manipulations. These will readily suggest themselves to a person who carries the above instructions into effect.

In graduating thermometers, the first step consists in taking the two fixed points. The freezing point is ascertained, by immersing, in thawing snow or ice, the ball and part of the stem; so that the mercury, when stationary, shall barely appear above the surface. At this place let a mark be made with a file. In taking the boiling point, considerable caution is required; and, for reasons which will afterwards be stated, attention must be paid to the state of the barometer, the height of which, at the time, should be precisely 29. 8. A tin vessel is to be provided, four or five inches longer than the thermometer, and furnished with a cover, in which are two holes. Through one of these, the thermometer stem must be passed (the bulb being within the vessel), so that the part at which the boiling point is expected, may be just in sight. The other hole may be left open; and the cover being fixed in its place, the vessel, containing a few inches of water at the bottom, is to be set on the fire. The thermometer will presently be wholly surrounded by steam; and when the mercury becomes stationary in the stem, its place must be marked. The scale of Fahrenheit is formed by transferring the intermediate space to paper by a pair of compasses, and dividing it into 180°, the lowest being called 32°, and the highest 212°. The scale of other countries, however, differs considerably; but these variations do not prevent the comparison of observations with different instruments, when the freezing and boiling points of water are agreed upon as fixed data. In the Appendix, rules will be given for converting the degrees of other scales to that of Fahrenheit.

III. *The dilutations and contractions of the fluid in the mercurial thermometer, are nearly proportional to the quantities of caloric, which are communicated to the same homogeneous bodies, or separated from them, so long as they retain the same form.*

Thus a quantity of caloric, required to raise a body 20° in temperature, by the mercurial thermometer, is nearly double that which is required to raise it 10°. Hence there appears to be a pretty accurate proportion between the increments or decrements of heat, and the increments and decrements of expansion in the mercury of a thermometer. On this principle, if equal quantities of hot and cold water be mixed together, and a thermometer be immersed in the hot water, and also in the cold, previously to the mixture, the instrument should point, after the mixture, to the arithmetical mean, or to half the difference of the separate heats, added to the less or subtracted from the greater. This will be proved to be actually the fact, by the following experiment. Mix a pound of water at 172° with a pound at 32°. Half the excess of the caloric of the hot water, will pass to the colder portion; that is, the hot water will be cooled 70°, and the cold will receive 70° of temperature; therefore $172-70$, or $32+70=102$, will give the heat of the mixture. To attain the arithmetical mean exactly, several precautions must be observed. (See Crawford on Animal Heat, p. 95, &c.)

The experiments of De Luc, however, have shown, that the ratio of expansion does not, *strictly*, keep pace with the actual increments of temperature; and that the amount of the expansion increases with the temperature. Thus if a given quantity of mercury, in being heated from 32 to 122°, the first half of the scale, be expanded 14 parts, in being raised from 122 to 212, the higher half, it will be expanded 15 parts.

From the enquiries of Mr. Dalton, it appears to follow that the irregularity of the expansion of mercury is considerably greater than has been stated by De Luc. By the common mercurial thermometer, we cannot ascertain the true rate of expansion in quicksilver; for it must be obvious that the expansion of the glass ball, in which it is contained, must considerably affect the result. If its capacity remained unaltered, we should then be able to determine the actual rate of expansion; but by an increase of temperature its capacity is enlarged, and space is thus found, within the ball, for the expansion of that mercury, which would otherwise be driven into the tube. By knowing the rate of expansion in glass itself we can correct this error; but a small error in this datum will lead us considerably wrong as to the true expansion of quicksilver. The *real* expansion of mercury in glass is *greater* than the *apparent*, by the expansion of the glass itself.

Making due correction for this circumstance, Mr. Dalton has been led to conclude from his experiments, that notwithstanding the apparent diversities of expansion in different fluids, they all actually expand according to the same law; *viz. that the quantity of expansion is as the square of the temperature from their respective freezing points, or from their points of greatest density.* If then a thermometer be constructed, with degrees corresponding to this law, they will be found to differ very considerably from those of the common mercurial thermometer, in which the space between freezing and boiling is divided into 180 equal parts. In the Appendix will be found a table showing the correspondence between the old scale, and the new one constructed on Mr. Dalton's principle.

IV. *Uncombined caloric has a tendency to an equilibrium.*—Any number of different bodies, at various temperatures, if placed under similar circumstances of exposure, all acquire a common temperature. Thus, if in an atmosphere at 60°, we place iron filings heated to redness, boiling water, water at 32°, and various other bodies of different temperatures, they will soon affect the thermometer in the same degree. The same equalization of temperature is attained, though less quickly, when a heated body is placed in the vacuum of an air-pump. The rate of cooling in air is to that *in vacuo*, the temperatures being equal, nearly as five to two.

II. *Motion of Free Caloric.*—1. *Its Radiation.*—2. *Its Passage through Solids and Fluids.*

Caloric escapes from bodies in two different modes.—Part of it finds its way through space, independently of other matter, and with immeasurable velocity. In this state it has been called, radiant heat, or radiant caloric.

RADIANT CALORIC exhibits several interesting properties.

1. *Its reflection.* (*a*) Those surfaces, that reflect light most perfectly, are not equally adapted to the reflection of caloric. Thus, a glass mirror, which reflects light with great effect when held before a blazing fire, scarcely returns any heat, and the mirror itself becomes warm. On the contrary, a polished plate of tin, or a silver spoon, when similarly placed, reflects, to the hand, a very sensible degree of warmth; and the metal itself remains cool. Metals, therefore, are much better reflectors than glass; and they possess this property, exactly according to their degree of polish.

(*b*) Caloric is reflected according to the same law that regulates the reflection of light. This is proved by an interesting experiment of M. Pictet; the means of repeating which may be attained at a moderate expense. Provide two thin reflectors (*a* and *b*, fig. 45), which may be 12 inches diameter, and segments of a sphere of nine inches radius. Parabolic mirrors are still better adapted to the purpose; but their construction is less easy. Each of these must be furnished, on its convex side, with the means of supporting it in a perpendicular position on a proper stand. Place the mirrors opposite to each other on a table, at the distance of from six to 12 feet. In the focus of one, let the ball of an air thermometer, *c*, or (which is still better) one of the balls of a differential thermometer, be situated; and in that of the other, suspend a ball of iron, about four ounces in weight, and heated below ignition, or a small mattras of hot water, *d*; having previously interposed a screen before the thermometer. Immediately on withdrawing the screen, the depression of the column of liquid, in the air thermometer, evinces an increase of temperature in the instrument. In this experiment, the caloric flows first from the heated ball to the nearest reflector; from this it is transmitted, in parallel rays, to the surface of the second reflector, by which it is collected into a focus on the instrument. This is precisely the course that is followed by radiant light; for if the flame of a taper be substituted for the iron ball, the image of the candle will appear precisely on that spot (a sheet of paper being presented for its reception) where the rays of caloric were before concentrated.

(*c*) When a glass vessel, filled with ice or snow, is substituted for the heated ball, the course of the coloured liquid in the thermometer will be precisely in the opposite direction; for its ascent will show, that the air in the ball is cooled by this arrangement. This experiment, which appears, at first view, to indicate the reflection of cold, presents, in fact, only the reflection of heat in an opposite direction; the ball of the thermometer being, in this instance, the hotter body.

(*d*) In Mr. Leslie's "Enquiry into the Nature, &c., of Heat," a variety of important experiments are detailed, which show the influence of covering the reflectors with various substances, or of mechanically changing the nature of their surfaces, on their power of returning caloric.

2. Caloric is *refracted*, also, according to the same law that regulates the refraction of light. This interesting discovery we owe to Dr. Herschell, whose experiments and apparatus, however, cannot be understood without the assistance of a plate. For this reason, I refer to his paper in the 90th vol. of the Philosophical Transactions, or in the 7th vol. of the Philosophical Magazine.

3. The nature of the *surface* of bodies has an important influence over their power of *radiating* caloric.

To exhibit this influence experimentally, let a canister of planished block tin, forming a cube of six or eight inches, be provided, having an orifice at the middle of its upper side, from half an inch to an inch diameter, and the same in height. This orifice is intended to receive a cap having a small hole, through which a thermometer is inserted, so that its bulb may reach the centre of the canister. Let one side of the canister be covered with black paint; destroy the polish of another side, by scratching it with sand paper; tarnish a third with quicksilver; and leave the fourth bright. Then fill the vessel with boiling water. The radiation of caloric from the blackened side is so much more abundant than from the others, as to be even sensible to the hand. Place it before a reflector (fig. 45), in lieu of the heated iron ball already described. The thermometer, in the focus of the second reflector, will indicate the highest temperature, or most copious radiation of caloric, when the blackened side is presented to the reflector; less when the tarnished or scratched side is turned towards it; and least of all from the polished side.

4. These varieties in the radiating power of different surfaces, are attended, as might be expected, with corresponding variations in the *rate of cooling*. If water in a tin vessel, all of whose sides are polished, cools through a given number of degrees in eighty-one minutes, it will descend through the same number in seventy-two minutes, if the surface be tarnished with quicksilver. Water, also, enclosed in a clean and polished tin ball, cools about twice more slowly than water in the same ball covered with oiled paper. Blackening the surface with paint, on the same principle, accelerates greatly the rate of cooling. These facts teach us, that vessels, in which fluids are to be long kept hot, should have their surfaces brightly polished;

and they explain, among other things, the superiority of metallic tea-pots over those of earthen ware.

5. Radiant caloric is *absorbed* with different facility by different surfaces. This is only stating, in other terms, that surfaces are endowed with various powers of reflecting caloric; since the power of absorbing caloric is precisely opposite to that of reflecting it. Hence the best reflectors of heat will absorb the least. It may be proper, however, to offer some illustrations of the principle under this form.

(*a*) Expose the bulb of a sensible thermometer to the direct rays of the sun. On a hot summer's day it will probably rise, in this climate, to 108°. (Watson's Essays, v. 193.) Cover it with Indian ink, and again expose it in a similar manner. During the evaporation of the moisture it will fall; but as soon as the coating becomes dry, it will ascend to 118°, or upwards, of Fahrenheit, or 10° higher than when uncovered with the pigment. This cannot be explained, by supposing that the black coating is gifted with the power of retaining caloric, and preventing its escape; because, from experiments already related, it appears, that a similar coating accelerates the cooling of a body to which it is applied.

(*b*) Colour has considerable influence over the absorption of caloric. This is shown by the following very simple experiment of Dr. Franklin.

On a winter's day, when the ground is covered with snow, take four pieces of woollen cloth, of equal dimensions but of different colours, *viz.* black, blue, brown, and white, and lay them on the surface of the snow, in the immediate neighbourhood of each other. In a few hours, the black cloth will have sunk considerably below the surface; the blue almost as much; the brown evidently less; and the white will remain precisely in its former situation. Thus it appears, that the sun's rays are absorbed by the dark coloured cloth, and excite such a durable heat, as to melt the snow underneath; but they have not the power of penetrating the white. Hence the preference, generally given to dark coloured cloths during the winter season, and to light coloured ones in summer, appears to be founded on reason.

(*c*) This experiment has been varied by Mr. Davy, in a manner which may be repeated at any season of the year. Take six similar pieces of sheet copper, each about an inch square, and colour the one white, another yellow, a third red, the fourth green, the fifth blue, and the sixth black. On the centre of one side of each piece, put a small portion of a mixture of oil and wax, or cerate, which melts at about 76°. Then expose their coloured surfaces, under precisely equal circumstances, to the direct rays of the sun. The cerate on the black plate will begin to melt perceptibly before the red; the blue next; then the green and the red; and, lastly, the yellow. The white will scarcely be affected, when the black is in complete fusion.

Caloric passes, also, but much more slowly, through solid and liquid bodies, which are then termed, CONDUCTORS of caloric.

1. Solid bodies convey heat in all directions, upwards, downwards, and laterally; as may be shown, by heating the middle of an iron rod, and holding it in different directions.

2. Some bodies conduct caloric much more quickly than others. Coat two rods, of equal length and thickness, the one of glass, the other of iron, with wax, at one end of

each only; and then apply heat to the uncoated ends. The wax will be melted vastly sooner from the end of the iron rod, than from the glass one; which shows, that iron conducts heat more quickly than glass.

Even the different metals possess very different powers of conducting caloric. An approximation to the degree in which they possess this property, may be attained by the following method, originally employed by Dr. Ingenhouz. Procure several solid cylinders, or rods, of the same size and shape, but of different metals. They may be six inches long, and one 4th of an inch in diameter. Coat them, within about an inch of one end, with bees-wax, by dipping them into this substance when melted, and allowing the covering to congeal. Let an iron heater be provided, in which small holes have been drilled, that exactly receive the clean ends of the cylinders. After heating it below ignition, insert the cylinders in their places. The conducting power may be estimated by the length of wax coating melted from each in a given time. According to the experiments of Dr. Ingenhouz, the metals may be arranged in the following order: Silver possesses the highest conducting power; next gold; then copper and tin, which are nearly equal; and, below these, platina, iron, steel, and lead, which are greatly inferior to the rest.

It is chiefly owing to the different conducting powers of bodies, that they affect us, when we touch them, with different sensations of cold. Thus, if we apply the hand in succession to a number of bodies (as a piece of wood, another of marble, &c.), they appear cold in very different degrees. And as this sensation is occasioned by the passage of caloric, out of the hand into the body which it touches, that body will feel the coldest, which carries away heat the most quickly; or which, in other words, is the best conductor. For the same reason, of two bodies which are heated to the same degree, and both considerably above the hand, the best conductor is the hottest to the touch. Thus the money in our pockets often feels hotter than the clothes which contain it.

3. Liquid and aëriform bodies convey heat on a different principle from that observed in solids, *viz.* by an actual change in the situation of their particles. That portion of the fluid, which is nearest to the source of heat, is expanded, and becoming specifically lighter, ascends, and is replaced by a colder portion from above. This, in its turn, becomes heated and dilated, and gives way to a second colder portion; and thus the process goes on, as long as the fluid is capable of imbibing heat.

(*a*) Take a glass tube, eight or 10 inches long, and about an inch in diameter. Pour into the bottom part, for about the depth of an inch, a little water tinged with litmus, and then fill up the tube with common water, pouring on the latter extremely gently, so as to keep the two strata quite distinct. If the upper part of the tube be first heated, the coloured liquor will remain at the bottom. But if the tube be afterwards heated at the bottom, the infusion will ascend, and will tinge the whole mass of fluid.

(*b*) Into a cylindrical glass jar, four inches diameter, and 12 or 14 deep, let a circular piece of ice be fitted 3½ inches thick, and of rather less diameter than the jar. Or water may be poured into the jar to the depth of 3½ inches, and allowed to congeal by exposure to a freezing atmosphere, or by surrounding it with a mixture of snow and salt. The ice is to be secured in its place by two slips of wood, crossing each other like two diameters of a circle, set at right angles to each other. Pour, over the cake of ice, water of 32° temperature, to the depth of two inches; and on its surface let there float a shallow circular wooden box, perforated with holes. From the cock of a tea-urn, filled with boiling water, and raised so that its spout may be above the top of the jar, suspend a number of moistened threads, the lower ends of which must rest on the surface of the box. By this arrangement, when the cock is turned, the hot water will trickle down the threads, and will have its fall considerably broken. It will then spread over the surface of the box, and pass through the perforated holes to the cold water beneath, over which it will float without mixing with it. Let the jar be thus completely filled with hot water. The ice will remain unmelted for several hours at the bottom of the vessel.

(*c*) Fill a similar jar with hot water; and, having provided a cake of ice, of equal size with the former one, let it be placed on the surface of the water. In about three minutes, the whole will be melted. Both these experiments are more striking, if the water, used for forming the cakes of ice, be previously coloured with litmus; for, in the latter experiment, the descending currents of cold water are thus made apparent.

(*d*) These experiments may be varied, by freezing, in the bottom of a tube one inch wide, a portion of water, about two inches in depth. Then fill the tube with water of the common temperature, and hold it inclined over an Argand's lamp, so that the upper portion only of the tube may be heated. When thus disposed, the water may be made to boil violently at the surface, and yet the ice will not be melted. But if the experiments be reversed, and (the ice floating on the surface) heat be applied to the bottom of the tube, the ice will be liquefied in a few seconds.

(*e*) Substituting water of the temperature of 41° for the boiling water used in experiment (*c*), Count Rumford found, that, in a given time, a much greater quantity of ice was melted by the cooler water. This appears, on first view, rather paradoxical. The fact, however, is explained by a remarkable property of water, *viz.* that when cooled below 40° it ceases to contract, and experiences, on the contrary, an enlargement of bulk. Water, therefore, at 40° (at the bottom of which is a mass of ice at 32°), is cooled by contact with the ice, and is expanded at the same moment. It therefore ascends, and is replaced by a heavier and warmer portion from above.

It is a consequence of the same property that the surface of a deep lake is sometimes covered with ice, even when the water below is only cooled to 40°; for the superficial water is specifically lighter than the warmer water beneath it, and retains its place, till it is changed into ice. This property of water is one of the most remarkable exceptions to the law, that bodies are expanded by an increase, and contracted by a diminution, of temperature.

From these facts, Count Rumford concluded, that water is a perfect non-conductor of caloric, and that it propagates caloric only in one direction, *viz.* upwards, in consequence of the motions which it occasions among the particles of the fluid. The count inferred also, that if these motions could be suspended, caloric would cease to pass through water; and, with the view of deciding this question, he made the following experiments, which admit of being easily repeated. A cylindrical tin vessel must previously be provided, two inches in diameter, and 2½ inches deep, having a moveable cover, perforated with a small aperture, for transmitting the stem of a thermometer, which is to be inserted so that its bulb may occupy the centre of the vessel.

(*f*) Fill this vessel with water of the temperature of the atmosphere; let the cover be put in its place; and let the whole apparatus, except the scale of the thermometer, be immersed in water, which is to be kept boiling over a lamp. Observe how long a time is required to raise the water from its temperature at the outset to 180°, and remove it from its situation. Note, also, how long it takes to return to its former temperature.

(*g*) Repeat the experiment, having previously dissolved in the water 200 grains of common starch. The thermometer will now require about half as long again to arrive at the same temperature. A similar retardation, and to a greater amount, is produced by the mixture of eider-down, cotton-wool, and various other substances, which are not chemically soluble in water, and which can diminish its conducting power in no other way than by obstructing the motion of its particles.

This inference, however, respecting the complete non-conducting power of water, has been set aside by the subsequent inquiries of Dr. Thomson and Mr. Murray, especially by a most decisive experiment of the latter. To establish the conducting power of water, it was justly deemed indispensable, that caloric should be proved to be propagated through that fluid downwards. This, on actual trial, it appeared to be; but it was objected, that the sides of the containing vessel might be the conductor. To obviate this objection, Mr. Murray contrived to congeal water into the form of a jar capable of holding liquids. This was separately filled with linseed oil and with mercury. At a proper distance below the surface, the bulb of a thermometer was placed; and on the surface of the liquid rested a flat iron vessel, containing boiling water. Under these circumstances, the thermometer invariably rose; and though it ascended only a very few degrees, yet it must be recollected, that the cooling power of the sides of the vessel would effectually prevent any considerable elevation of temperature. This experiment, in conjunction with others, decisively proves, that water is a conductor, though a slow or imperfect one, of caloric.

SECTION III.

Caloric the Cause of Fluidity.

I. *The temperature of melting snow, or of thawing ice, is uniformly the same at all times, and in all places.*—This may be ascertained by the thermometer, which will always, when immersed in liquefying ice or snow, point to 32° of Fahrenheit, whatsoever may be the height of the barometer, or the elevation, above the sea, of the place where the experiment is made. (Shuckburgh, Philosophical Transactions, lxix.)

II. *The sensible heat, or temperature of ice, is not changed by liquefaction.*—A thermometer in pounded ice stands at 32°, and at the very same point in the water which results from the liquefaction of ice.

III. *Yet ice, during liquefaction, must absorb much caloric.* — Expose a pound of water at 32°, and a pound of ice at 32°, in a room, the temperature of which is several degrees above the freezing point, and uniformly the same during the experiment. The water will arrive at the temperature of the room, several hours before the ice is melted; and the melted ice will have, as before its liquefaction, the temperature of 32°. Yet the ice must, during the whole of this time, have been imbibing caloric, because (according to Experiment IV. § 1.) a colder body can never be in contact with a warmer one, without receiving caloric from it. The caloric, therefore, which has entered the ice, but is not to be found in it by the thermometer, is said to have *become latent.* As it is the cause of the liquefaction of the ice, it is sometimes called *caloric of fluidity.*

IV. *The quantity of caloric that enters into a pound of ice, and becomes latent, during liquefaction, may be learned by experiment.*—To a pound of water, at 172°, add a pound of ice at 32°. The temperature will not be the arithmetical mean (102°), but much below it, *viz.* 32°. All the excess of caloric in the hot water has therefore disappeared. From 172° take 32°; the remainder, 140°, shows the quantity of caloric that enters into a pound of ice during liquefaction; that is, as much caloric is absorbed by a pound of ice, during its conversion into water, as would raise a pound of water from 32° to 172°.

It is from the property of its uniformly absorbing the same quantity of caloric for conversion into water, that ice has been ingeniously applied, by Lavoisier and Laplace, to the admeasurement of the heat, evolved in certain operations. Let us suppose the body (from which the caloric, evolved either by simple cooling or combustion, is to be measured) to be inclosed in a hollow sphere of ice, with an opening at the bottom. When thus placed, the heat, which is given out, will be all employed in melting the ice; and will produce this effect in direct proportion to its quantity. Hence the quantity of ice, which is converted into water, will be an accurate measure of the caloric, that is separated from the body submitted to experiment. In this way, Lavoisier ascertained that equal weights of different combustible bodies melt, by burning, very different weights of ice. The apparatus which he employed for this purpose, he has called the *calorimeter*. Its construction can scarcely be understood without the plate, which accompanies the description in his "Elements of Chemistry."

V. *Other examples of the absorption of caloric, during the liquefaction of bodies,* are furnished by the mixture of snow and nitric acid, or of snow and common salt, both of which, in common language, produce intense cold*.

1. Dilute a portion of nitric acid with an equal weight of water; and, when the mixture has cooled, add to it a quantity of light fresh-fallen snow. On immersing the thermometer in the mixture, a very considerable reduction of temperature will be observed. This is owing to the absorption, and intimate fixation, of the free caloric of the mixture, by the liquefying snow.

2. Mix quickly together equal weights of fresh-fallen snow at 32°, and of common salt cooled, by exposure to a freezing atmosphere, down to 32°. The two solid bodies, on admixture, will rapidly liquefy; and the thermometer will sink 32°, or to 0; or, according to Sir C. Blagden, to 4° lower. (Philosophical Transactions, lxxviii. 281.) To understand this experiment, it must be recollected, that the snow and salt, though at the freezing temperature of water, have each a considerable portion of uncombined caloric. Now,

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* The extraordinary powers of muriate of lime and snow, in generating cold, will be described hereafter.

salt has a strong affinity for water; but the union cannot take place while the water continues solid. In order, therefore, to act on the salt, the snow absorbs all the free caloric required for its liquefaction; and during this change, the free caloric, both of the snow and of the salt, amounting to 32°, becomes latent, and is concealed in the solution. This solution remains in a liquid state at 0, or 4° below 0 of Fahrenheit; but if a greater degree of cold be applied to it, the salt separates in a concrete form.

3. Most neutral salts, also, during solution in water, absorb much caloric; and the cold, thus generated, is so intense as to freeze water, and even to congeal mercury. The former experiment, however (*viz.* the congelation of water), may easily be repeated on a summer's day. Add to 32 drachms of water, 11 drachms of muriate of ammonia, 10 of nitrate of potash, and 16 of sulphate of soda, all finely powdered. The salts may be dissolved separately, in the order set down. A thermometer, put into the solution, will show, that the cold produced is at or below freezing; and a little water, in a thin glass tube, being immersed in the solution, will be frozen in a few minutes. Various other freezing mixtures are described in Mr. Walker's papers in the Philosophical Transactions for 1787, 88, 89, 95, and 1801. Of these the table, given in the Appendix, for which I am indebted to the obliging communication of the author, contains an arranged abstract.

4. Muriate of lime, when mixed with snow, produces a most intense degree of cold. This property was discovered some years ago by M. Lovitz, of St. Petersburgh, and has been since applied, in this country, to the congelation of mercury on a very extensive scale. The proportions which answer best, are about equal weights of the salt finely powdered, and of fresh-fallen and light snow. On

mixing these together, and immersing a thermometer in the mixture, the mercury sinks with great rapidity. For measuring exactly the cold produced, a spirit-thermometer, graduated to 50° below 0 of Fahrenheit, or still lower, should be employed. A few pounds of the salt are sufficient to congeal a large mass of mercury. By means of 13 pounds of the muriate, and an equal weight of snow, Messrs. Pepys and Allen froze 56 pounds of quicksilver into a solid mass. The mixture of the whole quantity of salt and snow, however, was not made at once, but part was expended in cooling the materials themselves.

On a small scale it may be sufficient to employ two or three pounds of the salt. Let a few ounces of mercury, in a very thin glass retort, be immersed, first in a mixture of one pound of each; and, when this has ceased to act, let another similar mixture be prepared. The second will never fail to congeal the quicksilver.

In plate iv. fig. 42, a very simple and cheap apparatus is represented, which I have generally employed to freeze mercury. The dimensions will be given in the description of the plates. (See Appendix.)

The salt thus expended may be again evaporated, and crystallized for future experiments.

The reader, who wishes for farther particulars respecting these experiments, is referred to the Philosophical Magazine, vol. iii. p. 76.

VI. *On the contrary, liquids, in becoming solid, evolve or give out caloric, or, in common language, produce heat.*

1. Water, if kept perfectly free from agitation, may be cooled down several degrees below 32°; but, on shaking it, it immediately congeals, and the temperature rises to 32°.

2. Expose to the atmosphere, when at a temperature below freezing (for example, at 25° of Fahrenheit), two equal quantities of water, in one only of which about a fourth of its weight of common salt has been dissolved. The saline solution will be gradually cooled, without freezing, to 25°. The pure water will gradually descend to 32°, and will there remain stationary a considerable time before it congeals. Yet while thus stationary, it cannot be doubted, that the pure water is yielding caloric to the atmosphere, equally with the saline solution; for it is impossible that a warmer body can be surrounded by a cooler one, without imparting caloric to the latter. The reason of this equable temperature is well explained by Dr. Crawford. (*On Heat*, p. 80.) Water, he observes, during freezing, is acted upon by two opposite powers: it is deprived of caloric by exposure to a medium, whose temperature is below 32°; and it is supplied with caloric, by the evolution of that principle from itself, *viz.* of that portion which constituted its fluidity. As these powers are exactly equal, the temperature of the water must remain unchanged, till the caloric of fluidity is all evolved.

3. The evolution of caloric, during the congelation of water, is well illustrated by the following experiment of Dr. Crawford:—Into a round tin vessel put a pound of powdered ice; surround this by a mixture of snow and salt in a larger vessel; and stir the ice in the inner one, till its temperature is reduced to + 4° of Fahrenheit. To the ice thus cooled, add a pound of water at 32°. One 5th of this will be frozen; and the temperature of the ice will rise from 4° to 32°. In this instance, the caloric, evolved by the congelation of one 5th of a pound of water, raises the temperature of a pound of ice 28°.

4. If we dissolve sulphate of soda in water, in the pro-

portion of one part to five, and surround the solution by a freezing mixture, it cools gradually down to 31°. The salt, at this point, begins to be deposited, and stops the cooling entirely. This evolution of caloric, during the separation of a salt, is exactly the reverse of what happens during its solution. (Blagden, Philosophical Transactions, lxxviii. 290.)

5. To a saturated solution of sulphate of potash in water, or of any salt that is insoluble in alcohol, add an equal measure of alcohol. The alcohol, attracting the water more strongly than the salt retains it, precipitates the salt, and considerable heat is produced.

SECTION IV.

Caloric the Cause of Vapour.

I. *Every liquid, when of the same degree of chemical purity, and under equal circumstances of atmospheric pressure, has one peculiar point of temperature, at which it invariably boils.*—Thus, pure water always boils at 212°, alcohol at 176°, and ether at 98° Fahrenheit; and, when once brought to the boiling point, no liquid can be made hotter, however long the application of heat be continued. The boiling point of water may be readily ascertained, by immersing a thermometer in water boiling over the fire. As there is some danger in applying heat directly to a vessel containing either ether or alcohol, the ebullition of these fluids may be shown, by immersing the vessel containing them in water, the temperature of which may be gradually raised. The appearance of boiling is owing to the formation of vapour at the bottom of the vessel, and its escape through the heated fluid above it. That the steam, which escapes, is actually formed at the bottom, and not at the top of the water, may be seen by boiling some water in a Florence flask, or other transparent vessel, over an Argand's lamp. The bubbles of vapour will all ascend from the bottom of the vessel.

II. *Steam has exactly the same temperature as boiling water.*—Let a tin vessel be provided, having two holes in its cover, one of which is just large enough to admit the stem of a thermometer. Fill it partly with water, and let the bulb of the thermometer be an inch or two above the surface of the water, leaving the other aperture open for the escape of vapour. When the water boils, the thermometer, surrounded by steam, will rise to 212°, which is precisely the temperature of the water beneath : yet water, placed on a fire, continues to receive heat, very abundantly, even when boiling hot; and as this heat is not appreciable by the thermometer, it must exist in the steam, in a latent state.

Perfectly formed steam is entirely invisible. We may satisfy ourselves of this by boiling strongly a small quantity of water in a flask; for complete transparency will exist in the upper part of the vessel. It is only when it begins to be condensed, that steam becomes visible. We have a proof also, of the same fact in the thick fogs, which are produced by a sudden transition from warm to cold weather; the vapour, which was imperceptible at the higher temperature, being condensed and rendered visible by the lower.

III. *The boiling point of the same fluid varies, under different degrees of atmospheric pressure.*—Thus, water, which has been removed from the fire, and ceased to boil, has its ebullition renewed, when it is placed under a receiver, the air of which is quickly exhausted by an air pump. Alcohol and ether, confined under an exhausted receiver, boil violently at the temperature of the atmosphere. In general, liquids boil *in vacuo*, with about 140° less of heat, than are required under a mean pressure of the atmosphere. (Black's Lectures, i. 151.) Even the ordinary variations in the weight of the air, as measured by the barometer, are sufficient to make a difference in the boiling point of water of about 5° between the two extremes*. On ascending considerable heights, as to the tops of mountains, the boiling point of water gradually falls on the scale of the thermometer. Thus on the summit of Mount Blanc, water was found by Saussure to boil at 187° Fahrenheit.

The influence of a diminished pressure in facilitating ebullition, may be inferred also from the following very simple experiment: Place, over a lamp, a Florence flask, about three 4ths filled with water; let it boil briskly during a few minutes; and, immediately on removing it from the lamp, cork it tightly, and suddenly invert it. The water will now cease to boil; but, on cooling the convex part of the flask by a stream of cold water, the boiling will be renewed. Applying boiling water from the spout of a tea-kettle to the same part of the flask, the water will again cease to boil. This renewal of the ebullition, by the application of cold (an apparent paradox), is owing to the formation of an imperfect vacuum over the hot water, by the condensation of steam; and the suspension of the boiling, on re-applying the heat, to the renewed pressure on the surface of the hot water, occasioned by the formation of fresh steam.

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* Sir G. Shuckburgh, in Philosophical Transactions, lxxix. 375.

From these facts, it may be inferred, that the particles of caloric are mutually repulsive, and that they communicate this repulsive tendency to other bodies in which caloric is contained. This repulsive power tends to change solids into fluids, and liquids into aëriform bodies, and is chiefly counteracted by the pressure of the atmosphere.

Were this counteracting cause removed, many bodies, which at present have a liquid form, would cease to be such, and would be changed into a gaseous state. Precisely the same effect, therefore, results from the prevalence of either of these forces. Add to certain liquids a quantity of caloric, in other words, place them in a high temperature, and they are immediately converted into gases: or, their temperature remaining the same, diminish the weight of the atmosphere; and the caloric, which they naturally contain, exerts its repulsive tendency with equal effect, and they are in like manner converted into gases. These facts are best shown by the following experiments on ether:

1. Ether, at the temperature of 104°, exists in the state of a gas. This may be shown by filling a jar with water of this temperature, and inverting it in a vessel of the same. Then introduce a little ether, by means of a small glass tube closed at one end. The ether will rise to the top of the jar, and, in its ascent, will be changed into gas, filling the whole jar with a transparent invisible elastic fluid. On permitting the water to cool, the ethereal gas is condensed, and the inverted jar again becomes filled with water.

2. Ether is changed into gas by diminishing the weight

of the atmosphere. Into a glass tube, about six inches long, and half an inch in diameter, put a tea-spoonful of ether, and fill up the tube with water; then, pressing the thumb on the open end of the tube, place it, inverted, in a jar of water. Let the whole be set under the receiver of an air pump, and the air exhausted. The ether will be changed into gas, which will expel the water entirely from the tube. On re-admitting the air into the receiver, the gas is again condensed into a liquid form.

IV. *On the contrary, by considerably increasing the pressure, water may be heated to above* 400° *Fahrenheit, without being changed into vapour.*—This experiment requires, for its performance, a strong iron vessel, called a Papin's digestor, a plate of which may be seen in Gren's Chemistry. That the boiling point of water, and the temperature of steam, are raised by an increased pressure, may be shown, however, by means of the small boiler, represented plate v. fig. 46, which will be found extremely useful in experiments on this subject. Its precise size, and directions for its construction, will be given in the Appendix.

On the cock *c* may be screwed, occasionally, a valve, loaded in the proportion of 14 pounds to the square inch. The boiler being rather more than half filled with water, and the perforated cap *d* being screwed into its place, the ball of the thermometer will be an inch or more above the surface of the water, and will indicate its temperature, as well as that of the steam, both being, necessarily, in all cases, precisely the same. Allowing the steam to escape through the cock *c*, before affixing the valve, the temperature of the steam, under a mean atmospheric pressure, will be 212°. When an additional atmosphere is added by the weighted valve, it will rise to above

240°; by a valve twice as heavy as the first, or loaded in the proportion of 28 pounds to the square inch, the temperature of the steam will be raised to nearly 270°. This is as far as it is safe to carry the experiment; but by substituting a strong iron vessel, the numbers have been obtained, which will be found in the form of a table, in the Appendix.

V. *The absorption of caloric, during evaporation, shown by experiment.*—Moisten a thermometer with alcohol, or with ether, and expose it to the air, repeating these operations alternately. The mercury of the thermometer will sink at each exposure, because the volatile liquor, during evaporation, robs it of its heat. In this way (especially with the aid of an apparatus described by Mr. Cavallo, in the Philosophical Transactions, 1781, p. 509), water may be frozen in a thin and small glass ball, by means of ether. The same effect may be obtained, also, by immersing a tube, containing water at the bottom, in a glass of ether, which is to be placed under the receiver of an air pump; or the ether may be allowed to float on the surface of the water. During the exhaustion of the vessel, the ether will evaporate rapidly; and, robbing the water of heat, will completely freeze it; thus exhibiting the singular spectacle of two fluids in contact with each other, one of which is in the act of boiling, and the other of freezing, at the same moment.

VI. *The fixation of caloric in water, by its conversion into steam, may be shown by the following experiments:* —1. Let a pound of water at 212°, and eight pounds of iron filings at 300°, be suddenly mixed together. A large quantity of vapour will be instantly generated; and the temperature of the mixture will be only 212°; but that of the vapour produced, is also not more than 212°; and the steam must therefore contain, in a latent or combined form, all the caloric which raised the temperature of eight pounds of iron filings from 212° to 300°.

2. The quantity of caloric, which thus becomes latent during the formation of steam, may be approximated, by repeating the following experiment of Dr. Black: He placed two cylindrical flat bottomed vessels of tin, five inches in diameter, and containing a small quantity of water at 50°, on a red hot iron plate, of the kind used in kitchens. In four minutes the water began to boil, and in twenty minutes the whole was boiled away. In four minutes, therefore, the water received 162° of temperature, or $40\frac{1}{2}°$ in each minute. If we suppose, therefore, that the heat continues to enter the water at the same rate, during the whole ebullition, we must conclude that $40\frac{1}{2}° \times 20° = 810°$ have entered the water, and are contained in the vapour.

It has been found by experiment that 75 pounds of Newcastle coal, or 100 pounds of coal of medium quality, applied in the best manner, are required for the vaporization of 12 cubic feet, or about $89\frac{3}{4}$ wine gallons, of water. A pound of coal, on the average, may be considered as equivalent to convert a gallon of water into vapour.

VII. *Water, by conversion into steam, has its bulk prodigiously enlarged, viz. according to Mr. Watt's experiments, about* 1800 *times.*—A cubic inch of water (or 252 grains) occupies, therefore, when converted into steam, the space of rather more than a cubic foot. Hence its specific gravity, under the ordinary pressure of the air, is to that of common air, nearly as 450 to 1000.

VIII. *On the contrary, vapours, during their conversion into a liquid form, evolve, or give out, much caloric.*—The heat given out, by the condensation of steam, is rendered apparent by the following experiment: Mix 100 gallons of water at 50°, with 1 gallon of water at 212°. The temperature of the water will be raised about $1\frac{1}{2}$°. Condense by a common still-tub, 1 gallon of water, from the state of steam, by 100 gallons of water, at the temperature of 50°. The water will be raised 11°. Hence, 1 gallon of water, condensed from steam, raises the temperature of 100 gallons of cold water $9\frac{1}{2}$° more than 1 gallon of boiling water; and, by an easy calculation, it appears, that the caloric imparted to the 100 gallons of cold water by 8 pounds of steam, if it could be condensed in 1 gallon of water, would raise it to 950°. (Black's Lectures, i. 169.)

For exhibiting the same fact, by means of a small apparatus, which may be placed on a table, and with the assistance only of a lamp, the boiler already described (fig. 46) will be found extremely well adapted. The right angled pipe *e* must be screwed, however, into its place, and must be made to terminate at the bottom of a jar, containing a known quantity of water of a given temperature. This conducting pipe and the jar should be wrapped round with a few folds of flannel. The apparatus being thus disposed, let the water in the boiler be heated by an Argand's lamp, with double concentric wicks, till steam issues in considerable quantity through the cock *c*, which is then to be closed. The steam will now pass through the right angled pipe into the water contained in the jar, which will condense the steam, and will have its temperature very considerably raised. Ascertain the augmentation of temperature and weight; and the result will show, how much a given weight of water has had its temperature raised by a certain weight of condensed steam. To another quantity of water, equal, in weight and temperature to that contained in the jar at the outset of the experiment, add a quantity of water at 212°, equal in weight to the condensed steam; it will be found, on comparison of the two resulting temperatures, that a given weight of steam has produced, by its condensation, a much greater elevation of temperature, than the same quantity of boiling water. This will be better understood by the following example, taken from actual experiment:

Into eight ounces of water, at 50° Fahrenheit, contained in the glass jar, *f*, fig. 46, steam was passed from the boiler, till the temperature of the water in the jar rose to 173°. On weighing the water, it was found to have gained $8\frac{1}{2}$ drachms; that is, precisely $8\frac{1}{2}$ drachms of steam had been condensed, and had imparted its heat to the water.—To facilitate the explanation of this experiment, it is necessary to premise the following remarks.

To measure the whole quantities of caloric contained in different bodies, is a problem in chemistry which has not yet been solved. But the quantities of caloric, added to, or subtracted from, different bodies (setting out from a given temperature) may, in many cases, be measured and compared with considerable accuracy. Thus, if, as has been already stated, two pounds of water at 120° be mixed with two pounds at 60°, half the excess of caloric in the hot water will pass to the colder portion; that is, the hot water will be cooled 30°, and the cold will receive 30° of temperature; and if the experiment be conducted with proper precautions, 90°, the arithmetical mean of the temperature of the separate part , will be the temperature of the mixture. If three pounds of water at 100° be mixed with one pound at 60°, we shall have the same quantity of heat as before, *viz.* four pounds at 90°. Hence, if the quantity of water be multiplied by the temperature, the product will be a comparative measure of the quantity of caloric which the water contains, exceeding the zero of the thermometer employed.

Thus, in the last example,

$3 \times 100 = 300$ = the caloric above zero in the first portion.
$1 \times 60 = 60$ = the caloric above zero in the second do.

The sum, 360 = the caloric above zero in the mixture.

Dividing 360 by 4, the whole quantity of water, we obtain 90°, the temperature of the mixture.

This method of computation may be conveniently applied to a variety of cases. Thus, in the foregoing experiment, $8\frac{1}{2}$ drachms of steam at 212°, added to 64 drachms of water at 50°, produced $72\frac{1}{2}$ drachms of water at 173°. Now,

$72\frac{1}{2} \times 173 = 12542\frac{1}{2}$ = whole heat of the mixture.
$64 \times 50 = 3200$ = heat of 64 drachms, one of the component parts.

$9342\frac{1}{2}$ = heat of $8\frac{1}{2}$ drachms, the other component part.

Therefore $9342\frac{1}{2}$ divided by $8\frac{1}{2} = 1099$, should have been the temperature of the latter portion (*viz.* $8\frac{1}{2}$ drachms), had none of its heat been latent: and $1099 - 212 = 887$ gives the latent heat of the steam. This result does not differ more than might be expected, owing to the unavoidable inaccuracies of the experiment, from Mr. Watt's determination, which states the latent heat of steam at 900°, or from that to 950°. (Black's Lectures, i. 174.) Lavoisier, with the aid of the calorimeter, makes it 1000° or a little more. (Ibid. 175.)

IX. *The same weight of steam contains, whatever may be its density, the same quantity of caloric; its latent heat being increased in exact proportion as its sensible heat is diminished; and the reverse.*—This principle, though scarcely admitting of illustration by any easy experiment, is one of considerable importance; and an ignorance of it has been the occasion of many fruitless attempts to improve the economy of fuel in the steam engine. The fact, so far as respects steam of lower density than that of 30 inches of mercury, was long ago determined experimentally by Mr. Watt. (Black, i. 190.) As the boiling point of liquids is known to be considerably reduced under a diminished pressure, it seemed reasonable to suspect that, under these circumstances, steam might be obtained from them with a less expenditure of heat. Water, Mr. Watt found, might easily be distilled in vacuo when in the temperature of only 70° Fahrenheit. But, by condensing steam formed at this temperature, and observing the quantity of heat which it communicated to a given weight of water, he determined that its latent heat, instead of being only 950°, was between 1200° and 1300°.

The same principle may be explained, also, by the following illustration, which was suggested to me by Mr. Ewart. Let us suppose that in a cylinder, furnished with a piston, we have a certain quantity of steam, and that it is suddenly compressed, by a stroke of the piston, into half its bulk. None of the steam will in this case be condensed; but it will acquire double elasticity, and its temperature will be considerably increased. Now if we either suppose the cylinder incapable of transmitting heat, or take the moment instantly following the compression before any heat has had time to escape, it must be evident that the sensible and latent heat of the steam, taken together before compression, are precisely equal to the sensible and latent heat taken together of the denser steam. But in the dense steam, the sensible heat is increased, and the latent heat proportionally diminished. The explanation of this fact will be furnished by a principle to be hereafter explained, that the capacities of elastic fluids for caloric are uniformly diminished by increasing their density.

The large quantity of caloric, latent in steam, renders its application extremely useful for practical purposes. Thus, water may be heated, as in the foregoing experiment, at a considerable distance from the source of heat, by lengthening the conducting pipe *e*. This furnishes us with a commodious method of warming the water of baths, which, in certain cases of disease, it is of importance to have near the patient's bed room; for the boiler, in which the water is heated, may thus be placed on the ground floor, or in the cellar of a house; and the steam conveyed by pipes into an upper apartment. Steam may also be applied to the purpose of heating or evaporating water, by a modification of the apparatus. Fig. 46, *g*, represents the apparatus for boiling water by the condensation of steam, without adding to its quantity; a circumstance occasionally of considerable importance. The steam is received between the vessel, which contains the water to be heated, and an exterior case; it imparts its caloric to the water, through the substance of the vessel; is thus condensed, and returns to the boiler by the perpendicular pipe. An alteration of the form of the vessel adapts it to evaporation (fig. 46, *h*). This method of evaporation is admirably suited to the concentration of liquids, that are decomposed, or injured, by a higher temperature than that of boiling water, such as medicinal extracts; to the drying of precipitates, &c. In the employment of either of these vessels, it is expedient to surround it with some slow conductor of heat. On a small scale, a few folds of woollen cloth are sufficient; and, when the vessel is constructed of a large size for practical use, this purpose is served by the brick work in which it is placed.

SECTION V.

Specific Caloric.

EQUAL weights of the *same* body, at the same temperature, contain the same quantities of caloric. But equal weights of *different* bodies, at the same temperature, contain unequal quantities of caloric. The quantity of caloric, which one body contains, compared with that contained in another, is called its *specific caloric;* and the power or property, which enables bodies to retain different quantities of caloric, has been called *capacity for caloric.* The method of determining the specific caloric, or comparative quantities of caloric in different bodies, is as follows:

It has already been observed, that equal weights of the *same* body, at different temperatures, give, on admixture, the arithmetical mean. Thus, the temperature of a pint of hot water and a pint of cold, is, after mixture, very nearly half way between that of the two extremes. But this is not the case, when equal quantities of *different* bodies, at different temperatures, are employed.

(*a*) If a pint of quicksilver at 100° Fahrenheit, be mixed with a pint of water at 40°, the resulting temperature will not be 70° (the arithmetical mean), but only 60°. Hence the quicksilver loses 40° of heat, which nevertheless raise the temperature of the water only 20°; in other words, a larger quantity of caloric is required to raise the temperature of a pint of water, than that of a pint of mercury, through the same number of degrees. Hence it is inferred, that water has a greater capacity for caloric than is inherent in quicksilver.

(*b*) The experiment may be reversed, by heating the water to a greater degree than the quicksilver. If the water be at 100°, and the mercury at 40°, the resulting temperature will be nearly 80°; because the pint of hot water contains more caloric, than is necessary to raise the quicksilver to the arithmetical mean.

(*c*) Lastly, if we take two measures of quicksilver to one of water, it is of no consequence which is the hotter; for the resulting temperature is always the mean between the two extremes; for example, 70°, if the extremes be 100° and 40°. Here, it is manifest, that the same quantity of caloric, which makes one measure of water warmer by 30°, is sufficient for making two measures of quicksilver warmer by the same number. Quicksilver has, therefore, a less capacity than water for caloric, in the proportion, when equal measures are taken, of one to two.

If, instead of equal *bulks* of quicksilver and water, we had taken equal *weights*, the disparity between the specific caloric of the mercury and water would have been still greater. Thus a pound of water at 100°, mixed with a pound of mercury at 40°, gives a temperature of 97½°, or 27½° above the arithmetical mean. In this experiment, the water, being cooled from 100° to 97½° has lost a quantity of caloric reducing its temperature only 2½°; but this caloric, communicated to the pound of mercury, has produced, in its temperature, a rise of no less than 57½°. Therefore, a quantity of caloric, necessary to raise the temperature of a pound of water 2½°, is sufficient to raise that of a pound of mercury 57½°; or, by the rule of proportion, the caloric, which raises the temperature of a pound of water 1°, will raise that of a pound of quicksilver about 23°. Hence it is inferred, that the quantity of caloric contained in water, is to that contained in the same *weight* of quicksilver as 23° to 1°. Or, stating the caloric of water at 1°, that of quicksilver will be $\frac{1}{23}$ part of 1°, or 0,0435*.

When this comparison is extended to a great variety of bodies, they will be found to differ very considerably in their capacities for caloric. The results of numerous experiments of this kind are comprised in a table of specific caloric. (See the Appendix.)

The capacities of bodies for caloric influence, considerably, the rate at which they are heated and cooled. In general, those bodies are most slowly heated, and cool most slowly, which have the greatest capacities for heat †. Thus, if water and quicksilver be set, in similar quantities, and at equal distances before the fire, the quicksilver will be much more rapidly heated than the water; and, on removal from the fire, it will cool with proportionally greater quickness than the water. By ascertaining the comparative rates of cooling, we may even determine, with tolerable exactness, the specific caloric of bodies; and particularly of one class (the gases), which are not easily compared in any other way. (See Leslie on Heat, chap. xxi.)

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* The above numbers, which differ from those commonly stated, are given on the authority of Mr. Dalton.

† See Martine, on Heat, page 74.

CHAPTER IV.

OF LIGHT.

THE laws of light, so far as they relate to the phenomena of its movement, and to the sense of vision, constitute the science of OPTICS; and are the objects, therefore, not of Chemistry, but of Natural Philosophy. Light, however, is capable of producing important chemical effects, and of entering into various chemical combinations. Its action is, for the most part, exerted in de-oxidizing bodies; and facts of this kind cannot be perfectly understood, until two important classes of bodies have been described, *viz.* those of oxides and of acids. In this place, therefore, I shall state only a few of its least complicated effects; and shall trace its agency on different bodies, as they become the objects of experiment in the sequel.

I. Light, in the state in which it reaches the organ of vision, it is well known, is not a simple body, but is capable of being divided, by the prism, into seven primary rays or colours, *viz.* red, orange, yellow, green, blue, indigo, and violet. These are refrangible in the above order, the red being least refrangible, and the violet most so. The image formed by the different rays, thus separated, constitutes the SOLAR SPECTRUM.

II. Heat and light are not present, in corresponding degrees, in different parts of the solar spectrum. With respect to the *illuminating* power of each colour, Dr. Herschell found that the red rays are far from having it in an eminent degree. The orange possess more of it than the red; and the yellow rays illuminate objects still more perfectly. The maximum of illumination lies in the brightest yellow or palest green. The green itself is nearly equally bright with the yellow; but from the full deep green, the illuminating power decreases very sensibly. That of the blue is nearly on a par with that of the red; the indigo has much less than the blue, and the violet is very deficient. (Philosophical Transactions, 1800, page 267.)

III. The *heating* power of the rays follows a different order.—If the bulb of a very sensible air thermometer be moved in succession, through the differently coloured rays, it will be found to indicate the greatest heat in the red rays; next in the green; and so on, in a diminishing progression, to the violet. The precise effects of the different rays, determined by Dr. Herschell's experiments, are as follows:

	The thermometer rose		
In the blue,	in 3	minutes from	55° to 56°
—— green,	in 3	——	54 to 58
—— yellow,	in 3	——	56 to 62
—— full red,	in 2½	——	56 to 72
— confines of red,	in 2½	——	58 to 73½

IV. When the thermometer is removed entirely out of the confines of the red rays, but with its ball still in the line of the spectrum, it rises even higher than in the red rays; and continues to rise, till removed half an inch beyond the extremity of the red rays. In this situation, quite out of the visible light, the thermometer rose in 2½ minutes from 61 to 79. The ball of the thermometer, employed for this purpose, should be extremely small*, and should be blackened with Indian ink. An air thermometer is better adapted than a mercurial one, to exhibit the minute change of temperature that ensues. These INVISIBLE HEAT-MAKING RAYS may be reflected by the mirror, and refracted by the lens, exactly in the same manner as the rays of light.

V. Beyond the confines of the spectrum on the other side, *viz.* a little beyond the violet ray, the thermometer is not affected; but in this place it is remarkable, that there are also invisible rays of a different kind, which exert all the chemical effects of the rays of light, and with even greater energy. One of the chemical properties of light, it will hereafter be stated, is, that it speedily changes, from white to black, the fresh-precipitated muriate of silver. (See chap. xviii. sect. 4.) This effect is produced most rapidly by the direct light of the sun; and the rays, as separated by the prism, have this property in various degrees. The blue rays, for example, effect a change of the muriate of silver in 15 seconds, which the red require 20 minutes to accomplish; and, generally speaking, the power diminishes as we recede from the violet extremity. But entirely out of the spectrum, and beyond the violet rays, the effect is still produced. Hence it appears, that the solar beams consist of three distinct kinds of rays: of those that excite heat, and promote oxidation; of illuminating rays; and of DE-OXIDIZING RAYS. A striking illustration of the different power of these various rays is furnished, by their effect on phosphorus. In the rays beyond the red extremity, phosphorus is heated, smokes, and emits white fumes; but these are presently suppressed, on exposing it to the de-oxidizing rays, which lie beyond the violet extremity.

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* Excellent thermometers for this purpose, and others requiring great sensibility, are made by Mr. Crichton, of Glasgow, and Mr. Cary, of London.

VI. There is an exception, however, as stated by Dr. Wollaston, to the de-oxidizing power of the rays above-mentioned. The substance, termed gum-guaiacum, has the property, when exposed to the light, of changing from a yellowish colour to green; and this effect he has ascertained to be connected with the absorption of oxygen. Now in the most refrangible rays, which would fall beyond the violet extremity, he found that this substance became green, and was again changed to yellow by the least refrangible. This is precisely the reverse of what happens to muriate of silver, which is blackened, or de-oxidized, by the most refrangible; and has its colour restored, or is again oxygenized, in the least refrangible rays.

VII. Certain bodies have the property of absorbing the rays of light in their totality; of retaining them for some time; and of again evolving them unchanged, and unaccompanied by sensible heat. Thus, in an experiment of Du Fay, a diamond exposed to the sun, and immediately covered with black wax, shone in the dark, on removing the wax, at the expiration of several months. Bodies, gifted with this property, are called SOLAR PHOSPHORI. Such are Canton's, Baldwin's, Homberg's, and the Bolognian phosphori, which will be described hereafter. To the same class belong several natural bodies, which retain light, and give it out unchanged. Thus snow is a natural solar phosphorus. So also is, occasionally, the sea when agitated; putrid fish have a similar property; and the glow-worm belongs to the same class. These phenomena are independent of every thing like combustion; for artificial phosphori, after exposure to the sun's rays, shine in the dark, when placed in the vacuum of an air-pump, or under water, &c., where no air is present to effect combustion.

VIII. From solar phosphori, the extrication of light is facilitated by the application of an elevated temperature; and, after having ceased to shine at the ordinary temperature, they again emit light when exposed to an increase of heat. Several bodies, which do not otherwise give out light, evolve it, or become phosphorescent, when heated. Thus, powdered fluate of lime becomes luminous, when thrown on an iron plate raised to a temperature rather above that of boiling water. The yolk of an egg, when dried, becomes luminous on being heated; and so also does tallow during liquefaction. To exhibit the last mentioned fact, it is merely necessary to place a lump of tallow on a coal, heated below ignition, making the experiment in a dark room.

IX. Attrition, also, evolves light. Thus, two pieces of common bonnet cane, rubbed strongly against each other in the dark, emit a faint light. Two pieces of borax have the same property much more remarkably.

X. Light is disengaged in various cases of chemical combination. Whenever combustion is a part of the phenomena, this is well known to happen; but light is evolved, also, in other instances, where nothing like combustion goes forwards. Thus, fresh prepared pure magnesia, added suddenly to highly concentrated sulphuric acid, exhibits a red heat.

XI. For measuring the relative intensities of light from various sources, an instrument has been contrived, called the PHOTOMETER. That of Count Rumford, described in the 84th volume of the Philosophical Transactions, being founded on optical principles, does not fall strictly within the province of this work. It is constructed on the principle, that the power of a burning body, to illuminate any defined space, is directly as the intensity of the light, and inversely as the square of the distance. If two unequal lights shine on the same surface at equal obliquities, and an opake body be interposed between each of them and the illuminated surface, the two shadows must differ in intensity or blackness; for the shadow formed by intercepting the greater light will be illuminated by the lesser light only; and, reversely, the other shadow will be illuminated by the greater light; that is, the stronger light will be attended with the deeper shadow. But it is easy, by removing the stronger light to a greater distance, to render the shadow, which it produces, not deeper than that of the smaller, or of precisely the same intensity. This equalization being effected, the quantity of light emitted by each lamp, or candle, will be as the square of the distance of the burning body from the white surface.

The photometer of Mr. Leslie is founded on a different principle, *viz.* that light, in proportion to its absorption, produces heat. The degree of heat produced, and consequently of light absorbed, is measured by the expansion of a confined portion of air. A minute description of the ingenious instrument contrived by Mr. Leslie with this view, may be seen in his work on Heat, or in the 3d vol. of Nicholson's 4to. Journal. In its construction, it bears a considerable resemblance to the differential thermometer, already described, page 114, and represented plate i. fig. 7. As both the balls of the latter instrument, however, are transparent, no change ensues in the situation of the coloured liquid when it is exposed to the variations of light. But, in the photometer, one of the balls is rendered opake, either by tinging the glass, or by covering it with a pigment; and hence this ball, absorbing the incident light which passes freely through the transparent one, the air included in it becomes warmer than that of the other ball, and, by its greater elasticity, forces the liquid up the opposite leg of the instrument. A graduated scale measures the amount of the effect; and a glass covering defends the photometer from being influenced by the temperature of the atmosphere.

CHAPTER V.

OF GASES.

SECTION I.

Of the Apparatus for Gases.

FOR performing the necessary experiments on gases, many articles of apparatus are essential, that have not hitherto been described. It may assist the student in obtaining the necessary instruments, if a few of the most essential be here enumerated. In this place, however, I shall mention such only, as are necessary in making a few general experiments on this interesting class of bodies.

The apparatus, required for experiments on gases, consists partly of vessels fitted for containing the materials that afford them, and partly of vessels adapted for the reception of gases, and for submitting them to experiment.

1. For procuring such gases as are producible without a very strong heat, glass bottles, furnished with ground stoppers and bent tubes, are sufficient (plate ii. fig. 18). Of these several will be required, of different sizes and shapes, adapted to different purposes. If these cannot be procured, a Florence flask, with a cork perforated by a bent glass tube, or even by a tin pipe, will serve for obtaining some of the gases.

Those gases that require, for their liberation, a red heat, may be procured, by exposing to heat the substance capable of affording them, in earthen retorts or tubes; or in a gun barrel, the touch-hole of which has been accurately closed by an iron pin. To the mouth of the barrel must be affixed a glass tube, bent so as to convey the gases where it may be requisite.

A very convenient apparatus, for obtaining such gases as cannot be disengaged without a red heat, is sold at the shops for philosophical apparatus in London. It consists of a cast-iron retort, to which a jointed metallic conducting tube is fitted by grinding; and by means of which the gas may be conveyed in any direction, and to any moderate distance. It is represented, as placed, when in actual use, between the bars of a common fire-grate (plate ix. fig. 85, *a*, *b*).

2. For receiving the gases, glass jars, of various sizes (figs. 21, 22, 23), are required, some of which should be furnished with necks at the top, fitted with ground stoppers. Others should be provided with brass caps, and screws, for the reception of air-cocks (fig. 22). Of these last (the air-cocks), several will be found necessary; and, to some of them, bladders, or elastic bottles, should be firmly tied, for the purpose of transferring gases. These jars will also be found extremely useful in experiments on the properties and effects of the gases. Some of them should be graduated into cubical inches.

To contain these jars, when in use, a vessel will be necessary, capable of holding a few gallons of water. This may either be of wood, if of considerable size; or, if small, of tin, japanned or painted. Plate iv. fig. 41, *ff* exhibits a section of this apparatus, which has been termed the pneumato-chemical trough, or pneumatic cistern. Its size may vary with that of the jars employed; and, about two or three inches from the top, it should have a shelf, on which the jars may be placed, when filled with air, without the risk of being overset. In this shelf should be a few small holes, to which inverted funnels may be soldered.

A glass tube, about 18 inches long, and three quarters of an inch diameter (fig. 24), closed at one end, and divided into cubic inches, and tenths of inches, will be required for ascertaining the purity of air by nitrous gas. It should be accompanied also with a small measure, containing about two cubic inches, and similarly graduated. For employing the solution of nitrous gas in liquid sulphate of iron (a happy invention of Mr. Davy, which leaves very little to be desired in eudiometry), glass tubes, about five inches long, and half an inch wide, divided decimally, are also necessary. Besides these, the experimentalist should be furnished with air funnels (fig. 19), for transferring gases from wide to narrow vessels.

An apparatus, almost indispensable in experiments on this class of bodies, is a GAZOMETER, which enables the chemist to collect and to preserve large quantities of gas, with the aid of only a few pounds of water. In the form of this apparatus there is considerable variety; but, at present, I have no other view than that of explaining its general construction and use. It consists of an outer fixed vessel *d* (plate iv. fig. 35), and an inner moveable one *c*, both of japanned iron. The latter slides easily up and down within the other, and is suspended by cords passing over pulleys, to which are attached the counterpoises, *ee*. To avoid the encumbrance of a great weight of water, the outer vessel *d* is made double, or is composed of two cylinders, the inner one of which is closed at the top and at the bottom. The space of only about half an inch is left between the two cylinders, as shown by the dotted lines. In this space the vessel *c* may move freely up and down. The interval is filled with water as high as the top of the inner cylinder. The cup, or rim, at the top of the outer vessel, is to prevent the water from overflowing, when the vessel *c* is forcibly pressed down, in which situation it is placed whenever gas is about to be collected. The gas enters from the vessel in which it is produced, by the communicating pipe *b*, and passes along the perpendicular pipe marked by dotted lines in the centre, into the cavity of the vessel *c*, which continues rising till it is full.

To transfer the gas, or to apply it to any purpose, the cock *b* is to be shut, and an empty bladder, or bottle of elastic gum, furnished with a stop-cock, to be screwed on *a*. When the vessel *c* is pressed down with the hand, the gas passes down the central pipe, which it had before ascended, and its escape at *b* being prevented, it finds its way up a pipe which is fixed to the outer surface of the vessel, and which is terminated by the cock *a*. By means of an ivory mouth-piece screwed upon this cock, the gas, included in the instrument, may be respired; the nostrils being closed by the fingers. When it is required to transfer the gas, into glass jars standing inverted in water, a crooked tube may be employed, one end of which is screwed upon the cock *b*; while the other aperture is brought under the inverted funnel, fixed into the shelf of the pneumatic trough. (See fig. 41, *c*.)

Several alterations have been made in the form of this apparatus; but they are principally such as add merely to its neatness and beauty, and not to its utility; and they render it less easy of explanation. The counterpoises *ee* are now, generally, concealed in the framing, and the vessel *c* is frequently made of glass.

When large quantities of gas are required (as at a public lecture), the gas-holder, plate iv. fig. 36, will be found extremely useful. It is made of tinned iron plate, japanned both within and without. Two short pipes, *a* and *c*, terminated by cocks, proceed from its sides, and another, *b*, passes through the middle of the top or cover, to which it is soldered, and reaches within half an inch of the bottom. It will be found convenient also to have an air-cock, with a very wide bore, fixed to the funnel at *b*. When gas is to be transferred into this vessel from the gazometer, the vessel is first completely filled with water through the funnel, the cock *a* being left open, and *c* shut. By means of a horizontal pipe, the aperture *a* is connected with *a* of the gazometer. The cock *b* being shut, *a* and *c* are opened, and the vessel *c* of the gazometer, fig. 35, gently pressed downwards with the hand. The gas then descends from the gazometer till the air-holder is full, which may be known by the water ceasing to escape through the cock *c*. All the cocks are then to be shut, and the vessels disunited. To apply this gas to any purpose, an empty bladder may be screwed on *a*; and water being poured through the funnel *b*, a corresponding quantity of gas is forced into the bladder. By lengthening the pipe *b*, the pressure of a column of water may be added; and the gas being forced through *a* with considerable velocity, may be applied to the purpose of a blow-pipe, &c., &c. The apparatus admits of a variety of modifications. The most useful one appears to me to be that contrived by Mr. Pepys, consisting chiefly in the addition of a shallow cistern (*e*, plate ix. fig. 85) to the top of the air-holder, and of a glass register tube *f*, which shows the height of the water, and consequently the quantity of gas, in the vessel. A more minute account of it will be given in the description of the ninth plate*.

The gazometer, already described, is fitted only for the reception of gases that are confinable by water; because quicksilver would act on the tinning and solder of the vessel, and would not only be spoiled itself, but would destroy the apparatus. Yet an instrument of this kind, in which mercury can be employed, is peculiarly desirable, on account of the great weight of that fluid; and two varieties of the mercurial gazometer have therefore been invented. The one, of glass, is the contrivance of Mr. Clayfield, and may be seen represented in the plate prefixed to Mr. Davy's Researches. In the other, invented by Mr. Pepys, the cistern for the mercury is of cast-iron. A drawing and description of it may be found in the 5th vol. of the Philosophical Magazine; but as neither of these instruments are essential to the chemical student, and as they are required only in experiments of research, I deem it sufficient to refer to the minute descriptions of their respective inventors.

For those gases that are absorbed by water, a mercurial trough is necessary. For the mere exhibition of a few experiments on these condensible gases, a small wooden trough, 11 inches long, two wide, and two deep, cut out of a solid block of mahogany, is sufficient; but for experiments of research, one of considerable size is required. (See plate iii. fig. 31, *ff*.)

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* Descriptions and figures of improved gas-holders may be seen also in the 13th, 24th, and 27th vols. of the Philosophical Magazine.

The apparatus, required *for submitting gases to the action of electricity*, is shown in plate ix. fig. 84; where *a* represents the knob of the prime conductor of an electrical machine; *b* a Leyden jar, the ball of which is in contact with it, as when in the act of charging; and *c* the tube standing inverted in mercury, and partly filled with gas. The mercury is contained in a strong wooden box *d*, to which is screwed the upright iron pillar *e*, with a sliding collar for securing the tube *c* in a perpendicular position. When the jar *b* is charged to a certain intensity, it discharges itself between the knob *a* and the small ball *i*, which, with the wire connected with it, may be occasionally fitted on the top of the tube *c*. The strength of the shocks is regulated by the distance between *a* and *i*.

By the same apparatus, inflammable mixtures of gases may be exploded by electricity. In this case, however, the jar *b* is unnecessary, a spark received by *i* from *a* being sufficient to kindle the mixture.

The method of *weighing gases* is very simple and easily practised. For this purpose, however, it is necessary to be provided with a good air-pump; and with a globe or flask, furnished with a brass cap and air-cock, as shown fig. 22, *b*. A graduated receiver is, also, required, to which an air-cock is adapted, as shown at fig. 22, *a*.

Supposing the receiver *a* to be filled with any gas, the weight of which is to be ascertained, we screw the cock of the vessel *b* on the transfer plate of an air-pump, and exhaust it as completely as possible. The weight of the exhausted vessel is then very accurately taken, even to a small fraction of a grain; and it is screwed upon the air-cock of the receiver *a*. On opening both cocks, the last of which should be turned very gradually, the gas ascends from the vessel *a*; and the quantity, which enters into the flask, is known by the graduated scale on *a*. On weighing the vessel a second time, we ascertain how many grains have been admitted. If we have operated on common air, we shall find its weight to be at the rate of about 31 grains to 100 cubical inches. The same quantity of oxygen gas will weigh about 34 grains, and of carbonic acid gas upwards of 47 grains.

In experiments of this kind it is necessary either to operate with the barometer at 30 inches, and the thermometer at 60° Fahrenheit, or to reduce the volume of gas employed to that pressure and temperature, by rules which are given in the Appendix. Great care is to be taken, also, not to warm any of the vessels by contact with the hands, from which they should be defended by a glove. On opening the communication between the receiver and the exhausted vessel, if any water be lodged in the air-cock attached to the former, it will be forcibly driven into the latter, and the experiment will be frustrated. This may be avoided by using great care in filling the receiver with water, before passing into it the gas under examination.

The specific gravity of any gas compared with common air is readily known, when we have once determined its absolute weight. Thus if 100 cubic inches of air weigh 31 grains, and the same quantity of oxygen gas weighs 34 grains, we say,

31 : 34 : : 1000 : 1103.

The specific gravity of oxygen gas will therefore be as 1103 to 1000. We may determine, also, the specific gravity of gases, more simply, by weighing the flask, first when full of common air, and again when exhausted; and afterwards by admitting into it as much of the gas under examination as it will receive; and weighing it a third time. Now as the loss between the first and second weighing is to the gain of weight on admitting the gas, so is common air to the gas whose specific gravity we are estimating. Supposing, for example, that by exhausting the flask it loses 31 grains, and that by admitting carbonic acid it gains 47; then

31 : 47 : : 1000 : 1516.

The specific gravity of carbonic acid is therefore 1516, air being taken at 1000. And knowing its specific gravity, we can, without any farther experiment, determine the weight of 100 cubic inches of carbonic acid; for as the specific gravity of air is to that of carbonic acid, so is 31 to the number required; or

1000 : 1516 : : 31 : 47.

One hundred inches of carbonic acid, therefore, will weigh 47 grains.

Previously to undertaking experiments on other gases, it may be well for an unpractised experimentalist to accustom himself to the dexterous management of gases, by transferring common air from one vessel to another of different sizes.

1. When a glass jar, closed at one end, is filled with water, and held with its mouth downwards, in however small a quantity of water, the fluid is retained in its place by the pressure of the atmosphere on the surface of the exterior water. Fill in this manner, and invert, on the shelf of the pneumatic trough, one of the jars, which is furnished with a stopper (fig. 23). The water will remain in the jar so long as the stopper is closed; but immediately on removing it, the water will descend to the same level within as without; for it is now pressed, equally upwards and downwards, by the atmosphere, and falls therefore in consequence of its own gravity.

2. Place the jar, filled with water and inverted, over

one of the funnels of the shelf of the pneumatic trough. Then take another jar, filled (as it will be of course) with atmospherical air. Place the latter with its mouth on the surface of the water; and on pressing it in the same position below the surface, the included air will remain in its situation. Bring the mouth of the jar beneath the funnel in the shelf, and incline it gradually. The air will now rise in bubbles, through the funnel, into the upper jar, and will expel the water from it into the trough.

3. Let one of the jars, provided with a stop cock at the top, be placed full of air on the shelf of the trough. Screw upon it an empty bladder; open the communication between the jar and the bladder, and press the former into the water. The air will then pass into the bladder, till it is filled; and when the bladder is removed from the jar, and a pipe screwed upon it, the air may be again transferred into a jar inverted in water.

4. For the purpose of transferring gases from a wide vessel standing over water, into a small tube filled with and inverted in mercury, I have long used the following contrivance of Mr. Cavendish. A tube, eight or 10 inches long, and of very small diameter, is drawn out to a fine bore, and bent at this end, so as to resemble the italic letter *l*. The point is then immersed in quicksilver, which is drawn into the tube till it is filled, by the action of the mouth. Placing the finger over the aperture at the straight end, the tube is next conveyed through the water, with the bent end uppermost, into an inverted jar of gas. When the finger is removed, the quicksilver falls from the tube into the trough, or into a cup placed to receive it, and the tube is filled with the gas. The whole of the quicksilver, however, must not be allowed to escape; but a column must be left, three or four inches long, and must

be kept in its place by the finger. Remove the tube from the water; let an assistant dry it with blotting paper; and introduce the point of the bent end into the aperture of the tube standing over quicksilver. On withdrawing the finger from that aperture which is now uppermost, the pressure of the column of quicksilver, added to the weight of the atmosphere, will force the gas from the bent tube into the one standing in the mercurial trough.

On every occasion, when it is necessary to observe the precise quantity of gas, at the commencement and close of an experiment, it is essential that the barometer and thermometer should exactly correspond at both periods. An increased temperature, or a fall of the barometer, augments the apparent quantity of gas; and a reduced temperature or a higher barometer diminishes its bulk. Another circumstance, an attention to which is indispensable in all accurate experiments, is that the surface of the fluid, by which the gas is confined, should be precisely at the same level within and without the jar. If the fluid be higher within the jar, the contained gas will be under a less pressure than that of the atmosphere, the weight of which is counterpoised by that of the column of fluid within. In mercury, this source of error is of very considerable amount; as any person may be satisfied by pressing down, into quicksilver, a tube partly filled with that fluid, and partly with air, for the volume of the air will gradually decrease, the deeper the tube is immersed.

In experiments on gases, it is not always possible, however, to begin and conclude an experiment at precisely the same temperature, or with the same height of the barometer; or even to bring the mercury within and without the receiver to the same level. In these cases, therefore, calculation becomes necessary; and, with the view of comparing results more readily and accurately, it is usual to reduce quantities of gas to the bulk they would occupy under one given pressure, and at a given temperature. In this country, it is now customary to assume as a standard 30 inches of the barometer, and 60° of Fahrenheit's thermometer; and to bring, to these standards, observations made under other degrees of atmospheric pressure and temperature. The rules, for these corrections, which are sufficiently simple, I shall give in the Appendix.

Of experiments illustrative of the nature of gases in general, it may be proper to mention one or two, that show the mode in which caloric exists in this class of bodies. In vapours, strictly so called, as the steam of water, caloric seems to be retained with but little force; for it quits the water when the vapour is merely exposed to a lower temperature. But, in gases, caloric is united by very forcible affinity, and no diminution of temperature, that has ever yet been effected, can separate it from some of them. Thus the air of our atmosphere, in the most intense artificial or natural cold, still remains in the aëriform state. Hence is derived one character of gases, *viz.* that they remain aëriform under almost all variations of pressure and temperature; and in this class are also included those aërial bodies, which, being condensed by water, require confinement over mercury. The following experiment will show, that the caloric, contained in gases, is chemically combined.

Into a small retort (plate ii. fig. 20, *b*) put an ounce or two of well dried common salt, and about half its weight of sulphuric acid. By this process, a great quantity of gas is produced, which might be received and collected over mercury. But, to serve the purpose of this experiment, let it pass through a glass balloon, *c*, having three openings, into one of which the neck of the retort passes, while, from the other, a tube *e* proceeds, which ends in a vessel of water, *f*, of the temperature of the atmosphere. Before closing the apparatus, let a thermometer *d* be included in the balloon, to show the temperature of the gas. It will be found, that the mercury, in this thermometer, will rise only a few degrees, whereas the water, in the vessel which receives the bent tube, will soon become boiling hot. In this instance, caloric flows from the lamp to the muriatic acid, and converts it into gas; but the heat, thus expended, is not appreciable by the thermometer, and must, therefore, be chemically combined. The caloric, however, is again evolved, when the gas is condensed by water. In this experiment, we trace caloric into a latent state, and again into the state of free or uncombined caloric.

A considerable part of the caloric, which exists in gases in a latent state, may be rendered sensible by rapid mechanical compression. Thus if air be suddenly compressed in the ball of an air-gun, the quantity of caloric liberated by the first stroke of the piston, is sufficient to set fire to a piece of the tinder called *amadou*. (Philosophical Magazine, xiv. 363.) A flash of light is said, also, to be perceptible at the moment of condensation. This fact has been applied to the construction of a portable instrument for lighting a candle. It consists of a common syringe, concealed in a walking stick. At the lower extremity, the syringe is furnished with a cap, which receives the substance intended to be fired, and which is attached to the instrument by a male and female screw. The rapid depression of the piston condenses the air, and evolves sufficient heat to set the tinder on fire. (Philosophical Magazine, xxxi. 130.)

For demonstrating the influence of variations of atmospheric pressure on the formation of gases, better experiments cannot be devised than those of Lavoisier. (See his Elements, chap. i.) But as some students, who have the use of an air-pump, may not possess the apparatus described by Lavoisier (the glass bell and sliding wire), it may be proper to point out an easier mode of showing the same fact. This proof is furnished by the experiment described, page 142, 3, in which ether is made to assume alternately an aëriform and liquid state, by removing and restoring the pressure of the atmosphere.

Gases, when once formed, undergo a considerable change of bulk by variations of external pressure. The general law, which has been established on this subject is, that *the volume of gases is inversely as the compressing force.* If, for example, we have a quantity of gas occupying 60 cubic inches under the common pressure of the atmosphere, they will fill the space of only 30 cubic inches, or one half under a double pressure; of 20 inches, or one 3d, under a triple pressure; of 15 inches, or one 4th, under four times the pressure; and so on.

The law of the dilatability of gases by heat has already been stated to be an enlargement of about $\frac{1}{180}$th part of their bulk for each degree of Fahrenheit's scale, between the freezing and boiling points of water.

The following table exhibits an arrangement of all the permanent gases hitherto known, classed according to the resemblance of their properties.

Class I.—*Gases with simple Bases.*

Oxygen gas.
Hydrogen gas.
Nitrogen gas.

CLASS II.—*Gases with compound Bases.*

1. *Compound combustible Gases.*

Name.	Composed of
1. Ammonia	Hydrogen and nitrogen.
2. Carbureted hydrogen .	Hydrogen and carbon.
3. Olefiant gas	Carbureted hydrogen and carbon.
4. Sulphureted hydrogen .	Hydrogen and sulphur.
5. Phosphureted hydrogen .	Hydrogen and phosphorus.
6. Arsenureted hydrogen .	Hydrogen and arsenic.
7. Potassureted hydrogen .	Hydrogen and potassium.

2. *Oxides in the state of Gas.*

1. Carbonous oxide . . .	Carbon and oxygen.
2. Nitrous oxide	Nitrogen and oxygen.
3. Nitric oxide	Nitrous oxide and oxygen.

3. *Acid Gases.*

1. Carbonic acid	Carbon and oxygen.
2. Sulphurous acid . . .	Sulphur and oxygen.
3. Nitrous acid	Nitric oxide and oxygen.
4. Nitric acid	Nitrous acid and oxygen.
5. Muriatic acid	Unknown. [illegible]
6. Oxy-muriatic acid . .	Muriatic acid and oxygen.
7. Hyper-oxymuriatic acid .	Oxymuriatic acid and oxygen.
8. Fluoric acid	Fluorium and oxygen.
9. Fluoboric acid . . .	Fluoric and boracic acids.

SECTION II.

Oxygen Gas.

We have no knowledge of the properties of oxygen in a state of complete separation. In the most simple form, under which we can procure it, it is combined with caloric, and probably with light and electricity, constituting oxygen gas.

I. *Oxygen gas may be procured from various substances.*

1. From the black oxide of manganese, heated to redness in a gun-barrel, or in an iron or earthen retort; or, from the same oxide, heated by a lamp in a retort or gas bottle, with half its weight of strong sulphuric acid.

2. From the red oxide of lead (the common red lead used by painters), heated either with or without sulphuric acid.

3. From various other oxides, as will be hereafter mentioned.

4. From nitrate of potash (common saltpetre) made red hot in a gun barrel, or in an earthen retort.

5. From oxygenized muriate of potash, heated in a small glass retort, over an Argand's lamp. The oxygen gas thus produced, is much purer than that obtained in any other mode, especially the last portions, which should be kept separate.

All these substances, after having yielded oxygen gas, are found considerably diminished in weight; and calculating each cubic inch of gas to be equal to one 3d of a grain, the loss of weight will be found pretty exactly equivalent to that of gas generated.

II. This gas has the following properties:

1. *It is not absorbed by water**; or, at least, is so sparingly absorbed, that, when agitated in contact with water, no perceptible diminution takes place.

2. *It is rather heavier than common air.*—Mr. Davy stated 100 cubic inches, at 55° Fahrenheit, and 30 inches of the barometer, to weigh 35.06 grains; and at the temperature of 60°, the same quantity would weigh 34.70. Messrs. Allen and Pepys have determined 100 cubic inches to weigh 33.82 grains, the barometer being 30, and thermometer 60°. By Biot and Arajo its specific gravity is stated to be 1.10359.

2. *All combustible bodies burn in oxygen gas with greatly increased splendour.*

(*a*) A lighted wax taper, fixed to an iron wire, and let down into a vessel of this gas, burns with great brilliancy, pl. iv. fig. 38. If the taper be blown out, and let down into a vessel of the gas while the snuff remains red hot, it instantly rekindles, with a slight explosion.

(*b*) A red hot bit of charcoal, fastened to a copper wire, and immersed in the gas, throws out beautiful sparks.

(*c*) The light of phosporus, burnt in this gas, is the brightest that can be in any mode produced. Let the phosphorus be placed in a small hemispherical tin cup, which may be raised by means of the wire stand, pl. ii. fig. 25, two or three inches above the surface of water contained in a broad shallow dish. Fill a bell-shaped receiver, having an open neck at the top, to which a compressed bladder is firmly tied, with oxygen gas; and, as it stands inverted in water, press a circular piece of pasteboard, rather exceeding the jar in diameter, over its mouth. When an assistant has set fire to the phosphorus, cover it instantly with the jar of oxygen gas, retaining the pasteboard in its place, till the jar is immediately over the cup. When this has been skilfully managed, a very small portion only of the gas will escape; and the inflammation of the phosphorus will be extremely brilliant. The expanded gas rises into the flaccid bladder, and is thus prevented from escaping into the room, and proving disagreeable by its suffocating smell.

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* In this as in several other instances, where a gas is said not to be absorbed by water, the assertion is not to be taken strictly, but merely as implying that only a minute and difficultly appreciable portion is absorbed. The precise proportion of each gas absorbed by water is stated in the Appendix, in the form of a table.

(*d*) Substitute, for the phosphorus in experiment *c*, a small ball formed of turnings of zinc, and in which about a grain of phosphorus is inclosed. Set fire to the phosphorus, and cover it expeditiously with the jar of oxygen. The zinc will be inflamed, and will burn with a beautiful white light. A similar experiment may be made with metallic arsenic, which may be moistened with spirit of turpentine. The filings of various metals may also be inflamed, by placing them in a small cavity, formed in a piece of charcoal, igniting the charcoal, and blowing on the part containing the metal a stream of oxygen gas.

(*e*) Procure some thin harpsichord wire, and twist it round a slender rod of iron or glass, so as to coil it up in a spiral form. Then withdraw the rod, and tie a little thread or flax round one end of the wire, for about one 20th of an inch; which end is to be dipped into melted sulphur. The other end of the wire is to be fixed into a cork; so that the spiral may hang vertically (fig. 39). Fill, also, with oxygen gas, a bottle capable of holding about a quart, and set it with its mouth upwards. Then light the sulphur, and introduce the wire into the bottle of gas, suspending it by the cork. The iron will burn with a most brilliant light, throwing out a number of sparks, which fall to the bottom of the bottle, and generally break it. This accident, however, may frequently be prevented by pouring sand into the bottle, so as to lie about half an inch deep on the bottom (see pl. iv. fig. 39). According to Mr. Accum (Nicholson's Journal, 8vo. i. 320), a thick piece of iron or steel, such as a file, if made sharp pointed, may be burnt in oxygen gas. A small bit of wood is to be stuck upon its extremity, and set on fire, previously to immersion in the gas.

(*f*) A little of Homberg's pyrophorus, a substance to be hereafter described, when poured into a bottle full of this gas, immediately flashes like inflamed gunpowder.

III. *During every combustion in oxygen gas, the gas suffers a considerable diminution.*—To exhibit this, experimentally, in a manner perfectly free from all sources of error, would require such an apparatus as few but adepts in chemistry are likely to possess. The apparatus required for this purpose is described in the 6th chapter of Lavoisier's Elements. The fact may, however, be shown, less accurately, in the following manner: Fill, with oxygen gas, a jar of moderate size, which has a neck and ground-glass stopper at the top. Then, with the assistance of a stand, formed of bent iron wire (pl. ii. fig. 25), place a shallow tin vessel, containing a bit of phosphorus or sulphur, three or four inches above the level of the water of a pneumatic trough. Invert the jar of oxygen gas, cautiously and expeditiously, over this cup, so as to confine it, with its contents, in the gas, and, pressing down the jar to the bottom of the trough, open the stopper. A quantity of gas will immediately rush out, and the water will rise to the same level within the jar as without. When this has taken place, set fire to the sulphur or phosphorus by a heated iron wire, and instantly put in the stopper. The first effect of the combustion will be a depression of the water within the jar; but when the combustion has closed, and the vessel has cooled, a considerable absorption will be found to have ensued.

Those persons who are possessed of a mercurial apparatus may repeat this experiment in a less exceptionable manner. On the surface of the quicksilver let a small hemispherical cup float, made of untinned sheet-iron; and, in order to keep it from the sides of the jar, it may rest on a wire-stand, shaped like the figure 43, plate iv. Let a jar, the height and diameter of which must be regulated by the size of the mercurial trough, be filled with oxygen gas over water, and be removed, by means of a piece of pasteboard, as before described, to the mercurial bath, inverting it dexterously over the tin cup. If the phosphorus had been previously set on fire, a large quantity of the gas, expanded by the heat, would have escaped, and would have prevented the accurate measurement of the absorption. After drying the surface of the mercury within the jar by blotting paper, a portion of the included gas must, therefore, be removed. This is done by an inverted syphon, one leg of which is to be introduced (in the same manner as is shown at fig. 41, *g*) within the jar before placing it over the mercury; and the gas will be forced through the open extremity of the other, when the jar is pressed down into the quicksilver. When the proper quantity has been expelled, remove the syphon. The cup, containing the phosphorus, will thus rest on the surface of the quicksilver within the jar, and above the level of the mercury without. The phosphorus is to be inflamed by passing a crooked iron wire, made red hot, through the quicksilver. On the first impression of the heat arising from its combustion, the included gas will be considerably expanded; but when the phosphorus has ceased to burn, a considerable absorption will be found to have taken place, the amount of which may be measured by ascertaining the height of the quicksilver within the jar, before and after the experiment. The quantity of phosphorus employed should be very small, and should not bear a greater proportion than that of 10 grains to each pint of gas; otherwise the combustion will go on so far as to endanger the breaking of the jar, by the approach of the inflamed phosphorus.

In this process, a white dense vapour is produced, which concretes on the inner surface of the jar in flakes. This substance has strongly acid properties; and, being formed by the union of oxygen with phosphorus, is termed the phosphoric acid.

The diminution of the volume of oxygen gas, by the combustion of other bodies, may be ascertained in a similar manner. When the substance employed is not easily set on fire, it is proper to enclose, along, and in contact with it, a small bit of phosphorus, the combustion of which excites sufficient heat to inflame iron-turnings, charcoal, &c. In the instance of charcoal, however, though that substance undergoes combustion, no absorption ensues; because, as will appear in the sequel, the product is a gas, occupying exactly the same bulk as the oxygen gas submitted to experiment.

IV. *All bodies, by combustion in oxygen gas, acquire an addition to their weight; and the increase is in proportion to the quantity of gas absorbed, viz. about one 3d* *of a grain for every cubic inch of gas.*—To prove this by experiment, requires also a complicated apparatus.

But sufficient evidence of this fact may be obtained by the following very simple experiment. Fill the bowl of a tobacco pipe with iron wire coiled spirally, and of known weight: let the end of the pipe be slipped into a brass tube, which is screwed to a bladder filled with oxygen gas: heat the bowl of the pipe, and its contents, to redness in the fire, and then force through it a stream of oxygen gas from the bladder. The iron wire will burn; will be rapidly oxydized; and will be found, when weighed, to be considerably heavier than before. When completely oxydized in this mode, 100 parts of iron wire gain an addition of about 30.

V. *Every substance, capable of union with oxygen, affords, by combustion, either an oxide, an acid, or an alkali.*—When a body, by being burnt in oxygen gas, affords a compound, which has none of those qualities that characterize acids or alkalis, we denominate this product an *oxide*. If we collect, for example, the iron wire, which was burned in the last experiment, we shall find that it has lost all its metallic qualities; and has become a brittle dark coloured substance totally destitute of lustre and of taste, and termed *an oxide of iron*. But if, instead of iron wire, we had burned a quantity of sulphur in oxygen gas, the result would have been that the water, which confined the gas, would have become acid or sour. Potassium on the contrary (one of the new metals discovered by Mr. Davy), would have yielded an alkali under the same circumstances. Hence the extensive class of combustible bodies may be subdivided into three orders, 1st, those which afford oxides by combustion; 2dly, those which yield acids; and 3dly, those which give alkalis. In many instances, however, a body is capable of passing through the intermediate state of an oxide, before it is converted either into an acid or an alkali.

VI. *Oxygen gas supports, eminently, animal life.*—It will be found that a mouse, bird, or other small animal, will live six times longer in a vessel of oxygen gas, than in one of atmospherical air of the same dimensions.

VII. *This effect seems connected with the absorption of oxygen by the blood.*—Pass up a little dark-coloured blood into a jar partly filled with oxygen gas, and standing over mercury. The gas will be in part absorbed, and the colour of the blood will be changed to a bright and florid red. This change to red may be shown, by putting a little blood into a common vial filled with oxygen gas, and shaking it up.

SECTION III.

Nitrogen or Azotic Gas.

AFTER separating, from any quantity of atmospherical air, all the oxygen which it contains, there remains a gas which was called by Lavoisier *azotic gas*, a name applied to it in consequence of its unfitness for supporting animal life; and derived from the Greek privative *a* and ζωὴ *vita*. This, however, as being merely a negative property, has since been deemed an improper foundation for its nomenclature; and the term NITROGEN gas has been substituted; because one of the most important properties of its base is, that by union with oxygen it composes nitric acid. By this appellation, therefore, I shall hereafter distinguish it.

Nitrogen has been hitherto considered as a simple or elementary body; but the recent experiments of Mr. Davy, which will be described in speaking of ammonia, suggest that it is a compound of hydrogen and oxygen.

I. Nitrogen gas may be procured, though not absolutely pure, yet sufficiently so for the purpose of exhibiting its general properties, in any of the following manners: 1. Mix equal weights of iron filings and sulphur into a paste with water, and place the mixture, in a proper vessel, over water, supported on a stand: then invert over it, a jar full of common air, and allow this to stand exposed to the mixture for a day or two. The air contained in the jar will gradually diminish, as will appear from the ascent of the water within the jar, till at last only about four 5ths of its original bulk will remain. The vessel containing the iron and sulphur must next be removed, by withdrawing it through the water; and the remaining air may be made the subject of experiment.

2. A quicker process, for procuring nitrogen gas, consists in filling a bottle, about one 4th, with the solution of nitrous gas, in liquid sulphate of iron, and agitating it with the air which fills the rest of the bottle. During the agitation, the thumb must be firmly placed over the mouth of the bottle; and, when removed, the mouth of the bottle must be immersed in a cup full of the same solution, which will supply the place of the absorbed air.

3. Atmospheric air, also, in which phosphorus has burned out, affords, when time has been allowed for the condensation of the phosphoric acid, tolerably pure nitrogen gas.

II. Nitrogen gas has the following properties:

1. It is not absorbed by water.

2. *It is a little lighter than atmospheric air,* 100 cubic inches being found by Mr. Davy to weigh 30.04 grains

under a pressure of 30 inches, and at the temperature of 55° Fahrenheit. At 60° Fahrenheit 100 inches weigh, therefore, 29.73 grains. According to Biot and Arajo, its specific gravity is 0.96913.

3. It immediately extinguishes a lighted candle, and all other burning substances.

Even phosphorus, in a state of active inflammation, is immediately extinguished when immersed in nitrogen gas. This is best shown by placing the burning phosphorus in a tin cup, raised by a stand above the surface of the water, and quickly inverting over it a jar filled with nitrogen gas.

4. It is fatal to animals that are confined in it.

5. When mixed with pure oxygen gas, in the proportion of four parts to one of the latter, it composes a mixture resembling atmospheric air in all its properties. Of this any one may be satisfied, by mixing four parts of azotic gas with one of oxygen gas, and immersing, in the mixture, a lighted taper. The taper will burn as in atmospherical air.

SECTION IV.

Atmospheric Air.

The air of our atmosphere, it appears, therefore, from the foregoing facts, is a mixture, or possibly a combination, of two different gases, *viz.* oxygen gas and azotic gas. The former of these two seems to be the only ingredient on which the effects of the air, as a chemical agent, depend. Hence combustible bodies burn in atmospheric air, only in consequence of the oxygen gas which it contains; and, when this is exhausted, air is no longer capable of supporting combustion*. Its analysis is most satisfactorily and beautifully demonstrated by the action of heated mercury, as explained by Lavoisier in the third chapter of his Elements of Chemistry. By exposure, during 12 days, to mercury heated in a retort; a given quantity of atmospheric air was found to be diminished in bulk, and to have lost its property of supporting combustion. The mercury, also, had suffered a considerable change; a part of it was no longer a shining fluid metal; but was changed into red scaley particles; and its weight was, also, a little increased. These red particles were collected, and distilled in a retort; by which operation a quantity of oxygen gas was evolved, precisely equal in bulk to what the air had lost in the first part of the experiment. These results afford the most satisfactory evidence, that the air of our atmosphere is composed of two distinct fluids. The one is capable of yielding its base to mercury; and, when separate, is eminently adapted to the support of combustion and of animal life; the other is not absorbable by mercury, and is destitute of both the latter important qualities.

The details of this method of analyzing atmospheric air I omit on account of the extreme tediousness of the process. Sufficient evidence of its composition may be obtained, however, much more expeditiously, by the following experiments.

I. Burn a little sulphur or phosphorus, in the manner described, sect. ii. substituting, for oxygen gas, common atmospherical air. The combustion will, in this instance, be less vivid; will cease sooner; and the absorption, when

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[*] Certain combustible bodies even cease to burn in atmospheric air, long before its oxygenous portion is consumed, for reasons that will hereafter be given.

the vessels have cooled, will be much less considerable than in the former case.

The phosphorus, however, will have absorbed the whole of the oxygen gas contained in the air submitted to experiment; and hence it may be employed for measuring the quantity of oxygen gas in a given bulk of atmospherical air. This may be accomplished, either by its slow or rapid combustion. Berthollet proposes (*Annales de Chimie*, xxxiv. 78) to expose a cylinder of phosphorus, fastened to a glass rod, in a narrow glass vessel, graduated into equal parts, and standing full of air over water. (See fig. 24). The phosphorus immediately begins to act on the included air; and in six or eight hours its effect is completed. The residuary azotic gas has its bulk enlarged about one 40th, by absorbing a little phosphorus; and, for this, allowance must be made in measuring the diminution.

In the eudiometer of Seguin, the rapid combustion of phosphorus is employed with the same view. A glass tube, open at one end only, about an inch in diameter, and eight or ten high, is filled with, and inverted in, mercury. A small bit of phosphorus, dried with blotting paper, is then introduced, and, by its inferior specific gravity, rises to the top of the tube, where it is melted, by bringing a red-hot poker near to the outer surface of the glass. When the phosphorus is liquefied, a measured portion of the air to be examined is admitted, by a little at once, into the tube. The phosphorus inflames at each addition, and the mercury rises. When all the air under examination has been added, the red-hot poker is again applied to ensure the completion of the process, and the residuary gas is transferred into a graduated measure, where its bulk is carefully ascertained. In this instance, about one 40th the volume of the residuary gas is to be deducted from the apparent quantity of azotic gas, because, in this case also, a small portion of phosphorus is dissolved by the latter, and occasions a trifling expansion. With this deduction, atmospheric air loses pretty accurately 21 parts out of each 100; and contains, therefore, 21 *per cent.* of oxygen, and 79 of azote by measure*.

II. The inferior fitness of atmospherical air, to that of oxygen gas, for supporting combustion, may be shown, also, by a comparative experiment with two candles. Provide a circular piece of lead, three inches diameter, and half an inch thick, from the centre of each of which proceeds a perpendicular iron wire, six or eight inches high; to the end of this wire fasten a piece of wax taper. Set the candle, supported by its stand, on the shelf of a pneumatic cistern; and place, also, the conducting pipe from the bladder (*e*, fig. 41), in the position shown by the figure; the cock *d*, however, being shut. Then, having the syphon *g* in the inverted position shown in the plate, sink the whole apparatus into the water. Part of the air in the jar *a* will escape through the syphon, and will be replaced by water. When we have left, in the jar, the proper quantity of air, the syphon must be removed, and the jar returned to its place. The level of the water will now be considerably higher within than without the receiver; and its height must be noted. On passing a succession of electrical sparks from the conducting wire to the bent pipe, and opening the cock *d* from the bladder filled with hydrogen gas, we shall have a small flame, which is to be extinguished as soon as, by its means, we have lighted the candle. The candle may be suffered to burn till it is extinguished; and the duration of its burning, and the diminution it occasions in the air, are to be noted. When the combustion is repeated in the same manner, but with the substitution of oxygen gas, it will be found to last considerably longer, and the diminution of volume in the gas will be much greater.

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* Various other methods of analyzing atmospherical air will be described in the course of the work References to them may be found in the Index, article Eudiometer.

The same fact may be demonstrated, but less accurately, by a simple apparatus. Provide two jars, each two inches diameter, and 12 inches long, and each having a neck at the top with a compressed bladder tied upon it. Fill one of the jars, leaving the bladder empty, with oxygen gas; and, at the same instant, with the aid of an assistant, invert both jars over the burning candles, keeping the oxygen gas in its place till the jar is inverted, by a piece of pasteboard. In the common air, the candle will soon be extinguished; but that confined in oxygen gas will burn with much greater splendour, and will continue burning long after the other is extinguished. On the first impression of the flame, a quantity of expanded gas will rise into each bladder, which is to be pressed out at the close of the experiment, in order that the absorption may be compared in both cases. The diminution in the jar of oxygen gas will be found greatly to exceed that of the common air.

III. Take two tubes, each a few inches long, closed at one end, and divided into 100 aliquot parts. Fill the one with atmospherical air, the other with oxygen gas, and invert them in two separate cups filled with a solution of sulphuret of potash. The sulphuret will ascend gradually within the tube of common air, till, after a few days, only about four 5ths of its original volume will remain; but, in that containing oxygen, it will ascend much higher, and, if the gas be pure, will even absorb the whole.

The explanation of this fact is, that liquid sulphuret of potash has the property of absorbing oxygen, but not nitrogen. It therefore acts on atmospheric air only as long as any oxygen gas remains, and may be employed as a means of ascertaining the quantity of this gas in the atmosphere at different times, and in distant places. An improved instrument, thus graduated, has lately been employed by Guyton as an *Eudiometer*[*]. (See Nicholson's Philosophical Journal, 4to. vol. i. page 268; or Tilloch's Philosophical Magazine, vol. iii. page 171.) But an apparatus, of much greater simplicity, and facility of application, is that of Professor Hope of Edinburgh, announced in Nicholson's Journal, 8vo. iv. 210. It consists of a small bottle, of the capacity of 20 or 24 drachms (fig. 28, pl. ii.), destined to contain the eudiometric liquid, and having a small stopper at *b*. Into the neck of the bottle a tube is accurately fitted, by grinding, which holds precisely a cubic inch, and is divided into 100 equal parts. To use the apparatus, the bottle is first filled with the liquid employed, which is best prepared by boiling a mixture of quicklime and sulphur with water, filtering the solution, and agitating it for some time in a bottle half filled with common air. The tube, filled with the gas under examination (or with atmospherical air, when the quality of this compound is to be ascertained), is next to be put into its place; and, on inverting the instrument, the gas ascends into the bottle, where it is brought extensively into contact with the liquid, by brisk agitation. An absorption ensues;

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[*] Other eudiometers will be described hereafter.

and, to supply its place, the stopper *b* is opened under water, a quantity of which rushes into the bottle. The stopper is replaced under water; the agitation renewed; and these operations are performed alternately, till no farther diminution takes place. The tube *a* is then withdrawn, the neck of the bottle being under water, and is held inverted in water for a few minutes; at the close of which the diminution will be apparent. Its amount may be measured by the graduated scale engraved on the tube.

To the eudiometer of Dr. Hope there are, however, a few objections. If the tube *a* and the stopper *b* are not both very accurately ground, air is apt to make its way into the instrument, to supply the partial vacuum, occasioned by the absorption of oxygen gas. This absorption, also, occasions a diminished pressure within the bottle; and, consequently, towards the close of each agitation, the absorption goes on very slowly. Besides, the eudiometric liquid is constantly becoming more dilute by the admission of water through *b*. To obviate all these difficulties, I have substituted, for the glass bottle, one of elastic gum, as shown by fig. 21, *b*. The tube *a* is accurately ground into a short piece of very strong tube of wider bore, as shown at *c*, the outer surface of which is made rough by grinding, and shaped as represented, that it may more effectually retain the neck of the elastic bottle when fixed by string. This instrument is used, in every respect, in the same way as Dr. Hope's. The only difficulty is, in returning the whole of the residuary gas into the tube; but the art of doing this will be acquired by practice.

An ingenious modification of the eudiometer, which enables us to measure an absorption of only $\frac{1}{1000}$th part of the gas employed, is described by Mr. Pepys, in the Philosophical Transactions for 1807, or Philosophical Magazine, vol. xxix.

IV. *Atmospheric air ministers to the support of animal life, only in consequence of the oxygen gas which it contains.*—Air, after having been received into the lungs, and again expired, is found to have lost considerably of its oxygenous part, *viz.* 10 to 12 *per cent.* It proves fatal to animals, however, long before this purer portion is wholly exhausted; and hence it appears, that a considerable portion of oxygen gas is even necessary to fit the air for supporting respiration. As the analysis of expired air requires an acquaintance with another gas, not hitherto described, *viz.* carbonic acid, its examination will be postponed to a future occasion.

V. *Atmospheric air is diminished in volume by animal respiration.*—This may be shown by repeating a very simple experiment, originally contrived by Mayow. He confined a mouse in a small glass jar, and tied the jar over, quickly and firmly, with moistened bladder. The heat of the animal first expanded the air, and rendered the bladder convex outwards; but when the animal after death had become cold, the bladder exhibited a hollow surface, proving that the air within was diminished in its bulk.

The exact amount of the diminution may be shown, by confining a mouse, over water, in a graduated jar, furnished with a stop-cock, and containing common air. As the heat of the animal, however, would occasion the expulsion of part of the air, it is expedient, on first depressing the jar into water, to open the cock, through which a part of the air will escape: the cock is then to be shut, and the height of the water within to be accurately noted. At first, the level will be depressed, in consequence of the expansion of the air by the warmth of the animal; but, after its death, a considerable diminution will be observed.

VI. The weight of a given volume of atmospheric air, at 60° Fahrenheit and 30 inches barometer, is said by Mr. Kirwan to be 30.92 grains. Mr. Davy states it, when under the same pressure, but at 55° Fahrenheit, to be 31.10 grains, from which may be deduced that, with the temperature and pressure assumed by Mr. Kirwan, 100 inches would weigh 30.78 grains.

SECTION V.

Hydrogen Gas.

THE most simple form, in which HYDROGEN has hitherto been obtained, is in that of a gas, or of union with caloric, and perhaps with electricity and light. From this combination we are not able to separate it, except by availing ourselves of the affinity of some other substance, in which case the hydrogen separates from the caloric, and forms, with the body, which has been added, a new combination. Of its nature, we know but little; but as it has not yet been resolved into any more simple state, it is still arranged among elementary bodies. From the recent experiments of Mr. Davy (which will be described under the article ammonia), it appears not improbable that hydrogen is a metallic body.

The most important compound of hydrogen, and the only one which will be noticed at present, is that which it affords by union with the base of oxygen gas. It is on its affinity for this base that all the properties depend, which are illustrated by the following experiments. Much of the force of this attraction, it will appear probable from the sequel, depends on its being in a state of opposite electricity to oxygen; for in common with all inflammable substances it is naturally in a state of *positive* electricity.

I. *To procure hydrogen gas*, let sulphuric acid, previously diluted with five or six times its weight of water, be poured on iron filings; or on small iron nails; or (what is still better) on zinc*, granulated by pouring it melted into cold water, and contained in a gas bottle or small retort. An effervescence will ensue, and the escaping gas may be collected in the usual manner.

II. This gas has the following properties:

1. *It remains permanent over water*, or is not absorbed in any notable proportion.

2. *It is inflammable.* This may be shown by the following experiments:

(*a*) Fill a small jar with the gas, and, holding it with the mouth downwards, bring the gas into contact with the flame of a candle. The air will take fire, and will burn silently with a lambent flame.

(*b*) Fill with this gas a bladder which is furnished with a stop-cock, and with a small pipe, of diameter less than that of a common tobacco pipe. Press the air out through the pipe, and, on presenting a lighted candle, the stream will take fire. If this apparatus cannot be procured, a very simple contrivance will answer the purpose: break off part of an eight-ounce vial, within an inch or two from the bottom, by setting fire to a string tied round it, and moistened with spirit of turpentine. The vial will then resemble a jar with an open neck at the top. Next bore a small hole, through a cork that fits the neck of the vial, and insert in it part of a common tobacco pipe, which may be fixed into the neck of the bottle, by a cement of resin and bees-wax. Then fill the bottle with water, and hold it, with the thumb pressed down on the aperture of the pipe, while hydrogen gas is passed into it. When the bottle is full of gas, remove the thumb, press the bottle down into the water, and, on the approach of a candle, the stream of air from the pipe will take fire.

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* Zinc may be purchased at the brass-founders or coppersmiths, under the name of speltre.

Persons, who are provided with the jars represented pl. ii. fig. 22, *a*, may screw to the cock a brass pipe with a small aperture. On pressing the jar, filled with hydrogen gas, into the water, and opening the cock, the gas will be forced out in a stream, which may be set on fire. On this principle are founded the artificial fireworks without smell or smoke. They consist of pipes, having variously sized apertures, some of which have a rotatory motion; but their precise construction it is impossible to describe, without very tedious details.

(*c*) In a strong bottle, capable of holding about four ounces of water, mix equal parts of common air and hydrogen gas. On applying a lighted candle, the mixture will burn, not silently, as in experiment (*a*), but with a sudden and loud explosion. If a larger bottle be used, it should be wrapped round with a handkerchief, to prevent the glass from doing any injury, in case the bottle should be burst*.

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* These experiments may also be made advantageously, by means of an apparatus sold under the name of the inflammable air-pistol.

(*d*) The same experiment may be repeated with oxygen gas, instead of atmospherical air; changing the proportions, and mixing only one part of oxygen gas with two of hydrogen. The report will be considerably louder. The bottle should be a very strong one, and should be wrapped round with cloth, to prevent an accident.

(*e*) The same experiment may be made over water, by means of the electric spark. Procure a strong tube, about three quarters of an inch diameter, and 12 inches long, closed at one end (plate ii. fig. 28). About a quarter or half an inch from the sealed end, let two small holes be drilled, opposite to each other, and into each of these let a brass conductor be cemented, so that the two points may be distant from each other, within the tube, about one 8th of an inch. An apparatus, serving the same purpose, and much more easily constructed, may be formed by hermetically sealing a piece of brass wire, or still better platina wire, into the end of a glass tube (fig. 29). With this conductor, an interrupted circuit may be formed, by introducing into the tube a longer wire, one end of which terminates one 10th of an inch from the upper one, while the other extends beyond the aperture of the tube. (See fig. 64.) Into this tube, standing over water, pass about half a cubic inch of a mixture of hydrogen and oxygen gases; in the proportion of two measures of the former to one of the latter. Hold the tube firmly, and pass an electric spark through the mixed gases. For relieving the shock, which is sometimes considerable on firing, an ingenious contrivance of Mr. Davy may be employed. It is described in the Philosophical Magazine, xxxi. 3. An immediate explosion will take place; after which the gases, if pure, and in the proper proportion, will be found to have disappeared entirely.

It is rarely, however, that the gases can be used in such a state of purity as to leave absolutely no residuum. To determine, indeed, the purity either of the oxygen or hydrogen gas employed, it is sometimes necessary so to adjust their proportions, that the whole mixture may not be condensed by firing. If, for example, we wish *to know the purity of a quantity of oxygen gas*, we are to use about three times its bulk of hydrogen. Let us suppose that 100 measures of oxygen are detonated with 300 of hydrogen gas, and that the total 400 is reduced by firing to 130; the diminution of volume will be 270. This number, divided by three, gives 90 for the quantity of oxygen; that is, the oxygen employed must have contained 10 *per cent.* of nitrogen or of some foreign gas not condensible by hydrogen.

If atmospherical air be employed, a diminution, though not equal in amount, will be produced by the union of the hydrogen with the oxygen gas contained in the air; and if a sufficient quantity of hydrogen gas be employed, the whole of the atmospheric oxygen will thus be removed. On this principle is founded the EUDIOMETER OF VOLTA, which may be constructed, by graduating either of the tubes already described, into equal parts. If, in one of these tubes, we mix 300 parts of common air, and 200 of pure hydrogen gas, there will remain, after the explosion excited by passing an electric spark between the two wires, about 305 measures. There will, therefore, have been a diminution of 195 measures, of which pretty exactly one 3d may be estimated to be pure oxygen. In this instance, therefore, 65 of oxygen have been lost by 300 of air, or 21 and a fraction *per cent.*

The general rule for ascertaining the purity of air by hydrogen gas, may be stated as follows: Add to three measures of the air under examination, two measures of pure hydrogen gas; inflame the mixture by electricity; observe the diminution when the vessel has cooled; and, dividing its amount by three, we obtain pretty nearly the quantity of oxygen gas which has been condensed.

In the reverse process, *i. e. in determining the purity of hydrogen gas*, we mix it with more oxygen gas than is required for saturation. Suppose that to 100 of hydrogen gas we add 100 of oxygen, and that 80 measures remain after detonation. The diminution will have been 120 measures; and, of these, two 3ds or 80 measures are hydrogen. Hence the inflammable gas, under examination, must contain 20 *per cent.* of some other gas, which is most probably nitrogen. In this way, we determine the proportions of hydrogen and nitrogen in any mixture composed of the two gases only.

(*f*) The diminution of hydrogen and oxygen gases, by the union of their bases, may be shown also by their slow combustion. Fill a tall jar with oxygen gas, and fill also, with hydrogen gas, a bladder furnished with a stop-cock, and with a long brass pipe, bent like the letter S, and drawn out to a fine point (plate iv. fig. 41). On pressing the bladder, a stream of gas will issue from the pipe, which may be set fire to, and brought cautiously under the tall inverted jar of oxygen gas. By this contrivance, the stream of hydrogen gas will be burnt in a confined portion of oxygen gas; and, on continuing the combustion a sufficient length of time, the water will be seen to rise gradually within the jar. On the first impression of the heat, indeed, a quantity of gas will escape from the jar, which will render it difficult to ascertain what degree of absorption has actually taken place. But this loss may be prevented, by using a jar with a neck at the top, to which a compressed bladder is firmly tied. The expanded air, instead of escaping through the water, will now fill the bladder at the top; and, when the experiment has closed, and the vessels have cooled, it may be ascertained, by pressing out the gas from the bladder, what quantity of oxygen gas has been consumed.

The same experiment may be more accurately and elegantly made, with the assistance of an apparatus, which I have described in the Philosophical Transactions for 1808. The description cannot be understood without the plate, which is there given, and which has been copied into the Philosophical Magazine, vol. xxxii. and Nicholson's Journal, vol. xxi. The fact may, also, be shown by substituting, for the bladder (*e*, fig. 41) a small gazometer, containing a measured quantity of hydrogen gas. Let the bent pipe be screwed on the cock of the gazometer; and over its open end, placed perpendicularly, invert a jar of oxygen gas. This jar must be provided at the top with a metallic conductor, screwed into a brass cap, as represented in fig. 41; which shows also the level of the water within the jar, attained by means of a syphon. After noting the height of the water within, let a rapid succession of electric sparks be passed between the two conductors; and, on opening the cock at this instant, the stream of oxygen gas will be inflamed. The end of the pipe must then be so far depressed, that the cement of the brass cap may not be melted by the flame; and the outer surface of the top of the vessel should be kept cool. When the gas is first lighted, the oxygen gas will be suddenly expanded; but, presently, a rapid diminution will go on, till the water rises above the end of the pipe and extinguishes the flame. If pure oxygen gas be employed, it will be found, after the experiment, uninjured in its quality, and will support the combustion of burning bodies as well as before.

When the above experiment is made, with the substitution of common air for oxygen gas, a diminution takes place, but much less considerable, *viz.* not amounting to one 6th of the original bulk of the gas.

(*g*) When a stream of hydrogen gas is burned under a tube, 18 or 24 inches long, a musical sound is produced. The experiment may be made in the following manner:

Into a glass bottle are put iron filings and sulphuric acid, diluted with five or six parts of water; and a cork is fitted into the neck, through which a glass tube is passed, having its upper extremity drawn out to a capillary bore. By setting fire to the hydrogen gas*, which escapes from this extremity, a continued current or jet of flame is produced, which is allowed to pass into a tube either of glass, earthenware, or metal. If the tube be not too large, the flame becomes smaller as it is depressed; and when the tube covers the flame to a considerable depth, very clear sounds are produced. But, on the contrary, if the tube be too narrow, the flame will be extinguished; and, in proportion as the tube is enlarged, the sound diminishes: so that there is a certain limit at which it totally ceases. The same happens when the tube is too long. The sounds may be raised at pleasure, by either using tubes of various figures or dimensions, or made of different substances. (See Nicholson's Journal, 8vo. i. 129, and iv. 23.)

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* The gas must not be inflamed, till it has been produced for some time, and has expelled all the common air of the bottle; otherwise an explosion will happen, and the bottle will be burst, with some danger to the operator.

(*h*) In a memoir lately read to the National Institute of France, M. Biot announces the important fact, that a mixture of hydrogen and oxygen gases may be made to explode by mechanical compression. A mixture of these two gases was introduced into a strong metallic syringe, furnished with a glass bottom, and a sudden stroke given to the piston. An extremely brilliant light appeared, accompanied with a loud detonation; and the glass bottom was forcibly driven out. The repetition of this experiment, it is obvious, must be attended with considerable difficulty and danger. (See Nicholson's Journal, xii. 212.)

The combustion of hydrogen and oxygen gases is successfully applied to the purpose of exciting an intense heat by the blow-pipe. The peculiar construction of the apparatus cannot be understood without a plate, which may be seen in the *Annales de Chimie*, tom. xlv. or in the 14th volume of the Philosophical Magazine. It may be sufficient here to state, that the gases are contained each in a separate gas-holder; that they are expelled by the pressure of a column of water obtained by lengthening the pipe *b*, fig. 36; and that their mixture does not take place, till they nearly reach the aperture of the pipe, at the extremity of which they are inflamed. This last precaution is of considerable importance, because a violent and dangerous explosion would otherwise happen. To guard the more effectually against this accident, it is advisable to affix a valve, opening outwards, in the pipe proceeding from each gas-holder, just before the junction of the two.

3. *Hydrogen gas has an unpleasant smell.*

4. *Though inflammable itself, it extinguishes burning bodies.*—Bring an inverted jar, filled with this gas, over the

flame of a candle; and suddenly depress the jar, so that the lighted wick may be wholly surrounded by the gas. The candle will immediately be extinguished.

5. *It is fatal to animals.*—This may be shown by confining, in the gas, a mouse, or other small animal.

6. *It is considerably lighter than atmospherical air.*—One hundred cubic inches, the barometer being 30 inches, and the thermometer 60°, weigh, according to Kirwan, 2.613 grains; according to Lavoisier, 2.372 grains; and according to Fourcroy, Vauquelin, and Seguin, 2.75 grains. Messrs. Biot and Arajo assign to it the specific gravity of 0.07321.

(*a*) Let a jar, filled with this gas, stand, for a few seconds, with its open mouth upwards. On letting down a candle, the gas will be found to have escaped.

(*b*) Place another jar of the gas inverted, or with its mouth downwards. The gas will now be found to remain in the jar, being prevented from escaping upwards by the bottom and sides of the vessel.

(*c*) Fill, with hydrogen gas, a bladder furnished with a stop-cock; and adapt to this a common tobacco pipe. Dip the bowl of the pipe into a lather of soap, and, turning the cock, blow up the lather into bubbles. These bubbles, instead of falling to the ground, like those commonly blown by children, will rise rapidly into the air. On this property of hydrogen gas, is founded its application to the raising of balloons.

(*d*) The experiment may be varied by filling the bladder with a mixture of two parts of hydrogen gas and one of oxygen gas. Bubbles, blown with this mixture, take fire on the approach of a lighted candle, and detonate with a loud report. It is proper, however, not to set them on fire, till they are completely detached from the bowl of the pipe; otherwise the contents of the bladder will be exploded, with considerable danger to the operator.

In this place a property of hydrogen gas may be described, which it possesses in common with all other aëriform bodies, *viz.* a tendency to diffusion through any other elastic fluid, with which it may be brought into contact. Common or inelastic fluids are capable of remaining in contact with each other for a long time without admixture. Thus if we half fill a wine glass with spirit of wine tinged with any colouring ingredient, and then, by means of the dropping tube, fig. 15, introduce under it a quantity of water, the spirit floats on the water, and the two surfaces remain perfectly distinct, provided we carefully avoid agitation or unequal changes of temperature. But this is not the case with elastic fluids or gases, which, it has been discovered by Mr. Dalton*, penetrate each other, and become thoroughly mixed under all circumstances. The fact, with respect to hydrogen and oxygen gases, may be proved by a very simple apparatus.

Provide two glass vials, each of the capacity of about an ounce measure, and also a tube open at both ends, 10 inches long and one 20th inch bore. At each end, the tube is to be passed through a perforated cork, adapted to the necks of the vials. Fill one of the bottles with hydrogen gas, and the other with oxygen gas; place the latter on a table with its mouth upwards, and into this insert the tube secured by its cork. Then, holding the hydrogen bottle with its mouth downwards, fit it upon the cork at the top of the tube. The two bottles, thus connected, are to be suffered to remain in this perpendicular position. After standing two or three hours, separate the vials, and apply a lighted taper to their mouths, when it will probably occasion an explosion in each. The hydrogen gas, though nearly 14 times lighter than the oxygen, must, therefore, have descended through the tube from the upper into the lower vial; and the oxygen gas, contrary to what might have been expected from its greater weight, must have ascended through the tube, and displaced the lighter hydrogen.

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* Manchester Memoirs, vol. i. new series.

Experiments of this kind, it has been shown by Mr. Dalton, may be extended to all the other gases; but to prove the effect, tests of a different kind are necessary, which require a previous knowledge of the properties of these gases. They tend to establish the conclusion, *that a lighter elastic fluid cannot remain upon a heavier without admixture.*

CHAPTER VI.

OF THE COMPOSITION, DECOMPOSITION, AND PROPERTIES OF WATER.

SECTION I.

Synthesis, or Composition, of Water.

IN chap. v. sect. v. it was stated, that oxygen and hydrogen gases, when fired over water, in the proper proportion, wholly disappear. To ascertain the nature of the product thus formed, however, the experiment must be repeated over mercury, in a similar manner, by means of the detonating tube (pl. ii. fig. 28). When this is done repeatedly, it is found that the product of the combustion is that well known fluid, water, which is thus proved to be composed of two elementary ingredients. The water, produced in this mode, is not, however, to be considered as a compound of the two gases, but only of their bases; for the light and caloric, which constituted the gases, escape, in considerable part, during the combustion. Every gas, it must be remembered, has at least two ingredients; the one, gravitating matter, which, if separate, would probably exist in a solid or liquid form; the other, an extremely subtile fluid, termed caloric. In the example before us, caloric (and perhaps electricity and light) is a common ingredient both of hydrogen and oxygen gases; but the two differ in having different bases. The basis of the one is called hydrogen, of the other oxygen; and water may, therefore, be affirmed to be a compound, not of hydrogen and oxygen *gases*, but of hydrogen and oxygen. This may be proved in two modes, by synthesis, *i. e.* by joining together its two elementary ingredients; and by analysis, in other words, by separating the constituent parts of water, and again exhibiting them in a distinct form.

I. Fill, with hydrogen gas, a bladder, furnished with a stop-cock and bent pipe (fig. 41, *c*), as in the last chapter. Then pour into a shallow earthen dish as much quicksilver as will about half fill it, and invert over this a glass bell, full of common air and perfectly dry. Expel the hydrogen gas through the pipe; light the stream, and bring it under the glass bell, by raising this, and depressing it into the mercury, as soon as the inflamed gas is introduced. A portion of air will escape, at first, in consequence of the rarefaction. As the combustion continues, water will form, and will condense on the sides of the glass. This water is produced by the union of hydrogen with the oxygen contained in atmospheric air.

II. Those persons who are not possessed of a sufficient quantity of quicksilver to repeat the above experiment, may substitute the following: procure a large glass globe, capable of holding three or four quarts, and having two openings, opposite to each other, which may be drawn out for a short distance, like the neck of a retort. Inflame the stream of hydrogen gas, and introduce it into the centre of the globe. The rarefied and vitiated air will ascend through the aperture of the globe, and a constant supply of fresh air will be furnished from beneath. By this combustion, a quantity of water will be generated, which will be condensed on the inner surface of the vessel.

III. A simple and ingenious apparatus, less costly than any other, intended for the purpose of exhibiting the composition of water, is made by Mr. Cuthbertson of London. It may be seen described and figured in Nicholson's Journal, 4to. vol. ii. page 235; or in the Philosophical Magazine, vol. ii. page 317[*]; and also in pl. iv. of this work, fig. 33.

In using this apparatus, however, instead of two glass receivers for the oxygen and hydrogen gases, standing inverted in a trough of water, I employ a couple of gazometers; and with this alteration, the experiment is more easily managed, as well as more striking. The apparatus, thus modified, consists of a large glass receiver or bottle *a* (pl. iv. fig. 34), with an opening at the bottom, into which is cemented a piece of brass, perforated with two holes. This brass piece is represented of a larger size in fig. 33; the aperture *a* conveying the oxygen gas, and *b* the hydrogen. Before commencing the experiment, the cock *c*, fig. 34, is screwed, by means of a collar-joint[†], to the cock *b* of the gazometer, fig. 35, containing oxygen gas; and to the cock *d*, by the same means, is affixed another gazometer, filled with hydrogen gas.

When it is intended to ascertain, accurately, the proportions of gases consumed and of water generated, the receiver *a*, previously weighed, is first exhausted by an air-pump, with which it may be connected by the female screw at *c*. The quantity of common air left in the receiver may be determined, by enclosing a gage within it. If the additional expense be not deemed an objection, it is advisable, that after exhausting the receiver, oxygen gas should be admitted; its contents be exhausted a second time; and again renewed by fresh oxygen from the gazometer, the quantity of which may be observed by the graduated scale. The receiver being thus filled with oxygen gas, and accurately closed by a cock at *c*, a succession of sparks is to be passed, from the prime conductor of an electrical machine, between the platina knob of the bent wire within the receiver, and the point of the brass cone. While the sparks are transmitted, the cock *d* is to be opened. A stream of hydrogen gas will immediately issue from the aperture at the point of the cone, and will be inflamed by the electric spark, as represented fig. 33. The cock *e* is now to be opened, and the size of the flame of hydrogen gas moderated by partly shutting the cock *d*. As the volume of hydrogen gas consumed is double that of the oxygen; and the pipe, which transmits it, is of less diameter than that conveying the latter, about twice the pressure is required to expel the hydrogen. This is given, by lessening, in that proportion, the weight of the counterpoises (*ee*, fig. 35) of the gazometer containing hydrogen.

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[*] In the same volume of the Philosophical Magazine an interesting account may be consulted of the principal experiments on the composition of water, accompanied by neat and perspicuous engravings of the apparatus employed in them.

[†] See pl. v. fig. 47; and the corresponding description of the structure of this joint, in the explanation of the plates at the end of the work.

During the combustion, the moveable vessel *c*, fig. 35, of each gazometer descends; and, by observing the graduated scales, it will be seen that the hydrogen vessel falls twice as quick as that which holds the oxygen gas. It is necessary to keep the receiver *a* cool by means of wet cloths; and, when this is done, the water, which is produced, will form into drops on the inside of the receiver, and collect at the bottom. At the conclusion of the experiment, the receiver is to be again weighed, and the increase noted. The quantity of gases consumed is to be observed, and their actual weight computed, by means of the table given in the Appendix. It will be found, that the weight of water produced is very nearly equal to that of the two gases expended; that is to say, excluding decimals, for every hundred grains of water generated in the receiver, eighty-five grains of oxygen gas, and fifteen grains of hydrogen gas (equal by measure to about 250 cubic inches of the former, and 500 of the latter), will have disappeared from the gazometers.

IV. By firing repeated portions of a mixture of oxygen and hydrogen gases over mercury, a sensible quantity of water will at last be produced.

SECTION II.

Analysis, or Decomposition, of Water.

The analytic experiments on water are of two kinds: 1st, Such as present us with one of its ingredients only, in a separate and distinct form; 2dly, Such as present us with its two component principles, the hydrogen and oxygen, mixed together in the state of gas.

I. Of the first kind are the following:

1. Procure a gun-barrel, the breech of which has been removed, so as to form a tube open at each end. Fill this with iron wire, coiled up in a spiral form. To one end of the barrel adapt a small glass retort, partly filled with water, and to the other a bent glass tube, the open end of which terminates under the shelf of the pneumatic cistern. Let the barrel be placed horizontally (or rather with that end, to which the retort is fixed, a little elevated) in a furnace, which has two openings in its body opposite to

each other. (Pl. iv. fig. 40.) Light a fire in the furnace; and, when the gun-barrel has become red-hot, apply a lamp under the retort. The steam of the water will pass over the red-hot iron, and will be decomposed. Its oxygen will unite with the iron; and its hydrogen will be obtained in the form of a gas. This is the readiest and cheapest mode of procuring hydrogen gas, when wanted in considerable quantity.

2. The same experiment may be repeated; substituting an earthen tube for a gun-barrel, and weighing the iron wire accurately, both before and after the experiment. The iron will be found to have gained weight very considerably; and, if attention be paid to the weight of the water that escapes decomposition, by an addition to the apparatus (fig. 40, *c*), and to the weight of the gases obtained, it will be found, that the weight gained by the iron, added to that of the hydrogen gas, will make up exactly the weight of the water that has disappeared. From experiments of this kind, conducted with the utmost attention to accuracy, as well as from synthetic experiments, it appears, that water is compounded of 85 *per cent.* oxygen, and 15 hydrogen, by weight, very nearly. But as hydrogen gas is eleven times lighter than common air, the proportion of gases, by volume, required to form water, is about two of hydrogen to one of oxygen gas. By the decomposition of every hundred grains of water, therefore, the iron employed gains 85 grains, and becomes oxidized; and 15 grains (equal to about 500 cubical inches) of hydrogen gas are obtained.

3. Water may be decomposed, in a similar apparatus, over charcoal instead of iron. The results, however, are different in this case, as will appear from a subsequent section.

4. Another mode of effecting the decomposition of water yet remains to be mentioned, in which not the hydrogen, but the oxygen, is obtained in a gaseous state. This is by the action of living vegetables; either entire, or by means of their leaves only. Fill a clear glass globe with water, and put into it a number of green leaves, from almost any tree or plant. A sprig or two of mint will answer the purpose perfectly well. Invert the glass, or place it, with its mouth downwards, in a vessel of water. Expose the whole apparatus to the direct light of the sun, which will then fall on the leaves surrounded by water. Bubbles of air will soon begin to form on the leaves, and will increase in size, till at last they rise to the top of the vessel. This process may be carried on as long as the vegetable continues healthy; and the gas, when examined, will prove to be oxygen gas, nearly pure. In this experiment, the hydrogen combines with the plant, to the nourishment and support of which it contributes, while the oxygen is set at liberty.

II. The processes, by which the elementary parts of water are separated from each other, and are both obtained in an aëriform state, as a mixture of hydrogen and oxygen gases, are dependent on the agency of electricity.

1. The first of these experiments requires for its performance the aid of a powerful electrical machine. This fact was the discovery of a society of Dutch chemists; and the principal circumstance, in the experiment, is the transmission of electrical shocks, through a confined portion of water. The apparatus employed, in this experiment of Messrs. Dieman and Van Troostwyk, is a glass tube, about one 8th of an inch diameter, and 12 inches long, one of the ends of which is sealed hermetically, a gold wire being inserted at this end, and projecting

about an inch and a half within the tube. About the distance of five 8ths of an inch from the extremity of this, another wire is to be fixed, which may extend to the open end of the tube. The tube is next to be filled with distilled water, and to be placed inverted in a vessel of the same. When thus disposed, electrical shocks are to be passed between the two ends of the wire, through the water; and, if these shocks be sufficiently strong, bubbles of air will be formed at each explosion, and will ascend till the upper part of the wire is uncovered by the water. As soon as this is effected, the next shock that is passed will set fire to the air, and the water will rise again in the tube, a very small quantity of gas remaining. Now, as hydrogen and oxygen gases, in a state of admixture, are the only ones that are capable of being inflamed by the electric shock; and as there is nothing in the tube, besides water, that can afford them in this experiment, we may safely infer, that the evolved hydrogen and oxygen gases arise from decomposed water.

2. An improved apparatus, exhibiting the same experiment, with less trouble to the operator, has been invented by Mr. Cuthbertson, and may be seen described and figured in Dr. Pearson's paper in the Philosophical Transactions for 1797, or in Nicholson's Journal, vols. i. and ii. 4to.

3. The decomposition of water by galvanic electricity is a process singularly adapted to demonstrate the fact in a simple and elegant manner. The manner of conducting it, as well as the results, will be fully explained, when we come to treat of the general principles of electro-chemical science.

SECTION III.

*Properties and Effects of Water**.

I. *Water contains air.*—This may be shown by placing a glass vessel of water under the receiver of an air-pump. During the exhaustion of the receiver, bubbles of air will be seen to ascend very plentifully. Much air escapes also from water, during ebullition, and may be collected by a proper apparatus. The same fact may also be exhibited, by filling a barometer tube, about 32 inches long, sealed at one end, with quicksilver, except about four inches, and the remainder with water. On inverting the open end of the tube in quicksilver, bubbles of air will be seen, in a short time, to rise from the water.

The kind of gas, extricated from the water of a spring at a considerable distance from the surface, I have made the object of experiment. (Philosophical Transactions, 1803.) From 100 cubic inches of the water, or about $3\frac{1}{2}$ wine pints, 4.76 cubic inches of gas were separated, of which 3.38 were carbonic acid gas, and 1.38 air of the same standard as that of the atmosphere.

Every gas is absorbed by water, deprived of all or the greatest part of its air by long boiling. The quantity, however, which water is capable of absorbing, varies considerably with respect to the different gases. Those, of which only a small proportion is absorbed, require violent and long continued agitation in contact with water. The following table has been drawn up by Mr. Dalton from the combined results of his own experiments and mine.

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* Whenever, in the course of this work, water is mentioned as an agent in any chemical operation, pure distilled water is to be understood.

Water absorbs, at the mean pressure and temperature of the atmosphere,

Of carbonic acid gas	its own bulk.
sulphuretted hydrogen	do.
nitrous oxide	do.
olefiant gas	$\frac{1}{8}$.
oxygen gas	$\frac{1}{27}$.
nitrous gas	do.
carburetted hydrogen	do.
carbonic oxide	$\frac{1}{64}$.
azotic gas	do.
hydrogen gas	do.

The principle, on which gases are retained by water, is still a matter of controversy. By Berthollet and the generality of chemists, it is ascribed to the exertion of a chemical affinity between the gas and the water; but it has been suggested by Mr. Dalton (and as appears to me with greater probability) that the fact may be better explained on mechanical principles. A statement of the argument may be seen in Mr. Dalton's "New System of Chemical Philosophy," or in two papers, which I have published in the eighth and ninth volumes of Nicholson's Journal.

II. *Water is contained in the air of the atmosphere, even during the driest weather.*—Expose to the air, in a shallow vessel, a little sub-carbonate of potash or common salt of tartar. In a few days it will have become moist, or *deliquidated.* On the same principle, water exposed to the air, in a shallow vessel, disappears, being dissolved by the atmosphere. The quantity of water, that may be extracted from 100 cubical inches of air, at 57° Fahrenheit is 0.35 of a grain; but, according to Clement and Desormes, at 54° Fahrenheit, only 0.236 of a grain can be detached by exposure to muriate of lime. The experiments, both of these chemists and of Mr. Dalton, concur in proving that at the same temperature, equal bulks of different gases give up the same quantity of water to deliquescent salts. The portion of humidity, which they thus abandon, has been called *hygrometric water*. Whether they contain a still farther quantity in a state of more intimate union and not separable by deliquescent substances, is still undetermined.

III. *Water dissolves a great variety of solid bodies.*—The substances, on which it exerts this effect, are said to be soluble in water; and there are various degrees of solubility. See chap. i, and the table in the Appendix.

IV. *During the solution of bodies in water, a change of temperature ensues.*—In most instances, an absorption of caloric (in other words, a production of cold) is attendant on solution, as in the examples given in chap. iii. sect. 2. But, in other cases, caloric is evolved, or heat is produced. Thus, common salt of tartar, during solution in water, raises the temperature of its solvent; and caustic potash, in a state of dryness, does the same still more remarkably. Both carbonated and pure potash, however, when crystallized, observe the usual law, and absorb caloric during solution. Now as their difference, in the crystallized and uncrystallized state, depends chiefly on their containing in the former, but not in the latter, water chemically combined, we may infer, that the cold, produced during the solution of salts, is occasioned by the conversion of the water, which exists in these bodies, from a solid to a liquid form.

V. *During the solution of salts in water, a quantity of air is disengaged.*—This air was partly contained mechanically in the salt, and partly in the water. That it does not arise entirely from the former source, is proved by varying the experiment in the following manner. Let an ounce or two of sulphate of soda be put into a vial, and pour on this as much water as will completely fill the bottle. The air contained in the pores of the salt will be thus disengaged; but only a small portion of the salt will be dissolved, agreeably to the principle laid down, chap. ii. 7. Let the vial be shaken, and the whole of the salt will disappear; a fresh portion of air being liberated during solution. The air, that now appears, is extricated from the water, in consequence of the affinity between the water and the salt being stronger than that between the water and the air. It is, therefore, a case of single elective affinity.

VI. *During the solution of bodies, the bulk of water changes.*—Take a glass globe, furnished with a long narrow neck (commonly termed a mattras, see fig. 4), and put into it an ounce or two of sulphate of soda. Then, add as much water as will fill the globe, and about three 4ths of the neck. This should be done with as little agitation as possible, in order that the salt may not dissolve, till required. Mark, by tying a little thread, or by a scratch with a file, the line where the water stands; and then agitate the mattras. The salt will dissolve; air will be set at liberty; and, during the solution, the water will sink considerably below its level. The contraction of bulk is owing to the diminution of temperature; and, when the water has regained its former temperature, it will also be found, that its bulk is increased by the addition of salt. The Bishop of Llandaff observed, that water exhibits a manifest augmentation of bulk, by dissolving only the two thousandth part of its weight of salt; a fact sufficiently decisive against that theory, which supposes pores in water capable of receiving saline bodies without an augmentation of volume.

VII. *Water has its solvent power increased, by diminishing the pressure of the atmosphere.*—Into a Florence flask, put half a pound of sulphate of soda; pour on it barely a pint of water, and apply heat so as to boil the water. The whole of the salt will be dissolved. Boil the solution for several minutes pretty strongly, so as to drive out the air; and cork the bottle tightly, immediately on its removal from the fire. To prevent more completely the admission of air, tie the cork over with bladder. As the vessel cools, an imperfect vacuum will be formed over the solution; for the steam, which arises during the ebullition, expels the air, and takes its place. The steam is condensed again, when the vessel cools. The solution, when perfectly cold, may be shaken without any effect ensuing, so long as the vessel is kept closely stopped; but, on removing the cork and shaking the vessel, the solution will immediately congeal, and heat will be produced. This experiment, besides the principle which it is peculiarly intended to illustrate, exemplifies also the general rule laid down, chap. iii. sect. 2. vi. *viz.* that caloric is always evolved, during the transition of bodies from a fluid to a solid state; and it furnishes a fact exactly the reverse of that in which cold is produced, or caloric absorbed, during the solution of salts.

VIII. It is unnecessary to add any thing to what has been already said in a former section, respecting the combination of caloric with water constituting steam; or to the history of the phenomena attending its conversion into ice; except that, during the latter change, its bulk is enlarged in the proportion of nine to eight, and that, in consequence of this expansion, water, during congelation, is capable of bursting the strongest iron vessels; and becomes specifically lighter. Hence, ice swims always on the surface of the water.

It is remarkable, that this enlargement of the bulk of water begins long before its temperature has descended to the freezing point, *viz.* at about 40° Fahrenheit. Let a thermometer bulb, and part of its tube, having a wide bore, be filled with water, tinged with a little litmus, which may be introduced by the same means as those already directed for filling with quicksilver. Immerse the thermometer in water of the temperature of 40°; and, when the included water may be supposed to have attained the same degree of heat, remove the instrument successively into water of the temperature of 36° and 32°. At each immersion, the water will rise in the tube. Bring its temperature again to 40°, and it will descend to the same point as before. Place it in water of 50°, and it will again be expanded. Precisely similar effects, therefore, appear to result, in these experiments, from two opposite causes; for the bulk of water is alike increased by reducing or raising its temperature. It is contended, however, by Mr. Dalton, that, in the apparent expansion by a lower temperature, there is a deception, arising from the contraction of the glass, which must lessen the capacity of the bulb, and force the water up the stem. The question is not yet decided; and is still contested by Mr. Dalton against the experiments of Count Rumford and of Dr. Hope. The former philosopher now contends, that water is of the greatest density at 36° of Fahrenheit, or 4° above its freezing point.

CHAPTER VII.

ON THE CHEMICAL AGENCIES OF COMMON AND GALVANIC ELECTRICITY.

THAT branch of natural science, which comprehends the phenomena of Galvanism, and the general principles under which they are arranged, is only of very recent origin. It was not till the year 1791, that Galvani, an Italian philosopher, being engaged in a course of experiments on animal irritability, observed accidentally the contractions, which are excited in the limbs of frogs, by applying a conductor of electricity between a nerve and a muscle. The theory, which he framed to account for this phenomenon, was that the different parts of an animal are in opposite states of electricity, and that the effect of the metal is merely to restore the equilibrium. The analogy, however, was afterwards shown to be without foundation, by Volta, who excited similar contractions by making a connection between two parts of a nerve, between two muscles, or between two parts of the same muscle; but to produce the effect two different metals were found to be essential. Hence he was led to infer that, by the contact of different metals, a small quantity of electricity is excited; and to the agency of this electricity, first upon the nerves, and through their mediation on the muscles, he ascribed the phenomena in question.

Several years elapsed, during which the action of galvanic electricity on the animal body, and the discussion of its cause, chiefly occupied the attention of philosophers. Early in 1800, the subject took a new turn, in consequence of the discovery by Signor Volta of the Galvanic Pile*; a discovery which has furnished us with new and important instruments of analysis, capable, if any such there are, of leading to a knowledge of the true elements of bodies. From this period, discoveries have multiplied with a rapidity and to an extent, which surpass any thing before known in the history of science; and the facts are now become so numerous, that an arrangement and classification of them seem to be preferable to an historical detail in the order of time. The method, which appears to me best calculated to give a distinct view of the subject, is to describe,

I. The construction of galvanic apparatus, and the circumstances essential to the excitement of this modification of electricity:

II. The facts, which establish its identity with the electricity excited by ordinary processes:

III. The agency of the electric or galvanic fluid in producing chemical changes:

IV. The theory, by which these changes, in the present state of our knowledge, are best explained: And

V. The hypotheses, which have been framed to account for the origin of the electricity, excited by galvanic arrangements.

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* Philosophical Transactions, 1800, or Philosophical Magazine, vii. 289.

SECTION I.

Of the Construction of Galvanic Arrangements.

For the excitation of ordinary electricity, it is well known that a class of substances are required, called *electrics*, by the friction of which the electric fluid is accumulated, and from which it may be collected by a different class of bodies termed *non-electrics* or *conductors*. When friction, for example, is applied to the glass cylinder of an electrical machine, the electric fluid flows to it from surrounding bodies, and from thence passes to the prime-conductor, in which it exists in a greater than natural quantity. All then that is effected, by the action of the machine, is a disturbance of the natural quantity of electricity in bodies, or a transfer of it from some to others, in consequence of which, while the latter acquire a redundance, the former become proportionally deficient in their quantity of electricity*.

The conditions necessary to the excitement of galvanic electricity are altogether different; for the class of bodies, termed electrics, have now no longer any share in the phenomena. All that is required is the simple contact of different conducting bodies with each other, one or more of which must be in a fluid state. Conductors of electricity have been divided into *perfect* and *imperfect*, the former comprehending the metals, plumbago and charcoal; the latter, or imperfect, including water, saline solutions, acids, and animal fluids.

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* This theory appears to me to have several advantages over the hypothesis, which supposes two different *kinds* of electricity, capable by combination of neutralizing each other. The only modification, I would propose, is a slight change of terms; and the adoption of those suggested by Mr. Cuthbertson, *viz. condensed* and *rarefied*, instead of *positive* and *negative* electricity.

The least complicated galvanic arrangement is termed A SIMPLE GALVANIC CIRCLE. It consists of three conductors, two of which must be of the one class, and one of the other class. In the following tables, constructed by Mr. Davy, some different simple circles are arranged in the order of their powers, the most energetic occupying the highest place.

TABLE OF SOME GALVANIC CIRCLES,

Composed of two perfect Conductors and one imperfect Conductor.

More oxidable substances.	Less oxidable substances.	Oxidating fluids.
Zinc.	With gold, charcoal, silver, copper, tin, iron, mercury.	Solutions of nitric acid in water, of muriatic acid and sulphuric acid, &c. Water holding in solution oxygene, atmospheric air, &c.
Iron.	.... gold, charcoal, silver, copper, tin.	
Tin.	.... gold, silver, charcoal.	
Lead.	.... gold, silver.	
Copper.	.... gold, silver.	Solution of nitrate of silver and mercury. Nitric acid, acetous acid.
Silver.	.... gold.	Nitric acid.

TABLE OF SOME GALVANIC CIRCLES,

Composed of two imperfect Conductors and one perfect Conductor.

Perfect conductors.	Charcoal. Copper. Silver. Lead. Tin. Iron. Zinc.	Imperfect conductors.	Solutions of hydrogenated alkaline sulphurets, capable of acting on the first three metals, but not on the last three.	Imperfect conductors.	Solutions of nitrous acid, oxygenated muriatic acid, &c. capable of acting on all the metals.

The powers of simple galvanic circles are but feeble; but they may be made sufficiently apparent by the following experiments.

1. When a piece of zinc is laid upon the tongue, and a piece of silver under it, no sensation is excited, so long as the metals are kept apart; but, on bringing them into contact, a metallic taste is distinctly perceived. In this case we have an example of the arrangement of two perfect conductors (the metals) with one imperfect one (the tongue, or rather the fluids which it contains). The metallic taste arises, in all probability, from the excitement of a small quantity of electricity by the contact of the metals, and its action on the nerves of the tongue.

2. A piece of zinc, immersed under water which is freely exposed to the atmosphere, oxidizes very slowly; but when placed in the same situation, in contact with a piece of silver, its oxidation is much more rapid. By immersing iron and silver (also in contact with each other) under diluted muriatic acid, the action of the acid upon the iron is considerably increased; and hydrogen gas is

evolved from the water, not only where it is in contact with the iron, but where it touches the silver. These facts explain, why, in the sheathing of ships, it is necessary to use bolts of the same metal which forms the plates; for if two different metals be employed, they both oxidate or rust very speedily, in consequence of their forming, with the water of the ocean, a simple galvanic circle.

Of compound Galvanic Circles or Batteries.

Galvanic batteries are formed by multiplying those arrangements, which compose simple circles. Thus if plates of zinc and of silver, and pieces of woollen cloth of the same size as the plates and moistened with water, be piled upon each other (fig. 77, pl. ix), in the order of zinc, silver, cloth; zinc, silver, cloth; and so on, for twenty or more repetitions, we obtain a galvanic battery termed, from its discoverer, the *Pile of Volta.* The power of such a combination is sufficient to give a smart shock, as may be felt by grasping in the hands, which should be previously moistened, two metallic rods, and touching with these the upper and lower extremities of the pile. The shock may be renewed at pleasure; until, after a few hours, the activity of the pile begins to abate, and finally ceases altogether.

The metals, composing a galvanic battery, may be more conveniently arranged in the form of a trough, a happy invention of Mr. Cruickshank. In a long and narrow wooden trough, made of baked wood, grooves are cut, opposite to and at the distance of about one 3d of an inch from each other; and into these are let down, and secured by cement, square plates of zinc and copper, previously united together by soldering. (See figs. 37 and 78.) The space, therefore, between each pair of plates, forms a cell for the purpose of containing the liquid, by which the combination is to be made active. The advantage of this contrivance, over the pile, is partly that it is much more easily put in order; but, besides this, it is a more efficient instrument. When constructed in the way which has been described, it affords an example of a galvanic combination *of the first kind*, formed by two perfect and one imperfect conductor. But it admits of being modified, by cementing, into the grooves, plates of one metal only, and filling the cells, alternately, with two different liquids, as diluted nitric acid and solution of sulphuret of potash. In this case, we have a battery *of the second order*, formed by the repetition of one perfect and two imperfect conductors. For all purposes of experiment, the first kind of arrangement is universally preferred.

Another modification of the apparatus, which may be called the *Chain of Cups*, was proposed by Volta, at the same time that he communicated his invention of the Pile; and, from the recent experiments of Mr. Children*, it appears to be a very useful and powerful one. It consists of a row of glasses (see fig. 75), such as wine glasses or small tumblers, for the purpose of containing any fluid that may be selected. Into each of these glasses is plunged a plate of zinc and another of copper, each not less than an inch square, which are not to touch each other. The plates of different cups are connected by metallic wires or arcs, in such a manner that the zinc of the first cup communicates with the copper of the second; the zinc of the second with the copper of the third; and so on through the whole row. The shock is felt on dipping the fingers of one hand into the fluid of the first cup, and those of the other hand into the last of the series. The superiority of this arrangement consists in both surfaces of each metallic plate being exposed to the action of the liquid; whereas, by soldering the plates together, one of the surfaces of each is protected from the liquid, and contributes nothing to the effect. The common trough has lately been made to combine this advantage, by dividing it into cells, not by plates of metal but by partitions of glass. Into each of these cells, filled with the proper liquid, a plate of each metal is introduced, but not so as to touch each other. A communication is then made, by a metallic arc, between the zinc plate of each cell and the copper one of the next, precisely as in the chain of cups. More lately the troughs themselves have been made of earthen ware, and the partitions of the same material; the apparatus being completed, in other respects, in the manner already described.

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* Philosophical Transactions, 1809, page 32.

The size of the plates has been varied from one or two inches to several feet. For ordinary purposes, plates of two inches square are sufficient; but for the repetition of Mr. Davy's experiments, not less than 100 pairs of plates, each four inches square, are required*. The enlargement of the size of the troughs, so as to contain this number, would be extremely inconvenient; and we may therefore combine the power of several troughs, by uniting the zinc end of the one with the copper end of the other, by the intervention of a metallic wire.

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* Some useful information respecting the number and size of plates, adapted to different purposes, is given by Mr. Singer in Nicholson's Journal, vol. xxiv. page 174.

It may be sufficient to add, in general terms, that every combination, which is capable of forming a simple galvanic circle, may, by sufficient repetition, be made to compose a battery. The combinations, also, which are most active in simple circles, are observed to be most efficient in compound ones. The foregoing tables of Mr. Davy express, therefore, the powers of compound as well as of simple arrangements.

To construct a battery *of the first order*, it is essential that a fluid be employed, which exerts a chemical action upon one of the metals. Pure water, entirely deprived of air, appears to be inefficient. In general, indeed, the galvanic effect is, within certain limits, proportional to the rapidity with which the more oxidable metal is acted upon by the intervening fluid. The fluid generally used is nitric acid, diluted with 20 or 30 times its weight of water. Mr. Children recommends a mixture of three parts fuming nitrous acid, and one sulphuric, diluted with thirty parts of water. Directions, also, respecting the best kind and density of acids, for producing galvanic electricity, are given by Mr. Singer. From his experiments it appears, that acid of different densities is required for different purposes.

The power of the apparatus has been found to be increased, when insulated by non-conductors; and when surrounded by an atmosphere of oxygen gas; not sufficiently, however, to make it necessary to resort to either of these expedients in ordinary cases. Oxygen gas disappears in this process, when carried on under a receiver; and, after all the oxygen is absorbed, the effect ceases, and is renewed by introducing a fresh portion. When the cells are filled with diluted nitric acid, the apparatus continues active, even under the exhausted receiver of an air-pump, or in an atmosphere of carbonic acid or nitrogen gases. But if the cells be filled with water only, all action is suspended, by placing it under any of these circumstances. Hence it appears that the oxidation of one or both of the metals, composing the trough, is essential to the excitement of galvanic electricity.

SECTION II.

On the mutual Relation of Electricity and Galvanism.

Is the influence, it may now be inquired, which is called into action in a way so different from that employed for the excitation of ordinary electricity, identical with it or of a different kind? This question will be decided by examining whether any of those phenomena, which are occasioned by the agency of the electric fluid, are produced also by that of galvanism; and we shall find the following striking resemblances.

1. The sensation, produced by the galvanic shock, is extremely similar to that which is excited by the discharge of a Leyden jar. Both influences, also, are propagated through a number of persons, without any perceptible interval of time.

2. Those bodies, which are conductors of electricity, are also conductors of the galvanic fluid, as the metals, charcoal, and a variety of liquids. Again, it is not transmitted by glass, sulphur, and the whole class of electrics, which do not convey ordinary electricity. Among liquids, those

only are conductors of electricity and galvanism, which contain oxygen as one of their elements*.

3. The galvanic fluid passes through air and certain other non-conductors, in the form of sparks; accompanied with a snap or report; and, like the electric fluid, it may be made to inflame gun-powder, phosphorus, and mixtures of hydrogen and oxygen gases. It has lately been found, also, by Mr. Children, that in the Voltaic apparatus there is, what is called in electricity, *a striking distance.* With a power of 1250 pairs of four inch plates, he found this distance to be one 50th of an inch, the thickness of a plate of air, through which the galvanic discharge is able to pass in the form of a spark. Increasing the number of plates, the striking distance will be greater; and the reverse when it is diminished.

4. The Voltaic apparatus is capable of communicating a charge to a Leyden jar, or even to a battery. If the zinc end of a pile (whether it be uppermost or the contrary) be made to communicate with the inside of a jar, it is charged positively. If circumstances be reversed, and the copper end be similarly connected, the jar is charged negatively†. The shocks do not differ from those of a jar or battery, charged to the same intensity by a common electrical machine.

5. Galvanism, even when excited by a single galvanic circle only (such as a piece of zinc, a similar one of copper, and a piece of cloth moistened with a solution of mu-

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* Cruikshank, in Nicholson's 4to. Journal, iv. 258.

† Cuthburtson's Practical Electricity and Galvanism, page 261. Volta, in Nicholson's Journal, 8vo. i. 140. Van Marum, in Philosophical Magazine, xii. 162.

riate of ammonia), distinctly affects the gold leaf of the condensing electrometer. If the zinc end be uppermost, and be connected directly with the instrument, the electricity indicated is positive; if the pin of the electrometer touch the copper, the electricity is negative. A pile consisting of sixty combinations produces the effect still more remarkably[*].

6. The chemical changes produced by galvanic and common electricity, so far as they have hitherto been examined, are precisely similar. These will form the subject of the following section.

SECTION III.

On the Chemical Agencies of Electricity and Galvanism.

The effects of the electric or galvanic fluids, in producing chemical decomposition, cannot be described, without introducing to the reader the names of several substances, with which, in the present state of his knowledge, he may be supposed to be unacquainted. This difficulty is unavoidable; for it is impossible to explain the general laws of electro-chemical action, without a variety of particular instances. In general, however, it will be found that a minute acquaintance with the bodies, which are brought in illustration, is by no means essential; and that it is sufficient to consider them as composed simply of two ingre-

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[†] Nicholson, 8vo. i. 139, and ii. 281; Cuthbertson, page 264.

dients, which are in opposite electrical states, and are subject to the laws of electrical attraction and repulsion.

The most simple chemical effect, produced alike by the agency of electricity and galvanism, is the *ignition* and *fusion* of *metals*. When a piece of watch pendulum wire is placed in the circuit of a common electrical battery, containing not less than three or four square feet of coating, at the moment of the discharge the wire becomes red-hot; but continues so only for a few seconds, no longer, indeed, than if it had been ignited in any other way*. The same effect may be produced by making a piece of wire the medium of communication between the opposite extremities of a galvanic trough; but, in this case, the heat continues sensibly longer, than when it is excited by an electrical explosion. Water, surrounding a wire so placed, may be made to boil briskly.

When the power of an electrical battery is increased, metallic wires, by transmitting the discharge through them, may be melted and dispersed in the form of smoke, or of an impalpable powder lighter than air. The galvanic discharge, also, is capable of fusing metallic wires; but being less violent, it does not scatter their particles to a distance. Even wire from the most infusible of the metals, platina, acquires a white heat, and melts into globules.

With a still more powerful electrical battery (one for example containing about 18 square feet) metallic wires are not only melted, but undergo absolute *combustion*.

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* On the quantity of coated surface required for igniting different lengths of wire, the reader may consult Mr. Cuthbertson's book, page 161, &c.

Lead and tin wire emit a yellow light, and copper and silver a green one. If the experiment be made on wire confined in a glass receiver, which contains a measured quantity of air, the bulk of the air, and its proportion of oxygen, are both found to be diminished*. The metals are converted into oxides of different colours; lead, tin, and zinc, into white oxides; platina, gold, silver, and copper, into oxides of a dark colour. The experiment may be pleasingly varied by passing the discharge through wires, stretched over panes of glass or sheets of paper, at a small distance from their surface. The metallic oxide which is produced is forcibly driven into the glass or paper; and produces beautiful figures, varying in colour with the metal employed†.

The combustion of metals may be effected, also, by galvanic electricity; but for this purpose the form of very thin leaves is preferable to that of wire. The plates, composing the galvanic trough, should, for this purpose, be not less than four inches square; and several troughs should be joined together, so as to form an aggregate of not less than 100 or 150 pairs of plates. The galvanic influence is to be conveyed by wires brought from each extremity of the arrangement, and placed in contact with the opposite surfaces of the leaf. For the protection of the fingers, the wires should be inclosed in glass tubes. When thus exposed, the metals burn, or rather deflagrate, with great brilliancy.

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* Cuthbertson, page 199.

† Cuthbertson, page 226, and Wilkinson's Elements of Galvanism, in the 9th plate of which these appearances are represented.

Gold emits a very vivid white light, inclining a little to blue, and leaves an oxide, whose colour verges towards that of mahogany. Copper presents similar phenomena.

The flame of silver is a vivid green, somewhat like that of a pale emerald, and the light is more intense than that of gold. Lead gives a vivid light of a dilute blueish purple. Tin a light similar to that of gold; and zinc a blueish white flame fringed with red*. In all these cases, provided the power be sufficiently strong, the deflagration is kept up, for some time, without intermission.

But a much more remarkable action is exerted by the electric and galvanic fluids, in disuniting the elements of several combinations. One of the first discoveries of the chemical agency of the pile was its power of decomposing water. Two pieces of any metallic wire are thrust through separate corks, which are fitted into the open ends of a glass tube in such a way, that the extremities of the wires, when the corks are in their places, may not be in contact, but may be at the distance from each other of about a quarter of an inch (see fig. 77, *a*). If the parts of the wire, which project from without the tube, be made to communicate, the one with the zinc or positive end, and the other with the copper or negative end, of a galvanic battery, a remarkable appearance takes place. The wire, connected with the zinc or positive end of the pile or trough, where it is in contact with the water, if of an oxidable metal, is rapidly oxidized; while from the negative wire a stream of small bubbles of gas arises. But if the wires employed be of a metal, which is not susceptible of oxidation, such as gold or platina, gas is then extricated from both wires, and, by a simple contrivance, may be separately collected. The apparatus for this purpose is shown by fig. 76, where the wires *p* and *n*, instead of being introduced into a straight tube, are inclosed in a syphon, and terminate before they reach the end, in which a small hole is to be ground. When a stream of galvanic electricity is made to act upon water thus confined, oxygen gas is found, at the close of the experiment, in the leg connected with the positive end of the battery, and hydrogen gas in that connected with the negative end; and in the proportions which, by their union, compose water. At an early period of the inquiry, it was found, however, by Mr. Cruickshank, that the water surrounding the positive wire became impregnated with a little acid; and that around the negative wire with a little alkali. If instead of water, we employ a metallic solution, the metal is revived round the negative wire *n*, and no hydrogen gas is liberated.

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* Philosophical Magazine, xi. 284, and xv. 96.

The gases constituting water, it was afterwards discovered by Mr. Davy, may be separately produced from two quantities of water, not immediately in contact with each other. The fact is of peculiar importance, from its resemblance to other more recent ones, which have led that distinguished philosopher to the discovery of the general laws of electro-chemical action. Two glass tubes (*p* and *n*, pl. ix. fig. 79), about one third of an inch diameter and four inches long, having each a piece of gold wire sealed hermetically into one end and the other end open, were filled with distilled water and placed inverted in separate glasses filled, also, with that fluid. The two glasses, *a* and *b*, were made to communicate, either by dipping the fingers of the right hand into one glass, and those of the left into the other, or by interposing fresh animal muscle, or a living vegetable, or even moistened thread, as shown at *c*. The gold wires, projecting from the sealed ends of these tubes, were then connected, the one with the positive, the other with the negative end of the trough. Gas was immediately evolved from both wires. At the close of the experiment, in the tube *p* oxygen gas was found; in the negative tube *n* hydrogen. The proportions by measure were, as nearly as possible, those which result from the decomposition of water, *viz.* two of hydrogen to one of oxygen gas*. Now if these gases arose, as they necessarily must, from the decomposition of the same portion of water, that portion of water must have been contained either in the tube *p* or in the tube *n*. In the former case, the hydrogen gas, found after the process in *n*, must have passed invisibly from *p* to *n*, through the intermediate substance *c*. Or, if the water was decomposed in *n*, then the reverse process must have happened with respect to the oxygen; and it must have been transmitted, in a like imperceptible manner, from *n* to *p*. Facts of this kind, evincing the transference of the elements of a combination, to a considerable distance, through intervening substances, and in a form that escapes the cognizance of our senses, however astonishing, it will appear from the sequel, are sufficiently numerous and well established.

Different chemical compounds require, for the disunion of their elements, galvanic arrangements of various powers and intensities. The decomposition of water is easily effected by a series of fifty pairs of plates, each one or two inches square. But for those which remain to be described, instruments of much greater power are necessary.

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* Nicholson's Journal, 4to. iv. 276.

The apparatus, employed in the masterly experiments of Mr. Davy, which have laid the groundwork of this new field of science, was extremely simple. In cases, where liquid substances were operated upon, he employed occasionally the agate cups *p* and *n*, fig. 80, each of which was capable of holding about sixty grains of water. They were connected together, as shown in the figure at *a*, by the fibres of a peculiar flexible mineral called amianthus; and into each was inserted a platina wire, the bent extremity of which is seen, in each figure, projecting above the cup. When the vessels were in actual use, the wire of *p* was connected with the zinc or positive end of a powerful galvanic series; and that of *n* with the copper or negative extremity.. For the agate cups two hollow gold cones were occasionally substituted (*p* and *n*, fig. 81) the wire projecting from *p* being connected with the positive, and that from *n* with the negative end of a trough or series of troughs. Solid bodies were submitted to the galvanic influence, either by immersing small pieces of them in the gold cones; or, at other times, by making the cups themselves of the substance intended to be decomposed. Or if it was desirable to preserve them from contact with water, they were laid on a small insulated dish of platina, with the inferior surface of which, immediately under the substance used, a wire from one end of the battery was connected, while the substance itself was made to communicate by another wire, with the opposite extremity of the apparatus.

When the gold cones were both filled with a solution of sulphate of potash (a salt composed of potash and sulphuric acid), after exposure, during a sufficient time, to a powerful galvanic arrangement, pure potash was found in the negative cone *n*, and sulphuric acid in the positive cone *p*. The decomposition was even quite complete; for the liquid in *n* contained no acid, and that in *p* no alkali.

The experiment was repeated with several other neutral salts*; and with the invariable result that the acid collected in the positive cone, and the alkali in the negative one. Strong solutions, or those in which the salt bore a considerable proportion to the water, were more rapidly acted upon than weak ones. Metallic salts were, also, decomposed. The acid appeared, as before, in the positive cone, and the metal was deposited, sometimes with a little oxide, in the negative one.

Salts, which are either insoluble, or very sparingly soluble, in water, had their elements disunited in the following manner. Cups were constructed of them, precisely resembling the gold cones, which, as the salts were hard and compact in their texture, was easily effected. These, after being filled with water, were connected, by platina wires, with the opposite ends of a galvanic battery, the vessels themselves communicating, as before, by means of moistened amianthus. At the conclusion of the experiment, sulphuric acid (when the cups were made of sulphate of lime) was found in the positive cup, and lime water in the negative one. Sulphate of strontites, fluate of lime, and sulphate of barytes were decomposed, though less easily, by the same expedient. In all these cases the acid element was found at the positive side, and the earthy one at the negative side, of the arrangement.

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* Minute directions for exhibiting the transfer of acid and alkali, by means of a power not exceeding thirty pairs of two inch plates, are given by Mr. Singer. (Nicholson's Journal, xxiv. 178.)

These facts evidently point out a transference of the elements of combinations from one electrified vessel or surface to another differently electrified. But the principle is made much more apparent by a little variation of the experiment. Thus, if solution of sulphate of potash be electrified in the positive cone *p*, water alone being contained in *n*, after a sufficient continuance of the electrical action *p* will be found to contain diluted sulphuric acid; and the potash will be discovered in the water of *n*. The alkali must necessarily, therefore, have passed, in an imperceptible form, along the connecting amianthus from the vessel *p* to the vessel *n*. Reversing the experiment, and filling *n* with solution of sulphate of potash, the alkali remains in this cone, and the acid is transferred to the opposite side *p*. In one experiment, in which nitrate of silver was placed in the positive cup, and pure water in the negative one, the whole of the connecting amianthus was covered with revived silver.

In the farther prosecution of the inquiry, Mr Davy succeeded in discovering a still more extraordinary series of facts. When an intermediate vessel (*i*, fig. 82) was placed between the positive and negative cups *p* and *n*, and was connected with both of them by moistened amianthus, it was found that acids may actually be made to pass from *n* to *p*, through the intermediate solution in *i*, without combining with it. Thus, solution of sulphate of potash being put into the negative cup *n*, solution of pure ammonia into *i*, and pure water into *p*, in half an hour sulphuric acid was found in the water of the positive cup, to have reached which it must have been transferred from *n* through the intermediate solution of ammonia. Muriatic acid, also, from muriate of soda, and nitric acid from nitrate of potash, were transferred from the negative to the positive side through an interposed solution of alkali. And contrariwise, alkalis and metallic oxides were transmitted from the positive to the negative side, through intervening solutions of acids.

It is necessary, however, that the solution, contained in the intermediate vessel *i*, should not be capable of forming an insoluble compound with the substance intended to be transmitted through it. Thus sulphuric acid, in its passage from sulphate of potash in the negative cup, through the vessel *i* containing a solution of pure barytes, is detained by the barytes, and falls down in the state of an insoluble compound with that earth.

Bodies, the composition of which is considerably more complicated, are, also, decomposed by galvanic electricity. Thus from certain minerals, containing acid and alkaline matter in only very minute proportion, these ingredients are separately developed. Basalt, for example (a kind of stone which, in 100 grains, contains only 3½ grains of soda and half a grain of muriatic acid), gave, at the end of ten hours, evident traces of alkali round the negative, and of acid round the positive wire. A slip of glass, also, negatively electrified in one of the gold cones, had soda detached from it, and sustained a loss of weight.

It may now be understood, why, by the agency of galvanism on water, alkali appears at the negative and acid at the positive wire. The fact was, for some time, not a little perplexing to Mr. Davy; till, at length, he ascertained that all water, however carefully distilled, contains neutral salts in a state of solution. From these impurities, the alkaline and acid elements are separated, agreeably to a law, which has already been explained. In the same way, also, the muriatic acid and alkali are accounted for, which some chemists have obtained by galvanizing what was before considered as pure water; a fact which has been urged in proof of the synthetical production of both those bodies. Absolutely pure water, it has been demonstrated by Mr. Davy, yields nothing but hydrogen and oxygen gases.

All the effects of galvanic arrangements, in producing chemical decompositions, it has been found, may be obtained by ordinary electricity. Its adaptation to this purpose was first successfully attempted by Dr. Wollaston*. The apparatus, which he employed, was similar to that already represented (fig. 77, *a*), excepting that the wires, instead of being exposed to the fluid, contained in the tube, throughout their whole length, were covered with wax, and the points only were laid bare. Or (what was found to answer still better) the wires were inclosed in capillary tubes, which were sealed at their extremities, and then ground away, till the points alone were exposed. The conducting wires, thus arranged, were then introduced into a tube, or other vessel containing the liquid to be operated on, and were connected, the one with the positive, the other with the negative conductor, of an electrical machine, disposed for positive and negative electricity †. When solution of sulphate of copper was thus electrized, the metal was revived round the negative pole. On reversing the apparatus, the copper was re-dissolved, and appeared again at the other wire, now rendered negative.

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* Philosophical Transactions, 1801.

† See Cuthbertson's Practical Electricity.

When gold wires, from $\frac{1}{700}$ to $\frac{1}{1500}$ of an inch in diameter, thus inclosed, were made to transmit electricity, a succession of sparks afforded a current of gas from water. When a solution of gold in nitro-muriatic acid was passed through a capillary tube; the tube then heated to drive off the acid; and afterwards melted and drawn out, it was found that the mere current of electricity, without sparks, evolved gas from water.

Mr. Davy has since proved that by a similar apparatus, solution of sulphate of potash is decomposed, potash appearing at the negative, and sulphuric acid at the positive pole*.

SECTION IV.

Theory of the Changes produced by Galvanic Electricity.

A FACT of considerable importance in explaining the phenomena that form the subject of the last section was discovered several years ago by Mr. Bennett, and has since been confirmed by the experiments of Volta and Davy. Different bodies, it is found, acquire, when brought into contact either by their whole surfaces or by a single point, different states with respect to their quantities of electricity. The best method of performing the experiment is to take two discs or plates, the one of copper, the other of zinc, each about four inches diameter and furnished with an insulating glass handle; to apply them for an instant to each other by their flat faces; and afterwards, to bring them separately into contact with the insulated plate of the condensing electrometer. The instrument indicates, by the divergence of its gold leaves, the electricity acquired by each of the plates, which in the zinc plate is shown to be positive, and in the copper plate negative*.

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* Philosophical Transactions, 1806.

. It had been established, also, by Mr. Davy, in 1801, that when a galvanic arrangement *of the second kind* is constructed, by alternating metallic plates with strata of different fluids, alkaline solutions always receive electricity *from* the metal, and acids on the contrary transmit it *to* the metal. When an arrangement, for example, is made of water, tin, and solution of potash, the current of electricity is from the tin to the alkali. But, in an arrangement of nitric acid, tin, and water, the circulation of electricity is from the acid to the tin. If then the alkali, after having acquired electricity from the metal, could be suddenly separated from the combination, there can be no doubt that it would be found in a positive state. For the contrary reason, the acid, having given electricity to the metal, must, if it could be detached, be found negative.

Still more satisfactory evidence has been since obtained of the electrical state of the acids and alkalis, by examining what kind of electricity they impart to an insulated metallic plate. Various dry acids, being touched on an extensive surface by a plate of copper insulated by a glass handle, the copper was found after contact to have become positively electrified, and the acid negatively. On the contrary, making the experiments with dry earths in a similar manner, the metal became negative. The alkalis gave less distinct results owing to their attraction for moisture. Bodies, moreover, possessing opposite electrical energies towards one and the same body, are found to possess them with regard to each other. Thus when lime and oxalic acid were brought into contact, the earth was found to be positive, and the acid negative. Sulphur appears to be in the positive state. Oxygen, judging from those compounds in which it is loosely combined, is negative; and hydrogen, by the same test, positive.

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* Volta, in Nicholson's Journal, 8vo. i. 136. Wilkinson, ii. 40, 50, 131. Cuthbertson, 267.

Now, if the common laws of electrical attraction and repulsion operate, as there is every reason to believe they must, among bodies so constituted, it will follow that hydrogen, the alkalis, metals, and oxides, being positively electrified, will be repelled by surfaces which are in the same state of electricity as themselves, and will be attracted by surfaces that are negatively electrified. And, contrariwise, oxygen, and the acids (in consequence of the oxygen they contain), being in a negative state, will be attracted by positive surfaces and repelled by negative ones.

To apply this theory to the simplest possible case, the decomposition of water, the hydrogen of this compound, being itself positively electrified, is repelled by the positive wire and attracted by the negative one; while, on the contrary, oxygen, being negative, is repelled by the negative wire, and attracted by the positive one. In the case of neutral salts, the negative acid is attracted by the positive wire; and the positively electrified alkali by the negative wire.

Thus then a power has been discovered, superior in its energy to chemical affinity, and capable either of counteracting it, or of modifying it according to circumstances. The chemical attraction between two bodies may be destroyed, by giving one of them an electrical state opposite to its natural one; or the tendency to union may be increased, by exalting the natural electrical energies.

All bodies, indeed, that combine chemically, so far as they have hitherto been examined, have been found to possess opposite states of electricity. Thus copper and zinc are in opposite states to each other; so are gold and mercury; sulphur and metals, acids and alkalis. Even after combination, it is thought by Mr. Davy not improbable, that bodies may still retain their peculiar states of electricity. If oxygen prevail, in any compound, over the combustible or positive base, the compound is negative, as in certain metallic oxides. But the combustible ingredient may be in such proportion, as to predominate, and to give to the compound a positive energy.

It is an interesting question, but one which can scarcely be determined in the present state of the science, whether the power of electrical attraction and repulsion be identical, as Mr. Davy has suggested, with chemical affinity; or whether it may not rather be considered, like caloric, as a distinct force, which only modifies that of chemical attraction. On the former hypothesis, two bodies, which are naturally in opposite electrical states, may have these states sufficiently exalted, to give them an attractive force superior to the cohesive affinity opposed to their union; and a combination will take place, which will be more or less energetic, as the opposed forces are more or less equally balanced. Again, when two bodies, repellent of each other, act upon a third with different degrees of the same electrical energy, the combination will be determined by the degree. Or, if bodies, having different degrees of the same electrical energy with respect to a third, have likewise different energies with respect to each other, there may be such a balance of attracting and repelling forces as to produce a triple compound.

This hypothesis, it is remarked by Mr. Davy, agrees extremely well with the influence of *mass*, which has been so well illustrated by Berthollet; for many particles, acting feebly, may be equal in effect to fewer acting more powerfully. Nor is it at all contradictory to the observed influence of caloric over chemical union; for an increase of temperature, while it gives greater freedom of motion to the particles of bodies, exalts also, their electrical energies. This Mr. Davy ascertained with respect to an insulated plate of copper and another of sulphur, when heated below 212° Fahrenheit; and at a still higher temperature these bodies, as is well known, combine with the extrication of heat and light, the usual accompaniments of intense chemical action.

On the supposition that electricity is a force, which only modifies the action of chemical affinity, we may regard it, when it promotes combination, as producing this effect by counteracting cohesive attraction. When it impedes combinations, or destroys those which are already formed, it probably acts as a force co-operating with elasticity.

SECTION V.

Theory of the Action of the Galvanic Pile.

Two theories have been framed to account for the phenomena of the Galvanic Pile, and of all similar arrangements. The first, originating with Volta, was suggested by the fact, which may be considered, indeed, as fundamental to it—that electricity is excited by the mere contact of different metals. When a plate of copper and another of zinc are made to touch by their flat surfaces, as was stated in the last section, the zinc, after separation, exhibits positive electricity, and the copper negative. It is natural, therefore, to conclude that a certain quantity of electricity has moved from the copper to the zinc. On trying other metals, Volta found that similar phenomena take place; and by a series of experiments he was led to arrange their powers in the following order, it being understood that the first gives up its electricity to the second; the second to the third; the third to the fourth, and so on.

Silver.
Copper.
Iron.
Tin.
Lead.
Zinc.

The metals, then, have been denominated by Volta, from this property, *motors* of electricity; and the process, which takes place, *electro-motion*, a term since sanctioned by the adoption of it by Mr. Davy.

It is on this transference of electricity from one body to another by simple contact, that Volta explains the action of the instrument discovered by himself, and of all similar arrangements. The interposed fluids, on his hypothesis, have no effect as chemical agents in producing the phenomena, and act entirely as conductors of electricity. Without disputing, however, the accuracy or value of the facts which suggested his theory, it is sufficient for its refutation that it is irreconcileable with other phenomena; and especially with the observation, that the chemical agency of the liquids, on the more oxidizable metal of galvanic arrangements, is essential to their sustained activity. It has been proved, indeed, that the phenomena begin and terminate with the oxidation; and that the energy of the pile bears a pretty accurate proportion to the rapidity of the process. Hence it seems, on first view, an obvious inference, that the oxidation of the metal is the primary cause of the evolution of electricity in galvanic arrangements.

But though the chemical agency of the fluids which are employed is now admitted, on all hands, to be essential to the excitement of this kind of electricity, yet it is by no means universally agreed that we are to consider it as the first in the order of phenomena. It has been suggested by Mr. Davy, as a correction of the theory of Volta, that the electro-motion, occasioned by the contact of metals, is the primary cause of the chemical changes; and that these changes are in no other way efficient, than as they restore the electric equilibrium. To explain this, let us suppose that in any three pairs of plates of a galvanic trough, the zinc plates *z* 1, *z* 2, *z* 3 (fig. 78), are in the state of positive, and the copper plates *c* 1, *c* 2, *c* 3, in that of negative electricity. The liquid, in any cell after the first, will be in contact, on the one side, with positively electrified zinc, and on the other with negatively electrified copper. And if the elements composing the fluid be themselves in different states of electricity, the negatively electrified element will be attracted by the zinc, and the positively electrified element by the copper. Thus when solution of muriate of soda in water is the fluid, the oxygen and the acid will pass to the zinc or positive plate, and the alkali to the copper one; while the hydrogen, having no affinity for copper, escapes. The electric equilibrium will be restored, but only for a moment; for, as the interposed fluid is but a very imperfect conductor of electricity, the zinc and copper plates will, by their electromotive power, again assume their states of opposite electricity; and these changes will go on, as long as any muriate of soda remains undecomposed. In a Voltaic arrangement, therefore, the electrical energies of the metals with respect to each other or to the substances dissolved in water, are the causes disturbing the equilibrium; and the chemical changes are the causes that restore it.

No theory of the galvanic pile, however, can be considered as complete, that does not account for the accumulation of electricity at the zinc end of the apparatus. On the theory that the oxidation of the zinc is the source of the evolved electricity, the fact has been ingeniously explained by Dr. Bostock. He takes it for granted that the electric fluid has an affinity for hydrogen; and supposes that the electricity, evolved at the surface of the first zinc plate, is carried, united to hydrogen, through the fluid of the cell to the opposite copper plate. Here the hydrogen and electricity separate; the former flies off in the state of gas, and the latter passes onwards to the next zinc plate. Being in some degree accumulated in this plate, it is disengaged by the action of the fluid in a more concentrated state than before. And in the same manner, by multiplying the number of pairs, it may be made to exist, in the zinc end of the pile, in any assignable degree of intensity.

On this theory, the electricity evolved is actually *generated* by the chemical action of the interposed fluids on every zinc plate of the series; and its accumulation is the aggregate of what is thus evolved. The concentration, which takes place at the zinc end of the arrangement admits, however, of being explained by the hypothesis of Volta, especially as modified by Mr. Davy. Taking the first cell as an example, the fluid interposed between the positive zinc plate z 1, fig. 78, and negative copper plate c 2, being itself a conductor of electricity, must in time produce an equilibrium between these two plates; but this can only be done by the passage of a certain quantity of electricity across the fluid. The absolute quantity of electricity will, therefore, be diminished in the first pair, and increased in the second. In like manner, the second zinc plate will give up part of its electricity to the third copper plate, and the second pair of plates will be deprived of part of its electricity. The electricity, thus lost by the second pair, it will regain from the first pair of plates. By multiplying, in this way, the number of plates, every successive pair, as we advance in the series, has a tendency to diminish the quantity of electricity in the first; and to have its own state of electricity proportionally exalted.

When a communication is made between two extremities of a series, for example between z^3 or its contiguous cell, and c^1, the opposite electricities tend to an equilibrium. The third pair gives up a share of its electricity to the first; and the intermediate pair, being placed between equal forces, remains in equilibrio. Hence, in every galvanic arrangement, there is a pair of plates at the centre, which is in its natural state of electricity. The effect of such a communication must necessarily be to reduce the pile to a state of inactivity, if there did not still exist some cause sufficient to destroy the equilibrium. On the hypothesis of Volta, this can be nothing else than the property of electro-motion in the metals, which originally produced its disturbance.

Such are the hypotheses that have been framed to explain the phenomena of the Voltaic pile. In the present state of the science, neither of them is entitled to be received as altogether satisfactory; and I have stated them rather with the view of exciting than of satisfying inquiry. On the theory of galvanic electricity, it only remains to point out its difference from the electricity developed by ordinary processes; and to explain the different effects, which are produced by varying the size of the plates in galvanic arrangements.

Though the identity of common and galvanic electricity appears to be sufficiently established, yet in some of their phenomena, which have already been described, there is a considerable difference. To explain these, it was long ago suggested by Mr. Nicholson*, that the electricity, excited by the common machine, is developed in much smaller quantity, but in a higher state of concentration or intensity, than the electricity of galvanism. Hence, its velocity is much more rapid; and hence it readily passes through plates of air and other non-conductors, that are scarcely permeable by galvanic electricity. By virtue of the same property it disperses the metals in the form of smoke;

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* See his Journal, 4to. iv. 244.

while the utmost effect of a Voltaic arrangement is to melt them into globules. By doubling the quantity of galvanic electricity, also, we ignite only a double length of metallic wire, and the ignition is more permanent; but the intensity of common electricity is such that by doubling its quantity we ignite four times the length of wire, and the effect is little more than momentary*.

The comparative quantities of electricity evolved by the common machine and by a Voltaic apparatus, have been made a subject of calculation by Mr. Nicholson. A pile consisting of 100 half crowns, with the same number of pieces of zinc, produces, he found, 200 times more electricity than can be obtained, in an equal time, from a 24 inch plate machine in constant action. Van Marum has, also, observed that a single contact of a Leyden jar or battery with a Voltaic pile charges it to the same degree, as six contacts with the prime conductor of a powerful machine.

It might naturally be expected that a proportion would be observed between the quantity of surface composing galvanic arrangements, and their power of action; and such, with some limitation, is the fact. With plates of the same size, the effect, generally speaking, is proportional to the number. But by enlarging the size, without increasing the number, neither the shock nor the power of decomposing water and other imperfect conductors, is proportionally increased. On the contrary, to obtain a great increase of effect in the combustion of metals, it is necessary to enlarge considerably the size of the plates. Thus 100 plates of four inches square produce, in this way, an incomparably greater effect, than the same surface divided into four times the number.

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* Cuthbertson, page 278.

The effect of multiplying the number of plates, it has already been observed, is that we obtain electricity of a higher intensity, and it has even been shown by Volta* that the proportion is, as nearly as can be judged, an arithmetical one. If for example, we have a certain intensity with 20 pairs, it is doubled by 40, trebled by 60, and so on. But by enlarging the size, without increasing the number, it has also been shown that we gain, not in intensity, which remains exactly the same, but in quantity. Now, for the combustion of metals what we principally want is a large quantity of electricity; for as they are perfect conductors, it finds a ready passage through them even when of low intensity. On the contrary, to find its way through fluids and other imperfect conductors, it must be evolved in a high state of concentration. The facts, therefore, accord sufficiently well with the explanation, to entitle it to be received as a probable hypothesis.

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* Nicholson's Journal, 8vo. i. 139.

CHAPTER VIII.

ALKALIS.

THE alkalis, in their pure state, are the products of chemical operations, which will be described in the sequel. They are distinguished by the following

General Qualities.

The properties, common to all the three alkalis, may be shown by those of a solution of pure potash.

(*a*) The alkalis change vegetable blue colours, as that of an infusion of violets, to green.

(*b*) They have an acrid and peculiar taste.

(*c*) They serve as the intermedia between oils and water.

(*d*) They corrode woollen cloth; and, if the solution be sufficiently strong, reduce it to the form of a jelly.

(*e*) They are readily soluble in water.

SECTION I.

Pure Potash and pure Soda.

Art. 1.—*Their Preparation and General Qualities.*

To prepare pure potash, dissolve any quantity of American or Dantzic pearlash in twice its weight of boiling water, and add the solution, while hot, to an equal weight of fresh quicklime, slaked with six times its weight of hot water. Boil the mixture in an iron kettle, and continue stirring during half an hour. Then separate the liquid alkali, either by filtering through calico or by subsidence; and boil it to dryness in a silver dish. Pour, on the dry mass, as much pure alcohol as is required to dissolve it; put the solution into a bottle, and let the insoluble part settle to the bottom. Then decant the alcoholic solution of potash; and distil off the alcohol in an alembic* of pure silver, furnished with a glass head. Pour the alkali, when in fusion, upon a silver dish, and, when cold, break it into pieces, and preserve it in a well-stopped bottle. If the distillation of the alcohol be not carried so far, the alkali will shoot, on cooling, into regular crystals.

From the recent electro-chemical researches of Mr. Davy, it appears that potash is not completely deprived of carbonic acid, by any process hitherto employed for its preparation. (Philosophical Transactions, 1808, 355.) Berthollet has shown, also, in the second volume of the *Mémoires d'Arcueil*, that solid potash, even after being kept some time in a state of fusion, still retains a considerable quantity of water, not less than 13½ parts in 100. This is proved, besides other methods, by mixing it with iron-filings, and exposing it to a red-heat, by which an enormous quantity of hydrogen gas is disengaged.

In the same mode may pure soda be prepared, substituting the carbonate of soda for the pearlash.

These alkalis have the following properties:

(*a*) They powerfully attract moisture from the atmosphere, or deliquiate.

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* The figure of an alembic may be seen in pl. i. fig. 2.

(*b*) They readily dissolve in water, and produce heat during their solution, if the fused alkalis be employed; but the crystallized alkalis generate cold, when dissolved.

(*c*) They are not volatilized by a moderate heat, and hence have been called fixed alkalis.

(*d*) When melted with silex, in proper proportions and by a sufficient heat, they form glass.

ART. 2.—*Analysis of the two fixed Alkalis.*

Though it had long been conjectured* that the fixed alkalis are not simple or elementary bodies, yet no distinct evidence had been obtained of their nature, until, in the year 1807, it was furnished by the splendid discoveries of Mr. Davy. From the facts which have been stated in a former section respecting the powers of electrical decomposition, it appeared to that philosopher a natural inference that the same powers, applied in a state of the highest possible intensity, might disunite the elements of bodies, which had resisted all other instruments of analysis. If potash, for example, were an oxide, composed of oxygen united to an inflammable base, it seemed to him probable, that when subjected to the action of opposite electricities, the oxygen would be attracted by the positive wire and repelled by the negative. At the same time, the reverse process might be expected to take place with respect to the combustible base, the appearance of which might be looked for at the negative pole.

In his first experiments, Mr. Davy failed to effect the decomposition of potash, owing to his employing the alkali in a state of aqueous solution, and to the expenditure of the electrical energy in the mere decomposition of water. In his next trials, the alkali was liquefied by heat in a platina dish, the outer surface of which, immediately under the alkali, was connected with the zinc or positive end of a battery consisting of 100 pairs of plates, each six inches square. In this state, the potash was touched with a platina wire proceeding from the copper or negative end of the battery; when instantly a most intense light was exhibited at the negative wire, and a column of flame arose from the point of contact, evidently owing to the development of combustible matter. The results of the experiment could not, however, be collected, but were consumed immediately on being formed.

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* See Philosophical Magazine, xxxii. 18, 62.

The chief difficulty in subjecting potash to electrical action is, that in a perfectly dry state it is a complete non-conductor of electricity. When rendered, however, in the least degree moist by breathing on it, it readily undergoes fusion and decomposition, by the application of strong electrical powers. For this purpose, a piece of potash, weighing from 60 to 70 grains, may be placed on a small insulated plate of platina, and may be connected, in the way already described, with the opposite ends of a powerful electrical battery, containing not less than 100 pairs of six inch plates. On establishing the connection, the potash will fuse at both places where it is in contact with the platina. A violent effervescence will be seen at the upper surface, arising, as Mr. Davy has ascertained, from the escape of oxygen gas. At the lower or negative surface, no gas will be liberated; but small bubbles will appear, having a high metallic lustre, and being precisely similar in visible characters to quicksilver. Some of these globules burn with an explosion and bright flame; while others are merely tarnished, and are protected from farther change by a white film, which forms on their surface*.

This production of metallic globules is entirely independent of the action of the atmosphere; for Mr. Davy finds that they may be produced *in vacuo*.

Pure soda gives similar results; but its decomposition demands a greater intensity of action. The quantity of soda should not exceed 15 or 20 grains; and the distance between the platina surfaces must be reduced from $\frac{1}{4}$ to $\frac{1}{8}$ or $\frac{1}{10}$ of an inch. The metal from soda does not, like that from potash, continue fluid at the temperature of the atmosphere; but speedily becomes solid, and bears a considerable resemblance to silver. When the electrical power is much increased, globules of the metal fly with great velocity through the air, in a state of vivid combustion, producing beautiful jets of fire.

To preserve these new substances, it is necessary to immerse them immediately in pure naphtha, a fluid which will be described in a subsequent part of the work. If they are exposed to the atmosphere, they are rapidly converted back again into the state of pure potash or pure soda. To prevent their oxidation still more effectually, Mr. Pepys has proposed to produce them under naphtha; and has contrived an ingenious apparatus for this purpose, which is described in the 31st volume of the Philosophical Magazine, page 241.

When the globules, obtained either from potash or soda, are exposed to the action of air over mercury in graduated

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[*] For the repetition of this experiment, very useful practical directions may be found in a paper by Mr. Singer—Nicholson's Journal, xxiv. 174.

glass tubes, an absorption of oxygen happens; and a crust of alkali is formed on the surface, which defends the interior from farther change. When heat is applied to the globules similarly confined, a rapid combustion ensues, attended with a brilliant white flame. The globules are found, after the experiment, converted into a white substance; which is potash when we have used those from potash, and soda when the globules from soda have been employed. In this process, oxygen is absorbed, and the weight of the alkali produced is found to exceed that of the globules consumed.

When either of these substances is thrown into water, a rapid disengagement of hydrogen gas takes place; and the oxygen of the water, uniting with the globules, regenerates alkali.

Nothing then can be more satisfactory than the evidence, furnished by these experiments, of the nature of the fixed alkalis. By the powerful agency of opposite electricities, each of them is resolved into oxygen and a peculiar base*. This base, like other combustible bodies, is repelled by positively electrified surfaces, and attracted by negative ones; and hence its own natural state of electricity must necessarily be positive. Again, by uniting with oxygen, these bases are once more changed into alkali, either slowly at ordinary temperatures; or with heat and light, if their temperatures be raised. We have the evidence, therefore, both of analysis and synthesis, that each of the fixed alkalis is a compound of oxygen with a peculiar inflammable basis.

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* The proportions of oxygen and base in each will be found at the end of the articles Potassium and Sodium.

But in what class of combustible bodies are we to arrange the alkaline bases? Some properties, common to both, have influenced Mr. Davy to place them among the metals, with which they agree in opacity, lustre, malleability, conducting powers as to heat and electricity, and in their qualities of chemical combination. The only property, which can be urged against this arrangement, is their extreme levity, which even exceeds that of water. But when we compare the differences, which exist among the metals themselves, this will scarcely be considered as a valid objection. Thus tellurium, which no chemist hesitates to consider as a metal, is only about six times heavier than the base of soda, while it is four times lighter than platina; thus forming a sort of link between the old metals and the bases of the alkalis.

In giving names to the alkaline bases, Mr. Davy has adopted that termination, which, by common consent, has been applied to other newly discovered metals, and which, though originally Latin, is now naturalized in our language. The base of potash he has called POTASSIUM, and the base of soda SODIUM; and these names have met with universal acceptation among chemical philosophers.

It is not, however, by electrical means only that the decomposition of the fixed alkalis has been accomplished. Soon after Mr. Davy's discoveries were known at Paris, Messrs. Gay Lussac and Thenard (*Annales de Chimie*, lxv. 325, or *Mémoires d'Arcueil*, ii. 299) succeeded in their attempts to decompose both the fixed alkalis, without the aid of a Voltaic apparatus, and merely by the intervention of chemical affinities. Their process, though it affords the alkaline bases of less purity, yields them in much larger quantity than the electrical analysis, *viz.* to the amount of nearly 400 grains. It consists in bringing the alkalis into contact with intensely heated iron, which, at this temperature, attracts oxygen more strongly than the alkaline base retains it.

The apparatus, used for obtaining potassium, differs very little from that which is commonly employed to decompose water by means of iron*. It consists of a common gun-barrel curved, and drawn out, at one end, to rather a smaller diameter, as represented in the 9th plate, fig. 83, *c*. To one end is adapted an iron tube *a*, of the capacity of two cubic inches, for containing the potash. At the bottom of this tube is a small hole, through which the potash gradually flows. To the opposite end of the gun-barrel a tube of safety *e* is to be cemented; and into this a sufficient quantity poured, either of mercury or of naphtha. Into the gun-barrel, 2½ parts of very clean iron turnings are to be introduced, and pushed on to the bent part *c*. The tube, carefully luted, is then to be placed in a small furnace nine or ten inches in diameter, and provided with a pair of double blast bellows, the pipe from which is shown at *f*. The next step is to insert the tube *a* in its place, after having put into it 1¾ parts of pure potash, deprived of as much water as possible by previous fusion. The whole apparatus should be perfectly dry, clean, and impervious to air.

A strong heat is now to be excited in the furnace *d*; and while this is doing, the tube containing the potash, as well as the opposite end of the barrel, should be kept cool by ice. When the barrel has attained a white heat, the potash in *a* may be melted by a small portable furnace. It will then flow, through the small hole, upon the iron turnings. A considerable quantity of hydrogen gas will be evolved by the decomposition of that portion of water, which the potash retains even after fusion, and which has been shown to exceed 13 *per cent*. When the production of this gas slackens, we may remove the small furnace from beneath the tube *a*, and increase the heat in the furnace *d*, in order to restore to the iron turnings at *c* the temperature proper for decomposing more potash. These operations may be repeated, alternately, till no more gas is produced; but last of all, the heat in the furnace should be strongly raised, in order to drive off some of the potassium, which strongly adheres to the iron turnings.

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* Hachette in Philosophical Magazine, xxxii. 89; and Mr. E. Davy, ditto, page 276.

When the furnace is quite cold, the safety tube *e* is to be removed, and its place supplied by an iron plug. If the end of the gun-barrel, projecting from this side of the furnace, has been kept carefully cooled during the experiment, the metal will be found adhering to it, in the form of brilliant laminæ. In order to extract it, the gun-barrel is to be cut at the commencement of the part which has been kept cool, where the greatest quantity will be found. Another portion will be found close to the plug, and this adheres so slightly to the gun-barrel, that the least effort serves to detach it. It is even partly oxidized by the air, which gains access during the cooling of the furnace; and when the whole is covered with naphtha, the oxidized part is detached in laminæ, exposing a white and brilliant metallic surface.

The *potassium*, which is condensed nearest the furnace, must be detached by a sharp chisel, and in the largest pieces we can possibly break off; for if it be in small molecules, it inflames in the air, even at very low temperatures. In the middle of the gun-barrel we shall find an amalgam of potassium and iron, which becomes green on exposure to the air, the potassium returning to the state of potash.

This operation has been successfully repeated in this country by Mr. Davy and others. When the iron turnings were very clean, the potash very dry and pure, and the whole apparatus free from foreign matters, the metal produced differed very little from that obtained by a Voltaic battery. Its lustre, ductility, and malleability were similar. Its point of fusion and specific gravity, however, were a little higher; for it required nearly 130° Fahrenheit, to render it perfectly fluid, and was to water as 796 to 1000 at 60° Fahrenheit. This Mr. Davy ascribes to contamination with a minute proportion of iron.

Charcoal, it has been asserted by Curaudau (Nicholson's Journal, xxiv. 37), may be employed, also, for the decomposition of the alkalis. To ensure success in the process, great attention, it appears, is necessary to the manipulations, which are fully described in the memoir of the inventor. The fact sufficiently explains an observation of Professor Woodhouse (in Nicholson's Journal, xxi. 290). A mixture of half a pound of soot and two ounces of pearlash, was exposed for two hours in a covered crucible to an intense heat. When the mixture became cold it was emptied upon a plate, and a small quantity of water poured upon it, when it immediately took fire. This could only be owing to the conversion of part of the potash into potassium.

ART. 3.—*Potassium.*

I. The base of potash, at 60° Fahrenheit, exists in small globules, which possess the metallic lustre, opacity, and general appearance of mercury; so that when a globule of mercury is placed near one of potassium the eye can discover no difference between them. At this temperature, however, the metal is only imperfectly fluid; at 70° it becomes more fluid; and at 100° its fluidity is so perfect, that several globules may easily be made to run into one.

By reducing its temperature, potassium becomes, at 50° Fahrenheit, a soft and malleable solid, which has the lustre of polished silver. At about the freezing point of water, it becomes hard and brittle, and exhibits, when broken, a crystallized texture, which, in the microscope, seems composed of beautiful facets of a perfect whiteness and high metallic splendor.

To be converted into vapour, it requires a temperature approaching that of a red heat; and, when the experiment is conducted under proper circumstances, it is found unaltered after distillation.

II. Potassium is a perfect conductor both of electricity and of heat.

III. Its specific gravity at 60° Fahrenheit, making some allowance for unavoidable errors in the experiment, is as 6 to 10, the latter number being assumed as that of water. Gay Lussac and Thenard make it between 8 and 9; but they probably operated on a less pure substance. Even in its solid form, it swims in naphtha, whose specific gravity is about 7½ to 10.

IV. Its combustibility has already been noticed. At the temperature of the atmosphere, it absorbs oxygen slowly; but if heated nearly to redness, or to its point of vaporization, it burns with a brilliant white flame and a very intense heat.

V. It appears to be susceptible of different degrees or stages of oxidizement. Thus by heating it to a point, below what is necessary for its inflammation, as to 400° Fahrenheit, and in a quantity of oxygen gas less than is required for its saturation, its tint is changed to a reddish brown. A similar compound may be formed by fusing together potassium and potash. In this case the potash gives up part of its oxygen to the potassium, and a compound is obtained, intermediate between the metallic base and the alkali.

VI. The action of potassium on water is attended with some beautiful phenomena. When it is thrown upon water exposed to the atmosphere, or when it is brought into contact with a drop of water, it decomposes the water with great violence; an instantaneous explosion is produced with a vehement flame; and a solution of pure potash is the result. The hydrogen gas, which is disengaged, appears to dissolve a portion of potassium; for, on escaping into the air, it forms a white ring of smoke, gradually enlarging as it ascends, like the phosphuretted hydrogen gas.

When water is made to act on the base of potash, atmospheric air being excluded, there is much heat and noise, but no luminous appearance; and the gas evolved is pure hydrogen. It is of importance to remember that each grain of potassium, by acting on water, detaches about 1.06 cubic inch of hydrogen gas.

If a globule of the base of potash be placed on ice, it instantly burns with a bright flame, and a deep hole is made in the ice filled with a fluid which is found to be a solution of potash.

The production of alkali, by the action of water on potassium, is most satisfactorily shown, by dropping a globule of the metal upon moistened paper, which has been tinged with turmeric. At the moment, when the globule comes into contact with the paper, it burns, and moves rapidly as if in search of moisture, leaving behind it a deep reddish brown trace, and acting upon the paper, exactly like dry caustic potash.

So strong, indeed, is the affinity of potassium for oxygen, that it discovers and decomposes the small quantities of water contained in alcohol and ether, even when carefully purified, and disengages, from both these fluids, hydrogen gas.

On naphtha colourless and recently distilled, potassium has very little power of action; but in naphtha, which has been exposed to the air, it soon oxidates, and alkali is formed, which unites with the naphtha into a brown soap, that collects round the globules.

VII. When thrown into the liquid mineral acids, the base of potash inflames, and burns on the surface; or, if kept beneath the surface, its effects are such as may be explained by its affinity for oxygen. In concentrated sulphuric acid, a white saline substance is formed, which is probably concentrated sulphuric acid surrounded by sulphur. At the same time a gas escapes which has the smell of sulphurous acid mixed with hydrogen gas. In nitrous acid, nitrous gas is disengaged, and nitrate of potash formed. In oxymuriatic acid gas, it burns vividly with bright scintillations, and muriate of potash is generated.

VIII. Potassium readily combines with the simple combustibles. To unite it with sulphur or phosphorus, it must be melted with these bodies under naphtha.

The *phosphuret of potassium* requires for its fusion a stronger heat than either of its constituents. It is of the colour of lead; and, when spread out, has a lustre similar to polished lead. By exposure to the air, or by rapid combustion, it forms phosphate of potash.

When potassium is fused with sulphur, in a vessel filled with the vapour of naphtha, a rapid combination ensues, accompanied with heat and light, and a disengagement of sulphuretted hydrogen. The result is a grey substance not unlike artificial sulphuret of iron.

IX. With mercury, potassium gives some extraordinary and beautiful results. The combination is very rapid, and is effected by merely bringing them into contact at the temperature of the atmosphere. The amalgam, in which the potassium is in least proportion, seems to consist of about 1 part in weight of basis and 70 of mercury. It is very soft and malleable; but by increasing the proportion of potassium, we augment, in a proportional degree, the solidity and brittleness of the compound.

In this state of division, potassium appears to have its affinity for oxygen considerably increased. By a few minutes' exposure to the air, potash is formed which deliquiates, and the mercury is left pure and unaltered. When a globule is thrown into water, it produces a rapid decomposition and a hissing noise; potash is formed; pure hydrogen disengaged; and the mercury remains free.

The fluid amalgam of potassium and mercury dissolves all the metals; and in this state of union, mercury even acquires the power of acting on platina.

Potassium unites, also, with gold, silver, and copper; and, when the compounds are thrown into water, this fluid is decomposed, potash is formed, and the metals are separated unaltered.

X. Potassium reduces all the metallic oxides when heated with them, even of those which most powerfully attract oxygen, such as oxides of iron. In consequence of this property it decomposes and corrodes flint and green glass by a very gentle heat; potash is generated with the oxyge taken from the metal, which dissolves the glass and exposes a new surface. At a red heat even the purest glass, formed merely of potash and silex, is acted upon. The alkali in the glass seems to give up a part of its oxygen to the potassium, and an oxide of potassium results, with a less proportion of oxygen than is necessary to constitute potash. The silex, also, it is probable, is partly de-oxidized.

From this summary of the action of potassium, it appears that all the most remarkable effects which it exhibits, are connected with its affinity for oxygen, which is sufficiently energetic to enable it to take oxygen from all other bodies. Hence the application of potassium to any substance is the best test of its containing oxygen, which, if present, it cannot fail to detect.

It was important to determine the proportions in which potassium and oxygen combine, when potash is re-generated. This Mr. Davy investigated by two different processes. The one consisted in ascertaining how much oxygen gas disappears by the action of a given quantity of potassium; the other how much hydrogen is disengaged from water by a known weight of the same substance. Dividing the bulk of the hydrogen gas by 2, he learned the quantity of oxygen which had been taken from the water.

The co-incidence of results, obtained by these different methods, is a strong confirmation of their accuracy. By the action of potassium on oxygen gas, it appeared, on an average, that

$$\text{Potash consists of} \begin{cases} 86.1 \text{ potassium.} \\ 13.9 \text{ oxygen.} \end{cases}$$

$$\overline{100}$$

By the agency of water, the proportions differed only by a small fraction, so that we may state in round numbers that the base is to the oxygen as six to one, or that

$$\text{Potash is composed of} \begin{cases} 86 \text{ potassium.} \\ 14 \text{ oxygen.} \end{cases}$$

$$\overline{100}$$

Potassureted Hydrogen Gas.

This name I would propose for the solution of potassium in hydrogen gas, which, it has already been stated, results from the action of potassium on water, and, as appears from Mr. Davy's experiments, may be formed, directly, by heating the metal in hydrogen gas. A large portion of potassium is thus dissolved; but the greater part precipitates on cooling. The properties of this gas have not been minutely investigated. It is spontaneously inflammable in the atmosphere; burns with a very brilliant light, which is purple at the edges; and throws off dense vapours of potash.

It does not appear that potassium has any power of absorbing and condensing hydrogen. Neither is the base of this gas at all acted upon by potash; a fact among many others, contradictory of the theory of Gay Lussac and Thenard, that potassium is a compound of hydrogen and potash.

Art. 4.—*Sodium.*

The base of soda agrees in many of its properties with the base of potash, and exerts on several bodies a precisely similar action, except that the results are compounds of soda instead of potash. Thus with nitric acid it affords nitrate of soda; with oxy-muriatic acid, muriate of soda. In this place, therefore, I shall describe only such of its properties as are peculiar to and characteristic of it.

I. Sodium, at common temperatures, exists in a solid form. It is white, opaque; and, when examined under a thin film of naphtha, has the lustre and general appearance of silver. It is exceedingly malleable, and much softer than any of the common metallic substances. When pressed upon by a platina blade with a small force, it spreads into thin leaves; and a globule of $\frac{1}{10}$th or $\frac{1}{12}$th of an inch in diameter is easily spread over a surface of a quarter of an inch. This property is not diminished, by cooling it to 32° Fahrenheit. Several globules, also, may, by strong pressure, be forced into one; so that the property of *welding*, which belongs to platina and iron at a strong heat only, is possessed by this substance at common temperatures.

II. It is lighter than water. As near as can be determined, its specific gravity is as 0.9348 to 1.

III. It is much less fusible than the base of potash. At 120° Fahrenheit, it begins to lose its cohesion, and it is a perfect fluid at about 180°. Hence it readily fuses under heated naphtha.

IV. Its point of vaporization has not been ascertained; but it remains fixed, in a state of ignition, at the point of fusion of plate glass.

V. When sodium is exposed to the atmosphere, it immediately tarnishes, and by degrees becomes covered with a white crust of soda, which deliquiates more slowly than that formed on potassium.

VI. It combines with oxygen, slowly and without luminous appearance, at all common temperatures. When heated, the combination becomes more rapid; but no light is emitted till it becomes nearly red hot. The flame, which it then produces is white; and it sends forth bright sparks, exhibiting a very beautiful effect. In common air, it burns with a similar colour to charcoal, but of much greater splendor.

VII. When thrown into water, it produces a violent effervescence and a loud hissing noise; it combines with the oxygen of the water to form soda; and hydrogen gas is evolved, which does not however, as in the case of potassium, hold any of the alkaline base in solution. Neither can sodium be made to dissolve in hydrogen gas, by being heated in contact with it.

When thrown into hot water, the decomposition is more violent, and in this case a few scintillations are generally observed at the surface of the fluid; but this is owing to small particles of the base, which are ejected from the water, sufficiently heated to burn in passing through the atmosphere.

VIII. Its action on alcohol, ether, volatile oils, and acids, is similar to that of potassium; but with nitric acid a vivid inflammation is produced.

IX. Sodium appears to be susceptible of different degrees of oxydation. Thus when it is fused with dry soda, a partition of oxygen takes place between the alkali and the base. A deep brown fluid is produced, which becomes a dark grey solid on cooling. This substance is capable of attracting oxygen from the atmosphere, and of decomposing water, by which it is again converted into soda. The same *oxide* of *sodium* is formed, by fusing this metal in tubes of plate glass.

X. There is scarcely any difference between the visible phenomena attending the action of the base of soda, and that of potash on sulphur, phosphorus, and the metals.

The sulphuret of sodium has a deep grey colour; the phosphuret resembles lead. Added to mercury in the proportion of $\frac{1}{40}$th, it renders that metal a fixed solid of the colour of silver, and the combination is attended with a considerable degree of heat. This amalgam seems, like that of potassium, to form triple compounds with other metals, and even with iron and platina, which remain united with the mercury, when it is deprived of the new metal by the action of air.

The proportions in which this base unites with oxygen to form soda were investigated by the methods already described in the article potassium. Of these the average result is that the base and oxygen exist in soda in the proportions of 7 to 2; hence

$$\text{Soda is composed of} \left\{ \begin{array}{l} 78 \text{ sodium.} \\ 22 \text{ oxygen.} \\ \hline 100 \end{array} \right.$$

◆

SECTION II.

Pure Ammonia.

ART. 1.—*Preparation and Qualities of Ammonia.*

I. Ammonia, in its purest form, subsists in the state of a gas. In order to procure it, one of the following processes may be employed.

(*a*) Mix together equal parts of muriate of ammonia and quicklime, each separately powdered; and introduce them into a small gas bottle or retort. Apply the heat of a lamp; and receive the gas, that is liberated, over mercury.

(*b*) To a saturated solution of ammonia in water, or the pure liquid ammonia, in a gas bottle, apply the heat of a lamp; and collect the gas, as in *a*.

II. This gas has the following properties:

(*a*) It has a strong and very pungent smell.

(*b*) It immediately extinguishes flame; and is fatal to animals. Before, however, a candle is extinguished, by immersion in this gas, the flame is enlarged, by the addition of another, of a pale yellow colour, which descends from the mouth of the jar to the bottom. If the flame of the candle be only in part immersed in the gas, this yellowish flame rises a few lines above the other.

(*c*) It is lighter than atmospheric air. Hence a jar filled with this gas, and placed with its mouth upwards, is soon found to change its contents for common air, which, being heavier, descends, and displaces the ammoniacal gas. One hundred cubic inches of this gas weigh 18.05 grains, barometer 30°, thermometer 55°; or 17.86 grains, the pressure being the same and the thermometer 60°. By the recent experiments of Messrs. Allen and Pepys (Philosophical Transactions, 1808, page 39), undertaken at the desire of Mr. Davy, 100 cubic inches of ammonia weigh 18.18 grains, barometer 30°, thermometer 60°. According to Gay Lussac, its specific gravity is to that of common air as 0.596 to 10; and hence (taking 100 cubic inches of air at 31 grains) 100 cubic inches of ammonia weigh 18.47 grains.

(*d*) It is not sufficiently inflammable to burn when in contact with common air.

(*e*) It is rapidly absorbed by water. A drop or two of water being admitted to a jar of this gas, confined over mercury, the gas will be immediately absorbed, and the mercury will rise, so as to fill the whole of the jar, provided the gas be sufficiently pure. Ice produces the same effect, in a still more remarkable manner. From Mr. Davy's experiments, it appears that 100 grains of water absorb 34 grains of ammoniacal gas, or 190 cubic inches. Therefore a cubic inch of water takes up 475 cubic inches of the gas.

Alcohol, also, absorbs several times its bulk; and affords a solution of ammonia in alcohol, which possesses the strong smell and other properties of the gas.

(*f*) Water, by saturation with this gas, acquires its peculiar smell; and constitutes what has been called liquid ammonia; or, more properly, solution of pure ammonia in water. The method of effecting this impregnation will be described hereafter; and processes will be given for obtaining the solution of ammonia in considerable quantity, which cannot conveniently be accomplished by the process described in experiment *e*. This solution again yields its gas on applying heat. (See I. *b*.)

ART. 2.—*Electrical Analysis of Ammonia.*

(1) Ammoniacal gas is decomposed by electricity. Into a glass tube, having a conductor sealed hermetically into one end (fig. 29) and standing inverted over mercury, pass about one tenth of a cubic inch of ammoniacal gas; and send through it a succession of electrical discharges from a Leyden jar. The arrangement of the apparatus, for this purpose, is shown in fig. 84, pl. ix. and is described in chap. v. sect. 1. When two or three hundred discharges have been transmitted, the gas will be found to have increased to almost twice its original bulk, and to have lost

its property of being absorbed by water. Mix it with a quantity of oxygen gas, equal to between one third and one half of its bulk, and pass an electric spark through the mixture. An explosion will immediately happen; and the quantity of gas will be considerably diminished. Note the amount of the diminution by firing; divide it by 3; and multiply the product by 2. The result shows the quantity of hydrogen gas in the mixed gases which have been generated by electricity; for two measures of hydrogen are saturated by one of oxygen gas.

Suppose, for example, that we expand 10 measures of ammonia to 18; and that, after adding 8 measures of oxygen gas, we find the whole (= 26 measures) reduced by firing to 6 measures; the diminution will be 20. Dividing 20 by 3 we have 6.66, which multiplied by 2 gives 13.32 measures of hydrogen gas from 10 of ammonia. Deducting 13.32 from 18, we have 4.68 for the nitrogen gas contained in the product of electrization. Therefore 10 measures of ammonia have been destroyed, and expanded into

13.32 measures of hydrogen gas.
4.68 ———— nitrogen gas.

According to the above proportions, 100 cubic inches of ammonia, which weigh about 18 grains, if they could be decomposed by electricity, would give about 133 cubic inches of hydrogen weighing 3.5 grains, and 46 of nitrogen weighing 14.4 grains, in all 17.9 grains, or one tenth of a grain less than the ammonia decomposed.

(2) Ammonia may be decomposed by transmitting the gas through a red-hot porcelain tube, placed horizontally in a furnace as represented fig. 40. The products are hydrogen and nitrogen gases.

(3) The decomposition of ammonia may also be easily shown, by galvanizing a saturated solution of ammonia in water. In this experiment a considerable quantity of gas is produced. Expose it over a solution of sulphuret of potash. A small part of it, being oxygen gas, will disappear. The remainder consists of hydrogen and nitrogen gases.

(4) In the Philosophical Transactions for 1809, I have described a property of ammonia, which forms the basis of a very easy and quick mode of analyzing that alkali. When mingled with oxygen gas it may, I find, be inflamed by the electric spark, precisely like a mixture of hydrogen and oxygen gases. To obtain accurate results, however, it is necessary to use less oxygen at first, than is sufficient to saturate the whole hydrogen of the alkali. This is easily calculated. If, for example, we take 10 measures of ammonia, we must use less oxygen than will saturate 13 or 14 measures of hydrogen gas, the quantity which exist in 10 of ammonia; and which require about 7 of oxygen gas. It will be advisable, therefore, not to add above 5 of oxygen. The whole 15 will probably after firing be reduced to about 9. To the remaining gas admit 4 or 5 measures more of oxygen; and on passing the electric spark again, a second explosion will happen, with a diminution of about 6 measures. But in the first explosion, the whole of the oxygen disappears, and it must therefore have saturated a quantity of hydrogen equal to 10 measures; besides which, two thirds of the second diminution $(6 \div 3 \times 2) = 4$ measures are owing to the condensation of hydrogen. Hence the whole hydrogen is $10 + 4 = 14$. The nitrogen, the whole of which exists in the product of the first detonation, is ascertained by deducting from it (*viz.* from 9 in the present instance) the second quantity

of hydrogen (4) which gives 5 for the nitrogen. These numbers may not, perhaps, be exactly obtained by experiment; and they are given merely as a general illustration of the process.

By experiments of this kind, I have determined that 100 measures of ammonia require, for saturating the hydrogen which they contain, between 67 and 68 of pure oxygen gas, and afford

Of hydrogen gas about	136	measures.
nitrogen gas ——	47	measures.
	183	

The results of this analysis furnish a good example of the condensation of the elements of gases which takes place on chemical union; for if we could by any means convert a mixture of 136 measures of hydrogen with 47 of nitrogen into ammonia, the new gas would only fill 100 measures. Simple mixture of these gases, however, even in the same proportions which are obtained by analyzing ammonia, is not sufficient to generate this alkali. The caloric, with which the hydrogen and nitrogen are respectively combined, opposes, by its elasticity, an obstacle to their union, and places them beyond the sphere of their mutual attractions. If these elements are presented to each other when one or both are deprived of part of their caloric, combination then takes place; and the composition of the volatile alkali is proved synthetically, as in the following experiment.

When iron filings, moistened with water, are exposed to nitrogen gas confined over mercury, the gas, after some time has elapsed, acquires the smell of volatile alkali. In this case, the iron decomposes the water and seizes its oxygen; while the hydrogen, at the moment of its liberation, unites with nitrogen and composes ammonia. This state of condensation, or absence of the quantity of caloric necessary to bring it into a gaseous form, has been called the nascent state of hydrogen; and the same term has been applied to the bases of other gases when in a similar state.

ART. 3.—*Decomposition of Ammonia by Potassium.*

Beside the hydrogen and nitrogen which, it has already been stated, are obtained by decomposing ammonia, it has lately been conjectured by Mr. Davy that this alkali contains, also, a small proportion of oxygen, not exceeding seven or eight parts in the hundred. The arguments, which he has brought in favour of this opinion, are derived chiefly from the following facts.

1. When he decomposed ammonia by electricity, the gases produced fell short, by nearly one eleventh, of the weight of the ammonia employed; in other words 100 grains of ammonia gave only about 91 grains of permanent gases. To obtain this result, however, several precautions are necessary, which are fully stated by Mr. Davy in the Philosophical Transactions for 1809, page 460.

2. By repeatedly transmitting ammoniacal gas (previously deprived, by passing it through a tube surrounded by a freezing mixture, of as much water as possible) over red-hot iron wire, the metal became superficially oxidized, and gained a very slight increase of weight. It is proper, however, to state that a similar experiment was made by Mr. Berthollet, junior, with different, or at least with equivocal, results. Besides, a very minute addition of oxygen might be furnished to the iron by the decomposition of a small portion of water, which ammoniacal gas, in common with all other gases, contains; and which would

scarcely be separated from it by the temperature of a freezing mixture.

It must be acknowledged, however, that the indirect evidence, in favour of the presence of oxygen as an element of ammonia, which is furnished by other experiments of Mr. Davy, is much stronger than that derived from the results of its analysis. These experiments even go so far as to suggest that ammonia may, like the fixed alkalis, be an oxide of a peculiar metal, or at least of some compound containing the elements of a metal. And, as hydrogen and nitrogen alone are obtained by the electrical analysis of ammonia, it will follow that the metal in question is either a compound of those two bases, or a component part of one of them. If this should be established, we shall obtain proof of a fact of the greatest novelty and curiosity, *viz.* the existence of a metal or a metallic oxide, whose natural state is that of an aëriform fluid.

To understand the general outline of these experiments, it may be necessary to premise that whenever mercury, after combination with another substance, retains in a great measure its characteristic properties, and forms what has been called an *amalgam*, we infer that the change has been produced by its union with a metal; for the metals are the only bodies which are capable of amalgamating with quicksilver. Now it was found, by M.M. Berzelius and Pontin of Stockholm, that when mercury, negatively electrified in the Voltaic circuit, is placed in contact with solution of ammonia, it gradually expands to four or five times its dimensions, and becomes a soft solid, which, at 70° or 80° Fahrenheit, has the consistence of butter. At the freezing temperature, it becomes firmer, and forms a crystallized mass, in which small shining facets appear. By this combination, it is very remarkable that mercury gains an addition of only about one twelve thousandth part of its weight; and yet has its specific gravity so much increased, that from being between 13 and 14 times heavier than water, it becomes only three times heavier. Its colour, lustre, opacity, and conducting powers remain unimpaired.

When this amalgam is exposed to the atmosphere, oxygen is absorbed; ammonia is reproduced; and the quicksilver is recovered in its metallic form. When thrown into water ammonia is also regenerated, and quicksilver separated, hydrogen gas being at the same time evolved. It appears, then, that in the formation of the amalgam, mercury combines with one or more of the elements of ammonia, and that in the subsequent oxidation of what is thus acquired by the ammonia, consists the process of regenerating alkali. In this view of the subject, there certainly appears great reason to believe that oxygen is one of the constituents of ammonia; but the facts are not sufficiently simple to furnish incontrovertible evidence.

On the supposition that the unknown substance, which amalgamates with the mercury, is of a metallic nature, Mr. Davy has proposed for it the term AMMONIUM. All attempts to detach it from this combination, and to exhibit it in a separate form, have hitherto failed; and it still remains an object of farther investigation.

It was reasonable to expect that oxygen, if contained in ammonia, would be discovered by the action of potassium; which, in that case, would be converted into potash. Accordingly it was found that by heating this new metal in ammoniacal gas, its surface soon became covered with a crust of potash. By continuing the heat, this crust soon disappeared; and the whole of the potassium was converted into a dark olive-coloured substance, and acquired an increase of about one third its original weight. At the same time, pure hydrogen gas was evolved, and constituted the only aëriform remainder. This last fact is peculiarly curious and important.

The olive-coloured substance, thus obtained, was heavier than water; was fusible at a heat a little above 212° Fahrenheit; and, if heated much higher, emitted gas. When immersed in water, ammonia was produced, and bubbles of hydrogen gas were disengaged. If moisture could have been employed in exactly the proper quantity, Mr. Davy believes that the ammonia regenerated would have been exactly equal to that, which disappeared in the first part of the process. By distillation *per se*, the fusible substance gave off a little undecomposed ammonia, and hydrogen and nitrogen gases in the proportion of 2½ measures to 1. The residue of this distillation was a black, shining, opaque, and brittle substance; which was highly inflammable when exposed to air at the ordinary temperature. When added to water, it effervesced violently; ammonia was generated; a little hydrogen evolved; and potash remained. In oxygen gas, it burned vividly; and a quantity of nitrogen resulted, less in bulk than ought to have been obtained from the ammonia decomposed; but no ammonia was regenerated. When submitted to distillation *per se*, in a platina tube (which was done with the expectation that nitrogen gas alone would have been evolved) a gas was obtained, of which one fifth only was nitrogen, and four fifths hydrogen, and which contained no ammonia. In the tube potash remained.

The explanation of these results is attended with considerable difficulty. The fusible substance, formed in the first part of the process, Mr. Davy is disposed to consider as an alloy of potassium with ammonium. When it is acted on by water, both these bodies are oxidated, and are consequently changed, the potassium into potash, and the ammonium into volatile alkali; while the hydrogen of the water is set at liberty. There are some circumstances, however, attending its production, which are not a little perplexing. Ammonia, by electrical analysis, yields both hydrogen and nitrogen gases. Yet by the action of potassium, hydrogen gas alone is evolved, and in quantity nearly equal to what the ammonia contains. What then becomes of the nitrogen? The most obvious conclusion appeared to be, that it must remain combined with the potassium, and would probably be developed by the action of heat. Yet, by distilling the fusible substance, a gas was evolved containing more hydrogen and less nitrogen, than exists in the product of the electrization of ammonia. For example, 11 cubic inches of ammonia give, by electrical decomposition, about 14 cubic inches of hydrogen gas and about 5½ of nitrogen. But by the agency of potassium on the same quantity of ammonia, six cubic inches (omitting fractions) of hydrogen gas were obtained in the first instance; and, by the distillation of the fusible substance, 10 cubic inches more, making in all 16. By these processes, only 3½ cubic inches of nitrogen gas were obtained. It appears then that two cubic inches of hydrogen were evolved, more than existed in the ammonia at the outset; and that there was a loss of two cubic inches of nitrogen. The most probable explanation of the fact is, that a part of the nitrogen was decomposed, into oxygen which united with the potassium, and hydrogen which appeared in a gaseous state.

Thus we acquire new views of the nature of nitrogen, which, if confirmed, will lead us to regard it as an oxide of hydrogen. It will follow also that hydrogen is the only ingredient of ammonia, which is essentially metallic, and capable of amalgamating with mercury; and thus even water itself will be considered as a metallic oxide.

By subsequent experiments on this subject*, Mr. Davy has established the accuracy of his former results. In the distillation of the fusible substance, he has found, moreover, that the proportion of the nitrogen to the hydrogen diminishes in the latter stages of the process. When the fusible substance is distilled with an addition of potassium, the proportion of nitrogen to hydrogen is still farther diminished; and a large quantity of potash is generated. The same result is the more remarkable, in proportion as the heat is quickly raised; while, by a slow and cautious distillation, the whole product of elastic matter is lessened, and the nitrogen bears a greater proportion to the hydrogen. The action, then, of fresh potassium on the fusible substance appears to consist in its decomposing a still farther proportion of nitrogen into hydrogen and oxygen.

Though these facts cannot be easily explained, except on the supposition that nitrogen is an oxide; yet (as is candidly acknowledged by Mr. Davy), in processes so delicate and complicated, and involving such numerous data, we cannot be certain that every source of fallacy has been avoided, and every circumstance observed and reasoned upon.

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* Philosophical Transactions, 1809, page 450.

CHAPTER IX.

EARTHS.

THE term earth was, till lately, employed to denote "a tasteless, inodorous, dry, brittle, and uninflammable substance, not more than five times heavier than water." This definition, however, was rendered imperfect by the discovery, that certain earths have a strong taste, and are readily soluble in water, which yet possess the other characters of earthy bodies. Some of the earths were therefore removed from this class, and arranged among the alkalis. The classification, however, which appears to me most eligible, is that which divides them into *earths* simply, and *alkaline earths;* the latter partaking of the characters both of earths and alkalis. The alkaline earths are Barytes, Strontites, Lime, and Magnesia. The earths are Silex, Alumine, Zircon, Glucine, and Yttria*.

Until the important æra of Mr. Davy's discoveries, the earths were, with respect to the state of our knowledge, simple or elementary bodies. Many conjectures, it is true, had been formed respecting their nature; and, among these, we find that their composition of oxygen and a metallic base had been suggested as a probable theory†.

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* The Agustine of Tromsdorff has been shown, by Berthollet, to be merely Phosphate of Lime. Nicholson's Journal, 8vo. vii. 117

† The reader may consult a history of opinion respecting the earths, in a note to Mr. Davy's paper, Philosophical Transactions, 1808.

Led by the analogy arising from his experiments on the alkalis, Mr. Davy, however, was the first to demonstrate what had before been only imagined; and to disunite, by the agency of strong electrical powers, the constituent principles of several of this class of bodies.

In this part of the investigation, difficulties were encountered which demanded great perseverance and more complicated processes. The affinity of the earthy bases for oxygen appears considerably to surpass that of the metals composing potash and soda; and it was found that simple exposure to the opposite electricities was not adequate to the separation of the principles, which compose the earths; or, at least, that the effect was too indistinct to furnish satisfactory evidence of their nature. Mr. Davy was, therefore, induced to electrify the earths, as he had formerly operated on potash, in contact with the oxides of known metals; with the expectation that the metallic base of the earth would unite with the metal contained in the oxide he employed, and form an alloy. Thus a mixture of barytes and red oxide of mercury might be expected to yield an alloy of mercury with the metallic base of barytes; and such, in fact, was the result of the experiment; for a solid amalgam adhered to the negative wire, which, when thrown into water, evolved hydrogen, leaving pure mercury, and a solution of barytes. Mixtures of lime, strontites, or magnesia with oxide of mercury gave similar amalgams, from which the respective alkalis were regenerated by the action of air or water; but the quantity obtained was too minute for investigation.

On the suggestion of Professor Berzelius of Stockholm, the earths were next electrified negatively in contact with mercury itself; and in this way amalgams were obtained from barytes, strontites, lime, and magnesia. These compounds of mercury with the metallic base of the earths decomposed water, and the earth, which had afforded them, was regenerated. Under naphtha, they might be preserved for a considerable time; but at length they became covered with a white crust.

To procure quantities of these amalgams sufficient for distillation, the earths were slightly moistened, and mixed with one third of red oxide of mercury: the mixture was then placed on a plate of platina, a cavity was made in the upper part of it to receive a globule of mercury of from 50 to 60 grains in weight, and the whole was covered with a thin film of mercury: lastly, the plate was made positive, and the mercury negative, by a proper communication with a battery of 500 pairs.

From these amalgams, the mercury was separated by distilling in small tubes of glass filled with the vapour of naphtha. Considerable difficulties, however, attended these operations, and after all Mr. Davy could, in no case, be absolutely certain, that there was not a small quantity of mercury in combination with the metals of the earths.

The proportion of oxygen and metal has not yet been ascertained in any of the earths; but the evidence from analysis of their composition is perfectly satisfactory, the inflammable base appearing uniformly at the negative surface in the Voltaic circuit, and the oxygen at the positive surface.

The decomposition of the other earths, alumine, silex, zircon, and glucine was not effected by the same means, that had been applied successfully to the alkaline earths. Combinations of potash and alumine, and of potash and silex, were electrified with the hope that the bases of these earths would be obtained in the state of an alloy with potassium. Soda and zircon were similarly treated. In all these cases, the phenomena indicated that some portion of the several earths had been decomposed; but in too minute a quantity to examine the properties of their bases.

Lastly, potassium, amalgamated with about one third its weight of mercury, was electrified negatively under naphtha, in contact with the four earths, which were last enumerated. The potassium generated was thrown into water, and the alkali produced saturated with acetic acid. Now if any metal had thus been obtained from the earths, it would exist in the form of an alloy with potassium; both metals would be oxydized by the water; the potassium would reproduce potash, and the other metal the earth which gave it origin, which earth would be dissolved by the solution of potash, and would reappear on adding an acid. The general tenor of the results gave great reason to conclude that alumine, silex, glucine, and zircon are, like the alkaline earths, metallic oxides; but the evidence it must be confessed, is not altogether unequivocal.

SECTION I.

Barytes.

Barytes may be obtained in a state of purity, by the calcination of its carbonate or nitrate, in a manner which will be hereafter described. (See chap. xi. sect. 2. art. iv.) It exhibits, when pure, the following properties.

I. Barytes, in a pure form, has a sharp caustic taste; changes vegetable blue colours to green; and serves as the intermedium between oil and water. In these respects, it bears a strong resemblance to alkalis.

II. When exposed to the flame of the blow-pipe on charcoal, it melts; boils violently; and forms small globules, which sink into the charcoal. After being kept in fusion in a crucible during ten minutes, it still, according to Berthollet, contains 9 *per cent.* of water. This, however, is true only of barytes which has been obtained from the carbonate, by a process to be described hereafter. Barytes, procured by decomposing the nitrate of that earth, is not fusible, and appears to contain little if any water. (Nicholson's Journal, xxiii. 281.)

III. If a small quantity of water be added to recently prepared barytes, it is absorbed with great rapidity; prodigious heat is excited; and the water is completely solidified, a sort of hard cement being obtained. A little more water converts this mass into a light bulky powder; and, when completely covered with water, the barytes is dissolved. Boiling water should be employed for this purpose; unless sufficient temperature has been produced, by the sudden addition of the whole quantity necessary for solution.

IV. When the solution, prepared with boiling water, is allowed to cool slowly, it shoots into regular crystals. These have the form of flattened hexagonal prisms, having two broad sides, with two intervening narrow ones; and terminated, at each end, by a quadrangular pyramid.

V. The crystals are so soluble, as to be taken up, when heated, merely by their own water of crystallization. When exposed to a stronger heat, they swell, foam, and leave a dry white powder, amounting to about 47 parts from 100 of the crystals. This again combines with water with great heat and violence. At 60° of Fahrenheit, an ounce measure of water dissolves only 25 grains of the crystals, *i. e.* they require for solution, 17½ times their weight of water. Exposed to the atmosphere, they effloresce, and become pulverulent.

VI. When added to spirit of wine, and heated in a spoon over a lamp, they communicate a yellowish colour to its flame.

VII. The specific gravity of this earth, according to Fourcroy, is 4, but Hassenfratz states it at only 2.374. The former account, however, is the more probable. All its combinations have considerable specific gravity; and hence its name is derived, *viz.* from the Greek word ϐαρὺς signifying heavy.

VIII. Barytes does not unite with any of the alkalis.

Of the base of barytes, or barium.—The base of barytes was obtained by Mr. Davy by distilling its amalgam, obtained in the mode, which has already been described in the section on the earths in general.

The residuum of this distillation was a white metal of the colour of silver. At the ordinary temperature of the air it remained a solid; but became fluid at a heat below redness. It did not rise in vapour, till heated nearly to redness, and then acted violently upon the glass.

When exposed to the air, this substance rapidly tarnished, and fell into a white powder, which was barytes. When this process was conducted in a small portion of air, the oxygen was found absorbed; and the nitrogen remained unaltered. A portion of it thrown into water acted upon it with great violence, and sunk to the bottom, producing barytes and evolving hydrogen gas.

The quantities obtained were too minute for an examination either of its physical or chemical qualities. It sunk rapidly in water, and even in sulphuric acid, though surrounded by globules of hydrogen equal to two or three times its volume. Hence it is probable that it cannot be less than four or five times as heavy as water. It was flattened by pressure, but required considerable force for this effect.

SECTION II.

Strontites.

I. STRONTITES (called also Strontia, from Strontian in Scotland, the place where it was first discovered) resembles barytes in many of its properties; and all that is included in the first three paragraphs of the last section may be applied, also, to this earth.

II. Like barytes, strontites is readily soluble in boiling water; and the solution, on cooling, affords regular crystals; but the shape of these differs considerably from that of barytic crystals. The crystals of strontites are thin quadrangular plates; sometimes square, oftener parallelograms; not exceeding in length, and not reaching in breadth, a quarter of an inch. Sometimes their edges are plain, but they oftener consist of two facets, meeting together, and forming an angle like the roof of a house. They adhere to each other in such a manner as to form a thin plate, of an inch or more in length, and half an inch in breadth. Sometimes they assume a cubic form.

III. These crystals undergo, by the action of heat, much the same changes as those of barytes; and leave only about 32 *per cent.* of the dry earth. One part of the crystals requires about 51½ of water at the temperature of 60° for solution, but boiling water takes up half its weight.

IV. Boiling alcohol, with the addition of these crystals, burns with a blood red flame.

V. Strontites does not combine with alkalis. Barytes has no affinity for it; for no precipitation ensues, on mixing the watery solutions of the two earths.

From the preceding enumeration of its characters, it appears that strontites differs from barytes in the form of its crystals, which contain also more combined water, and are less soluble than those of barytes; and also in affording, with alcohol, a flame of different colour. These distinctions were deduced by Dr. Hope, from his excellent series of experiments on the two earths. (See Edinburgh Transactions, vol. iv). Other circumstances of distinction, derived from the properties of their respective salts, will be stated hereafter.

Of the base of strontites or strontium.—We know but little of the base of strontites. All that Mr. Davy observes respecting it is, that it sinks in sulphuric acid, and exhibits the same characters as barium, excepting that it produces strontites by oxidation.

SECTION III.

Lime.

I. *Its external qualities.*—These may be exhibited in common quicklime, such as is employed for the purposes of building or agriculture. In the same state, it is sufficiently pure for demonstrating its chemical properties; but, when used for purposes of the latter kind, it should be fresh burnt from the kiln.

II. *Relation of lime to water.*

(*a*) Lime absorbs water very rapidly with considerable heat and noise. This may be shown by sprinkling a little water on some dry quicklime. The above-mentioned phenomena will take place, and the lime will fall into powder, which has been called *hydrat of lime*. In this compound, the lime is to the water as 23 to 8. Some care, however, is necessary in its preparation, lest more water should be added, than is essential to its constitution. It affords a very convenient form of keeping lime, for occasional use in a laboratory; for the hydrat may safely be preserved in glass bottles, which are almost constantly broken by the earth in its perfectly dry state.

The degree of heat, produced by the combination of lime with water, is sufficient to set fire to some inflammable bodies; and when a large quantity of lime is suddenly slaked in a dark place, even light, according to Pelletier, is sometimes evolved. The caloric, which is thus set at liberty, is doubtless that contained in the water, and essential to its fluidity. By combination with lime, water passes to a solid state, and probably even to a state of much greater solidity than that of ice. Hence, during this change, it evolves more caloric than during conversion into ice; and hence even ice itself, when mixed with quicklime, in the proportion of one to two, enters into a combination which has its temperature raised to 212°. When a sufficient quantity of water has been added to reduce lime into a thin liquid, this is called *milk* or *cream of lime.*

Lime is, in some degree, convertible into vapour by combination with water. When a piece of moistened paper, stained with the juice of the violet, is held in the steam, which arises from lime suddenly slaked, its colour is changed from blue to green. Hence the smell which is perceived during the slaking of lime.

(*b*) Lime absorbs moisture from the atmosphere, and falls gradually into powder.

(*c*) Lime is very sparingly soluble in water, *viz.* in the proportion of about 1 to 500; and, when thus dissolved, forms what has been termed *lime-water*. This solution tastes strongly of lime, turns vegetable blues to green, and unites with oil, forming an imperfect soap. To prepare the solution, lime is to be slaked to a thin paste, and a sufficient quantity of boiling water afterwards added. The mixture is to be stirred repeatedly, the lime allowed to settle, and the clear liquor decanted for use. It must be preserved in closely stopped vessels.

(*d*) Lime does not combine, in any notable proportion, with the alkalis or earths already described.

The base of lime, to which Mr. Davy has given the name of *calcium*, he has never been able to examine, exposed to air or under naphtha. In the case, in which he was able to distil the quicksilver from it to the greatest extent, the tube unfortunately broke whilst warm; and, at the moment that the air entered, the metal, which had the colour and lustre of silver, instantly took fire, and burnt, with an intense white light, into lime.

SECTION IV.

Magnesia.

MAGNESIA possesses the properties of an alkali, but in a considerably less degree than any of the foregoing earths. Its characters are as follow:

I. When perfectly pure, it is entirely destitute of taste and smell.

II. No heat is excited by the affusion of water, and only a very small proportion, not exceeding a 2000dth its weight, of the earth is dissolved. Magnesia appears, however, to have some affinity for water; for when moistened, and afterwards dried, its weight is found increased in the proportion of 118 to 100.

III. Magnesia changes to green the blue colour of the violet; but the watery solution of magnesia, when filtered through paper, does not produce a similar effect. In this respect, it differs from lime.

IV. It is not dissolved by liquid alkalis, nor by alkaline earths; and in the dry way, it has no affinity for barytes or strontites.

The base of magnesia, for which Mr. Davy has proposed the term *magnium,* is but imperfectly known. In the attempts to distil its amalgams, the metal seemed to act upon the glass, even before the whole of the quicksilver was distilled from it. In one experiment, in which Mr. Davy stopped the process before the mercury was entirely driven off, it appeared as a solid, having the same whiteness and lustre as the other metals of the earths. It sunk rapidly in water, though surrounded by globules of gas, and produced magnesia. In the air, it quickly changed, becoming covered with a white crust, and falling into a white powder, which proved to be magnesia.

SECTION V.

Silex.

I. Siliceous earth, or silex, may be obtained tolerably pure from flints by the following process:—Procure some common gun-flints, and calcine them in a crucible in a low red-heat. By this treatment they will become brittle, and easily reducible to powder. Mix them, when pulverized, with three or four times their weight of carbonate of potash, and let the mixture be fused in a strong red-heat, in a crucible. The materials must bear only a small proportion to the capacity of the crucible; and the heat must at first be very moderate, and slowly increased. Even with this precaution, the mass, on entering into fusion, will be apt to overflow; and must be pressed down as it rises, by an iron rod. When this effervescence has ceased, let the heat be considerably raised, so that the materials may be in perfect fusion during half an hour, and pour the melted mass on a copper or iron dish. We shall thus obtain a compound of alkali and siliceous earth. Dissolve this in water, filter the solution, and add to it diluted sulphuric or muriatic acid. An immediate precipitation will ensue, and, as long as this continues, add fresh portions of acid. Let the precipitate subside, pour off the liquor that floats above it, and wash the sediment with hot water, till it comes off tasteless. Then dry it.

Silex, obtained by this process, though pure enough for the following experiments, may still contain a portion of alumine. To separate the latter earth, boil the precipitate with diluted sulphuric acid, to which a little sulphate of potash may be added. The alumine will thus be dissolved; and the silex may be freed from the solution of alum by repeated washings with water. Even silex, however, that has been most carefully washed, still gives traces of potash on the application of electro-chemical powers. (Davy.)

II. Siliceous earth, as thus obtained, has the following qualities:

(*a*) It is perfectly white and tasteless.

(*b*) When mixed with water, it does not form a cohesive mass like alumine, and has a dry and harsh feel to the fingers.

(*c*) It is insoluble in water. Yet, when fresh precipitated, water has the property of retaining in solution about one thousandth of its weight*. That silex, however, is dissolved in water by processes of nature, can scarcely be doubted, when it is considered, that it is found, in considerable quantities, in a crystallized form.

(*d*) It is not acted on by any acid, except the fluoric.

(*e*) When prepared in the foregoing manner, and very minutely divided, silex is taken up by a solution of pure potash, or of soda, but not by ammonia. In the aggregated state of flints, however, it is perfectly insoluble in this way by alkaline solutions, an excellent illustration of the principle that a very minute division of solid bodies, by presenting a greater surface to the action of fluids, facilitates solution.

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* See Klaproth's Contributions, vol. i. page 399, 400.

(*f*) When mixed with an equal weight of carbonate of potash, and exposed to a strong heat in a furnace, it forms a glass, insoluble in water, and identical in all its properties with the glass commonly manufactured. It is owing to the siliceous earth which it contains, that glass is decomposed by the fluoric acid. Glass, however, has occasionally other ingredients, besides the two that have been mentioned.

(*g*) With a larger proportion of alkali, as three or four parts to one of silex, this earth affords a compound called, by Dr. Black, *silicated alkali.* This compound, formed by the process which has been just described, is soluble in water, and affords a good example of the total change of the properties of bodies by chemical union; for, in a separate state, no substance whatever is more difficult of solution than silex. The solution of silicated alkali was formerly termed, *liquor silicum,* or *liquor of flints.* Acids seize the alkali, and precipitate the silex, which is even separated by mere exposure to the atmosphere, in consequence of the absorption of carbonic acid by the alkali.

(*h*) Barytes, or strontites, and silex combine together, in a manner similar to the union of this earth with alkalis; but the combination has not been applied to any useful purpose.

(*i*) When a solution of silex in potash is mingled with one of barytes, of strontites, or of lime in water, or of alumine in alkali, a precipitation ensues. Hence silex may be inferred to have an affinity for all these earths, in the humid way.

Base of silex.—In his attempts to obtain the base of silex, or *silicium,* in a state of separation, Mr. Davy has hitherto been unsuccessful; though the results of his experiments leave little room to doubt that this earth is, like the rest, a metallic oxide.

Silex, he found, when diffused through water, and placed in the electrical circuit, manifested no particular attraction either for the positive or negative pole. Hence he infers that its elements, as to their electrical energies, are nearly in an equilibrium.

Mr. Davy's first attempts to decompose silex were founded on the presumption that it might possibly be a compound of some unknown acid and earth; but the results gave no countenance to this opinion. He next electrified silex and potash united by fusion; and obtained a metal, which, when immersed in water, gave both silex and potash. By fusion with silex, also, potassium seemed to be changed into protoxide; but the results were by no means satisfactory.

Potassium, amalgamated with about one third of mercury, was negatively electrified under naphtha by a strong power, with the expectation of obtaining an alloy of that metal and the base of silex. The potassium, when afterwards thrown into water, generated potash which was combined with silex. It is possible, however, that this might be owing to the combination of silex itself, and not of its base, with potassium. And on the whole it must be confessed, that the evidence of the nature of silex falls far short of that which we possess respecting the alkalis and alkaline earths.

SECTION VI.

Alumine.

I. *Alumine may be obtained* free from other earths, but still combined with carbonic acid, by precipitating a solution of alum in water by the crystallized carbonate of potash. To secure its complete purification from sulphuric acid, Guyton advises that the precipitate be re-dissolved in nitric acid, that nitrate of barytes be cautiously added to the solution, till it no longer occasions milkiness, and that the alumine be afterwards precipitated, or separated from the nitric acid by heat. (*Annales de Chimie,* xxxii. 64.) Electro-chemical analysis, however, in this as in many other instances, shows the imperfection of the common methods of separating bodies from each other; for the most carefully prepared alumine yields the metals of soda and potash, when negatively electrified in contact with mercury. (Davy, Philosophical Transactions, 1808.)

II. Alumine has the following properties:

1. It is destitute of taste and smell.

2. When moistened with water, it forms a cohesive and ductile mass, susceptible of being kneaded into a regular form. It is not soluble in water; but retains a considerable quantity.

3. It does not affect blue vegetable colours.

4. It is dissolved by the liquid fixed alkalis, and is precipitated by acids unchanged. In ammonia, it is very sparingly soluble. It is not soluble in alkaline carbonates.

5. Barytes and strontites combine with alumine, both by fusion and in the humid way. In the first case, the re-

sult is a greenish or bluеish coloured mass. In the second, two compounds are formed. The first, containing an excess of alumine, is in the state of an insoluble powder; the other, having an excess of the alkaline earth, remains in solution. Alumine may be united, by fusion, with the fixed alkalis, and with most of the earths.

6. Alumine, as will be afterward shown, has a strong affinity for colouring matter.

7. Alumine has the property of shrinking considerably in bulk, when exposed to heat, and its contraction is in proportion to the intensity of the heat applied. On this property is founded the *pyrometer of Wedgwood*, which measures high degrees of heat, by the amount of the contraction of regularly shaped pieces of china clay. The pieces of clay are small cylinders, half an inch in diameter, flattened on the under surface, and baked in a low red-heat. The contraction of these pieces is measured, by putting them between two fixed rulers of brass or porcelain, twenty four inches long, half an inch distant from each other at one end, and three tenths of an inch at the other. The rulers are divided into 240 equal parts, called degrees, which commence at the wider end; and each of which is equal to 130° of Fahrenheit. When the clay piece is fixed in its place, before exposure to heat, it is stationary at the first degree, which indicates about 1077° of Fahrenheit. After being strongly heated, in a small case which defends it from the fuel, its bulk is diminished, and it slides down, between the converging rulers, till stopped by their approach. The number on the graduated scale, opposite to the upper end of the piece, indicates the degree of heat to which it has been exposed. In the Appendix, rules may be found for re-

ducing the degrees of Wedgwood's pyrometer to those of Fahrenheit's thermometer.

Almost every thing that has been said respecting the base of silex is true, also, of that of alumine; for Mr. Davy attempted the decomposition of the two earths by much the same processes. All that his results afford, is a strong presumption that alumine is a metallic oxide; but its base, *aluminum*, has not been yet obtained in such a state as to make it a fit object of investigation.

SECTION VII.

Zircon.

I. This earth was discovered by Klaproth in the year 1789, in a precious stone from the island of Ceylon, called Jargon or Zircon; and has since been detected in the hyacinth. It may be obtained by the following process:

Reduce the hyacinth to fine powder, which may be done in an agate mortar, after previously igniting the stone, and plunging it into cold water, to render it brittle. Mix the powder with nine times its weight of pure potash; and project it, by a spoonful at once, into a red-hot crucible, taking care not to add fresh portions till the former ones are melted. When the whole is in fusion, increase the heat for an hour, or an hour and a half. When the crucible has cooled, break it, and detach its contents; reduce them to powder, and boil them with distilled water. Let the insoluble part subside; decant the clear liquor.

and wash the sediment with water, till the washings cease to precipitate muriated barytes. On the residuum pour muriatic acid to excess, and boil it during a quarter of an hour; filter the liquor, and evaporate to dryness in a leaden vessel. Re-dissolve the dry mass; filter again, and precipitate the zircon with carbonate of soda. The carbonate of zircon is thus obtained, from which the carbonic acid may be expelled by calcination.

II. Zircon has the following properties:

1. It has the form of a fine white powder, which has somewhat the harsh feel of silex, when rubbed between the fingers. It is entirely destitute of taste and smell.

2. It is insoluble in water; yet it appears to have some affinity for that fluid, for it retains, when slowly dried after precipitation, one third its weight, and assumes a yellow colour and slight transparency, like that of gum Arabic.

3. It is insoluble in pure liquid alkalis; nor does it even combine with them by fusion; but it is soluble in alkaline carbonates. In the foregoing process, therefore, the carbonate of soda should not be added to excess.

4. Exposed to a strong heat, zircon fuses, assumes a light grey colour; and such hardness, on cooling, as to strike fire with steel, and to scratch glass.

5. Its action on other earths has not been fully investigated.

6. It is precipitated from its solutions in acids by triple prussiate of potash. (Klaproth, ii. 214.)

III. The base of zircon, or *zirconium*, is still unknown to us, though investigated by Mr. Davy in the same manner as the base of silex.

SECTION VIII.

Glucine.

I. THIS earth was discovered by Vauquelin, in the year 1798. He obtained it from the aqua marine or beryl, a precious stone of a green colour, and very considerable hardness, which is found crystallized in Siberia. Glucine has since been detected in the emerald of Peru, and in the gadolinite. The following process may be employed to separate it from the beryl:

Let the stone, reduced to a fine powder, be fused with three times its weight of pure potash. To the fused mass add a quantity of water, and afterwards diluted muriatic acid; which last will effect a complete solution. Evaporate the solution to dryness, re-dissolve the dry mass, and add carbonate of potash so long as any precipitation ensues. Dissolve the precipitate in sulphuric acid, add a little sulphate of potash; and, on evaporation, crystals of alum will be obtained. By this process the alumine is detached. The residuary liquor, which yields no more crystals, contains the glucine, and a small portion of alumine. Add a solution of carbonate of ammonia to excess; this will throw down the alumine, and the glucine will remain dissolved by the superabundant carbonate. When this solution is evaporated to dryness, and moderately heated, the alkaline carbonate is expelled, and a carbonate of glucine remains, in the proportion of 16 parts from every 100 parts of the stone.

II. Glucine has the following properties:

1. It is a fine white and soft powder, resembling alumine in its sensible properties; and, like that earth, adhering to the tongue.

2. It has no action on blue vegetable colours.

3. It does not harden, or contract, like alumine, by heat; and is infusible.

4. It is insoluble in water, but forms with it a ductile paste.

5. It is soluble in liquid potash and soda, but not in the solution of pure ammonia. In these respects it agrees with alumine.

6. Glucine is soluble in carbonate of ammonia; a property distinguishing it from alumine.

7. It appears, like alumine, to have an affinity for colouring matter.

8. With the different acids it forms combinations, which have a sweet and rather astringent taste. Hence its name has been derived from γλυκυς, signifying sweet.

9. It is not precipitated by triple prussiate of potash.

III. We have no knowledge of the base of glucine. When obtained, its proper denomination will be *glucium*.

SECTION IX.

Yttria, or Ittria.

1. This earth was discovered in 1794, by Professor Gadolin, in a stone from Ytterby in Sweden, and its title to the character of a peculiar earth rests, also, on the unquestionable authority of Klaproth and Vauquelin, both of whom have made it the subject of experiment. The following process for obtaining it, is described by Vauquelin in the 36th volume of the *Annales de Chimie*, page 150.

Fuse the pulverized stone (called Gadolinite) in the manner already described, with twice its weight of potash; wash the mass with boiling distilled water, and filter. The filtered solution, which has a beautiful green colour, yields, during evaporation, a black precipitate of oxide of manganese. When this has ceased to appear, allow the liquor to stand; decant the clear part, and saturate with nitric acid. Let the insoluble part be, also, digested with extremely dilute nitric acid, which will take up the soluble earths only, and will leave, undissolved, the silex and oxide of iron. Let the two portions be mingled together, and evaporated to dryness; then re-dissolved and filtered: by which means any remains of silex and oxide of iron are separated. To obtain the yttria from the nitric solution, it would be sufficient, if no other earth were present, to precipitate it by carbonate of ammonia; but small portions of lime, and of oxide of manganese, are still present along with it. The first is separated by a few drops of carbonate of potash; and the manganese, by the cautious addition of hydro-sulphuret of potash. The yttria is then to be precipitated by pure ammonia, washed abundantly with water, and dried. It amounts to about 35 *per cent.* of the stone.

II. Yttria has the following properties:

1. It is perfectly white; but it is difficult to preserve it free from a slight tinge of colour, owing to its contamination with oxide of manganese.

2. It has neither taste nor smell; and it is smooth to the touch, like alumine.

3. It is insoluble in water, and infusible by heat.

4. It is very ponderous; its specific gravity being 4.842.

5. It is not attacked by pure alkalis; and, in this re-

spect it differs from glucine and alumine, both which are abundantly soluble in fixed alkalis.

6. Like glucine, it is soluble in carbonate of ammonia, but five or six times less so than that earth; or, in other words, of equal quantities of glucine and yttria, the latter requires for solution five or six times more of the carbonate of ammonia than the former.

7. It is soluble in most acids; and is precipitated by pure alkalis, by barytes, and by lime.

8. From these solutions it is also precipitated by the oxalic acid, and by oxalate of ammonia, in a state resembling precipitated muriate of silver. Prussiate of potash throws it down in small white grains passing in a short time to pearl grey; phosphate of soda in a white gelatinous form; and infusion of galls in brown flocculi.

9. Yttria, which has been a long time exposed to the action of fire, gives out oxygenized muriatic acid, when dissolved in common muriatic acid; thus manifesting one property of a metallic oxide. (Nicholson's Journal, xviii. 77.)

III. The base of yttria did not form an object of Mr. Davy's investigations.

CHAPTER X.

OF ACIDS IN GENERAL.

THE term *acid* is assigned to all bodies, that possess one or more of the following properties.

1. The acids have a peculiar taste, which is expressed, in common language, by the term sourness. This is very different, as to its degree, in different acids. In some it is so intensely strong, that they cannot be applied to the tongue without producing pain; and it characterizes them, even when diluted with several hundred times their weight of water. The sourness of others is such only, as to render them agreeable to the palate.

2. The acids redden blue vegetable colours; and they possess this property even when very greatly diluted. Hence blue vegetable infusions, or papers stained with them, become tests of the presence of uncombined acids. A single drop of sulphuric acid is capable of reddening a large quantity of water coloured with litmus, or with syrup of violets.

3. The acids combine chemically with alkalis, earths, and metallic oxides; and totally destroy the peculiar qualities of those bodies. Let a few ounce-measures of water be tinged blue with syrup of violets; add a few drops of solution of potash, and the colour will be changed to green; then drop in, very slowly and cautiously, sulphuric acid much diluted, and the blue colour will be restored. At this point, neither the acid nor the alkali is in excess; or they are said to *neutralize* each other. In this state, therefore

one of the most remarkable qualities, both of the acid and alkali, is destroyed; and, on farther examination, it will be found that all the other characteristics of the components are no longer apparent in the compound.

It is not necessary, however, in order to entitle a body to rank among the acids, that it should possess all the qualities which have been enumerated. The prussic acid, for example, is neither sour to the taste, nor does it redden blue vegetable colours; but yet, from its properties of chemical combination, enumerated under the third head, it is arranged among the acids. Other bodies, again, are excluded from this class, which possess, partly, the characters of acids. Thus sulphuretted hydrogen changes the blue colours of vegetables to red; and combines chemically with alkalis and earths.

All the acids, which have hitherto been decomposed, agree in containing oxygen, which has been considered as the general principle of acidity. The universality of this inference, however, has been limited by Mr. Davy's discovery that alkalis, also, have oxygen for one of their elements. Notwithstanding this limitation, it is of importance to have ascertained the fact; because, as will afterwards appear, many of the most remarkable qualities of the acids depend on their containing oxygen. The peculiar properties of each acid are derived from the combustible base, with which the oxygen is united; and so, also, is its denomination. Thus sulphur, when oxygenated, affords sulphuric acid; carbon carbonic acid; and so of the rest. But the same combustible base admits of being combined with different proportions of oxygen; and the compounds, thus generated, are distinguished by a very different train of qualities. Sulphur, for example, when combined with the full proportion of oxygen, which it is capable of uniting with, affords a very dense and corrosive acid called the sulphur*ic*; when oxygenated in a less degree, it yields a penetrating and suffocating gas called the sulphur*ous* acid. By these two terminations, the degrees of oxygenation are distinguished. Thus we have the phosphor*ic* and phosphor*ous* acids, the nitr*ic* and nitr*ous;* the termination *ic* denoting an acid with its full proportion of oxygen. In some cases, a combustible base, which affords an acid when fully oxygenated, constitutes only an oxide when combined with a less quantity of oxygen. Carbon, for example, affords carbonous oxide and carbonic acid, but, so far as is hitherto known, no intermediate product.

The following table exhibits the compounds, which result from the oxygenation of some of the principle combustible bases.

100 parts of	united with	Result.
Carbon . .	. . 260 oxygen . .	carbonic acid.
	. . 132½ —— . .	carbonous oxide.
Sulphur . .	. . 130 oxygen . .	sulphuric acid.
	. . 92 —— . .	sulphurous acid.
Nitrogen .	. . 227 oxygen . .	nitric acid.
	. . 190 —— . .	nitrous acid.
	. . 114 —— . .	nitric oxide.
	. . 57 —— . .	nitrous oxide.
Phosphorus	. . 114 oxygen . .	phosphoric acid.
	. . 76 —— . .	phosphorous acid.
Muriatic acid	. . 30 oxygen . .	oxy-muriatic acid.

The composition of muriatic acid is still unknown. That of fluoric and boracic acids, it will appear from the sequel, has been determined in this country by Mr. Davy, and in France by Thenard and Gay Lussac. But we are not yet fully acquainted with the properties of their bases; nor are we able to assign the proportions in which these bases unite with oxygen.

It is in consequence of the oxygen, which they contain, that several of the acids are decomposed by inflammable substances, with the disengagement of intense heat and light; or that the acids are (as they have been well termed by Dr. Thomson) *supporters of combustion.* This property belongs most remarkably to the acids, in which oxygen is most weakly combined. Thus the nitric and oxy-muriatic acids retain their oxygen so feebly, that many inflammable substances, when merely introduced into them at the ordinary temperature of the atmosphere, take fire and burn with vehemence. The oxy-muriatic acid even supports the combustion of many bodies, which are not inflammable in the atmosphere, such as gold and other perfect metals.

Every acid, with a few exceptions, is capable of uniting with each individual of the classes of alkalis, earths, and metallic oxides. In these compounds, the separate qualities of the component principles are in many instances no longer apparent, and hence they have been called *neutral salts.* In every salt, then, there are present two distinct ingredients. The acid, of whatever kind it may be, has been denominated, by Lavoisier, the *salifying principle;* and the body, with which the acid is combined, whether an alkali, an earth, or an oxide of any of the common metals, the *salifiable base*, or simply *the base.* The salts, formed by any individual acid with all these different bases, may be considered as a genus or class; and may be distinguished by a generic name, expressive, in part, of their composition. This general name is taken from that of the acid. The combination of sulphuric acid, for instance, with any base, is called a *sulphat* or *sulphate;* of phosphoric acid a *phosphate;* and so of the rest. The name of the individual salt is derived from that of the base. Thus we have the *sulphat of potash*, the *sulphat of soda*, &c. But sulphur, phosphorus, and other bodies, it has already been observed, are susceptible of different degrees or stages of oxygenation; and afford, in these different stages, acids which are characterized by a peculiar train of properties. The compounds, also, which result from the union of two different acids, having the same combustible base, with alkalis and earths, are altogether different from each other. The salt, for example, which sulphur*ic* acid affords with potash, is wholly unlike that which results from the combination of sulphur*ous* acid with the same base. It was necessary, therefore, to distinguish the compounds of the more oxygenated from those of the less oxygenated acid; and this has been done by changing the termination from *at* or *ate* to *ite*. Thus the salts, formed with sulphurous and phosphorous acid, are called sulphites and phosphites; as sulphite of potash, phosphite of soda, &c. To this rule there is the exception of the compounds of muriatic acid. No inferior state of oxygenation being known with respect to this acid, its compounds are termed muriates; and those of the same acid with an excess of oxygen, oxygenized muriates, or hyper-oxygenized muriates, or (as is now most usual for the sake of brevity) oxy-muriates.

Hitherto, however, we have considered the compounds of acids with their respective bases only in the state of neutral compounds, in which neither the acid, nor the base predominates. But we have several instances, in which a neutral compound is susceptible of uniting with an additional quantity of acid or of base, and thus of acquiring an entirely new set of properties. Common cream of tartar, for example, is a salt in which acid properties are rendered apparent by its taste and action on vegetable colours, and which requires ten times more water for solution than the neutral compound of the same acid and base. To distinguish this and other similar salts, the epithet *acidulous* was first proposed; but, for the sake of brevity, it has now become customary to prefix the Latin preposition *super*. Thus we have the *super*-tartrate of potash; the super-sulphate of potash; &c. On the contrary, when the base is predominant, we denote the deficiency of acid by the preposition *sub*, as sub-carbonate of potash, sub-borate of soda, &c. If it should, hereafter, be proved (which is extremely probable) that, consistently with the views of Mr. Dalton, bodies unite atom to atom singly, or if either is in excess that it exceeds by a ratio to be expressed by some simple multiple of the number of its atoms, we may then adopt the nomenclature proposed by Dr. Wollaston. Thus he has suggested the terms binoxalate and quadroxalate of potash, to denote the compounds of oxalic acid; the proportions in the former being supposed to be two particles of potash to two of acid, and in the latter two of potash to four of acid.

There are a few instances of salts with compound bases; and in cases of this kind it is customary to annex to the generic name those of both the bases. Thus, for example, we have the tartrate of potash and soda, the phosphate of ammonia and magnesia, or as it sometimes called ammoniaco-magnesian phosphate.

In no part of chemistry is the advantage of the new nomenclature more strongly experienced, than in the class of neutral salts; for the number of these compounds is susceptible of being multiplied to an immense extent. If the knowledge of the name did not lead to that of the compound, no powers of memory would be adequate to retain them. But by changing the arbitrary titles, formerly assigned to them, for names expressive of their composition, we are furnished with a kind of artificial memory, which renders their recollection perfectly easy. Thus for the terms butter of antimony, sugar of lead, and Glauber's salt, are now substituted the more appropriate epithets of muriate of antimony, super-acetate of lead, and sulphate of soda.

Of those acids, which are supporters of combustion, a few retain the same property even in combination. Nitrate of potash, it is well known, enters into active inflammation with charcoal, sulphur, and other combustible bodies. This is owing to the quantity of oxygen which the nitric acid contains, and which is less strongly attracted by the nitrogen than by the newly added body.

Other properties, general to the class of salts, have already been described in the section on cohesion; especially their solubility, and their crystallization. I proceed, therefore, to the history of the different acids, and of the compounds which they yield with the several alkaline and earthy bases. Under each head, I shall first enumerate the properties of the base of the acid; and its combinations with such other combustible bodies, as may already have been introduced to the reader's notice.

CHAPTER XI.

CARBONIC ACID AND ITS BASE.—CARBONATES.—BINARY COMPOUNDS OF CARBON.

THE bodies, which form the subject of this chapter, will be described in the following order.

I. CARBON, and its various modifications.

II. The compound of carbon and oxygen, in its highest stage of oxygenation, constituting CARBONIC ACID; and, connected with it, the class of salts called CARBONATES.

III. The OXIDE OF CARBON, or CARBONOUS OXIDE, a compound containing less oxygen than carbonic acid.

IV. The various combinations of carbon and hydrogen, termed CARBURETTED HYDROGEN.

SECTION I.

Carbon and Charcoal.

IT had long been admitted as an established truth, chiefly on the evidence of the experiments of Guyton (*Annales de Chimie*, xxxi.), that the diamond is the only form of pure carbon; and that charcoal is a compound of carbon and oxygen, or an oxide of carbon. The important experiments of Messrs. Allen and Pepys have suggested, however, that the diamond and charcoal, though so widely remote from each other in external characters, are, as to their chemical nature, identically the same; and that the difference between them, in all probability, results merely from the respective states of aggregation of their particles.

Some doubt, it must be confessed, has been thrown on this conclusion by the recent experiments of Mr. Davy, who, by the powerful agency of galvanic electricity, has discovered both in the diamond and in charcoal very minute quantities of foreign matter.

Small pieces of willow charcoal, which had been intensely ignited, were subjected by Mr. Davy to the action of a powerful Voltaic battery, in a Toricellian vacuum; every precaution being taken to exclude moisture, both from the mercury and the charcoal. A flame issued from it of a most brilliant purple colour; and formed, as it were, a conducting chain of light of nearly an inch in length. At the same time, elastic matter was rapidly formed, some of which was permanent, and in the proportion of a cubic inch from eight grains of charcoal. It was inflammable when mingled with oxygen gas, and gave carbonic acid; and hence contained both hydrogen and carbon.

Charcoal, heated with potassium, seemed to combine with it; but there was no evidence that the potassium had acquired the smallest quantity of oxygen by this treatment; for it evolved from water as much hydrogen gas, as an equal weight of pure potassium.

This was not the case, however, with the diamond. When heated with potassium, it lost somewhat of its weight; and became covered with a greyish crust, which, when separated by washing, appeared as a fine powder of a black colour. The potassium seemed to be in some degree oxidized; for it gave a diminished product of hydrogen gas, when made to act on water.

From these experiments it appears that charcoal contains a little hydrogen; and the diamond a very minute proportion of oxygen, both however, too trifling to be discovered by any difference in the results of their combustion.

No remarkable effects were obtained by igniting charcoal intensely in contact with nitrogen gas. Carbonic acid was not formed, and all that took place was the evolution of a portion of carburetted hydrogen.

To obtain charcoal free from contamination, pieces of oak, willow, hazle, or other woods, deprived of the bark, may be buried in sand in a crucible, which is to be exposed, covered, to the strongest heat of a wind-furnace.

For purposes of accuracy, charcoal must be used when recently prepared. and before it has had time to become cold; or if it cannot be had fresh made, it must be heated again to redness under sand in a crucible.

From 100 parts of each of the following woods, Messrs. Allen and Pepys obtained the annexed quantities of charcoal; *viz.* from fir, 18.17; lignum vitæ, 17.25; box, 20.25; beech, 15; oak, 17.40; mahogany, 15.75.

Charcoal has the following properties:

1. In its aggregated state it is black, perfectly insipid, and free from smell; insoluble in water, brittle, and easily pulverized. In close vessels, and entirely secured from contact with air, it is unchanged by any degree of heat. A gas, however, may be collected from it by distillation, which consists of hydrogen and carbon, and perhaps a little oxygen. Berthollet has found, also, in the aëriform products of its distillation, a considerable proportion of nitrogen. (*Mémoires d'Arcueil*, ii. 484.)

2. Charcoal has the singular property of absorbing gases,

without alteration. Fill a jar with common air, or any other gas, and place it over dry mercury: take a piece of charcoal, red-hot from the fire, and plunge it in the mercury of the bath: when cold, let it be passed into the vessel of gas, without bringing it into contact with the atmosphere. A considerable diminution of the gas will be speedily effected.

Count Morozzo has given the following table of the quantities of different gases absorbed, in the foregoing manner, by charcoal. In each experiment, he employed a piece of that substance 1 inch long and ¾ of an inch diameter. The receiver containing gas was 12 inches long and 1 inch diameter.

Gas absorbed.	*Inches.*	*Gas absorbed.*	*Inches.*
Atmospheric	$3\frac{1}{2}$	Nitrous	$6\frac{5}{6}$
Carbonic acid . . .	11	Hydrogen	$2\frac{1}{11}$
Ammonia	11	Oxygen	$2\frac{1}{6}$
Muriatic acid . . .	11	Sulphurous acid . .	$5\frac{1}{2}$
Sulphuretted hydrog. .	11		

3. From the experiments of Rouppe (*Annales de Chimie*, xxxii. 1), it appears, that if charcoal, which has imbibed oxygen gas, be brought into contact with hydrogen gas, water is generated.

4. Charcoal, by long exposure to the atmosphere, absorbs one twentieth of its weight, three fourths of which are water. (Clement and Desormes).

The charcoal of different woods, Messrs. Allen and Pepys found to increase very differently in weight; that from fir gaining, by a week's exposure, 13 *per cent.*; that from lignum vitæ, in the same time, 9.6; from box, 14; beech, 16.3; oak, 16.5; mahogany, 18. The absorption goes on most rapidly during the first 24 hours; and by much the largest part of what is absorbed consists of water merely.

5. Charcoal resists the putrefaction of animal substances. A piece of flesh-meat, which has begun to be tainted, may have its sweetness restored by rubbing it daily with powdered charcoal; and may be preserved sweet for some time by burying it in powdered charcoal, which is to be renewed daily. Putrid water is also restored by the application of the same substance; and water may be kept unchanged at sea, by charring the inner surface of the casks which are used to contain it. (Lovitz, *Annales de Chimie*, xiv.) It produces, also, a remarkable effect in destroying the taste, odour, and colour of many vegetable and animal substances. Common vinegar, by being boiled on it, is rendered perfectly limpid. Rum and other varieties of ardent spirit, which are distinguished by peculiar colours and flavours, lose both by maceration with powdered charcoal. The colour of litmus, indigo, and other pigments, dissolved or suspended in water, is destroyed. Putrid animal fluids, also, are completely deprived of their odour.

6. Charcoal is a very slow conductor of caloric. The experiments of Guyton have determined, that caloric is conveyed through charcoal more slowly than through sand. in the proportion of three to two. Hence powdered charcoal may be advantageously employed to surround substances which are to be kept cool in a warm atmosphere; and also to confine the caloric of heated bodies.

SECTION II.

Combustion of Carbon.

If a small piece of charcoal be exposed red-hot to the common atmospheric air, it exhibits scarcely any signs of combustion, and soon becomes cold. But if a similar piece, heated to about 802° Fahrenheit or nearly to redness, be introduced into a receiver filled with oxygen gas, it continues to burn with greatly increased splendour, and with bright scintillations. If the charcoal be pure, and its proportion rightly adjusted, it is entirely consumed. When the quantity burnt is considerable, a manifest production of water takes place, and the inner surface of the glass vessel becomes covered with moisture, which disappears, however, on standing. This portion of water owes its origin to the union of oxygen with the hydrogen which, it appears from Mr. Davy's experiments, and from the results of its distillation, all charcoal contains.

The diamond, also, which was formerly considered as an incombustible substance, may be consumed by a sufficiently intense heat, even in atmospherical air. The Florentine academicians, in the year 1694, appear first to have ascertained this fact, by exposing diamonds to the focus of a powerful burning lens. Their experiment has been repeated by subsequent chemists, with various modifications. It has been found by Sir George Mackenzie that diamonds burn, when exposed on a muffle, to the temperature of about 14° Wedgwood. In oxygen gas they are consumed more rapidly, their surface first becoming black. By effecting their combustion in this way, Guyton thought he had determined that the diamond, in an equal weight, contains more real carbon than common charcoal. His experiments, however, have not been confirmed by those of subsequent chemists. One fact, however, has been contributed on this subject by Guyton, which is of considerable value. The diamond, he first ascertained, is destroyed when thrown into red-hot and melted nitre; and this property, it will afterwards appear, has been successfully applied by Mr. Tennant to the determination of the nature of the diamond, and of the proportion of ingredients in carbonic acid.

To collect the products of the combustion of carbon, requires rather a complicated apparatus. Lavoisier burnt charcoal in a known quantity of oxygen gas, which was confined by mercury, setting fire to the charcoal by a bent iron wire heated to redness. (Elements of Chemistry, pl. iv. fig. 3.) Messrs. Allen and Pepys collected the products of the combustion of charcoal and of the diamond, by burning them separately in a platina tube, set horizontally in a charcoal furnace, and connected, at each extremity, with a mercurial gazometer. An idea of this arrangement will best be obtained by imagining that to each end of the tube *c*, fig. 40, the pipe *b* of a gazometer, like that shown, fig. 35, is connected. At the outset of the experiment, one of the gazometers was filled with a known quantity of the purest oxygen gas, and the other was empty. The tube was then made red-hot; and the gas, being forced alternately from one gazometer to the other, was repeatedly brought into contact with the red-hot charcoal or diamond. The volume of the gas was found to be entirely unaltered; but it had received an addition to its weight, precisely equal to what the charcoal or diamond, on weighing, was ascertained to have lost; and it was partly converted into a gas, totally different in its properties from oxygen gas, and called carbonic acid. It appears, therefore, that oxygen gas, by conversion into carbonic acid, undergoes neither expansion nor condensation.

SECTION III.

Carbonic Acid.

From the quantity of charcoal or diamond consumed, and the quantity of oxygen converted into carbonic acid, it was easy to infer the proportion of carbon and oxygen in the new compound. Reducing these to centesimal proportion, for every 28 or 29 grains of the combustible base which had disappeared, 100 grains of carbonic acid (=about 201 cubic inches) were generated; and it is remarkable that these proportions agree exactly with those originally stated by Lavoisier. The same quantity of carbonic acid resulted, also, from the combustion of between 28 and 29 grains of diamond. Hence it may be inferred, that the actual quantity of carbon in equal weights of diamond and charcoal is precisely the same; and that charcoal is not, as has hitherto been supposed, an oxide of carbon. If this inference required confirmation, it is furnished by its agreement with Mr. Tennant's experiments on the combustion of the diamond, published in 1797. Two grains and a half of diamond, this philosopher found, when consumed in a tube of gold by means of nitre, gave nine grains of carbonic acid, which, in 100 parts should contain, therefore, as nearly as possible, 28 parts of diamond or carbon. The mean of a number of

Messrs. Pepys and Allen's experiments gave the following statement of the composition of carbonic acid:

100 parts by weight consist of . . .	28.60 carbon 71.40 oxygen.
	100

It is remarkable, also, that these numbers are precisely those which result from the experiments of Clement and Desormes. (*Annales de Chimie*, xxxix. 42.)

In addition to the proofs of the constitution of carbonic acid, derived from its synthesis, we have also the evidence of its analysis, which may be effected by several processes.

1. By passing a succession of electrical discharges through a quantity of carbonic acid gas confined over mercury, I have found that the gas is separated into oxygen, and a gas called carbonous oxide, which consists of oxygen united with a larger proportion of carbon, than exists in carbonic acid. When the carbonic acid, which escapes decomposition, is washed out by solution of potash, an electric spark inflames the residuary mixture; the oxygen and carbonous oxide again uniting, and re-composing carbonic acid. (Philosophical Transactions, 1809, page 448.)

2. When a mixture of carbonic acid and hydrogen gases is electrified, the hydrogen combines with part of the oxygen in the acid, and reduces it to the state of carbonous oxide.

3. When potassium is heated in carbonic acid gas, Mr. Davy has found that it inflames; part of it is oxidated at the expence of the acid; and part of it unites with the charcoal, which is precipitated.

4. By simply heating phosphorus in carbonic acid gas, no step is made towards the decomposition of the latter.

But by applying phosphorus to some of the combinations of carbonic acid, the phosphorus is oxygenated, and carbon appears in a separate form. The original discovery of this fact is due to Mr. Tennant (Philosophical Transactions, 1791, page 182), and the details of the experiment have been ably followed up by Dr. Pearson (Philosophical Transactions, 1792, page 289).

To exhibit this fact, provide a tube of very thin glass about one third of an inch wide, and 18 or 20 inches long, sealed at one end. Coat it, within about an inch of the sealed extremity, with a lute of sand and clay; and when this is dry, put into it as much purified phosphorus, in small pieces, as will fill the uncoated part. Then cover the phosphorus with carbonate of lime, or carbonate of soda which has been deprived of its water of crystallization. Let the part of the tube, which contains the carbonate, be made red-hot by means of a portable furnace; and, at this moment, apply heat to the part containing the phosphorus, sufficient to melt and raise it into vapour. The vapour of the phosphorus, coming into contact with the red-hot carbonate, will decompose the carbonic acid; and charcoal will be found in the residue of the process, in the form of a very light and black powder.

To procure carbonic acid, sufficiently pure for the exhibition of its properties, the combustion of charcoal is far from being the best process. The student may, therefore, have recourse to another, the rationale of which he will not, at present, understand; but which will be explained afterwards. Into a common gas bottle, put a little powdered marble or chalk, and pour on this sulphuric acid, diluted with five or six times its weight of water. A gas will be produced, which those, who have an opportunity, may receive over mercury; but a mercurial apparatus is not absolutely essential, since the gas may be collected over water, if used immediately when procured. Carbonic acid may, also, be separated by heat alone, from carbonate of lime. For this purpose, coarsely powdered chalk or marble may be put into the iron vessel *a*, fig. 85, which may be connected, by means of the conducting pipe *b*, with a gazometer. The receiving cylinder of the latter, after a sufficiently long continuance of heat to the bottle *a*, will be filled with carbonic acid gas. Its properties are the following:

(*a*) *It extinguishes flame.*—Set a vessel, filled with the gas, with its mouth upwards, and let down a lighted candle. The candle will instantly be extinguished.

A person, says Dr. Priestley, who is quite a stranger to the properties of this kind of gas, will be agreeably amused with extinguishing lighted candles, or blazing chips of wood, on its surface. For the smoke readily unites with this kind of air; so that little or none of it escapes into the atmosphere. It is remarkable, that the upper surface of this smoke, floating in the fixed air, is smooth and well defined; whereas the lower surface is exceedingly ragged, several parts hanging down to a considerable distance within the body of the carbonic acid, and sometimes in the form of balls, connected to the upper stratum by slender threads, as if they were suspended. The smoke is also apt to form itself into broad flakes, exactly like clouds. Making an agitation in this air, the surface of it (which still continues exactly defined) is thrown into the form of waves; and if, by this agitation, any of the carbonic acid be thrown over the sides of the vessel, the smoke, which is mixed with it, will fall to the ground, as if it was so much water.

(*b*) *It is fatal to animals.*—Put a mouse, or other small animal, into a vessel of the gas, and cover the vessel, to prevent the contact of common air. The animal will die in the course of a minute or two.

By means of this gas, butterflies, and other insects, the colours of which it is desirable to preserve, for the purpose of cabinet specimens, may be suffocated better than by the common mode of killing them with the fumes of sulphur.

(*c*) *This gas is heavier than common air.*—According to Mr. Davy, 100 cubic inches, at 55° Fahrenheit, and 30 inches of the barometer, weigh 47.5 grains; and at 60°, with the same pressure, would weigh 47.11. Messrs. Allen and Pepys have lately determined that 100 cubic inches, at 60° Fahrenheit, and 30 inches barometer, weigh 47.26 grains. Its specific gravity, according to Biot and Arajo, is 1.5196; and hence, if 100 cubic inches of atmospheric air weigh 31 grains, the same bulk of carbonic acid gas should weigh 47.1 grains.

To show the superior specific gravity of this gas in a general way, the following experiment will be sufficient. Let a long glass tube, proceeding from a gas bottle containing powdered marble and dilute sulphuric acid, be twice bent at right angles; let the open end of the longer leg reach the bottom of a glass jar, perfectly dry within, and standing with its mouth uppermost. The carbonic acid will expel the common air from the jar, because it is heavier.—This superior gravity may be farther shown as follows: When the jar is perfectly filled with the gas (which may be known by a lighted candle being instantly extinguished when let down into it), take another jar, of rather smaller size, and place at the bottom of it a lighted taper, supported by a stand: then pour the contents of the first-mentioned jar into the second, as if you were pouring in water. The candle will be instantly extinguished, as effectually as if it had been immersed in water.

It is owing to its superior gravity, that carbonic acid gas is often found at the bottom of deep wells and of mines, the upper part of which is entirely free from it. Hence the precaution, used by the sinkers of wells, of letting down a candle before they venture to descend in person.

(*d*) *Carbonic acid gas is absorbed by water.*—Fill partly a jar with this gas, and let it stand a few hours over water. An absorption will gradually go on, till at last none will remain. This absorption is infinitely quicker when agitation is used. Repeat the above experiment, with this difference, that the jar must be shaken strongly. A very rapid diminution will now take place. In this manner water may be charged with rather more than its own bulk of carbonic acid gas; and it acquires, when thus saturated, a very brisk and pleasant taste. This impregnation is most commodiously effected by an apparatus, sold in the glass shops, under the name of Nooth's machine.

The influence of pressure, in occasioning water to absorb a larger quantity of carbonic acid, may be illustrated by an apparatus, which I have described in the Philosophical Transactions for 1803, but which cannot be understood without the engraving that accompanies it. From a long series of experiments with this apparatus, I have deduced, as a general law, that water takes up the same volume of compressed carbonic acid gas, as of gas under ordinary pressure. And since the space, occupied by any gas, is inversely as the compressing force, it follows that the quantity of gas, forced into water, is directly as the pressure. Thus, if water, under common circumstances, takes up an equal bulk of carbonic acid, under the pressure of two atmospheres, it will absorb twice its bulk; under three atmospheres three times its bulk, and so on.

(*e*) *From water, thus impregnated, carbonic acid is again set at liberty, on boiling the water, or by exposing it under the receiver of an air-pump.*—During exhaustion, the gas will escape so rapidly, as to present the appearance of ebullition; and will be much more remarkable than the discharge of air from a jar full of common spring water, confined, at the same time, under the receiver, as a standard of comparison.

(*f*) *Carbonic acid is expelled from water by freezing.*—If the impregnated water be rapidly congealed, by surrounding it with a mixture of snow and salt, the frozen water has more the appearance of snow than of ice, its bulk being prodigiously increased by the immense number of air bubbles. When water, thus congealed, is liquefied again, it is found, by its taste, and other properties, to have lost nearly the whole of its carbonic acid.

(*g*) *Carbonic acid gas, when combined with water, reddens vegetable blue colours.*—This may be shown by dipping into water, thus impregnated, a bit of litmus paper, or by mixing, with a portion of it, about an equal bulk of the infusion of litmus. This fact establishes the title of the gas to be ranked among acids. When an infusion of litmus, which has been thus reddened, is either heated, or exposed to the air, its blue colour is restored, in consequence of the escape of the carbonic acid. This is a marked ground of distinction from most other acids, the effect of which is permanent, even after boiling.

(*h*) *Carbonic acid gas precipitates lime water.*—This character of the gas is necessary to be known, because it affords a ready test of the presence of carbonic acid whenever it is suspected. Pass the gas, as it proceeds from the materials, through a portion of lime water. This, though perfectly transparent before, will instantly grow milky:—Or, mix equal measures of water saturated with carbonic acid and lime water. The same precipitation will ensue. By means of lime water, the whole of any quantity of carbonic acid existing in a mixture of gases, may be entirely removed.

(*i*) *By the application of the test (h), it will be found, that carbonic acid is generated in several cases of combustion.*—1. Let the chimney of a small portable furnace, in which charcoal is burning, terminate, at a distance sufficiently remote to allow of its being kept cool, in the bottom of a barrel provided with a moveable top, or of a large glass vessel having two openings. A small jar of lime water being let down into the tub or vessel, and agitated, the lime water will immediately become milky. The gas will also extinguish burning bodies, and prove fatal to animals that are confined in it. Hence the danger of exposure to the fumes of charcoal, which, in several instances, have been known to be fatal. These fumes consist of a mixture of carbonic acid and nitrogen gases with a very small proportion of oxygen gas. 2. Fill the pneumato-chemical trough with lime water, and burn a candle, in a jar filled with atmospheric air, over the lime water till the flame is extinguished. On agitating the jar, the lime water will become milky. The same appearances will take place, more speedily and remarkably, if oxygen gas be substituted for common air. The carbonic acid, thus formed during combustion, by its admixture with the residuary air, renders it more unfit for supporting flame, than it otherwise would be from the mere loss of oxygen. Hence, if a candle be burnt in oxygen gas, it is extinguished long before the oxygen is totally absorbed, because the admixture of carbonic acid with oxygen gas, in considerable proportion, unfits it for supporting combustion. Whenever any substance, by combustion in oxygen gas or common air over lime water, gives a precipitate, soluble with effervescence in muriatic acid, we may confidently infer that it contains carbon.

(*k*) *The respiration of animals is another source of carbonic acid.*—On confining an animal, in a given portion of atmospheric air, over lime water, this production of carbonic acid is evinced by a precipitation. The same effect is also produced more remarkably in oxygen gas. The production of carbonic acid, by respiration, may be proved, also, by blowing the air from the lungs, with the aid of a quill, through lime water, which will immediately grow milky. The carbonic acid, thus added to the air, unfits it for supporting life, not merely by diminishing the proportion of oxygen gas, but apparently by exerting a positively noxious effect. Hence a given quantity of air will support an animal much longer, when the carbonic acid is removed as fast as it is formed, than when suffered to remain in a state of mixture. It has been found, that an atmosphere, consisting of oxygen gas and carbonic acid, is fatal to animals, though it should contain a larger proportion of oxygen than the air we commonly breathe.

(*l*) *Carbonic acid retards the putrefaction of animal substances.*—This may be proved, by suspending two equal pieces of flesh meat, the one in common air, the other in carbonic acid gas, or in a vessel through which a stream of carbonic acid is constantly passing. The latter will be preserved untainted some time after the other has begun to putrefy.

(*m*) *Carbonic acid gas exerts powerful effects on living vegetables.*—These effects, however, vary according to the mode of its application.

Water, saturated with this gas, proves highly nutritive, when applied to the roots of plants. The carbonic acid is decomposed, its carbon forming a component part of the vegetable, and its oxygen being liberated in a gaseous form.

On the contrary, carbonic acid, applied as an atmosphere, by confining a living vegetable in the undiluted gas over water, is injurious to the health of the plant, especially in the shade. M. Saussure, *jun.* found, that a proportion of carbonic acid in common air, greater than one eighth, is always injurious to vegetation; but that in this proportion it promotes the growth of plants, and is manifestly decomposed.

Carbonic acid is susceptible of combination with alkalis, earths, and metals, and forms an order of compounds, termed *carbonates*. At present, however, we shall only attend to the results of its union with alkalis, and earths. In the compounds of carbonic acid with these bases, and especially with the alkalis, the qualities of the base still predominate. The alkaline carbonates, for example, retain the taste, though in a less degree, which characterizes their bases; and change blue vegetable colours to green. Ammonia, also, preserves in a great measure its odour and volatility. By combination with the earths, however, carbonic acid produces a more perfect neutralization of their properties.

SECTION IV.

Carbonates.

Art. 1.—*Sub-carbonate, and Carbonate of Potash.*

(*a*) *Carbonic acid gas is very abundantly absorbed by a solution of pure potash.*—The simplest mode of showing this fact is the following: Fill a common phial with carbonic acid gas over water; and, when full, stop it by applying the thumb. Then invert the bottle in a solution of pure potash, contained in a cup, and rather exceeding in quantity what is sufficient to fill the bottle. The solution will rise into the bottle, and, if the gas be pure, will fill it entirely. Pour out the alkaline liquor, fill the bottle with water, and again displace it by the gas. Proceed as before, and repeat the process several times. It will be found, that the solution will condense many times its bulk of the gas; whereas water combines only with its own volume.

This experiment may be made, in a much more striking manner, over mercury, by passing into a jar, about three fourths filled with this gas, a comparatively small bulk of a solution of pure potash, which will condense the whole of the gas. If dry potash be substituted in this experiment, no change will ensue; which proves, that solution is essential to the action of alkalis on this gas.

(*b*) *The changes effected in the alkali may next be examined.*—It will be found, after having absorbed as much carbonic acid as it is capable of condensing, to have lost much of its corrosive and penetrating taste, and will no longer destroy the texture of woollen cloth; but it still turns to green the blue infusion of vegetables. Before its absorption of this gas, no remarkable change ensued on mixing it with diluted sulphuric acid; but if this, or almost any other acid, be now added, a violent effervescence will ensue, arising from the escape of the gas that had been previously absorbed. If the mixture be made in a gas bottle, the gas, that is evolved, may be collected, and will be found to exhibit every character of carbonic acid.

For experimental purposes, sub-carbonate of potash may be obtained from crystals of tartar (acidulous tartrate of potash) calcined in a crucible; then lixiviated with water; and evaporated to dryness. By this treatment, the salt yields about one third its weight of sub-carbonate. Or it may be mixed with about an eighth of purified nitrate of potash, and wrapped up in paper in the form of cones, which may be placed on an iron dish, and set on fire. The residuary mass is to be lixiviated, and evaporated as before directed. Or purified nitrate of potash may be mixed with a fourth of its weight of powdered charcoal, and projected into a red-hot crucible, the contents of which are to be poured, when in fusion, into an iron dish. The sub-carbonate, thus obtained, amounts to rather less than one half the nitre which has been employed. Even when thus prepared, it is apt to contain some impurities, consisting chiefly of a minute proportion of sulphate and muriate of potash, from which it is extremely difficult entirely to free it.

(*c*) In this state of sub-saturation with carbonic acid, potash generally occurs in the arts. The potash and pearl-ash of commerce, are *sub-carbonates* of potash, of different degrees of purity. The quantity of carbonic acid, contained in these alkalis, may be learned by a very simple experiment. Put one or two hundred grains of the alkali into a Florence flask, and add a few ounce-measures of water. Take also a phial filled with dilute sulphuric acid, and place this, as well as the flask, in one scale. Balance the two, by putting weights into the opposite scale; and, when the equilibrium is attained, pour gradually the acid into the flask of alkali, till an effervescence no longer ensues. When this has ceased, the scale containing the weights will be found to preponderate. This shows that the alkali, by combination with an acid, loses considerably of its weight; and the exact amount of the loss may be ascertained, by adding weights to the scale containing the flask and phial, till the balance is restored.

(*d*) As it is sometimes of importance to know what proportion of real alkali a given weight of potash or pearlash contains, it may be proper to point out how this information may be acquired. The strength of the alkali is in proportion to the quantity of any acid required to saturate it. Thus, if an ounce of one kind of potash requires, for saturation, a given quantity of sulphuric acid, and an ounce of another kind requires twice that quantity, the latter is twice as strong as the former. In order, however, to obtain a sufficiently accurate standard of comparison, it will be necessary to employ, constantly, an acid of the same strength. This may be effected, though not with absolute uniformity, yet sufficient for ordinary purposes, by diluting the common oil of vitriol of commerce to the same degree. For example, let the standard acid consist of one part of sulphuric acid and five of water. The strength of an alkali will be learned, by observing what quantity of this acid a given quantity of alkali requires for saturation. For this purpose, put half an ounce of the alkali, or any other definite weight, into a jar with a few ounces of water, and filter the solution; weigh the dilute acid employed, before adding it to the alkali; then pour it in gradually, till the effervescence ceases, and till the colour of litmus paper, which has been reddened with vinegar, ceases to be restored to blue. When this happens, the point of saturation will be attained. Weigh the bottle, to ascertain how much acid has been added, and the loss of weight will indicate the strength of the alkali. Another less accurate mode of determining the strength is founded on the following property of carbonate of potash.

(*e*) *Sub-carbonate of potash dissolves very readily in water, which, at the ordinary temperature, takes up more than its own weight.*—Hence, when an alkali, which should consist almost entirely of sub-carbonate of potash, is adulterated, as very often happens, with substances of little solubility, the fraud may be detected by trying how much of one ounce will dissolve in two or three ounce-measures of water. In this way I have detected an adulteration of one third. There are certain substances of ready solubility, however, which may be used in adulterating ashes, as common salt for example; and, when this is done, we must have recourse to the above-mentioned test for the means of discovery.

(*f*) Sub-carbonate of potash, when exposed to the atmosphere, attracts so much moisture, as to pass to a liquid state. This change is termed *deliquescence*.

(*g*) When submitted, in a crucible, to a high temperature, a part only of its carbonic acid is expelled. The alkali runs into fusion, and the fused mass still effervesces on the addition of an acid.

Carbonate of Potash.

Carbonate of potash, in the states which have been already described, is far from being completely saturated with acid. This sufficiently appears from its strongly alkaline taste. It may be much more highly charged with carbonic acid, by exposing a solution of one part of the sub-carbonate in three of water to streams of carbonic acid gas, in a Nooth's machine, or other apparatus; or by the process to be described in art. 3, *g*. When a solution of alkali, after this treatment, is very slowly evaporated, it forms regular crystals. In this state the alkali constitutes the crystallized carbonate of potash, which contains *per cent.*, according to Pelletier, 40 parts of potash, 43 of acid, and 17 of water. It has, therefore, a much larger proportion of water and of acid than the common sub-carbonate, 100 parts of which are composed of 70 parts of alkali, 23 of acid, and 5 of water. According to Dr. Wollaston (Philosophical Transactions, 1808) the quantity of acid in the carbonate is exactly double that in the sub-carbonate. This he proves by disengaging the carbonic acid, from each, by a stronger acid, such as the sulphuric. One part of the carbonate, thus treated, is found to give twice as much carbonic acid as the sub-salt.

In these conclusions of Dr. Wollaston, Berthollet for the most part coincides. (*Mémoires d'Arcueil*, ii. 470.) According to the proportions which he has given, 500 grains of crystallized carbonate of potash are deprived, by the action of dilute sulphuric acid, of 189 grains of carbonic acid. But 1000 grains of the same salt, reduced to sub-carbonate, by calcination in a heat below redness, lose, on adding sulphuric acid, as nearly as possible, the same quantity, *viz.* 190 grains. Between these two extremes, however, Berthollet contends that there exist carbonates with intermediate proportions, which are susceptible of crystallization.

The CARBONATE OF POTASH differs from the sub-carbonate in the following particulars.

1. In the greater mildness of its taste. Though still alkaline, yet it may be applied to the tongue, or taken into the stomach, without exciting any of that burning sensation, which is occasioned by the sub-carbonate.

2. It is unchanged by exposure to the atmosphere.

3. It assumes the shape of regular crystals. The form of these crystals is a four-sided prism, with dihedral triangular summits, the facets of which correspond with the solid angles of the prism.

4. It requires, for solution, four times its weight of water at 60°; and, while dissolving, absorbs caloric. Boiling water dissolves five sixths of its weight; but, during this solution, the salt is partly decomposed, as is manifested by the escape of carbonic acid gas. The quantity thus separated amounts, according to Berthollet, to about $\frac{1}{10}$th the weight of the salt.

5. By calcination in a low red-heat, the portion of carbonic acid, which imparts to this salt its characteristic properties, is expelled, and the salt returns to the state of a sub-carbonate.

(*k*) Carbonate of potash, in all its forms, is decomposed by the stronger acids; as the sulphuric, nitric, and muriatic, which unite with the alkali, and set the gas at liberty.

This may be shown by pouring, on the carbonate contained in a gas bottle, any of the acids, and collecting the gas by a proper apparatus.

ART. 2.—*Carbonate of Soda.*

The combination of carbonic acid with soda, like that of the same acid with potash, exists in two very different states, *viz.* in those of a *sub-carbonate,* and of perfectly saturated *carbonate.* 1. The substance, generally met with in the shops, under the name of soda, affords an example of the *sub-carbonate.* This salt continues dry when exposed to the atmosphere, and even gives up a part of its water of crystallization, the crystals losing their transparency and something of their weight. They have the form of decahedrons, which are composed of two four-sided pyramids, applied base to base, and having their apices truncated. They frequently, also, present large transparent flat rhomboidal prisms. Water, of the temperature of 60°, takes up half its weight; and boiling water rather more than its own weight. In 100 parts, this salt contains, according to Bergman, 16 carbonic acid, 20 soda, and 64 water.

The *carbonate of soda* may be formed either by the process described in the foregoing article *h*, or by that which will be given in the following article, *g*. It is with considerable difficulty brought to crystallize, and affords generally a shapeless mass. Its taste is considerably milder than that of the sub-carbonate of soda, and even than that of neutral carbonate of potash. In 100 parts, it contains, according to Klaproth, 39 acid, 33 soda, 23 water.

In its relation to carbonic acid, Berthollet has observed that soda follows a similar law to potash, the *carbonate* containing twice the quantity of acid which exists in the *sub-carbonate.* The former, however, retains its acid more strongly than the similar salt with base of potash, and is less easily changed into sub-carbonate by heat. One hundred parts of soda are neutralized by 132½ carbonic acid *.

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* *Annales de Chimie,* lxv. 318.

ART. 3.—*Sub-carbonate and Carbonate of Ammonia.*

Ammonia, in its pure state, exists in the form of a gas, permanent over mercury only: and carbonic acid has, also, the form of an aërial fluid. But, when these two gases are mixed together over mercury in proper proportions (*viz.* one measure of carbonic acid to two or three of alkaline gas), they both quit the state of gas, and are entirely condensed into a white solid body. The compound thus formed, it appears from the recent experiments of Gay Lussac, is the *sub-carbonate* of ammonia; for the two gases, he finds, cannot by simple mixture, be made to unite in the proportions necessary to neutralize each other. To effect this, it is necessary to expose a solution of ammonia in water to carbonic acid gas, in which case the affinity of the water concurs in overcoming the elasticity of the acid gas.

Those persons who are not possessed of a mercurial trough, may compose the sub-carbonate of ammonia, in the following manner:—Provide a globular receiver, having two open necks opposite each other. Into one of these introduce the neck of a retort, containing carbonate of lime and dilute sulphuric acid, from which a constant stream of carbonic acid will issue. The inner surface of the globe will remain perfectly unclouded. Into the opposite opening, let the mouth of a retort be introduced, containing the materials for ammoniacal gas. (Chap. vii. sect. 2.) The inner surface of the globe will now be covered with a dense crust of carbonate of ammonia

The sub-carbonate of ammonia may also be formed, by passing, into a jar three fourths filled with carbonic acid over mercury, a solution of pure ammonia, which will instantly effect an absorption of the gas. The ordinary mode of producing it for useful purposes will be described hereafter.

(*b*) Sub-carbonate of ammonia retains, in a considerable degree, the pungent smell of the pure volatile alkali. It is, also, unlike the other sub-carbonates, volatilized by a very moderate heat, and evaporates without entering previously into a liquid state. The vapour that arises may be again condensed in a solid state; affording an example of *sublimation*. This may be shown, by applying heat to the sub-carbonate of ammonia in a retort, to which a receiver is adapted. The sub-carbonate will rise, and be condensed in the receiver in the form of a white crust.

(*c*) This sub-carbonate does not attract moisture from the air, but, on the contrary, loses weight.

(*d*) Sub-carbonate of ammonia, like those of potash and soda, converts vegetable blue colours to green, as the pure alkalis do.

(*e*) It is soluble in rather less than twice its weight of cold water, or in an equal weight of boiling water. At the latter temperature, however, it is partly decomposed, and a violent effervescence ensues.

(*f*) In composition it varies considerably, according to the temperature in which it has been formed. Thus, sub-carbonate of ammonia, which has been produced in a temperature of 300° Fahrenheit, contains 50 *per cent.* of alkali; while carbonate formed at 60° contains only 20 *per cent.* By Gay Lussac (*Mémoires d'Arcueil*, ii. 214), the sub-carbonate of ammonia is said to consist of

Ammonia .	43.98
Carbonic acid	56.02
	100

(*g*) It is decomposed by pure potash and pure soda; and by the sub-carbonates of those alkalis, which attract its carbonic acid, and expel the alkali. Hence it has been recommended, by Berthollet, to employ this salt for the full saturation of potash with carbonic acid, which may be accomplished by the following process.

To a filtered solution of four pounds of pearlash in four quarts of water, add one pound of carbonate of ammonia, reduced to powder; and stir the mixture at intervals, till the carbonate of ammonia is entirely dissolved. Filter the liquor, and put it into a retort, which may be set in a sand-bath, and be connected with a receiver. A very gentle heat is to be applied; so as to distil off about half a pint of the liquor, which will consist of a solution of carbonate of ammonia in water. The liquor in the retort may either be allowed to cool in it, or be transferred into a flat evaporating dish of Wedgwood's ware. When cold, crystals of the carbonate of potash will probably be formed; otherwise another portion must be distilled off, and this must be repeated till the crystals appear; separate the first crystals that are formed; and, on repeating the distillation and cooling, fresh sets will appear in succession. A considerable portion of the solution, however, will refuse to crystallize. This may be boiled to dryness, and applied to the purposes of sub-carbonate of potash. The crystals of carbonate of potash may be washed with a small quantity of cold water, and dried on blotting paper; or, if they are required of great purity, they may be dissolved in cold water, and re-crystallized, using the gentlest heat possible in evaporating the solution.

The NEUTRAL CARBONATE OF AMMONIA was formed by Berthollet, by impregnating a solution of sub-carbonate with carbonic acid gas. According to his experiments, it is composed of

28.19	ammonia,
71.81	carbonic acid.
100	

From the known specific gravity of those two bodies, Gay Lussac has calculated that the neutral carbonate consists of exactly equal quantities by *measure* of the two gases.

Art. 4.—*Carbonate of Barytes.*

I. Pure barytes has a very powerful affinity for carbonic acid.

1. Let a solution of pure barytes be exposed to the atmosphere. It will soon be covered with a thin white pellicle; which, when broken, will fall to the bottom of the vessel, and be succeeded by another. This may be continued, till the whole of the barytes is separated. The effect arises from the absorption of carbonic acid, which is always diffused through the atmosphere, and which forms with barytes a substance, *viz.* carbonate of barytes, much less soluble than the pure earth.

2. Blow the air from the lungs, by means of a quill, a tobacco pipe, or glass tube, through a solution of barytes. The solution will immediately become milky, for the same reason as before.

3. With a solution of pure barytes, mingle a little water, impregnated with carbonic acid. An immediate precipitation of carbonate of barytes will ensue.

4. Barytes has so strong an affinity for carbonic acid, as even to take it from other bodies. To a solution of a

small portion of carbonate of potash, of soda, or of ammonia, add the solution of barytes. The barytes will separate the carbonic acid from the alkali, and will fall down in the state of a carbonate. By adding a sufficient quantity of a solution of barytes in hot water, the whole of the carbonic acid may thus be taken from a carbonated alkali; and the alkali will remain perfectly pure.

II.—1. Carbonate of barytes is nearly insoluble in water, which, at 60°, does not take up more than $\frac{1}{4304}$ part, or, when boiling, about $\frac{1}{2304}$. Water impregnated with carbonic acid dissolves a considerably larger proportion.

2. Carbonate of barytes is perfectly tasteless, and does not alter vegetable blue colours. It acts as a violent poison.

3. The combination of carbonic acid with barytes may either be formed artificially, as in the manner already described, and by other processes, to be detailed in the sequel, in which case it is termed, the *artificial carbonate:* or it may be procured, ready formed, from the earth, and is then called the *native carbonate.* It is not, however, a very common production of nature. The largest quantity, hitherto discovered, is in a mine, now no longer worked, at Anglezark, near Chorley, in Lancashire.

4. The native and artificial carbonates differ in the proportion of their components. The former contains, in 100 parts, 20 acid and 80 barytes. The artificial, according to Pelletier, consists of 22 acid, 62 earth, and 16 water. Strictly speaking they are both sub-carbonates; but as we are unacquainted with any other compound of barytes and carbonic acid, they may be allowed to retain the accustomed name of carbonate. The latest analysis of this compound

by Mr. Aikin and Mr. James Thomson*, fix its composition as follows:

	Carb. acid.	*Barytes.*
From Mr. Aikin's experiments .	21.67	78.33
—— Mr. Thomson's	21.75	78.25

5. Carbonate of barytes is decomposed by an intense heat; its carbonic acid being expelled; and the barytes remaining pure. The artificial carbonate is most readily decomposed; but the native one is generally employed for obtaining pure barytes, because it may be had in considerable quantity. The process, which I have found to answer best, is nearly that of Pelletier. Let the native carbonate be powdered, and passed through a fine sieve. Work it up with about an equal bulk of wheaten flour into a ball, adding a sufficient quantity of water. Fill a crucible of proper size, about one third its height, with powdered charcoal; place the ball on this; and surround and cover it with the same powder, so as to prevent its coming into contact with the sides of the crucible. Lute on a cover; and expose it, for two hours, to the most violent heat that can be raised in a wind furnace. Let the ball be removed when cold. On the addition of water, it will evolve great heat, as already described (chap. viii.), and the barytes will be dissolved. The filtered solution, on cooling, will shoot into beautiful crystals.

6. Carbonate of barytes is decomposed by the sulphuric, nitric, muriatic, and various other acids, which detach the carbonic acid, and combine with the earth.

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* Nicholson's Journal, xxii. xxiii.

ART. 5.—*Carbonate of Strontites.*

The relation of strontites to carbonic acid resembles, very closely, that of barytes; and all the experiments, directed to be made with the solution of the latter earth, may be repeated with that of strontites, which will exhibit similar appearances.

The carbonate of strontites requires for solution 1536 parts of boiling water. It is found native at Strontian in Argyleshire; and may, also, be prepared by artificial processes, which will be afterwards described. From this carbonate pure strontites may be prepared, by treating it in the same manner as was directed for the calcination of carbonate of barytes.

ART. 6.—*Carbonate of Lime.*

1. Lime has a strong attraction for *carbonic acid*, but not when perfectly dry.

(*a*) If a piece of dry quicklime be passed into a jar of carbonic acid gas over mercury, no absorption ensues. But invert a bottle, filled with carbonic acid gas, over a mixture of lime and water of the consistence of cream, and a rapid absorption will be observed, especially if the bottle be agitated.

(*b*) Let a jar or bottle, filled with carbonic acid, be brought over a vessel of lime water. On agitating the vessel, a rapid diminution will ensue, and the lime water will become milky.

(*c*) Leave a shallow vessel of lime water exposed to the air. A white crust will form on the surface, and this, if broken, will fall to the bottom, and be succeeded by another. This is owing to the absorption of carbonic acid gas from the air by the lime, which is thus rendered insoluble in water.

(*d*) Lime, when exposed to the atmosphere, first acquires moisture, and then carbonic acid; and, in a sufficient space of time, all the characters distinguishing it as lime disappear.

(*e*) Lime has an extremely strong affinity for carbonic acid, which enables it to take this acid from other substances. Thus carbonates of alkalis are decomposed by lime. Slake a given quantity of lime into a paste with water, and add half its weight of carbonate of potash or soda. Boil the mixture, for half an hour, in an iron kettle, and separate the liquid part by filtration or by subsidence. The carbonic acid combines with the lime, and the alkali is obtained in a state of solution perfectly free from carbonic acid. This is the ordinary mode of depriving the alkalis of carbonic acid.

(*f*) Lime, when saturated with carbonic acid, must necessarily form carbonate of lime. We are unacquainted, however, with this salt, and it is chiefly by a process of reasoning that Berthollet has shown it must consist of 100 parts of lime united to 150.6 carbonic acid; whereas in the sub-carbonate 100 parts of lime are combined with only half that quantity. (*Mémoires d'Arcueil*, ii. 478.) Of this, common chalk may be taken as a fair sample; and in all sub-carbonates of this earth, we find the characters of insipidity and insolubility in water. Calcareous spar, marble, stalactites, lime-stone, and chalk, are all varieties of sub-carbonate of lime.

(*g*) Carbonate of lime is decomposed by a strong heat. If distilled in an earthen retort, carbonic acid gas is obtained, and lime remains in the retort in a pure or caustic state. By this process it loses about 45 *per cent.*

The very curious and important experiments of Sir James Hall have proved, that when the escape of the carbonic acid is prevented by strong pressure, carbonate of lime is fusible in a heat of about 22° of Wedgwood's pyrometer. (Nicholson's Journal, xiii. xiv.) And Mr. Bucholz has lately fused this substance, by the sudden application of a violent heat, without additional compression. (Nicholson's Journal, xvii. 229.)

(*h*) Carbonate of lime is decomposed by the stronger acids. Put some chalk into a gas bottle, and pour on it diluted sulphuric acid. The sulphuric acid will unite with the lime, and the carbonic acid will be set at liberty.

By a comparison of this experiment with the preceding one (*g*), we may learn the proportion of carbonic acid and water contained in any carbonate of lime. Let 100 grains of the carbonate be put into a Florence flask, with an ounce or two of water; place this in the scale of a balance; and in the same scale, but in a separate bottle, about half an ounce of muriatic acid. Add the muriatic acid to the carbonate as long as any effervescence is produced, and then blow out the disengaged carbonic acid, which remains in the flask, by a pair of bellows. Ascertain, by adding weights to the opposite scale, how much has been lost; suppose it to be forty grains; this shows the quantity of carbonic acid disengaged. Calcine another 100 grains in a covered crucible. It will lose still more of its weight; because, besides its carbonic acid, all the water is expelled which it may contain. Let this loss be stated at 45 grains; the former loss deducted from this (45 – 40), or 5 grains, shows the quantity of water in 100 of the carbonate.

(*i*) Carbonate of lime, though scarcely dissolved by pure water, is soluble in water saturated with carbonic acid. The most striking method of showing this is the following: Add to a jar, about one fourth filled with lime water, a very small quantity of water saturated with carbonic acid. An immediate milkiness will ensue, because the carbonic acid forms with the lime an insoluble carbonate. Add gradually more of the water, impregnated with carbonic acid, shaking the jar as these additions are made. At last the precipitate is re-dissolved. Hence it appears that lime, with a certain proportion of carbonic acid, is insoluble, and, with a still larger, again becomes soluble in water.

(*k*) The carbonate of lime, dissolved by an excess of carbonic acid (*i*), is again separated, when this excess is driven off. Thus boiling, which expels the superabundant acid, precipitates the carbonate. Caustic, or pure alkalis, also produce a similar effect.

Art. 7.—*Carbonate of Magnesia.*

I. Pure magnesia does not attract carbonic acid with nearly the same intensity as lime. Hence magnesia may be exposed, for a long time, to the air, without any important change in its properties, or much increase of weight. The carbonate of magnesia, used in medicine, and for experimental purposes, is prepared by a process to be described in the sequel. In this state, however, it is not entirely saturated with carbonic acid, and is rather a sub-carbonate.

II. The saturated carbonate may be obtained, by passing streams of carbonic acid gas through water, in which the sub-carbonate is kept mechanically suspended. The solution yields, when evaporated, small crystals, which are transparent hexagonal prisms, terminated by hexagonal planes. These crystals have no taste, and are soluble in 48 parts of cold water; whereas the sub-carbonate requires at least ten times that quantity. The crystallized carbonate contains *per cent.* 50 acid, 25 earth, and 25 water; the sub-carbonate 34 acid, 45 earth, and 21 water.

III. The carbonate of magnesia is decomposed by the same agents as the carbonate of lime. It yields its carbonic acid, however, in a much more moderate heat.

IV. Lime has a stronger affinity than magnesia for carbonic acid. Hence, if lime water be digested with carbonate of magnesia, the lime is precipitated in the state of an insoluble carbonate.

ART. 8.—*Carbonate of Glucine.*

Glucine appears to have a considerable affinity for carbonic acid; for, when precipitated from acids by pure alkalis, and dried in the air, it becomes effervescent. The carbonate of glucine is white, insipid, insoluble, and very light. It contains about one fourth its weight of carbonic acid, which it loses by exposure to a low red-heat.

The carbonate of silex does not exist, and those of zircon, alumine, and yttria, have no peculiarly interesting properties.

SECTION V.

Gaseous Oxide of Carbon, or Carbonous Oxide.

THIS combination of carbon with oxygen contains a less proportion of oxygen than is found in carbonic acid.

Its discovery was announced in Nicholson's Journal, for April, 1801, by Mr. Cruickshank, and in the 38th volume of the *Annales de Chimie*, by Clement and Desormes, whose experiments are continued in the 39th volume, page 26. The Dutch chemists, however, in volume 43, object to its being considered as a distinct gas, and regard it merely as a carburet of hydrogen. But their objections do not appear sufficiently strong to prevent the acknowledgement of the gaseous oxide as a new and peculiar species.

It may be procured by any of the following processes:

1. By the distillation of the white oxide of zinc with one eighth its weight of charcoal, in an earthen or glass retort; from the scales which fly from iron in forging, mixed with a similar proportion of charcoal; from the oxides of lead, manganese, or, indeed, of almost every imperfect metal, when heated in contact with powdered charcoal. It may also be obtained from the substance which remains after preparing acetic acid from acetate of copper.

2. From well dried carbonate of barytes or of lime (common chalk), distilled with about one fifth of charcoal; or with rather a larger proportion of dry iron or zinc filings, which afford it quite free from hydrogen.

3. By transmitting carbonic acid gas over charcoal ignited in a porcelain tube. The acid gas combines with an additional dose of charcoal; loses its acid properties; and is converted into the carbonous oxide. An ingenious apparatus, contrived by M. Baruel, and extremely useful for this and similar purposes, is described, and represented by a plate, in the 11th volume of Nicholson's Journal.

The last product of the distillation is the purest, but still contains carbonic acid, which must be separated by washing the gas with lime liquor.

Its properties are as follow:

(*a*) It has an offensive smell.

(*b*) It is lighter than common air, in the proportion of 966 to 1000. One hundred cubical inches weigh 30 grains, the temperature being 55° Fahrenheit, and pressure 29.5 (Cruickshank); or at temperature 60°, and barometer 30°, 100 cubic inches weigh 30.19 grains. Its specific gravity from calculation, according to Gay Lussac, is .96782.

(*c*) It is inflammable, and, when set fire to, as it issues from the orifice of a small pipe, burns with a blue flame. When mixed with common air, it does not explode like other inflammable gases, but burns silently with a lambent blue flame. It detonates, however, with oxygen gas.

(*d*) When a stream of this gas is burnt, in the manner described in speaking of hydrogen gas, no water is condensed on the inner surface of the glass globe, a proof that the gaseous oxide contains no hydrogen. Berthollet, indeed, still contends, in opposition to most chemists (and among others to Gay Lussac) that hydrogen is one of the elements of this gas.

(*e*) It is sparingly soluble in water; is not absorbed by liquid caustic alkalis; nor does it precipitate lime water.

(*f*) It is extremely noxious to animals; and fatal to them if confined in it. When respired for a few minutes, it produces giddiness and fainting.

(*g*) When 100 measures of carbonous oxide are fired over mercury in a detonating tube, with 45 of oxygen gas, the total 145 are diminished to 90, which consist entirely of carbonic acid. Proportions, differing a little from these, have been stated by Berthollet, *viz.* that 100 measures of carbonous oxide are saturated by 50 measures of oxygen, and give 100 of carbonic acid.

(*h*) It is not expanded by electric shocks, nor does it appear to undergo any change by electrization.

(*i*) When the carbonous oxide, mingled with an equal bulk of hydrogen gas, is passed through an ignited tube, the tube becomes lined with charcoal. In this temperature, the hydrogen attracts oxygen more strongly than it is retained by the charcoal, and forms water. It was found, also, by Gay Lussac to be decomposed by the action of potassium, which combines with the oxygen and precipitates charcoal.

According to Mr. Cruickshank, it contains *per cent.* about 70 oxygen, and 30 carbon by weight; or the former is to the latter as 21 to 8.6, or as 21 to 9. Gay Lussac, however, makes it to consist of

43	charcoal,
57	oxygen.
100	

This determination is probably nearer the truth than Mr. Cruickshank's; for, calculating on the statement which he has given, of the proportion of carbonic acid resulting from its combustion, and assuming that this acid contains 28 *per cent.* of carbon, 100 parts of carbonous oxide ought to consist of 40 charcoal and 60 oxygen.

SECTION VI.

Combination of Carbon with Hydrogen, forming Carburetted Hydrogen Gas, or Hydro-Carburet.

I. Of this combination there appear, on first view, to be several distinct varieties, consisting of carbon and hydrogen, united in various proportions, and obtained by different processes.

1. When the vapour of water is brought into contact with red-hot charcoal (by means of an apparatus similar to that represented, fig. 40), two different products are obtained. The oxygen of the water, uniting with the carbon, constitutes carbonic acid; and the hydrogen of the water dissolving, at the moment of its liberation, a portion of charcoal, composes carburetted hydrogen gas*. The carbonic acid may be separated from the hydro-carburet, by agitating the gas, which has been produced, in contact with lime and water, mixed together, so as to be of the consistence of cream.

2. By stirring, with a stick, the mud that is deposited at the bottom of ditches or stagnant pools, bubbles of gas ascend to the surface, and may be collected in an inverted bottle of water, to the mouth of which a funnel, also inverted, is fixed.

3. By submitting coal to distillation, in an iron or coated glass retort, a large quantity of gas, besides a portion of tar, is produced. The latter may be received in an intermediate vessel; and the gas must be well washed with lime liquor. The first product only may be reserved as a specimen of coal gas; for, as the distillation proceeds, its density becomes gradually less; till, at length, the gas which is produced at the close of the operation is only about half as heavy as that evolved at first. The quantity of gas, also, which is produced from a given weight of coal, is so variable in different kinds of this mineral, that it is

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* In Nicholson's Journal, xi. 68, I have stated my reasons for believing that this gas is not pure hydro-carburet.

scarcely possible to state any general average. From 120 pounds avoirdupois of the sort of coal called *Wigan Cannel*, about 340 cubic feet of gas may be obtained, of which half a cubic foot *per* hour is equal to a mould candle of six to the pound, burning during the same space of time.

4. Let a porcelain tube, coated with clay, be fixed horizontally in a furnace, in the manner represented, fig. 40. To one end let a retort be luted, containing an ounce or two of ether or alcohol; and, to the other, a bent tube, which terminates under the shelf of the pneumatic trough. A gas will be disengaged, on igniting the tube, and transmitting, through it, the alcohol or ether in vapour, which, when washed with lime-liquor, is the carburetted hydrogen.

5. A fifth mode of obtaining hydro-carburet, consists in distilling, in a glass retort, with a gentle heat, three parts of concentrated sulphuric acid, and one part of alcohol. The mixture assumes a black colour and thick consistence; and bubbles of gas are disengaged, which may be collected over water. For reasons which will afterwards be stated, this gas has been named the olefiant gas.

II. These varieties of carburetted hydrogen gas all agree in being inflammable; but they possess this property in various degrees, as is evinced by the variable brightness of the flame, which they yield when set on fire. They may be inflamed as they proceed from the orifice of a small pipe, or from between two concentric cylinders of sheet-iron or copper, placed at the distance of a small fraction of an inch from each other. On this principle, an Argand's lamp may be constructed, for burning the gases, which will issue from that space, commonly occupied by the wick.

1. When burned in either of these modes, there is a manifest gradation in the density and brightness of the flame. The gas from charcoal burns with a faint blue light, not suited to the purpose of illumination; that from ether or alcohol with more brilliancy; but still short of that with which the coal gas burns, when recently prepared. The olefiant gas surpasses them all, in the quantity of light evolved by its combustion.

2. If these gases be burned in a vessel of oxygen gas over lime-water, by means of a bladder and bent brass pipe (pl. iv. fig. 41), two distinct products are obtained, *viz.* water and carbonic acid. That water is produced, may be shown by burning a very small stream of this gas under a long funnel-shaped tube open at both ends. The formation of carbonic acid is evinced, by the copious precipitation of the lime-water in the foregoing experiment.

3. The composition of each of the above gases is learned by firing it, in a detonating tube over mercury, with a known quantity of oxygen gas; and observing the nature and quantity of the products. These products are carbonic acid and water. The former may be exactly measured; but the water is generated in such small quantity, that it can only be computed. The following table shows the results of a few experiments of this kind.

Kind of Gas.	Measures of Oxygen Gas required to saturate 100 Measures.	Measures of Carbonic Acid produced.
Pure hydrogen gas	50 to 54	
Gas from charcoal	60	35
——— coal	190	97.5
——— stagnant water	200	100
Olefiant gas	300	200

Now since, for the formation of each measure of carbonic acid gas, in the foregoing experiments, an equal volume of oxygen gas is required, we may learn, by deducting the number in the third column from the corresponding one in the second, what proportion of oxygen has been spent in the saturation of the hydrogen of each hydro-carburet. Thus, for example, in burning the gas from coal, 100 measures of oxygen have been employed in forming carbonic acid; and the remaining 90 in saturating hydrogen. But 90 measures of oxygen are sufficient to saturate 180 of hydrogen gas; and a quantity of hydrogen must therefore be contained in 100 measures of gas from coal; which, expanded to its usual elasticity, would occupy 180 measures.

4. The gases vary in their solubility by water, the olefiant gas being absorbed in the largest proportion, *viz.* one eighth the bulk of the water; the gas from stagnant water one sixty-fourth; and the others in still less quantity.

5. They vary also in density or specific gravity. Common air being 1000, the olefiant gas is 909; the gas from ether or alcohol 520; and from moistened charcoal 480. The specific gravity of gas from coal, and of that from stagnant water, I am informed by Mr. Dalton, are the same, *viz.* 666, or as 2 to 3.

6. The only distinct and well characterized species of hydro-carburet appear to me to be the olefiant gas; and the gas from stagnant water, yielding by combustion, an equal bulk of carbonic acid. Of these, the other varieties appear to be only mixtures. The reasons for this opinion I have stated in Nicholson's Journal, volume xi. page 68. It is proper, however, to add that they have not been satisfactory to Berthollet, who, in the 2d volume of the *Mémoires d'Arcueil*, has since argued that hydrogen, carbon,

and oxygen may combine in unlimited proportions, and form a great variety of combustible gases, to which he thinks the term *oxy-carburetted hydrogen* more appropriate. On the contrary, it is contended by Mr. Dalton and myself, that oxygen is not an essential element of these compounds; and that the hydrogen and carbon, of which alone they consist, are capable of uniting in but few proportions. The controversy, which would be unseasonable in an elementary work, I shall probably resume on some future occasion.

CHAPTER XII.

SULPHUR.—SULPHURIC ACID.—SULPHATES.—BINARY COMPOUNDS OF SULPHUR.

IN describing sulphur and its compounds, I shall take them in the following order.

I. SULPHUR in its uncombined state.

II. Sulphur united with its full proportion of oxygen, constituting SULPHURIC ACID; and the compounds of this acid with alkalis and earths, termed SULPHATES.

III. Sulphur united with a less proportion of oxygen, composing SULPHUROUS ACID; and the compounds of this acid, called SULPHITES.

IV. The compounds of sulphur with alkalis and earths, termed SULPHURETS.

V. The combination of sulphur and hydrogen, named SULPHURETTED HYDROGEN; and the compounds, which it forms with alkaline and earthy bases, called HYDRO-SULPHURETS.

VI. The compound of sulphuretted hydrogen with a still farther dose of sulphur, composing SUPER-SULPHURETTED HYDROGEN; and its compounds with different bases called HYDROGURETTED SULPHURETS, or sometimes SULPHURETTED-HYDRO-SULPHURETS.

SECTION I.

SULPHUR.

I. The sulphur, which occurs as an article of commerce, is a mineral production, and is brought to this country chiefly from Sicily. That which is obtained in our own island, is generally of very inferior quality, and contains a portion of the metal, from combination with which it has been separated. It is met with under two different forms, that of a compact solid, which has generally the shape of long rolls or sticks; and that of a light powder called *flowers of sulphur*. In general, the latter may be considered as most pure; but the two varieties, it will presently appear, are readily convertible into each other by the modified application of heat.

II. Sulphur is readily fused and volatilized. When heated to 170° of Fahrenheit, it begins to evaporate, and to produce a very disagreeable smell; at 185° or 190° it begins to melt; and at 220° is completely fluid. If the heat be rapidly increased, it loses its fluidity, and becomes firm, and of a deeper colour. It regains its fluidity, if we reduce the temperature; and this may be repeated at pleasure, in close glass vessels, if the changes of heat be not slow; otherwise it is volatilized.

III. If, after being melted, it be suffered to cool, it congeals in a crystalline form, but so confusedly, that we cannot define the shape of the crystals, farther than that they are slender interlaced fibres. If a large mass be kept fluid below, while it congeals at the surface, the crystallization there is much more distinct. When sulphur in complete fusion is poured into water, it becomes tenacious like wax, and may be applied (as is done by Mr. Tassie) to take impressions from engraved stones, &c. These impressions are quite hard, when the sulphur has become cold.

IV. At the temperature of about 290° Fahrenheit, sulphur is converted into vapour; and if this operation be conducted in close vessels, the volatilized sulphur is again collected in a solid form. What remains has been called *sulphur vivum*. This affords an example of the process of *sublimation*, which differs from distillation, in affording a solid product, while the latter yields a condensed liquid. In this mode, sulphur may, in part, be purified; and its purification is completed, by boiling it repeatedly in distilled water; then in twice or thrice its weight of nitro-muriatic acid, diluted with one part of distilled water; and, finally, by washing it with distilled water, till this comes off tasteless, and incapable of changing the blue colours of vegetables.

V. When flowers of sulphur are digested in alcohol, no union takes place; but if the two bodies be brought into contact, when both are in a state of vapour, they enter into chemical union. This may be shown by an ingenious experiment of La Grange, the apparatus for performing which is represented in the first plate of his "Manual." Into a glass alembic (see the plates to this work, fig. 2) put a little sulphur; over this suspend, by a couple of strings, a small bottle filled with alcohol; and apply a receiver to the pipe of the alembic, the head being put into its place. Lute the junctures, and apply a gentle heat to the alembic. The sulphur will now be raised in vapour; and the vapour surrounding the bottle of alcohol, the latter will be volatilized, and will meet in this state the fumes of sulphur. A combination will take place between the two bodies, and sulphurized alcohol will pass into the receiver, On pouring this preparation into water, the sulphur will be precipitated.

VI. Though it had already been suspected (chiefly from the experiments of M. Berthollet, junior, described in sect. 6, art. 4, of this chapter) that sulphur contains hydrogen, yet the first unequivocal evidence of the fact was furnished by Mr. Davy. In considering the analytical powers of the Voltaic apparatus, it occurred to him, that though sulphur, being a non-conductor, could not be expected to yield its elements to the attractive and repulsive powers of opposite electricities, yet that the intense heat, which is thus produced, might possibly effect some alteration in it, and tend to separate any elastic matter it might contain. On this idea, a bent glass tube, having a platina wire hermetically sealed into its upper extremity, was filled with sulphur. The sulphur was melted by heat; and a proper connection being made with the Voltaic apparatus of 500 double plates, each six inches square and highly charged, a most intense action took place. A very brilliant light was emitted; the sulphur soon entered into ebullition; elastic matter was evolved in great quantities; and the sulphur, from being of a pure yellow, became of a dark reddish brown tint. The gas was found to be sulphuretted hydrogen, or hydrogen gas holding sulphur in solution; and its quantity, in about two hours, was more than five times the volume of the sulphur employed.

Another proof of the presence of hydrogen in sulphur is derived from the action of potassium; for these two bodies act upon each other with great energy, and evolve sulphuretted hydrogen, with intense heat and light.

Lastly, when dry sulphur is burned in dry oxygen gas, Mr. Davy is of opinion that, beside sulphuric acid, a portion of water is also formed.

Another ingredient of sulphur, it appears probable from the experiments of the same chemist, is oxygen. When potassium is made to act on sulphuretted hydrogen, it burns with a brilliant light; and a small quantity of potash is formed, which unites with a portion of the sulphur. But to the production of this potash, oxygen is absolutely necessary, and since the hydrogen could not have furnished it, no other source can be assigned but the sulphur which the gas holds in solution.

An additional argument, in favour of the presence of oxygen in sulphur, is derived from the following fact. Potassium and sodium, in their perfectly neutralized state, evolve a certain quantity of hydrogen from water, in consequence of their affinity for oxygen. But whenever they are partially oxidized, before being added to water, they do not evolve so much; because the affinity, by which they effect the decomposition of water, is already in part satisfied Now it is remarkable that potassium, after being combined with sulphur, detaches a less proportion of hydrogen from water than an equal weight of the pure metal; and the more we increase the proportion of the sulphur to the potassium the less hydrogen does the latter evolve, probably because it has already acquired a greater share of oxygen from the sulphur.

Sulphur, then, in its common state, appears to be a combination of small quantities of oxygen and hydrogen with a large quantity of a peculiar basis. The nature of this basis, however, though it has been an object of investigation with Mr. Davy, is but imperfectly known. He attempted to obtain it in a separate form, from the compound of potassium and sulphur, in which we may naturally suppose it to exist. Muriatic acid detached a substance of a dark grey colour, harsh to the touch; which, at common temperatures, had no smell; but, when heated, emitted the peculiar odour of sulphur. Though certainly differing from sulphur in its ordinary state, yet in its chemical properties it approached so nearly to that body, that we must still consider the exhibition of the base of sulphur in a pure form, as not satisfactorily accomplished.

VII. Sulphur is inflammable, and appears susceptible of two distinct combustions, which take place at different temperatures*. At 140° or 150° Fahrenheit, it begins sensibly to attract oxygen; and if the temperature be raised to 180° or 190°, the combination becomes pretty rapid, accompanied by a faint blue light. But the heat evolved is scarcely sensible; at least it is so weak, that the sulphur may thus be burnt out of gunpowder, and the powder be rendered useless without inflaming it. At a temperature of 300°, its combustion, though still feeble compared with that of some other bodies, is much more active, and accompanied with a redder light. When set on fire in oxygen gas, it burns with a very beautiful and brilliant light; but of a given quantity of oxygen gas, it is not possible to condense the whole by this combustion, for reasons which hereafter will be stated. The product of these combustions, when examined, will be found to be sulphuric acid†.

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* For an account of the oxides of sulphur, see Dr. Thomson's paper in Nicholson's Journal, vi. 101.

† Much sulphurous acid is also generated in these processes.

SECTION II.

Sulphuric Acid.

I. Though this is not the mode in which sulphuric acid is ordinarily prepared (which will be afterwards described in chap. xiii. sect. 5), yet it will be proper for the chemical student to examine the result of this combustion, on account of the simplicity of the process. Let the glass bell, under which sulphur has been burnt, be rinced out with a little water. This water will have an acid taste, will turn vegetable blue colours red, and will effervesce with carbonated alkalis. It is therefore an acid; and as it is composed of sulphur and oxygen, it is termed the sulphuric acid. The properties of this acid must be exhibited by a portion of that usually found in the shops. They are as follows:

(*a*) Sulphuric acid has a thick and oily consistence; as may be seen by pouring it from one vessel into another.

(*b*) In 100 parts, it contains, according to Chenevix, 61½ sulphur and 38½ oxygen. Berthollet, however, has lately stated (*Mémoires d'Arcueil*, ii.), that 17.846 parts of sulphur, oxygenized by nitric acid, give a quantity of sulphuric acid which affords 127.515 parts of sulphate of barytes. Hence it follows that 100 parts by weight of sulphur are convertible into 230.79 parts of real sulphuric acid (=about 292 parts of density 1.85); and sulphuric acid must consist of

43.28	sulphur
56.72	oxygen
100.	

This last statement accords so nearly with those of Klaproth, Richter, and Bucholz (which all agree in fixing the sulphur in sulphuric acid at between 42 and 43 *per cent.*), that it is probably the correct one. It will be found, also, if we assume, as appears now to be decisively proved, the sulphate of barytes to contain 33½ *per cent.* of acid, and not 23½ as estimated by Mr. Chenevix, that the results of this last mentioned chemist differ very little from the one, which I have considered as most nearly approaching the truth.

(*c*) Sulphuric acid is nearly twice as heavy as water. This will appear from weighing a small phial filled with the acid, and afterwards filled with distilled water. The same result will be more conveniently obtained, by making the comparison in the long-necked bottle. (Pl. i. fig. 14.) The specific gravity of the acid, as prepared for purposes of commerce, is generally 1.845, containing, in this state, according to Kirwan,

79	real acid
21	water
100.	

It does not appear that the specific gravity of rectified sulphuric acid can be carried so high as 1.850*. When of the specific gravity 1.417, it contains, according to Berthollet, 38.58 *per cent.* of real acid; the remainder being water.

(*d*) In a pure state, it is perfectly limpid and colourless.

(*e*) When mixed suddenly with water, considerable

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[*] Transactions of the Irish Academy, iv. 88.

heat is produced. Four parts, by weight, of concentrated sulphuric acid, and one of water, when suddenly mixed together, each at the temperature of 50° Fahrenheit, have their temperature raised to 300°. This rise of temperature takes place, because the affinity or capacity of the compound of sulphuric acid and water for caloric, is less than that of the acid and water separately. A diminution of bulk also ensues; that is, one measure of acid and one of water do not occupy the space of two measures, but about $\frac{1}{75}$th less. Owing to the heat produced by its admixture with water, the dilution, for ordinary purposes, should be conducted very gradually; and the acid should be added to the water by small portions at once, allowing each portion to cool before a fresh addition is made. On the principle of its attraction for water is to be explained, also, the rapid increase of weight which the acid acquires when exposed to air. In one day, three parts of sulphuric acid, exposed to the atmosphere, are increased in weight one part; and one ounce, by twelve months exposure, has been found to gain an addition of $6\frac{1}{4}$.

(*f*) A perfectly pure sulphuric acid remains quite limpid during dilution. The sulphuric acid, however, commonly found in the shops,[*] under the name of oil of vitriol, on admixture with water, deposits a white powder, in considerable quantity, consisting of various impurities, but chiefly of sulphate of lead.

(*g*) The boiling point of sulphuric acid is considerably higher than that of water, and even approaches to that of mercury. Bergman stated its boiling point to be 540° Fahrenheit, but Mr. Dalton has determined it to be 590°.

(*h*) Sulphuric acid, by a sufficient reduction of its temperature, may be frozen; and under favourable circumstances, it assumes a regular crystalline form, a considerable degree of solidity or hardness, and a density exceeding that which it possessed in a fluid state. From the experiments of Mr. Keir* it follows that there is a certain point of specific gravity (*viz.* 1780 to 1000), at which the sulphuric acid most readily congeals; and when of this degree of strength it requires even a less degree of cold than is sufficient to freeze water, its congelation taking place at 45° Fahrenheit. From the specific gravity of 1786 on the one hand to 1775 on the other, it freezes at 32° Fahrenheit. It is singular that it remains congealed at a temperature higher than that originally required for freezing it. Acid, for example, which did not become solid till its temperature was reduced to 32°, remains frozen at 45°. When of the specific gravity of 1843, or as nearly as possible of that of commerce, it was found by Mr. Macnab† to freeze at −15° Fahrenheit; but this acid, mixed with rather more than half its weight of water required for congelation the temperature of −36° Fahrenheit.

(*i*) To purify sulphuric acid, it must be distilled in a glass retort, placed in the sand-bed of a reverberatory furnace. This process is a very difficult one; and an inexperienced chemist should, therefore, not attempt it. To those, however, who have the means of repeating the process, and sufficient experience in chemical operations, the following instructions may be useful; especially as it is indispensable, in all experiments of research, to employ an acid thus purified.

The furnace, in which this process is conducted, should have a contrivance for supporting a sand-bath within it at a proper height; and an opening in the side, for transmitting the neck of the retort. (Pl. vii. fig. 62, 63.) The retort must be coated with clay and sand over its whole body, and also over that part of the neck which is exposed to the fire. It is then to be placed, the coating being previously dry, in the sand-bath, about one half filled with sulphuric acid; and a receiver must be applied, but not luted on. The fire must now be lighted, and raised with extreme caution. The first portion that comes over, amounting to about one sixth, consists chiefly of water, and may be rejected. This is followed by the concentrated acid; and, at this period, there is great risk that the neck of the retort will be broken, by the contact of the condensed acid, which has a very high temperature, and which frequently cracks the glass, as effectually as the application of a red-hot iron. The fire must be regulated by the register-door of the ash-pit, so that several seconds may elapse between the fall of the drops into the receiver. The process may be continued as long as any acid is condensed. The retorts, employed for this purpose, should be most attentively annealed; and it is advisable, that the operator should anneal them himself, by first heating them in an oven, and then allowing the oven to cool as slowly as possible.

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* Philosophical Transactions, lxxvii. 267.

† Philosophical Transactions, lxxvi. 241.

Sulphuric acid may be less perfectly purified by diluting it with an equal weight of water, allowing the impurities to settle, decanting the clear liquor, and evaporating it to the proper degree in a glass retort.

(*k*) Sulphuric acid is decomposed at the temperature of the atmosphere, by inflammable substances, and acquires a dark colour. The addition of a little brown sugar, or a drop of oil, to a portion of the acid, imparts to it a brownish hue, which in time changes to black. Hence this acid should always be kept in bottles with glass stoppers; for a small bit of straw or cork, if dropped into a considerable quantity of sulphuric acid, changes it in the manner that has been pointed out.

(*l*) In high temperatures; sulphuric acid is still farther decomposed by combustible bodies.

1. Hydrogen gas, brought into contact with sulphuric acid, in a state approaching ignition, decomposes it, and water and sulphurous acid are formed. This, however, is a most dangerous and difficult process, which it is not advisable to repeat.

2. According to Gay Lussac, sulphuric acid is decomposed by a high temperature, and is resolved into two parts by measure of sulphurous acid gas, and one of oxygen gas. This experiment is best performed by passing the acid through a red-hot tube of glass or porcelain.

3. Sulphur, by being boiled in sulphuric acid, partly deoxygenates it, and converts a portion of it into sulphurous acid, which comes over in a gaseous state.

4. Into a glass retort put such a quantity of sulphuric acid as will fill about one fourth part of it, and add a small portion of powdered charcoal. On applying the heat of a lamp, gas will be produced very abundantly. Let this gas be conveyed by a tube fixed to the mouth of a retort, and bent in the proper manner, into an inverted jar of water; or, if it can be had, into an inverted jar of quicksilver in a mercurial apparatus. During this operation, the carbon attracts part of the oxygen of the sulphuric acid, and forms carbonic acid gas. But the sulphur is not entirely disoxygenated; and a compound is therefore formed of sulphur and oxygen, containing less oxygen than the sulphuric acid. This compound exists in the state of a gas, and its properties may next be examined. To avoid, however,

the complication which the admixture of carbonic acid with this new product introduces into the experiment, it may be proper to prepare it in a mode less objectionable, but the *rationale* of which cannot at present be explained. This consists in dissolving one part, by weight, of quicksilver in two of sulphuric acid, and boiling the mass to dryness, in the bottom of a broken Florence flask. The dried mass, still remaining in the retort, is next to be distilled in a strong sand-heat; a glass globe being interposed between the retort and the receiving mercurial trough, to condense any sulphuric acid that may escape decomposition. (See pl. iii. fig. 31.) The gas thus obtained is termed, conformably to the principles of the new nomenclature, *sulphurous acid*.

SECTION III.

Sulphurous Acid Gas.

Sulphurous acid may be formed, also, by burning sulphur at a low temperature under a glass bell; and if slips of linen cloth, dipped in a solution of potash, be exposed to the vapour, the alkali forms a combination with the sulphurous acid, which may afterwards be washed off and evaporated. The dry salt, distilled with liquid tartaric acid, gives sulphurous acid gas.

Its properties are the following:

(*a*) It has a pungent and suffocating smell, exactly resembling that which arises from burning sulphur.

(*b*) It is above twice heavier than atmospherical air.

According to Kirwan, its specific gravity is 2.265, common air being 1; or, according to Gay Lussac, from calculation 2.30314. At the temperature of 60°, the barometer being 30 inches, 100 cubic inches of this gas are stated by Mr. Kirwan to weigh 70.215 grains. By Monge and Clouet it is said, if exposed at the same time to a temperature of 31° Fahrenheit, and to great pressure, to assume a fluid state.

(*c*) It extinguishes burning bodies, and kills animals.

(*d*) It has the property of whitening or bleaching silk, and of giving it lustre.

(*e*) Of sulphurous acid, water absorbs 33 times its bulk, or one eleventh of its weight, caloric is evolved, and the solution at 68° has the specific gravity 1.0513. From this fluid it is again separated, like carbonic acid, by the application of heat; but not by congelation.

(*f*) This watery solution does not redden infusion of litmus, as acids in general do, but totally destroys its colour. Hence its use in bleaching several vegetable and animal products. It restores the colour of syrup of violets, which has been reddened by other acids. (Nicholson's Journal, xvii. 303.)

(*g*) Sulphuric acid, saturated with this gas, which may be effected by passing the gas through the acid, acquires a strong smell, a yellowish brown colour, smokes when exposed to the air, and has the property of assuming a solid form, by a moderate reduction of its temperature. When distilled, the first product, which is a compound of the two acids, assumes a solid form. It has been called *glacial sulphuric acid.*

(*h*) Sulphurous acid is again converted to the state of sulphuric, by restoring oxygen to it.

1. A mixture of oxygen and sulphurous acid gases, both

perfectly dry, and standing over mercury, is not diminished by remaining in contact with each other during some months; but if a small quantity of water be added, the mixture begins to diminish, and sulphuric acid is formed. The same gases in a state of mixture, driven through a red-hot porcelain tube, afford sulphuric acid. The proportions required for mutual saturation are two measures of sulphurous acid and one of oxygen gas.

2. To a portion of water saturated with this gas, add a little oxide of manganese, a substance that contains much oxygen loosely combined. The pungent smell of the water, and the other characteristics of sulphurous acid, will soon disappear.

3. Sulphurous acid gas is condensed into sulphuric acid by admixture with nitrous gas, and also by oxy-muriatic acid gas.

(*i*) When the temperature of sulphurous acid gas is greatly reduced, by surrounding it with a mixture of snow and muriate of lime, it is changed into a liquid.

(*k*) If sulphurous acid gas and fresh muriate of tin are brought into contact over mercury, the volume of the gas is speedily diminished, sulphur is deposited, and the simple muriate becomes an oxygenized muriate of tin. (Accum.)

Both these acids are susceptible of combination with alkalis.

(*l*) It is decomposed, when submitted to the heat of ignition, in contact with certain combustible bodies. Thus, when a mixture of sulphurous acid and hydrogen gases are driven through a red-hot porcelain tube, the oxygen of the acid combines with the hydrogen, and forms water, and sulphur is obtained in a separate form. The sulphurous acid is decomposed, also, when transmitted over red-hot charcoal; and, as appears from Gay Lussac's experiment, by potassium.

From the testimony of the same chemist we learn that 100 parts of sulphur, to become sulphurous acid, unite with 95 oxygen; and hence this acid must be composed of

Sulphur	51.3
Oxygen	48.7
	100

Dr. Thomson had formerly determined that 100 sulphurous acid consist of

Sulphuric acid	82
Sulphur .	18
	100

And if we admit that 100 sulphuric acid contain 43 of sulphur, then 82 must contain 35.26, which, added to 18, make 53.26 sulphur in 100 of sulphurous acid, numbers not very remote from those of Gay Lussac.

* * *

SECTION IV.

Combination of Sulphuric Acid with Alkalis.

ART. 1.—*Sulphate of Potash.*

THIS salt may be formed by saturating the carbonate of potash with sulphuric acid, and crystallizing the solution. Its properties are the following:

(*a*) It crystallizes in small six-sided prisms, terminated by six-sided pyramids with triangular faces. Its specific gravity, according to Hassenfratz, is 2.0473.

(*b*) It has a bitter taste.

(*c*) It decrepitates, or crackles, when thrown on a red-hot iron, or on red-hot coals, and is volatilized by a strong heat, first running into fusion. By a low red-heat it loses very little of its weight, not more than one and a half or two *per cent.*

(*d*) Water, at 60° of Fahrenheit, takes up only one sixteenth of its weight; but boiling water dissolves one fifth, or by continuing the application of heat even one fourth.

(*e*) One hundred parts contain 30.21 acid, 64.61 alkali, and 5.18 water. The determination of Berthollet differs considerably from this; for, setting aside the water of crystallization, he makes the salt to consist of

51.38 potash
46.62 acid

100

(*f*) This sulphate is decomposed, in high temperatures, by carbon. Mix any quantity of the salt with one fourth of its weight of charcoal finely powdered, and expose the mixture, in a crucible, to a strong heat. The carbon will unite with the oxygen of the sulphuric acid, and will escape in the state of a gas. What remains is a compound, hereafter to be described, of sulphur and potash.

Super-sulphate of Potash.

When to a saturated solution of sulphate of potash in boiling water, we add an excess of sulphuric acid, the first crystals, which are formed, contain a considerable excess of sulphuric acid, not less in the whole, according to Berthollet*, than 55.8 *per cent.* By continuing to evaporate the solution, we obtain successive quantities of crystals, which hold less and less acid in combination. Thus the second set, according to the same chemist, contain only 49.5 *per cent.* of acid.

This salt has an intensely sour taste and a powerful action on blue vegetable colours. One part is soluble in two of water at 60°, and in less than an equal weight at 212°. It is insoluble in alcohol.

ART. 2.—*Sulphate of Soda.*

(*a*) This salt forms regular octahedral crystals, of a prismatic or cuneiform figure; the two terminating pyramids of which are truncated near their basis.

(*b*) It has a more bitter taste than the preceding, and melts more easily in the mouth.

(*c*) It swells upon a heated iron, in consequence of the loss of its water of crystallization, and a white powder is left, amounting to only about 36 parts from 100 of the original salt.

(*d*) By exposure to the atmosphere, it effloresces, and loses weight.

(*e*) It is very soluble in water, three parts of which, at 60° of temperature, dissolve one of the salt; and boiling water dissolves its own weight.

(*f*) It contains *per cent.*, according to Kirwan, 23.52 acid, 18.48 base, 58 water.

(*g*) It is decomposed by charcoal like the preceding salt, and a compound remains of sulphur and soda.

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* *Mémoires d'Arcueil,* ii. 480.

ART. 3.—*Sulphate of Ammonia.*

(*a*) The sulphate of ammonia forms long flattened prisms with six sides, terminated by six-sided pyramids.

(*b*) It slightly attracts moisture from the air.

(*c*) It has a cool bitter taste.

(*d*) Two parts of water, at 60°, take up one of the salt, and boiling water dissolves its own weight. During solution, it produces cold; and also when mingled with powdered ice, or with snow.

(*e*) The sublimed salt has an excess of acid; a portion of the base being expelled by the application of heat.

(*f*) It contains, *per cent.*, 55 acid, 14 ammonia, and 31 water.

(*g*) It liquefies, by a gentle heat, and is volatilized. If a stronger heat be applied, it is decomposed. (See Mr. Hatchett's paper in Philosophical Transactions, 1796, or Davy's Researches.)

(*h*) The pure fixed alkalis, potash, and soda, seize the sulphuric acid, and set at liberty the alkali. Hence a strong smell of ammonia arises on the admixture of pure soda or potash with this salt.

ART. 4.—*Sulphate of Barytes.*

Barytes has a powerful affinity for sulphuric acid; and the combination of these two bodies may be effected with great facility.

(*a*) To a solution of pure barytes, add sulphuric acid. A white precipitate will appear, which is the sulphate of barytes.

(*b*) The same compound is formed, by adding sulphuric acid to carbonate of barytes, or to a solution of muriate or nitrate of barytes.

(*c*) The sulphate of barytes is one of the most insoluble substances that chemistry presents, requiring for its solution 43,000 times its weight of water.

(*d*) Barytes has a stronger affinity than any other body for sulphuric acid.

(*e*) Owing to these properties the solution of pure barytes, and of the nitrate and muriate of barytes, are excellent and very sensible tests of sulphuric acid, and of all its combinations. Let a single drop of sulphuric acid fall into a wine quart of pure distilled water. On adding a few drops of one of the foregoing solutions of barytes, a precipitation will ensue.

(*f*) Sulphate of barytes is decomposed by carbonate of potash. Boil the powdered sulphate with a solution of twice or three times its weight of carbonate of potash. The carbonic acid will pass to the barytes, and the sulphuric to the potash.

(*g*) By this process, carbonate of barytes may be procured, for the purpose of preparing the pure earth, and its various salts, when the native carbonate cannot be had in sufficient abundance. The sulphate is found, in considerable quantities, accompanying lead ore, in Derbyshire and other parts of England, where it is known by the names of cawk, ponderous spar, &c. When applied to the purpose of obtaining the carbonate of barytes, it is to be mixed with three or four times its weight of sub-carbonate of potash, and boiled with a proper quantity of water for a considerable time, in an iron kettle, stirring it, and breaking down the hard lumps, into which it is apt to run, by an iron pestle. It is then to be washed with boiling water, as long as this acquires any taste. On the addition of dilute muriatic acid, a violent effervescence will ensue, and a considerable portion of the earth, probably along with some metals, will be dissolved. To the saturated solution, add a small portion of pure ammonia. This will throw down any metals that may be present; and the barytes may afterwards be precipitated in the state of a carbonate, by a solution of carbonate of potash. Let the precipitated earth be well washed with distilled water; and, if the pure barytes is to be obtained from it, let it be treated as directed, chap. x.

(*h*) Sulphate of barytes is also decomposed when ignited with powdered charcoal, which abstracts the oxygen of the sulphuric acid, and leaves a combination of sulphur and barytes. From this, the barytes may be removed by muriatic acid, as already directed, and the muriatic solution be decomposed by carbonate of potash.

(*i*) The sulphate of barytes, when decomposed by charcoal, affords one variety of solar phosphorus. This phosphorus has been called, from the place where the sulphate is found from which it was first prepared, the Bolognian phosphorus. The native sulphate, powdered after being ignited, and finely sifted, is to be formed into a paste with mucilage of gum arabic, and divided into cylinders or pieces of one fourth of an inch in thickness. These, after being dried in a moderate heat, are to be exposed to the temperature of a wind furnace, placed in the midst of the charcoal. When the fuel is half consumed, it must be replenished, and suffered to burn out. The pieces will be found, retaining their original shapes, among the ashes, from which they may be separated by the blast of a pair of bellows. They must be preserved in a well-stopped phial.

This phosphorus, after being exposed a few minutes to the sun's rays, shines in the dark sufficiently to render visible the dial of a watch. This property is lost by repeated use, in consequence of the oxygenation of the sulphur; but may be restored by a second calcination.

(*k*) Sulphate of barytes, when artificially formed and calcined, contains in 100 parts,

	Base.	Acid.
According to Klaproth . . .	66.55 .	33.45
———— Mr. A. Aikin* .	66.04 .	33.96
———— Mr. J. Thomson†	66.96 .	33.04
———— Berthollet ‡ . .	66.50 .	33.50

The determination of Berthollet, being nearly a mean of the three last, may be considered as sufficiently accurate. The results of other chemists, and especially of Thenard and Chenevix, are so remote from the above, that it is probable some source of fallacy has insinuated itself into their experiments. The native sulphate, according to Klaproth, is composed of one third acid and two thirds base. (Contributions, i. 377.)

ART. 5.—*Sulphate of Strontites.*

I. This salt resembles, very nearly, the sulphate of barytes. It may be formed in a similar manner, by pouring the solution of pure strontites into diluted sulphuric acid, or into the solution of an alkaline sulphate; for it has a stronger affinity than any of the alkalis for sulphuric acid. It is soluble in 3840 parts of boiling water.

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* Nicholson's Journal, xxii. 301.

† Nicholson's Journal, xxiii. 174.

‡ *Mémoires d'Arcueil,* ii.

II. The sulphate of strontites is also found native in considerable quantities; chiefly at Aust Passage, and at other places in the neighbourhood of Bristol. As the native carbonate is now becoming scarce, this compound may be advantageously employed for procuring artificial carbonate of strontites. The process is precisely similar to that already prescribed for decomposing the sulphate of barytes. (Art. 4, *g.*)

According to a considerable majority of the chemists who have analyzed it, it consists of

42	acid
58	strontites
100	

Vauquelin, however, has stated that it is composed of 46 acid and 54 base.

Art. 6.—*Sulphate of Lime.*

I. The sulphate of lime is formed, by adding to the carbonate a sufficient quantity of sulphuric acid; and by gently calcining the residue, to expel the redundancy of the latter acid. It is also found native, in great abundance, under the names of gypsum, plaster of Paris, &c.

II. It has the following properties:

1. It is insipid and free from smell.

2. It is difficultly soluble, requiring 500 times its weight of cold water, or 450 of hot water.

3. It is fusible by a moderate heat. When sulphate of lime, which has been dried at 160° Fahrenheit, is exposed to a low red-heat, it loses 22 *per cent.* of its weight, consisting entirely of water. After calcination, it absorbs water rapidly, and forms a good cement.

4. It is decomposed by carbonates of alkali, a double exchange of principles ensuing. Hence the milkiness which ensues on adding carbonate of potash to most spring waters; the carbonate of lime, which is generated, being less soluble than the sulphate. Hence, also, hard waters, which always contain sulphate of lime in solution, curdle soap, the alkali of which is detached by the sulphuric acid, and the oil is set at liberty.

5. It is decomposed by ignition with charcoal, which separates the oxygen of the sulphuric acid, and leaves a combination of lime with sulphur.

By dissolving 100 grains of calcined sulphate of lime in boiling distilled water, and adding muriate of barytes, I obtained a precipitate, which, when well washed, dried, and calcined in a low red-heat, weighed 175.9. Hence 100 parts of calcined sulphate of lime must contain very nearly

$$\begin{array}{l} 59 \text{ sulphuric acid} \\ 41 \text{ lime} \\ \hline 100 \end{array}$$

This determination very nearly agrees with that of Mr. James Thomson, who has given, instead of the above numbers, 58 acid, and 42 base. (Nicholson's Journal, xxiii. 182.)

ART. 7.—*Sulphate of Magnesia.*

I. When highly concentrated sulphuric acid is suddenly added to fresh prepared and pure magnesia, prodigious heat and vapour are excited, and are accompanied frequently with an extrication of light. This appearance was first observed by Westrumb.

II. If the carbonate of magnesia be added to diluted sulphuric acid, the carbonic acid is expelled, and a solution of sulphate of magnesia is formed, which crystallizes on cooling. Crystals of sulphate of magnesia may also be procured in the shops, under the name of Epsom salt.

III. These crystals have the following properties:

1. They have the form of small quadrangular prisms, surmounted by quadrangular pyramids with dihedral summits.

2. At the temperature of 60°, this salt is soluble in an equal weight of cold water, and in three fourths its weight of boiling water, which thus receives an addition of one fourth of its bulk.

3. It effloresces in the air, and is slowly reduced to powder. When exposed to a low red-heat, it undergoes the watery fusion, but is not volatilized. It loses, however, about one half its weight, which is water of crystallization.

4. One hundred grains of sulphate of magnesia, deprived, by calcination in a low red-heat of its water of crystallization, afforded me 200 grains of sulphate of barytes when precipitated by the muriate of the latter earth. Hence 100 grains of dry sulphate of magnesia are composed of 67 acid and 33 magnesia, and the crystallized salt, supposing it to contain half its weight of water, will consist in 100 parts of

50	water
33.5	sulphuric acid
16.5	magnesia

5. Its solution is precipitated by carbonates of potash and of soda; but not by carbonate of ammonia, unless heat is applied. The carbonate of magnesia of the shops

is prepared, by mixing together concentrated and hot solutions of carbonate of potash and sulphate of magnesia. The sulphate of potash, thus formed, is removed by copious washing with water, and the carbonate of magnesia is then dried. The proportions employed are filtered solutions of equal weights of the two salts, each in its own weight of boiling water. One hundred parts of the desicated sulphate give about 71 of sub-carbonate of magnesia, or about 31.6 of the pure earth.

6. When a dilute solution of carbonate of soda is mixed with a dilute solution of sulphate of magnesia, and the sub-carbonate which is formed, if any, is separated by filtration, crystals of carbonate of magnesia, after some time, shoot in the liquid, containing a larger proportion of carbonic acid, and already described, chap. x. sect. 4. When solution of pure ammonia is added to that of sulphate of magnesia, part of the earth is precipitated. The rest remains in solution, and, by evaporation, a triple salt is formed consisting of sulphuric acid, magnesia, and ammonia, and called *ammoniaco-magnesian sulphate.*

ART. 8.—*Sulphate of Alumine.*

The properties of this salt may be exhibited by those of the common alum of commerce; though, as will afterwards appear, alum is not merely a combination of this earth with sulphuric acid; but is a triple salt, composed either of sulphuric acid, alumine and potash; or of sulphuric acid, alumine, and ammonia. It has the following characters.

(*a*) It has a sweetish astringent taste.

(*b*) It dissolves in water, five parts of which, at 60°, take up one of the salt, but hot water dissolves about three fourths of its weight.

(*c*) This solution reddens vegetable blue colours; which proves the acid to be in excess.

(*d*) When mixed with a solution of carbonate of potash, an effervescence is produced by the uncombined acid, which also prevents the first portions of alkali, that are added to a solution of sulphate of alumine, from occasioning any precipitate.

(*e*) On a farther addition of alkali, the alumine is precipitated.

(*f*) Sulphate of alumine, when heated, swells up, loses its regular form, and becomes a dry spongy mass; but, according to Vauquelin (*Annales de Chimie*, xxxvii. 91), the whole of its acid cannot thus be expelled.

(*g*) The combination of sulphuric acid with alumine is incapable of crystallizing without an admixture of sulphate of potash, which forms a constituent of all the alum of commerce. According to Vauquelin, 100 parts consist of 49 dry sulphate of alumine, 7 sulphate of potash, and 44 water. Or 100 grains are composed of

Acid . .	30.52
Alumine .	10.50
Potash . .	10.40
Water . .	48.58
	100.00

The acid, perhaps, in the above estimate, is rated a little too low; for alum, when precipitated by muriate of barytes, gives, as nearly as possible, an equal weight of sulphate of barytes, 100 grains of which contain 33.5 of sulphuric acid, the quantity, therefore, present in 100 parts of alum.

Common alum may be brought to the state of a saturated salt by boiling its solution with pure and fresh precipitated alumine. The compound has the form of a white and tasteless powder, very sparingly soluble in water, not crystallizable, and not fusible when heat is applied.

(*h*) Alum is decomposed by charcoal, which combines with the oxygen of the sulphuric acid, and leaves the sulphur attached to the alumine. A combination of alumine, sulphur, and charcoal, forms the *pyrophorus of Homberg*. To prepare this, equal parts of powdered alum and brown sugar are melted over the fire, and are kept stirring till reduced to dryness. The mixture, when cold, is to be finely powdered, and introduced into a common phial, coated with clay, to which a glass tube, open at each end, is to be luted, to allow the escape of the gases that are produced. The phial must then be set in the fire, surrounded by sand, in a crucible. Gas will issue from the open end of the tube, and may be inflamed by a lighted paper. When this ceases to escape, the crucible may be removed from the fire, and a little moist clay pressed down upon the open end of the tube, to prevent the access of air to the contents of the phial. When cold, the tube may be removed, and a cork substituted in its place. The principal difficulty in the process, is to stop it precisely at the period, when the pyrophorus is formed; for if the heat be continued longer, the sulphur will be sublimed, and the preparation spoiled.

The pyrophorus thus formed is a black and light powder, which instantly takes fire when poured out of the bottle into the air, and inflames suddenly in oxygen gas. Sulphate of potash appears to be essential to its production, and hence the sulphuret of potash is a necessary ingredient. From the recent discoveries of Mr. Davy, it appears not improbable that this pyrophorus may contain sulphuret of potassium. The principal part of the phenomena, however, are owing to the combustion of an extremely light and finely divided charcoal.

Art. 9.—*Sulphate of Glucine.*

Glucine combines readily with sulphuric acid, both in its pure and carbonated state. The resulting salt is extremely soluble; insomuch that, when evaporated, it assumes the form of a syrup, without crystallizing. Its taste is sweet, and rather astringent. It is decomposed entirely in a high temperature, the earth being left in a state of purity. It is also destroyed by ignition with charcoal. It does not yield its earthy ingredient to any of the acids; but is decomposed by all the alkalis and earths, alumine excepted.

Art. 10.—*Sulphate of Zircon.*

To effect the combination of zircon with any acid, this earth should be fresh precipitated; for, after being dried, it enters with difficulty into union.

The salt, resulting from the union of sulphuric acid with zircon, is white, insoluble, and without taste. It is decomposed by a high temperature, which expels the acid, and leaves the zircon pure. It is not changed by other acids, but yields its sulphuric acid to the alkalis, and to most of the earths.

ART. 11.—*Sulphate of Yttria.*

Sulphuric acid readily dissolves yttria, and caloric is evolved during the process. As the solution goes on, the sulphate crystallizes in small brilliant grains, which have a sweetish taste, but less marked than that of the sulphate of glucine. Their colour is a light amethyst red. They require 30 parts of water, of the temperature of 60°, for solution, and give up their acid when exposed to a high temperature. They are decomposed by oxalic acid, prussiate of potash, infusion of galls, and phosphate of soda.

SECTION V.

Sulphites.

I. THE combination of sulphurous acid with alkaline and earthy bases, may be effected by passing the gas, as it proceeds from the materials (sect. 2, *h*), through the base, dissolved or diffused in water. An intermediate vessel may be placed, as represented, fig. 30 and 31, to condense any sulphuric acid that may pass over; and the solution of the alkali or earth may be contained in a bottle with two necks. Pure potash, soda, or ammonia, are readily kept in solution: but barytes or strontites must be dissolved in boiling water; and the bottle containing them must be surrounded with hot water, while the gas is transmitted through the solution. The solution, when saturated with gas, may be evaporated; and this is best done in an alembic, covered with its capital, because the salts of this class are changed by the action of the atmosphere.

II. The sulphites have no peculiarly interesting properties, that can entitle them to minute and specific description, in a work devoted solely to the students of chemical science. I shall enumerate, therefore, only the principal ones; and refer, for farther information, to the 2d and 24th volumes of the *Annales de Chimie*, and to Dr. Thomson's memoir in Nicholson's Journal, vi. 94. Their general qualities are the following:

1. They have a disagreeable taste and smell, resembling that of the fumes of burning sulphur.

2. When heated, they emit sulphurous acid and water, and then sulphur, which, on the application of an inflamed substance, takes fire, and burns violently.

3. Exposed to the atmosphere, in a state of solution, they absorb oxygen, and are slowly changed into sulphates.

4. When added to nitric acid, red fumes arise, and the salts become sulphates. Oxygenized muriatic acid produces the same effect. Concentrated sulphuric acid expels sulphurous acid gas, which may be collected over mercury.

5. When perfectly pure, sulphites are not precipitated by a solution of pure barytes or strontites, or by any of the salts with base of either of those earths. If a precipitation ensue, it indicates the presence of a portion of sulphate.

SULPHITE OF POTASH crystallizes in the form of lengthened rhomboidal plates, or of needles, which have sometimes a slight yellowish tinge. It has a pungent and sulphurous taste, and is soluble in an equal weight of cold, or in less than an equal weight of boiling water. At the temperature of 300° Fahrenheit, it loses only about 2 *per cent.;* but when more strongly heated, the salt is decomposed, and sustains a loss of about 22 *per cent.*, of which 15 are sulphurous acid, 5 sulphur, and 2 water. When thrown into a red-hot crucible, a blue flame arises from it, and its weight is diminished in the proportion which has just been stated. When its solution is exposed to the air, the salt slowly attracts oxygen, and is converted into sulphate of potash. This change goes on more rapidly in oxygen gas; or when it is mixed with any substance holding oxygen in loose combination, as nitric or oxymuriatic acid. It contains, in 100 parts, from Dr. Thomson's experiments,

43.5 acid
54.5 potash
2 water
———
100

SULPHITE OF SODA forms compressed tetrahedral prisms with dihedral summits. It requires for solution less than its own weight of boiling water, or four times its weight of cold water. It effloresces in the air, but much less perfectly than the sulphate of soda. It is composed, according to Dr. Thomson, of

31 acid
18 soda
51 water
———
100

SULPHITE OF AMMONIA crystallizes in hexahedral prisms terminated by pyramids with the same number of sides, or in rhomboidal prisms with trihedral summits. It is soluble in an equal weight of cold water, or in less than an equal weight of boiling water. It deliquiates in the atmosphere, and absorbing oxygen is changed into sulphate of ammonia, which becomes dry.

SULPHITE OF BARYTES, like the salt formed by uniting the same base with sulphuric acid, is almost insoluble. When united with an excess of sulphurous acid, however, (which may be done by dissolving the white powder, that is first formed, in liquid sulphurous acid,) it forms a crystallizable salt, still of sparing solubility. The solution of this salt may be advantageously used to purify the solution of any sulphite from sulphuric acid, which it precipitates in the state of an insoluble sulphate of barytes.

Sulphite of lime is also insoluble, but may be crystallized by being first dissolved in liquid sulphurous acid. In this state it requires 800 parts of water for solution.

Sulphite of magnesia differs from the sulphate of this earth in being vastly less soluble in water, of which it requires 20 parts at the common temperature. Hot water takes up a farther portion, which is deposited on cooling.

Sulphite of alumine is not crystallizable, but has the form of a white soft powder, insoluble in water, but soluble in an excess of acid. It consists, according to Dr. Thomson, of

32	acid
44	alumine
24	water
100	

SECTION VI.

Binary Compounds of Sulphur.—1st, With Alkalis—2d, With Hydrogen.

ART. 1.—*Sulphurets.*

I. THE combination of sulphur, with the fixed alkalis and earths, may be formed by fusing together, in a covered crucible, equal parts of sulphur and the respective alkali or earth with which it is to be combined. The compound is to be poured, when in a state of fusion, into an iron dish, or upon a smooth stone; and preserved in a well-closed bottle. These compounds have, for the most part, a reddish brown or liver colour; and hence were formerly called *hepars* or *livers* of sulphur. They may be formed, also, by fusing the alkaline or earthy sulphates with powdered charcoal; but in this case, the sulphuret generally contains a portion of carbonic acid, and also of charcoal.

Sulphuret of lime, when intended for the purpose of *Canton's phosphorus*, is best prepared, by placing, in a crucible, alternate strata of calcined and pounded oyster-shells and sulphur; exposing them to a moderate heat; and then confining them in a bottle with ground stopper. Or, according to the original directions of Canton, three parts of oyster-shells, calcined for about an hour and pulverised, are to be mixed with one of sulphur, and rammed tightly into a crucible, which is to be kept red-hot for about an hour. The compound, when cold, has the properties already assigned to the Bolognian phosphorus.

II. Sulphurets have the following properties:

(*a*) In a moist state they emit an offensive smell, and have a disagreeable taste.

(*b*) They change to green the colour of violets, in the same manner as uncombined alkalis.

(*c*) They blacken the skin, silk, and other animal substances.

(*d*) They are decomposed by all acids. Into a Nooth's machine put a weak solution of sulphuret of alkali, and pass through it streams of carbonic acid gas. In the course of a few days, the sulphur will be precipitated, and a carbonate of alkali will be obtained. This decomposition ensues, instantly, on adding, to a solution of sulphuret, any of the stronger acids, as the sulphuric, nitric, or muriatic; and we obtain a compound of the alkali with the respective acid employed.

(*e*) The liquid sulphurets absorb oxygen gas. This may be shown by the experiments already described (chap. v.). If the change thus effected be examined, it will be found that the oxygen has combined with the sulphur, and formed sulphuric acid, which, uniting with the alkali, has composed the sulphate of potash.

(*f*) If dilute muriatic acid be poured on the solution of sulphuret of potash or soda, a violent effervescence will ensue, and a very offensive gas be disengaged. This gas may be collected over water. It is termed sulphuretted hydrogen gas.

ART. 2.—*Sulphuretted Hydrogen Gas.*

I. This gas may be obtained in the foregoing manner, or from a mixture of three parts by weight of iron filings and one of sulphur, previously melted together in a covered crucible. A portion of the fused mass may be put into a gas bottle, and diluted sulphuric or muriatic acid poured on it, which will extricate the sulphuretted hydrogen gas.

Sulphuretted hydrogen gas may be formed, also, by simply heating sulphur in a retort along with hydrogen gas. By thus dissolving sulphur, the bulk of the hydrogen gas is not at all altered.

When metals are heated with sulphur, at the moment of combination sulphuretted hydrogen gas is evolved.

II. Its properties are the following:

(*a*) Its smell is extremely offensive, resembling that of putrefying eggs.

(*b*) It is inflammable, and burns either silently or with an explosion, according as it is previously mixed, or not, with oxygen gas or atmospheric air. During this combustion, water results from the union of the hydrogen with the oxygen, and sulphuric and sulphurous acids from that of the oxygen and sulphur.

When three parts of sulphuretted hydrogen are mingled with two of nitrous gas, the mixture, on being inflamed, burns with a yellowish green flame.

(*c*) It tarnishes silver, mercury, and other polished metals, and instantly blackens white paint.

(*d*) It is absorbed by water, which takes up its own bulk, and thus acquires the peculiar smell of the gas. It is this gas which gives to the Harrogate, and some other natural waters, their disagreeable odour.

(*e*) Water, saturated with this gas, reddens the infusion of violets, in this respect producing the effect of an acid. From this and other properties, some of the German chemists have proposed for it the name of *hydrothionic acid.*

(*f*) Water impregnated with sulphuretted hydrogen, when exposed to the atmosphere, becomes covered with a pellicle of sulphur. Sulphur is even deposited when the water is kept in well-closed bottles.

(*g*) On the addition of a few drops of nitric or nitrous acid to the watery solution, sulphur is instantly precipitated. In this case the oxygen of the acid combines with the hydrogen of the gas, and the sulphur is separated.

(*h*) This gas, as will afterwards appear, is decomposed by mixture with oxygenized muriatic acid gas; and sulphur is precipitated.

(*i*) It is decomposed also when kept in a state of mixture with atmospheric air, the oxygen of which combines with the hydrogen, and forms water, while the sulphur is precipitated.

(*k*) A succession of electric explosions throws down sulphur from it, and the volume of the gas remains unaltered.

(*l*) When sulphuretted hydrogen gas and sulphurous acid gas are mingled together, the hydrogen of the former unites with the oxygen of the latter, and the sulphur of both is precipitated.

(*m*) It is decomposed when passed over ignited charcoal, and is converted into carburetted hydrogen gas.

(*n*) Sulphuretted hydrogen, according to Thenard, is composed, in 100 parts, of 29 hydrogen, and 71 sulphur. One hundred cubic inches weigh 38.17. Mr. Kirwan (Philosophical Transactions, 1786) states the weight of the same quantity, at the temperature of 60°, and under 30 inches pressure, at only 34.286 grains. From his experiments, also, 100 grains should consist of

89	sulphur
11	hydrogen
100	

This result differs so considerably from Thenard's, that one or both may be suspected of inaccuracy.

From the late experiments of Mr. Davy (Philosophical Transactions), it appears that sulphuretted hydrogen contains a volume of hydrogen equal to its own. He considers 100 cubical inches as weighing 35 grains; from which it may be concluded that 100 parts contain about 6,5 hydrogen and 94.5 sulphur.

(*o*) It precipitates, both in the state of a gas and of watery impregnation, all metallic solutions, excepting those of iron, nickel, cobalt, manganese, titanium, and molybdena.

(*p*) It is copiously absorbed by alkalis, and by all the earths, excepting alumine and zircon. These alkaline and earthy combinations are termed hydro-sulphurets.

Some very curious phenomena were observed by Mr. Davy in his experiments on the action of potassium on sulphuretted hydrogen, which seem to show that either the gas itself, or the sulphur dissolved in it, may contain oxygen as one of its elements.

Potassium burns brilliantly when heated in sulphuretted hydrogen, and its inflammability is destroyed in the process, and it becomes a sulphuret, which from its properties would appear slightly oxygenated.

Art. 3.—*Hydro-Sulphurets.*

In its union with alkaline and earthy bases, sulphuretted hydrogen seems to perform, in a great measure, the functions of an acid; and presents, therefore, an important exception to the doctrine of acidification; for, in this instance, a body, which, if it contain oxygen at all, contains it only in very minute proportions, possesses some of the most important characters of an acid, *viz.* the property of changing vegetable blues to red, and of uniting with alkalis.

I. The hydro-sulphurets may be formed, by transmitting sulphuretted hydrogen gas, as it issues from the materials that afford it, through a solution of the alkaline or earthy base. Or the base, when insoluble, must be kept suspended in water by mechanical agitation.

II. The hydro-sulphurets have several qualities common to the whole genus.

1. They are all soluble in water, and the recent solution is colourless. By exposure to the air, however, it first becomes green, or greenish yellow, and deposits sulphur on the sides of the vessel. The glass bottle, containing the solution, becomes black on its inner surface, in consequence of the combination of sulphur with the oxide of lead contained in the glass.

2. After long exposure to the atmosphere, the solution entirely loses its colour, and again becomes perfectly

limpid. When examined, it is found to consist of a combination of sulphuric acid with the peculiar base of the hydro-sulphuret. This is owing to the absorption of oxygen, which all hydro-sulphurets take from the atmosphere; the formation of a sulphite; and the conversion of this, by farther oxygenation, into a sulphate. Hence, when confined in contact with a limited quantity of atmospherical air, hydro-sulphurets effect a diminution of volume; and may be employed to ascertain its proportion of oxygen. They entirely absorb pure oxygen gas.

3. When an acid is poured on any of the hydro-sulphurets, the sulphuretted hydrogen gas is disengaged and no sulphur is deposited. This non-precipitation of sulphur distinguishes hydro-sulphurets, both from sulphurets, and hydrogaretted sulphurets. The acid employed should be one which strongly retains its oxygen, such as the sulphuric or muriatic; otherwise it will probably be decomposed. A hydro-sulphuret, which has been a few days exposed to the air, yields, by this treatment, sulphurous acid gas, along with sulphuretted hydrogen.

4. The solutions of hydro-sulphurets precipitate all metallic solutions. They also precipitate alumine and zircon from their solutions, but no other earths.

5. The hydro-sulphurets are, for the most part, susceptible of a regularly crystallized form.

HYDRO-SULPHURET OF POTASH forms large transparent crystals not unlike in size those of sulphate of soda, but having the shape of four-sided prisms, acuminated with four planes; or of six-sided prisms acuminated by six planes. It is deliquescent, and affords a thick syrupy liquor, which gives a green colour to the skin. It dissolves readily in water and alcohol, producing cold. When dilute acids are added to the solution, a brisk effervescence is excited, but no sulphur is deposited.

HYDRO-SULPHURET OF SODA is a compound, which derives some importance from its being produced along with carbonate of soda, in several processes for decomposing the sulphate of soda*. It is transparent at first, colourless, and crystallized in four-sided prisms acuminated by four planes. It has an acrid and alkaline taste, which soon becomes very bitter. Its solution is colourless, but tinges the skin or paper green. It effervesces briskly with acids, and no sulphur is deposited unless the nitric or oxymuriatic acids are added, which oxidate the hydrogen, and throw down sulphur.

Vauquelin has proposed to distinguish these two sulphurets, which so closely resemble each other, by the following test. The hydro-sulphuret of potash, when added to a solution of sulphate of alumine, occasions a crystallization of alum; but that of soda has no such action.

HYDRO-SULPHURET OF AMMONIA may be formed by the direct mixture of sulphuretted hydrogen and ammoniacal gases over mercury. But, for all practical uses, it is better prepared by putting a solution of pure ammonia into the middle of a Nooth's machine, and passing through it streams of sulphuretted hydrogen gas, till the liquid acquires a yellowish colour. In this state, it constitutes the hepatized ammonia, so strongly recommended as a remedy of diabetes.

HYDRO-SULPHURETS OF BARYTES AND STRONTITES are crystallized salts, having a white silky lustre, and readily soluble in water*.

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* *Annales de Chimie*, lxiv. 59.

HYDRO-SULPHURET OF LIME, formed by transmitting sulphuretted hydrogen through water, in which lime is kept mechanically suspended, composes a crystallizable salt, soluble in water; and having the general properties of hydro-sulphurets.

ART. 4.—*Super-Sulphuretted Hydrogen, and Hydroguretted Sulphurets.*

Super-sulphuretted hydrogen is obtained, when hydrosulphuret of potash is poured, by little and little, into muriatic acid. A very small portion only of gas escapes; and while the greater part of the sulphur separates, one portion of it combines with the sulphuretted hydrogen; assumes the appearance of an oil; and is deposited at the bottom of the vessel. Or dissolve sulphur in a boiling solution of pure potash; and into a phial, containing about $\frac{1}{3}$ its capacity of muriatic acid, of the specific gravity 1.07, pour about an equal bulk of the liquid hepar. Cork the phial, and shake it; the hydroguretted sulphur gradually settles to the bottom. Its properties are the following:

1. If gently heated, sulphuretted hydrogen gas exhales from it; the super-sulphuret loses its fluidity; and a residue is left, consisting merely of sulphur.

2. It combines with alkalis and earths; and forms with them a class of substances called hydroguretted sulphurets.

There are, therefore, three distinct combinations of sul-

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[*] See *Annales de Chimie,* lxii. 181.

phur and its compounds with alkalis and earths. The first consist, simply, of sulphur, united with an alkaline or earthy base, and are called strictly *sulphurets*. The second are composed of sulphuretted hydrogen, united with a base, and are called *hydro-sulphurets*. The third contain super-sulphuretted hydrogen, attached to a base, and constitute *hydroguretted sulphurets*.

The sulphurets can exist, as such, only in a dry state; for the moment they begin to dissolve in water, a decomposition of that fluid commences; sulphuretted hydrogen is formed; and this, uniting with an additional dose of sulphur, composes super-sulphuretted hydrogen. This last, uniting with the base, forms an hydroguretted sulphuret. Hence the sulphurets are partly changed, by solution, into hydroguretted sulphurets. The solution, however, still contains a proportion of sulphur so considerable, that we may consider it as composed, in part, of liquid sulphuret. Hence the affusion of an acid throws down a large quantity of sulphur. A distinguishing character, also, of solutions of this kind is that sulphur is precipitated by passing through them sulphuretted hydrogen gas.

According to Proust, if red oxide of mercury be added to solutions of the kind which have just been described, the sulphuretted hydrogen is removed, and what remains is a pure liquid sulphuret, from which acids precipitate sulphur alone, without any effervescence.

II. The hydroguretted sulphurets are also formed by boiling, along with a sufficient quantity of water, the alkaline, or earthy base, with flowers of sulphur. Thus a solution of pure potash, pure soda, or of barytes or strontites, may be changed into an hydroguretted sulphuret. To prepare this compound, with base of lime, the powdered earth, mixed with sulphur, may be boiled with a proper quantity of water, and the solution filtered. The hydroguretted sulphuret of ammonia (which base cannot, in strictness, owing to its liquid form, compose a true sulphuret) may be prepared as follows: Mix together, in a mortar, three parts of hydrat of lime, one part of muriate of ammonia, and one of flowers of sulphur. Introduce the mixture into a retort, and apply a receiver. Begin the distillation with a gentle heat. The first liquor, that comes over (long known under the name of *Boyle's Fuming Liquor*), has a light yellow tinge, and emits fumes; the second has a deeper colour, and is not fuming. When the latter begins to appear, the fire may be raised.

Another method of forming, by a very simple process, the hydroguretted sulphurets, consists in digesting, in a gentle heat, a hydro-sulphuret with powdered sulphur, an additional portion of which is thus dissolved by the sulphuretted hydrogen.

Hydroguretted sulphurets have the following properties:

1. They have a deep greenish-yellow colour; an acrid and intensely bitter taste; and an excessively offensive smell.

2. They deposit sulphur when kept in close vessels; become much more transparent and lighter coloured; and less offensive to the smell.

3. They rapidly absorb oxygen from the atmosphere, and from oxygen gas. Hence their employment in eudiometry. (See chap. v. sect. 4.)

4. On the addition of dilute sulphuric, or muriatic, or of certain other acids, they are decomposed. Sulphuretted hydrogen gas is evolved, and sulphur is precipitated.

HYDROGURETTED SULPHURETS OF POTASH AND OF

SODA differ very little from each other. They may be formed by boiling solutions of pure potash or soda with flowers of sulphur. When very concentrated, they have a deep reddish brown colour, a nauseous taste, a disagreeable odour, and a soapy feel, tinging the cuticle black. When exposed to the air, a thin pellicle of sulphur forms upon their solutions, which, by sufficiently long exposure, are changed into sulphates. When an acid is suddenly added, sulphur is thrown down, which becomes, when washed with sufficient water and dried, almost white, and constitutes what has been called precipitated sulphur, milk of sulphur, or magistery of sulphur.

HYDROGURETTED SULPHURET OF AMMONIA may be formed by the process already described, or by digesting hydro-sulphuret of ammonia with sulphur, a portion of which is dissolved.

HYDROGURETTED SULPHURET OF BARYTES is obtained by boiling crystals of pure barytes with one fourth their weight of sulphur and sufficient water. Two compounds are formed, *viz.* an hydroguretted sulphuret, which has a red colour and remains in solution; and colourless crystals, which are supposed to be a hydrosulphuret of barytes. Strontites forms similar compounds.

HYDROGURETTED SULPHURET OF LIME is formed by boiling hydrate of lime with a third its weight of sulphur and ten times its weight of water. The compound has a deep orange colour, and is of importance from its application to eudiometrical purposes.

Liquid Sulphuretted Hydrogen.

The existence of a liquid carburet of sulphur, or compound of carbon and sulphur, was for some time acknowledged, chiefly on the authority of Clement and Desormes, who gave that name to a substance originally discovered by Lampadius. It has since however been attentively investigated by Berthollet junior*, and has been proved by him to be, as Lampadius originally supposed it, a combination of hydrogen and sulphur, without any proportion of carbon essential to its constitution.

To prepare this substance, a coated porcelain tube, partly filled with small pieces of charcoal, may be disposed in a furnace as represented fig. 40, *c c*, one end being placed higher than the other. To this end may be adapted a glass tube, open at both ends, containing small bits of sulphur; and, to the other end, by means of an adopter, is to be fixed a curved tube, passed into water contained in a two-necked bottle. The part of the tube, containing the charcoal, may then be made red-hot; and, when this happens, the bits of sulphur are to be pushed forwards, one by one, by means of a wire, carefully excluding air. As soon as the sulphur comes into contact with the charcoal, bubbles of gas will be produced in great abundance, and a vapour will appear, which will condense, under the water in the bottle, into a liquid. This liquid has several remarkable properties.

1. It is nearly transparent like water; its specific gravity is about 13 to 10; and, when shaken in a phial, it adheres, in drops, to the side like oil. Its smell is similar to that of sulphuretted hydrogen, but has more pungency. It is extremely volatile; and, when dropped upon the hand, produces a sense of cold by its rapid evaporation. When admitted to a confined portion of atmospheric air, it enlarges the bulk of the air; and the mixture burns with a

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[*] *Annales de Chimie,* lxi. 127.

blue flame, but without explosion. If oxygen gas be substituted in this experiment, the detonation is very violent.

2. By its combustion, it is entirely consumed, without the production of carbonic acid. Oxy-muriatic acid throws down sulphur. With alkaline solutions, it forms hydroguretted sulphurets. By distillation at a low temperature, it yields sulphuretted hydrogen gas; and a compound remains in the retort, which becomes solid by cooling.

There can be no doubt, then, that this singular fluid is composed solely of hydrogen and sulphur; but the precise proportions remain to be determined by more accurate experiments. It appears, indeed, from the observations of Berthollet junior, that hydrogen and sulphur are capable of uniting in several proportions, and that the density of the compound is regulated, probably, by the quantity of hydrogen which it contains. The greater the proportion of hydrogen, the more elastic is the compound; while the predominance of sulphur gives it a form more or less approaching to solidity.

CHAPTER XIII.

COMBINATION OF NITROGEN WITH OXYGEN, CONSTITUTING NITRIC ACID,—NITROUS GAS,—NITROUS OXIDE,—AND COMPOUNDS OF NITRIC ACID WITH ALKALIS.

WHEN nitrogen and oxygen gases are mingled together, in whatsoever proportions we employ, no combination ensues. The result is a simple mixture of the two gases, which do not, like inelastic fluids, separate on standing, but remain, on the contrary, diffused through each other for any length of time. This is the case with the air of our atmosphere; and it is fortunate that such a provision of nature exists, since the mixture contains the elements of several combinations which, if actually formed, would be fatal to animal and vegetable life. When either one or both of these elements, however, in any mixture of the two, is in a condensed state, or deprived of part of that caloric which keeps their particles at a distance from each other, they unite and form compounds, distinguished by very striking properties. According to the proportions in which the oxygen and nitrogen exist in these compounds, their qualities undergo a remarkable variation, so that from two simple or elementary bodies, variously united, we have several compounds, totally unlike each other in external qualities, as well as in their chemical relations.

Before describing the compounds of oxygen and nitrogen in detail, it will contribute to perspicuity to take a general view of the whole. Some of them exist essentially in an aëriform state, and are capable of uniting with water and other liquids in only small proportion. Others, again, combine with water to such an extent, that the liquid form is the only one under which they occur to our observation. When entirely deprived of water, however, they are all gaseous bodies.

In a series of the compounds of nitrogen, founded on their proportion of oxygen, they occupy (excluding atmospherical air from the number) the following order, the last containing the largest proportion of oxygen—nitrous oxide—nitric oxide or nitrous gas—nitrous acid or nitrous vapour—nitric acid—and oxy-nitric acid*. The two first are sparingly soluble in water; but the three last unite with it largely, and form liquid compounds of decidedly acid taste and quality.

The following table exhibits the composition of three of these bodies, the calculation being founded on the experiments of Mr. Davy.

	Proportion of Elements by Weight.		Proportion by Measure.	
	nitrogen.	oxygen.	nitr^{n} gas.	ox. gas.
Nitrous oxide . .	63.30	36.70	100	49.5
Nitrous gas . . .	44.05	55.95	100	108.9
Nitric acid . . .	29.50	70.50	100	204.7

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* I include this compound on the authority of Mr. Dalton, whose arguments for its existence will appear in the second part of his "New System of Chemical Philosophy," now in the press. They are derived, chiefly, from the proportions in which nitrous and oxygen gases unite,—proportions resulting from his own experiments and differing considerably from those assigned by Gay Lussac.

That the numbers in the fourth column are not precisely simple multiples of those in the third, Gay Lussac imputes to the unavoidable inaccuracies attendant on delicate processes for determining the constitution of gaseous bodies. In one instance, this was proved experimentally; for M. Berard, by the combustion of potassium in 100 measures of nitrous oxide gas, obtained exactly 50 measures of nitrogen. Hence, if the table were corrected, it would stand as follows:

	Measures of Nitrogen	 Oxygen.
Nitrous oxide consists of . .	100	50
Nitrous gas	100	100
Nitric acid	100	200

Or (which is the same thing) we may consider nitric acid as constituted of 200 measures of nitrous gas, and 100 measures of oxygen.

To these three compounds, Gay Lussac has added a fourth, which he thinks entitled to rank as a distinct combination. It will be described, hereafter, under the name of nitrous acid vapour. According to his views, it consists, by measure of

Nitrous gas.	Oxygen gas.
300	100

And as nitrous gas itself contains half its bulk of oxygen gas, nitrous acid vapour must be formed by the union of 150 measures of nitrogen with 250 measures oxygen gases. And it will, therefore, consist by weight of

34.507	nitrogen
65.493	oxygen.

In all these compounds, the elements are in a state of condensation, except in nitrous gas, in which the nitrogen and oxygen are supposed by Gay Lussac to be precisely in the same state, as in nitrogen and oxygen gases. With respect to the other compounds, the condensation, he apprehends, is exactly that of the oxygen gas. For example, as 100 measures of nitrogen gas and 50 measures of oxygen gas go to form nitrous oxide, the condensation is 50 measures; and we obtain only 100 of nitrous oxide. On the same principle 100 measures of nitrogen gas and 200 of oxygen gas should form only 100 of nitric acid gas.

The observation of these remarkable coincidences, between the proportion of elements in the compounds of nitrogen with oxygen, and that of numbers which are simple multiples of each other, has led Gay Lussac to examine how far the law applies to other gaseous bodies. In many other instances, he has found that the facts concur to establish the general conclusion, that aëriform bodies unite in measures, which are simple multiples of each other; and that when all inaccuracies are avoided in the experiments, they do not combine in any proportions intermediate between these. The contraction of volume, produced by their combination, follows also, he supposes, whenever it occurs, a general law; in the compounds of oxygen for example, the condensation is that of the oxygen gas.

These deductions, however, require, for their establishment, a broader groundwork of facts than the state of the science at present supplies. They derive, it may be observed, additional probability from their coincidence with the views of Mr. Dalton, to which, indeed, they bear so strong a resemblance, that it is scarcely possible to suppose they were not suggested by the principles developed in his New System of Chemical Philosophy.

SECTION I.

Nitric Acid.

I. THE direct combination of nitrogen and oxygen, affording a decisive synthetic proof of the nature of this acid, may be effected by passing the electric shock through a mixture of nitrogen and oxygen gases. The experiment is an extremely laborious one, and requires, for its performance, a powerful electrical machine; but those who are disposed to repeat it, may proceed as follows:

Let the tube, fig. 29, be filled with, and inverted in, mercury. Pass into it a portion of atmospherical air, or an artificial mixture of nitrogen and oxygen gases, in the proportion of one of the former to two of the latter.—Let an iron wire, lengthened out with one of platina, be introduced within the tube, so that the latter metal only may be in contact with the mixed gases; and let the end of this wire be distant about one fourth of an inch from the extremity of the upper one. When the apparatus is thus disposed, pass a series of electric sparks or shocks through the gases for several hours. The mixture will be diminished in bulk; will redden litmus-paper when enclosed in it; and will exhibit distinctly the smell of nitrous acid. If the experiment be repeated, with the addition of a few drops of solution of potash, in contact with the gases, we shall obtain a combination of nitric acid with potash.

This interesting experiment on the generation of nitric acid we owe to Mr. Cavendish, who discovered the fact in the year 1785. (Philosophical Transactions, lxxv.) The proportions, which he found to be necessary for mutual saturation, were five parts of oxygen gas and three of common air, or seven parts of oxygen gas to three of nitrogen gas. No evolution either of light or heat attends this combination, which is very slowly and gradually effected.

Another synthetic proof of the production of nitric acid will be stated under the article *nitrous gas*. It is furnished by the generation of nitrous gas, and ultimately of nitric acid, when ammonia is brought into contact with the black oxide of manganese.

For all purposes of utility or experiment, however, nitric acid is prepared in a different manner, *viz.* by the decomposition of nitrate of potash, in a way which will presently be described

II. The analysis of the acid may be obtained by driving its vapour through a red-hot porcelain tube (fig. 40, *cc*), and receiving the generated gases, which prove to be a mixture of nitrogen and oxygen gases.

III. The nitric acid has the following properties:

(*a*) It is heavier than water, in the proportion of 1.5 to 1.

In its heaviest form, however, it still contains a portion of water. Pure nitric acid may be considered as a gaseous body, of the specific gravity, compared with common air, of 2440: one hundred cubic inches at 55° Fahrenheit and under 30 inches pressure, weigh, according to Mr. Davy, 76 grains; or corrected to the temperature of 60° Fahrenheit, they weigh 75.21 grains. The liquid acid consists of this gas condensed by water, of which it contains various proportions, as is shown by the following table of Mr. Davy.

TABLE

Of the Quantities of True Nitric Acid in Solutions of different Specific Gravities.

100 Parts Nitric Acid of Specific Gravity		True Acid*	Water
1.5040	contain	91.55	8.45
1.4475		80.39	19.61
1.4285		71.65	28.35
1.3906		62.96	37.04
1.3551		56.88	43.12
1.3186		52.03	47.97
1.3042		49.04	50.96
1.2831		46.03	53.97
1.2090		45.27	54.73

(*b*) Pure gaseous nitric acid, according to the same ingenious chemist, is composed in 100 grains, of $29\frac{1}{2}$ azote, and $70\frac{1}{2}$ oxygen. This is probably a nearer approximation to the truth than any other that has been hitherto stated, and differs but little from the proportions deducible from the experiments of Cavendish, *viz.* 27.8 nitrogen to 72.2

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* The quantities of oxygene and azote in any solution, may be thus found.—— Let $a =$ the true acid, x the oxygen, and y the azote.

$$\text{Then } x = \frac{2.389\,a}{3.389} \text{ and } y = \frac{a}{3.389}$$

oxygen. Lavoisier had previously obtained results which are considerably remote from these, and which make the nitric acid to consist of 20 nitrogen and 80 oxygen.

(*c*) The watery solution is perfectly limpid and colourless.

(*d*) It gives a yellow stain to the skin.

(*e*) It boils at 248° Fahrenheit, and may be distilled over, without any essential change.

(*f*) Nitric acid may be frozen by the application of a sufficiently low temperature. Like sulphuric acid, there is a certain point of density, at which it most readily congeals. Mr. Cavendish has described this, not by its specific gravity, but by the quantity of marble which it is capable of dissolving. When it takes up $\frac{418}{1000}$ths of its weight, the acid freezes at 2° below 0 Fahrenheit. When considerably stronger and capable of dissolving $\frac{561}{1000}$ths, it required cooling to -41.6; and when so much diluted as to take up only $\frac{272}{1000}$ths, it did not congeal till cooled to -40.3. (Philosophical Transactions, 1788.)

(*g*) It absorbs moisture from the atmosphere; and hence it increases in weight, and diminishes in specific gravity, by exposure to the air.

(*h*) When two parts of the acid are suddenly diluted with one of water, an elevation of temperature is produced to about 112° Fahrenheit. When more water is added to this diluted acid, its temperature is reduced. Snow or ice added to the acid is instantly liquefied, and an intense degree of cold is produced.

(*i*) It becomes coloured by exposure to the sun's light, passing first to a straw colour, and then to a deep orange. This effect is produced by the union of the light of the sun with oxygen, in consequence of which the proportion of the acidifying principle to the nitrogen is diminished.

By exposing it to the sun's rays in a gas bottle, the bent tube of which terminates under water, oxygen gas may be procured.

(*k*) This acid retains its oxygen with but little force.—Hence it is decomposed by all combustible bodies, which are oxygenized by it, with more or less rapidity in proportion to their affinity for oxygen.

1. When brought into contact with hydrogen gas at a high temperature, a violent detonation ensues. This experiment, therefore, requires great caution. 2. Poured on perfectly dry and powdered charcoal, it excites the combustion of the charcoal, which becomes red-hot, and emits an immense quantity of fumes. 3. It also inflames essential oils (as those of turpentine and cloves), when suddenly poured on them. In these experiments, the acid should be poured out of a bottle, tied to the end of a long stick; otherwise the operator's face and eyes may be severely injured. 4. Nitric acid is decomposed, by boiling it in contact with sulphur, which attracts the oxygen, and forms sulphuric acid.

(*l*) The acid is also decomposed by metals; as iron, tin, zinc, copper, &c., and with different phenomena, according to the affinity of each metal for oxygen. This may be seen, by pouring some strong nitric acid on iron filings, or powdered tin. Violent heat, attended with red fumes, will be produced, and the metals will be oxydized.

(*m*) If the action of metals on nitric acid be more moderately conducted, a new product is obtained in a gaseous state. Dilute some nitric acid with an equal weight of water, and dissolve, in this, some turnings of copper, or a portion of quicksilver, applying heat, if necessary.—This must be done in a gas bottle, and the product received, over water, is nitrous gas, or nitric oxide.

SECTION II.

Nitrous Gas, or Nitric Oxide.

The properties of this gas are the following:

(*a*) It is permanent over water; but it is absorbed in small quantity, when agitated with water which has been recently boiled, and has become cold. (See the table in the Appendix.) This solution, according to La Grange (vol. i. page 131), is converted, by long keeping, into nitrate of ammonia, in consequence of the decomposition of the water.

Nitrous gas is rather heavier than common air, 100 cubic inches at 55°, barometer 30°, weigh 34.26, or at 60° Fahrenheit 33.80 grains. (Davy.)

(*b*) When well washed with water, it is not acid. It will be found not to redden the colour of litmus. This may be shown by introducing a piece of paper, tinged with that substance, into a jar of nitrous gas, standing inverted over water. To accomplish this, the paper should be fastened to the end of a glass rod or a piece of stick. The colour will remain unchanged.

(*c*) It extinguishes flame, and is fatal to animals. Homberg's pyrophorus, however, is inflamed by it; and charcoal and phosphorus, introduced into it when in a state of actual combustion, continue to burn vehemently.

(*d*) Mingled with hydrogen gas, it imparts a green colour to its flame. It does not, however, explode with hydrogen in any proportion, nor with any of the varieties of carburetted hydrogen. But, when mixed with ammonia, an electric spark produces a detonation, as I have shown in the Philosophical Transactions for 1809. The proportions required for mutual saturation are about 120 measures of nitrous gas to 100 of ammonia.

(*e*) When mixed with oxygen gas, red fumes arise; heat is evolved; a diminution takes place; and if the two gases be in proper proportion, and perfectly pure, they disappear entirely. Nitrous acid, at the same time, is regenerated.

(*f*) The same appearances ensue, less remarkably, with atmospheric air; and the diminution is proportionate to the quantity of oxygen gas which it contains. On this principle, of its condensing oxygen, but no other gas, is founded the application of nitrous gas to the purpose of *eudiometry*, or of ascertaining the purity of air. The sources of error, in its employment in this mode, have hitherto been considered such, as to forbid our relying implicitly on the results which it may afford. Learning, however, from Mr. Dalton, that he constantly employs nitrous gas in determining the purity of air, and with perfect satisfaction as to the accuracy of his results, I have obtained from him the following communication. It may be necessary to premise, that for applying nitrous gas to this purpose, two tubes will be found convenient, shaped like fig. 24; each from three to four tenths of an inch in diameter; eight or nine inches long, exclusive of the funnel-shaped part; and accurately graduated into minute aliquot parts. What these parts are, is of no consequence. Hundredth parts of a cubical inch give rather too large divisions of the scale; but if each of these be divided into two, the scale will be sufficiently small. If the tube employed be not long enough to comprise 100 of these parts, the experiment may be made on 50 parts only of atmospherical air; and the results, multiplied by 2, will give the proportion in 100 parts.

" To use nitrous gas accurately in eudiometry, it is " only requisite to take both gases in a dilute state, " namely, containing three or four times their bulk of " azotic gas (which atmospheric air naturally does), or of " any other gas not acted upon by nitrous or oxygen gases. " In this case, if an excess of one gas be used, the other " is, in a few minutes, entirely taken up, and in a constant proportion; whatsoever may be the form of the " vessel, or the manner of mixing the gases. The proportion is 1 of oxygen to 1.7 of nitrous, so that 10-27ths " of the diminution over water are oxygen, and 17-27ths " nitrous gas. It is proper, as soon as the greater part of " the diminution has ensued, to transfer the mixture " through water into a graduated vessel, without using any " agitation.

" If pure nitrous gas be admitted to pure oxygen gas " in a narrow eudiometer tube, so that the oxygen gas " is uppermost, the two unite very nearly in the same " uniform proportion as above. If, on the other hand, " the nitrous be the upper gas, a much less quantity of it " disappears, *viz.* 1.24 nitrous to one oxygen. If undiluted nitrous gas be admitted to pure oxygen gas in a " wide vessel over water, the whole effect takes place " immediately; and one measure of oxygen will condense " 3.4 nitrous gas.

" To render this rule more intelligible, an example " may be necessary. Let 100 measures of common air " be admitted to 100 measures of a mixture of nitrous gas, " with an equal proportion of azotic or hydrogen gas. " After standing a few minutes in the eudiometer, there

"will be found 114 measures. The loss 56 being di-"
"vided by the common divisor, 2.7, gives 21 nearly for
"the oxygen gas present in 100 measures of common
"air."

To these directions I may add, that when atmospherical air is the subject of experiment, it is scarcely necessary to dilute the nitrous gas, with any other gas, previously to its use. If a number of experiments be made, it will be proper, in all cases, to let the gases remain together a definite time (say 10 minutes) before noting the diminution; and it is needless to transfer them into another vessel. If the mixed gas, under examination, contain much more oxygen, than is present in atmospherical air, then it is proper to dilute the nitrous gas with an equal bulk of hydrogen gas; and, in this case, the narrower the tube, in which the experiment is made, the more accurate will be the result.

Subsequent experience, however, has convinced me that the method, proposed by Mr. Dalton, though sufficiently correct, when applied to a mixture of the same, or nearly the same, standard as the atmosphere, cannot be relied on when the proportion of oxygen is either considerably greater or less. In the former case, the process gives too great a diminution, sometimes indeed to such an extent as to indicate more oxygen gas than the whole amount of what was submitted to experiment. When the air, on which we are operating, is of an inferior standard to the atmosphere, we do not learn its full proportion of oxygen. Notwithstanding these objections, however, the method has still considerable value, since it may be applied to determine the proportion of oxygen in some mixed gases, to which other eudiometrical tests are not applicable; for example, to mixtures of hydro-carburet and oxygen gases.

The application of nitrous gas to eudiometrical purposes, it has lately been contended by Gay Lussac, is susceptible of perfect accuracy, provided certain precautions be observed, which he has pointed out, and which were suggested by his theoretical views of the constitution of these gases. A narrow tube he finds to be unfit for an eudiometer, his object being to form nitrous acid gas, which is but slowly absorbed by water. Instead therefore of a tube, we must take a wide vessel, such as a small tumbler glass; and to 100 parts of atmospheric air, previously measured, we must pass 100 measures of nitrous gas. A red fume will appear, which will soon be absorbed without agitation, and in half a minute, or a minute at most, the absorption will be complete. Pass the residuum into a graduated tube, and it will be found, almost invariably, that 84 measures have disappeared. Dividing this number by 4, we have 21 for the quantity of oxygen condensed.

By a series of experiments on mixtures of oxygen and nitrogen gases in various proportions, Gay Lussac found that this eudiometrical process may be depended upon, whether the oxygen exceed or fall short considerably of the proportion contained in atmospheric air.

(*g*) The generation of an acid, by the admixture of nitrous gas with common air or oxygen gas, may be shown by the following experiment. Paste a slip of litmus-paper within a glass jar, near the bottom; and into the jar, filled with and inverted in water, pass as much nitrous gas, previously well washed, as will displace the water below the level of the paper. The colour of the litmus will remain unchanged; but, on passing up atmospheric air or oxygen gas, it will be immediately reddened.

(*h*) That the peculiar acid, thus produced, is the nitrous, will appear from the following experiment. Into a jar, filled with and inverted in mercury, pass a small quantity of a solution of pure potash; and, afterward, measures of oxygen and nitrous gases, separately, and in proper proportion. On removing the solution from the jar, exposing it for some time to the atmosphere, and afterward evaporating it, crystals of nitrate of potash will be formed, a salt which is ascertained to be formed of potash and nitrous acid.

(*i*) Nitrous gas is absorbed by nitric acid, which, by this absorption, is considerably changed in its properties.—Pass the gas, as it issues from the materials that afford it, through colourless nitric acid. The acid will undergo successive changes of colour, till at last it will become orange-coloured and fuming. In this state it is called nitrous acid, because it contains a less proportion of oxygen than the colourless nitric acid.

According to Dr. Priestley, 100 parts of nitric acid, of the specific gravity of 14 to 10, absorb, in two days, 90 parts by weight of nitrous gas *. When about seven parts of gas have been absorbed, the acid acquires an orange colour; when 18 have been absorbed, it becomes green; and the whole quantity, which it is capable of condensing, changes it into a liquor, which emits an immense quantity of red fumes. The gas, thus absorbed, is mostly separated again on dilution with water.

(*k*) The nitrous gas, thus absorbed, is expelled again by a gentle heat. This may be shown by gently heating the acid coloured in Experiment *i*, till it again becomes limpid. In this experiment light should be excluded.

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* Priestley on Air, 2d edition, i. 383. In the experiment alluded to, one fifth of an ounce-measure of nitric acid absorbed 130 ounce-measures of nitrous gas. There is reason, however, to suspect some inaccuracy in the experiment; for, according to Mr. Davy, 100 parts of nitric acid, after having absorbed all the nitrous gas which it is capable of condensing, hold only nine or between nine and ten parts in combination.

(*l*) Nitrous gas is decomposed by exposure to bodies that have a strong affinity for oxygen. Thus, iron filings decompose it, and become oxydized, affording a proof of the presence of oxygen in this gas.

During this process, water, ammonia, and nitrous oxide are generated. Sulphuret of potash, &c., have a similar effect. Mixed with sulphurous acid, it is decomposed, and this acid is changed into the sulphuric. (Nicholson's Journal, xvii. 43.) Potassium was found by Gay Lussac to burn vividly in it, and to afford a ready way of effecting its analysis.

(*m*) Nitrous gas is absorbed by the green sulphate and muriate of iron*, which do not absorb nitrogen gas. To ascertain, therefore, how much nitrogen gas a given quantity of nitrous gas contains, let it be agitated in a graduated tube with one of these solutions. This information is necessary, previously to deducing, from its effects on atmospheric air, the proportion of oxygen gas.

From the important use which is now made by some chemists of this solution of nitrous gas in eudiometry, it may be proper to describe the mode of its preparation.

Dissolve as much of the green sulphate of iron in water as the water will take up, or dissolve iron filings in sulphuric acid, diluted with five or six parts of water, leaving an excess of the iron, in order to ensure the perfect saturation of the acid. Fill a wide-mouthed bottle with this solution, invert it in a cupful of the same, and into the inverted bottle receive the nitrous gas from nitric acid and quicksilver, shaking the inverted bottle frequently. The colour of this solution will change to black, and the production of gas and the agitation are to be continued, till the absorption can be carried no farther. The impregnated solution should be preserved in a number of small bottles, not holding more than an ounce each. The most commodious method of applying this solution, is by means of Dr. Hope's eudiometer, already described. (Chap. v. sect. 4.)

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* For an account of these salts, see chap. xviii. sect. 6.

(*n*) A very interesting experiment, affording a synthetic proof of the constitution of nitrous gas, we owe to Dr. Milner, of Cambridge*. Into an earthen tube, about 20 inches long and three fourths of an inch wide, open at both ends, put as much coarsely powdered manganese as is sufficient nearly to fill it. Let this be placed, horizontally, in a furnace, having two openings opposite to each other (fig. 40). To one end of the earthen tube adapt a retort, containing a strong solution in water of pure ammonia, and to the other a bent glass tube, which may terminate in a two-necked empty bottle. To the other neck of the bottle, lute a glass tube, bent so as to convey any gas that may be produced, under the shelf of the pneumatic trough. Let a fire be kindled in the furnace; and, when the manganese may be supposed to be red-hot, drive over it the vapour of the ammonia. The alkali will be decomposed; its hydrogen, uniting with part of the oxygen which is combined with the manganese, will form water; while its nitrogen, uniting with another portion of the oxygen, will constitute nitrous gas. The gas, thus generated, may be collected by the usual apparatus.

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* Philosophical Transactions, 1789.

(*o*) Nitrous gas is stated, by Mr. Davy, to consist of 56 oxygen and 44 nitrogen.

(*p*) Another fact, showing the mutual relation of ammonia and of the compounds of nitrogen, was discovered some years ago by Mr. Wm. Higgins*. Moisten some powdered tin (which is sold under this name by the druggists) with strong nitric acid; and, when the red fumes have ceased to arise, add some quick-lime or solution of pure potash. A strong smell of ammonia will be immediately produced.

In this experiment, the tin, at the same instant, attracts the oxygen both of the nitric acid and of the water. Hydrogen and nitrogen are consequently set at liberty; and, before they have assumed the gaseous state, these two bases combine, and constitute ammonia. The ammonia, thus generated, unites with a portion of undecomposed nitric acid; and is disengaged from this combination by potash or lime, which render it evident to the smell.

SECTION III.

Gaseous Oxide of Nitrogen.—Nitrous Oxide of Davy.

I. This compound, also consisting of oxygen and nitrogen, but in different proportions from those of nitrous gas, may be obtained by several processes.

(*a*) By exposing common nitrous gas for a few days to iron filings, or to various other bodies strongly attracting oxygen, this gas is changed into the nitrous oxide.

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* See his Comparative View of the Phlogistic and Antiphlogistic Theories, 2d edition, page 300, note.

Some nicety and experience are required to suspend the decomposition before it has gone too far; in which case nitrogen gas is obtained. The sulphite of potash, being incapable of decomposing nitrous oxide, is best adapted to the conversion of nitrous gas into that elastic fluid. The process, in all cases, may be suspended, when about two thirds the original bulk of the gas are left.

(*b*) By dissolving zinc, or tin, in nitric acid, diluted with five or six times its weight of water. Zinc, during this solution, disengages nitrous oxide till the acid begins to exhibit a brownish colour, when the process must be suspended, as nitrous gas is then formed. But by neither of these processes is the gas obtained sufficiently pure for exhibiting its qualities. To procure it in a state of purity, the following process is the best adapted.

(*c*) To nitric acid, diluted with five or six parts of water, add carbonate of ammonia, till the acid is saturated. Then evaporate the solution; and, to supply the waste of alkali, add, occasionally, a little more of the carbonate. Let the solution be evaporated by a very gentle heat. The salt obtained, when the solution has cooled, is next to be put into a glass retort, and distilled with a sand-heat, not exceeding 440 Fahrenheit *. The heat of an Argand's lamp is more than sufficient, and requires cautious regulation. The salt will presently liquefy, and must be kept gently simmering, avoiding violent ebullition. The gas may be collected over water, and allowed to stand a few hours before it is used, during which time it will deposit a white cloud, and will become perfectly transparent.

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* From the observations of Mr. Sadler (Nicholson's Journal, xv. 286), it appears that the purity of the nitrate of ammonia is of considerable importance; and that its adulteration with muriate diminishes the quantity, and impairs the quality of the gas.

A gazometer is best adapted for its reception; because all danger is then avoided of an absorption of the water of the trough into the retort; and because the gas is brought into contact with a much smaller surface of water, which has the property of absorbing a considerable proportion of the gas. On this account, water, which has been once used to confine the gas, may be kept for the same purpose.

The changes that take place, during the conversion of nitrate of ammonia into nitrous oxide, are the following: Nitric acid is composed of oxygen and nitrous gas; ammonia, of hydrogen and nitrogen. In a high temperature, the nitrous gas combines with an additional dose of nitrogen, and forms nitrous oxide; while the oxygen of the decomposed nitric acid unites with the hydrogen of the ammonia, and forms water.

The gas, thus obtained, was termed, by the Society of Dutch chemists, gaseous oxide of azote; but, for the sake of brevity, and as more conformable to the nomenclature of other compounds of nitrogen, I shall use, with Mr. Davy, the name of nitrous oxide*.

In order to ascertain whether nitrous oxide be adulterated with either common air or oxygen gas, we may mix equal measures of the gas under examination, and of nitrous gas. If any diminution ensue, the presence of one of these may be suspected; and the amount will show which of them is contained in it. Nitrous gas, however, is a much more common contamination; for it is generated, along with nitrous oxide, whenever the temperature of the salt is raised too high. Its presence may be detected, either by a diminution on the admixture of oxygen gas; or by an absorption being effected, on agitating the gas with a solution of green sulphate of iron, which has no action on pure nitrous oxide. (See also page 418, *g*.)

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* For a full account of this gas, consult Mr. Davy's Researches, Chemical and Philosophical. London. Johnson, 1800.

II. This gas has the following properties:

(*a*) It is considerably heavier than common air. About 55° Fahrenheit and 30 inches pressure, 100 cubic inches weigh 50.20 grains, or under the same pressure, and at 60° Fahrenheit 49.68. (Davy.)

(*b*) A candle burns in it with a brilliant flame and crackling noise. Before its extinction, the white inner flame becomes surrounded with an exterior blue one.

(*c*) Phosphorus, introduced into it in a state of inflammation, burns with increased splendour.

Phosphorus, however, may be melted, and sublimed in this gas, without alteration. It may even be touched with red-hot iron wire, without being inflamed; but when a wire, intensely heated, or made white-hot, is applied, the phosphorus burns, or rather detonates, with prodigious violence.

(*d*) Sulphur, introduced into nitrous oxide when burning with a feeble blue flame, is instantly extinguished; but when in a state of an active inflammation, it burns with a vivid and beautiful rose-coloured flame.

(*e*) Red-hot charcoal burns in it more brilliantly than in the atmosphere.

(*f*) Iron wire burns in this gas with much the same appearance as in oxygen gas, but for a shorter period. Potassium was found by Berard to inflame and burn in it with much activity. After the combustion, 100 measures left 50 of nitrogen gas.

(*g*) Nitrous oxide is rapidly absorbed by water that has been previously boiled, about one thirtieth the original bulk of the gas remaining uncondensed. (See table.) A quantity of gas, equal to considerably more than half the bulk of the water, may be thus made to disappear. This property furnishes a good test of the purity of nitrous oxide; for the pure gas is almost entirely absorbed by boiled water, which has cooled without the access of air. The water employed should exceed the gas three or four times in bulk.

(*h*) Water, that has been saturated with this gas, gives it out again, unchanged, when heated.

(*i*) The impregnated water does not change blue vegetable colours.

(*k*) It has a distinctly sweet taste, and a faint, but agreeable, odour.

(*l*) It is not diminished by admixture with either oxygen or nitrous gas.

(*m*) A mixture of this gas with hydrogen gas detonates loudly, on applying a lighted taper, or passing an electric spark.

When the proportion of hydrogen is nearly equal to that of nitrous oxide, or as 39 to 40, nitrogen gas only remains after the explosion; but when the proportion of hydrogen is smaller, nitric acid is also generated. It forms, also (as I have shown, Philosophical Transactions, 1809, page 444), a combustible mixture with ammoniacal gas, 100 measures of the latter requiring for saturation 130 measures of nitrous oxide.

(*n*) Nitrous oxide is not absorbed by alkalis; but if it be brought into contact with them, when in a nascent state, or before it has assumed the condition of a gas, it then enters into combination with alkaline bases. Thus, when a mixture of sulphite of potash and pure potash is exposed to nitrous gas, the latter is disoxygenized by the sulphite, and changed into nitrous oxide, which unites with the alkali. We obtain, therefore, a mixture of sulphate of potash with a compound of nitrous oxide and alkali, the former of which may be separated by priority of crystallization. The latter is composed of about three parts of alkali, and one of nitrous oxide. It is soluble in water; has a caustic taste, of peculiar pungency; and converts vegetable blues to green. Powdered charcoal, mingled with it, and inflamed, burns with bright scintillations. The nitrous oxide is expelled from fixed alkalis by all acids, even by the carbonic.

(*o*) Animals, when wholly confined in this gas, die speedily.

(*p*) One of the most extraordinary properties of this gas is exhibited by its action on the human body, when received into the lungs. When thus employed, it does not prove fatal, because, when received into the lungs, it is mixed and diluted with the atmospherical air present in that organ. To administer the gas, it may be introduced into an oiled silk bag or clean bladder, furnished with a stop-cock, and may be breathed repeatedly from the bag and back again, as long as it will last. The sensations that are produced vary greatly in persons of different constitutions; but, in general, they are highly pleasurable, and resemble those attendant on the pleasant period of intoxication. Great exhilaration, an irresistible propensity to laughter, a rapid flow of vivid ideas, and an unusual fitness for muscular exertion, are the ordinary feelings it produces. These pleasant sensations, it must be added, are not succeeded, like those accompanying the grosser elevation from fermented liquors, by any subsequent depression of nervous energy.

(*q*) From the experiments of Mr. Davy, it appears that 100 parts, by weight, of this gas, contain 36.7 oxygen, and 63.3 azote, or, excluding decimals, 37 oxygen and 63 azote.

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SECTION IV.

Nitrous Acid.

It has been a subject of controversy whether an acid, entitled to this denomination, and holding the same relation to the nitric, which the sulphurous bears to the sulphuric, has really existence. That the acid, obtained from nitre, has different states of oxygenation, and contains a less quantity of oxygen in proportion to the depth of its colour, is generally admitted. But it has been contended that we are to consider all these varieties as nitric acid, holding in combination variable proportions of nitrous gas; and the principal argument in favour of this theory is that the substance, occasioning the colour, may be separated by the mere application of heat. Mr. Davy has given the following table, showing the proportion of nitrous gas in nitrous acid of different colours.

		100 parts by weight contain		
	Sp. Gr.	Real Acid.	Nitr. Gas.	Water.
Pale yellow	1.502	90.5	1.2	8.3
Bright do.	1.50	88.94	2.96	8.1
Dark orange	1.480	86.84	5.56	7.6
Light olive	1.479	86	6.45	7.55
Dark olive	1.478	85.4	7.1	7.50
Bright green	1.476	84.8	7.76	7.44
Blue green	1.475	84.6	8	7.40

Mere dilution with water is sufficient to vary these colours. Thus the dark orange coloured acid, by dilution, passes through the shades of blue, olive, and bright green. Nitric acid, also, by absorbing nitrous gas, has its specific gravity diminished. Colourless acid, for example, when rendered of pale yellow, becomes lighter in the proportion of 1.51 to 1.502.

It has lately, however, been argued by Gay Lussac*, and apparently with much probability, that the nitrous acid is as much a distinct and peculiar compound, as any other of the combinations of nitrogen and oxygen. It is formed, he observes, whenever we mix oxygen and nitrous gases in such proportion, that the nitrous gas predominates, *viz.* about one measure of the former to four of the latter. It is of no consequence which is first added; for the result is invariably a condensible red vapour, containing by measure, 100 oxygen gas, and 300 nitrogen; or by weight

34.507 nitrogen
65.493 oxygen
———
100

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* *Mémoires d'Arcueil,* ii.

To form liquid nitrous acid, nothing more is necessary than to saturate water with this vapour. The water becomes first green, then blue, and finally an orange colour more or less deep. The latter may be brought to the state of green or blue by adding more or less water. Hence the colour depends merely on the circumstance of density.

Admitting the correctness of these views, it will follow that when nitrous gas is transmitted through nitric acid, the latter is partly de-oxygenated; and the nitrous gas, acquiring oxygen, becomes nitrous acid. Hence the ordinary acid of nitre is a compound of nitric and nitrous acids, and not of nitric acid and nitrous gas. Its colour will be deep in proportion to the quantity of nitrous acid which it holds in combination.

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SECTION V.

Nitrates.

Art. 1.—*Nitrate of Potash.*

I. A DIRECT synthetic proof of the composition of this salt may be obtained by saturating nitric acid with potash, either pure or in a carbonated state. The solution, on evaporation, yields crystals of nitrate of potash, or nitre.

For the purposes of experiment, however, the nitrate of potash, which may be met with in the shops, and which is an abundant product of nature, may be employed on account of its greater cheapness. The nitre, which is met with as an article of commerce, is brought to this country, chiefly from the East Indies. When it arrives, it is a very impure salt, containing besides other substances, a considerable proportion of muriate of soda. In this state it is called rough nitre. For the purposes of chemistry, it requires to be purified by solution in water and re-crystallization; and it then obtains the name of refined nitre or refined saltpetre.

II. This salt has the following properties:

(*a*) It crystallizes in prismatic octahedrons, generally constituting six-sided prisms, terminated by two-sided summits. Exclusive of the water of crystallization, Thenard has determined that it consists of

59.5 potash
40.5 nitric acid

But as potash itself, in the driest form under which we can procure it, still contains water. Berthollet has given the following proportions as those of nitrate of potash.

51.38 potash
48.62 acid
100

(*b*) For solution, it requires seven times its weight of water at 60° of Fahrenheit; and boiling water takes up its own weight.

These are the proportions assigned by Bergman; but La Grange asserts, that of water, at the ordinary temperature, nitrate of potash requires only three or four times its weight for solution; and half its weight of boiling water. (*Manuel*, 1st edition, i. 243.)

(*c*) By the application of a moderate heat it fuses, and being cast in moulds, forms what is called Sal Prunelle.

(*d*) If a red-heat be applied, nitrate of potash is decomposed in consequence of the destruction of its acid. By distilling it in an earthen retort, or in a gun-barrel, oxygen gas may be obtained in great abundance, one pound of nitre yielding about 12,000 cubic inches, of sufficient purity for common experiments, but not for purposes of accuracy.

(*e*) Nitrate of potash, that has been made red-hot, seems to contain an acid less highly oxygenated than the common nitric acid, and having a weaker affinity for alkalis. For if acetic acid be poured on nitre that has been thus treated, the nitrous acid is expelled in red fumes, whereas common nitre is not at all affected by acetic acid.

(*f*) Nitrate of potash is rapidly decomposed by charcoal in a high temperature. This may be shown, by mixing two parts of powdered nitre with one of powdered charcoal, and setting fire to the mixture in an iron vessel under a chimney.—The products of this combustion, which may be collected by a proper apparatus, are carbonic acid and nitrogen gases. Part of the carbonic acid also remains attached to the residuary alkali, and may be obtained from it on adding a stronger acid.

This residue was termed, by the old chemists, *clyssus* of nitre.

(*g*) Nitrate of potash is also decomposed by sulphur, and with different results according to the temperature and proportions employed.

1. Mix powdered nitre and sulphur, and throw the mixture, by a little at a time, into a red-hot crucible. The sulphur will unite with the oxygen of the nitric acid, and form sulphuric acid; which, combining with the potash, will afford sulphate of potash. The production of the latter salt will be proved by dissolving the mass remaining in the crucible and crystallizing it, when a salt will be ob-

tained exhibiting the characters described, chap. xii. sect. 4.

2. Mix a portion of sulphur with one sixth or one eighth its weight of nitrate of potash; put the mixture into a tin cup, and raise it, by a proper stand (fig. 25), a few inches above the surface of water, contained in a flat shallow dish. Set fire to the mixture, and cover it with a bell-shaped receiver. In this case, also, sulphuric acid will be formed; but it will not combine, as before, with the alkali of the nitre, which alkali is present in sufficient quantity to absorb only a part of the acid produced. The greater part of the acid will be condensed on the inner surface of the glass bell, and by the water, which will thus become intensely acid. The operation may be repeated three or four times, using the same portion of water. When the water is partly expelled, by evaporation in a glass dish, concentrated sulphuric acid remains, which has been formed by the union of the oxygen of the nitre, and that of the atmospherical air, with the sulphur submitted to experiment. By a process of this kind, but on a large scale, and in extensive leaden chambers, the sulphuric acid of commerce is prepared. The dilute acid, resulting from the union of the condensed vapour of the burning materials, with the stratum of water at the bottom of the chamber, is first boiled down in part in shallow leaden vessels, and is then transferred into glass retorts, where it is farther concentrated by the continued application of heat.

In a memoir of Clement and Desormes, published in Nicholson's Journal, xvii. 41, it is proved, that the nitre does not furnish above one tenth part of the oxygen, required for the conversion of sulphur into sulphuric acid, and that the rest of the oxygen is derived from the atmospherical air of the chamber. Sulphurous acid, they suppose, is in the first instance formed by the combustion of the sulphur; and, at the same moment, nitrous gas is evolved from the de-oxygenation of the nitric acid contained in the saltpetre. This nitrous gas, uniting with the atmospheric oxygen, composes *nitrous acid gas*, which has the property of converting sulphurous into sulphuric acid, and of returning, at the same time, to the state of nitrous gas. The same process is repeated, and thus the same portion of nitrous gas acts repeatedly as an intermedium between the sulphur, previously changed into sulphurous acid, and the atmospheric oxygen.

(*h*) A mixture of three parts of powdered nitre, two of carbonate of potash, or common salt of tartar, and one part of sulphur, all accurately mixed together, forms the *fulminating powder*, which explodes with a loud noise, when laid on an iron heated below redness.

(*i*) A mixture of five parts of powdered nitre, one part of sulphur, and one of powdered charcoal, composes *gunpowder*. The materials are first very finely powdered separately, then mixed up together, and beaten with a wooden pestle, a sufficient quantity of water being added to prevent an explosion. The mixture is afterward granulated, by passing through sieves, and dried very cautiously*.

Process for preparing Nitric Acid.

Nitrate of potash is decomposed by sulphuric acid, which combines with the potash, and expels the nitric acid. Put into a glass retort, which may be either tubulated or not, four parts of nitrate of potash, reduced to a coarse powder and pour upon it three parts of concentrated sulphuric acid. Apply a tubulated receiver, of large capacity, between which, and the retort, an adopter may be interposed; these junctures being luted with a mixture of pipe-clay, sifted sand, and cut tow or flax. —To the tubulure of the receiver, a glass tube may be fixed by means of the fat lute, and may terminate in another large receiver, containing a small quantity of water. If the operator wishes to collect the gaseous products also, this second receiver should be provided with a tubulure, to which a bent pipe may be luted, terminating under one of the inverted funnels in the shelf of the pneumatic trough. Apply heat to the retort, through the intervention of a sand-bath. The first product that passes into the receiver, is generally of a red colour, and of a smoking quality. These appearances gradually diminish; and if the materials used were clean, the acid will come over pale, and even colourless. Afterwards it gradually reassumes a red colour, and smoking property; which appearances go on increasing till the end of the operation; and the whole product, mingled together, has either a yellow or an orange colour, according to the temperature employed.

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* On the preparation of gunpowder, and the theory of its detonation, consult Nicholson's Journal, xxiii. 277.

The proportions recommended in the new London Pharmacopœia for the preparation of nitric acid, are two pounds of nitrate of potash, deprived by heat of its water of crystallization, and two pounds of sulphuric acid. These are directed to be mixed in a glass retort, and distilled in a sand-bath, until a red vapour arises. The acid in the receiver is to be mixed with an ounce of nitrate of potash, and again distilled in a similar manner. After the second distillation its specific gravity is 1.500. The Edinburgh Pharmacopœia makes it 1.550, and Mr. Kirwan, at 60° Fahrenheit, 1.5543, and when of this density he considers it as composed of 73.54 real acid in 100, and the rest water.

The following table, showing the results of several experiments to determine the best proportions for preparing nitric acid, is given by Dr. Powell in his translation of the Pharmacopœia of the London College of Physicians.

Quantity of dried Nitre.	Sulphuric Acid.	Colour of Acid obtained.	Sp. Gr.	Weight.	Relative Value.
6	3	Red	1.53	3	21
6	6	White	1.50	4	29
60	29	Red	1.456	30	19

When the proportions were 6 of nitre, and 3 of sulphuric acid, there was no redundant acid in the caput mortuum.

The acid, which first passes over, has the greatest specific gravity. In an experiment of Dr. Perceval of Dublin, the product was taken in three portions; the first of which had the specific gravity of 1.494, the second of 1.485, and the third of 1.442. (Transactions of the Irish Academy, iv. 37.)

In the large way, and for purposes of the arts, it is usual to substitute earthen or cast-iron retorts, made extremely thick, for those of glass. An earthen head is adapted, and this is connected with a range of proper condensors. The strength of the acid is varied also, by putting more or less water into the receiver.

Nitric acid, obtained by this process, is never perfectly pure. It contains, generally, both sulphuric and muriatic acids; the former of which is indicated by a white precipitate, on adding a solution of nitrate of barytes; and the latter, by a milkiness produced by nitrate of silver. The sulphuric acid may be separated, either by a second distillation from a portion of very pure nitre, equal in weight to one eighth of that originally employed, or by adding nitrate of barytes; allowing the precipitate to settle; decanting the clear liquid, and distilling it. Muriatic acid is separated by the addition of nitrate of silver. An immediate milkiness ensues, and fresh additions must be made of nitrate of silver, as long as it occasions this appearance. Then allow the precipitate to subside; decant the clear liquid, and re-distill it; leaving one eighth or one tenth in the retort. The product will be pure nitric acid. Nitrate of lead may be substituted for nitrate of silver. (See Nicholson's Journal, xi. 134.)

The nitric acid may also be obtained free from muriatic acid, if a perfectly pure nitrate of potash be employed for distillation. This purification is, in a great measure, effected by repeated solution of the nitre, in boiling distilled water, and crystallization. Finally, re-dissolve the crystals in warm distilled water, and add nitrate of silver as long as any precipitate appears. Allow this to settle, or separate it by filtration. The next product of crystals will be perfectly pure. The precipitated silver must not be thrown away, but must be washed with distilled water, dried, and preserved for the recovery of the silver, by a process to be hereafter described.

Nitric acid obtained in this manner is deficient also in another respect; for it is not perfectly oxygenated, but holds in solution a considerable quantity of nitrous gas, and hence is in the state rather of nitrous than nitric acid. To convert the former into the latter, put the acid into a retort, to which a receiver is applied, the two vessels not being luted, but joined merely by paper. Apply a very gentle heat for several hours to the retort, changing the receiver as soon as it becomes filled with red vapours. The nitrous gas will thus be expelled, and the acid will remain in the retort in a state of purity, and as limpid and colourless as water. It must be kept in a bottle secluded from the light.

One hundred parts of nitrate of potash, according to La Grange, yield by this process 43 of acid, or, according to my experience, above 50. This, however, is not the whole of what was contained in the salt: for a part is decomposed by the temperature necessary to the operation. Accordingly, a large quantity of oxygen gas is disengaged during the distillation, and may be collected by an obvious addition to the apparatus.

In the retort, there remains a compound of potash with more sulphuric acid than is essential to its saturation, or a super-sulphate of potash. On submitting this to a pretty strong heat, the excess of sulphuric acid is expelled; and the residue, dissolved and evaporated, affords crystallized sulphate of potash.

ART. 2.—*Nitrate of Soda.*

I. This salt may be formed, by saturating carbonate of soda with nitric acid; or by distilling common salt with three fourths its weight of nitric acid. When the former process is adopted, the solution must be evaporated, till a pellicle appears on its surface, and then allowed to cool. Crystals will be produced, having the shape of rhomboids, or rhomboidal prisms.

II. These crystals have a taste like that of saltpetre, but more intense. They are soluble in three parts of water at 60°, and in less than an equal weight of boiling water. They attract moisture from the atmosphere. In other respects, in the means by which their decomposition is effected and its results, they agree with the nitrate of potash. The only use of nitrate of soda is, perhaps, that which has been suggested by Proust, who has found it to be more economical in the making of fire-works than nitrate of potash. (Nicholson's Journal, xv. 262.)

ART. 3.—*Nitrate of Ammonia.*

I. The most simple mode of preparing this salt is by adding carbonate of ammonia to dilute nitric acid, till saturation has taken place. If the liquor be evaporated, by a heat between 70° and 100°, to a certain extent, it shoots, on cooling, into crystals, having the shape of six-sided prisms, terminated by long six-sided pyramids. Evaporated at the temperature of 212°, it yields, on cooling, thin fibrous crystals; and when the evaporation is carried so far, that the salt immediately concretes on a glass rod by cooling, it then forms a compact and shapeless mass.

II. The solubility of this salt varies, according to the temperature in which it has been formed. When in crystals, it requires twice its weight of water, for solution, or half its weight of boiling water. It deliquiates, in all its forms, when exposed to the atmosphere.

III. The most important property of this salt is the one which has been already described, *viz.* of yielding, when decomposed by heat, the nitrous oxide. One pound of the compact kind gives, by careful decomposition, nearly five cubic feet of gas, or rather more than 34 doses; so that the expense, estimating the salt at 5*s*. 10*d*. the pound, is about 2*d*. for each dose.

IV. In a temperature of 600° this salt explodes, and is entirely decomposed. Hence it was formerly called *nitrum flammans*.

V. Its composition varies according to the mode of its preparation, and is stated by Mr. Davy as follows:

Prismatic.	Fibrous.	Compact.	
69.5	72.5	74.5	acid
18.4	19.3	19.8	ammonia
12.1	8.2	5.7	water
100	100	100	

Art. 4.—*Nitrate of Barytes.*

Nitrate of barytes may be prepared, by dissolving either the artificial or native carbonate in nitric acid, diluted with eight or ten parts of water. If the artificial carbonate be employed, it should be previously well washed with distilled water, till the washings cease to precipitate nitrate of silver. A solution of nitrate of barytes, mixed with one of silver, should continue perfectly transparent. On evaporation, it yields regular octahedrons, often adhering to each other in the form of stars; and sometimes it is obtained in small brilliant plates. It requires for solution 12 times its weight of water at 60°, and three or four parts of boiling water. It is not altered by exposure to the air. In a red-heat, its acid is decomposed, and the earth remains pure. This furnishes another method of procuring pure barytes; but the heat must not be carried too far, otherwise the barytes is apt to vitrefy with the crucible. The residue, on the addition of water, dissolves with great heat and noise, and the solution, on cooling, yields crystals of pure barytes.

Nitrate of barytes is composed, in 100 parts, according to Clement and Desormes, of 60 base, and 40 acid and water. Mr. James Thomson states its composition to be

59.3	barytes
40.7	acid and water

ART. 5.—*Nitrate of Strontites.*

This salt may be obtained in the same manner as the nitrate of barytes, with which it agrees in most properties. The solubility of its crystals, however, differs considerably; for they are dissolved by their own weight of water at 60°, or by little more than half their weight of boiling water. When applied to the wick of a candle, or added to boiling alcohol, they communicate to the flame a deep blood-red colour. They are decomposed by a high temperature, and afford pure strontitic earth.

ART. 6.—*Nitrate of Lime.*

This salt is found abundantly in the cement of old buildings, which have been long inhabited by man or other animals. To prepare it artificially, nitric acid, diluted with five or six parts of water, may be saturated with carbonate of lime. When this solution is boiled down to the consistence of syrup, and exposed in a cool place, long prismatic crystals are formed, resembling, in their disposition, bundles of needles diverging from a common centre. These crystals are readily soluble in water, of which, at 60°, they require two parts, and boiling water dissolves an equal weight. They deliquiate speedily, when exposed to the air; and are decomposed at the temperature of ignition.

When a solution of nitrate of lime is evaporated to dryness in an earthen vessel, then fused for five or ten minutes in a crucible, and poured, while in fusion, into an iron pot previously heated, the congealed mass forms *Baldwin's phosphorus.* It must be broken into pieces, and preserved in a well-stopped phial. These pieces, after having been exposed to the sun for a few hours, emit in the dark a beautiful white light, affording one variety of solar phosphorus.

ART. 7.—*Nitrate of Magnesia.*

This compound may be prepared, by dissolving carbonate of magnesia in diluted nitric acid. The solution, when evaporated, yields crystals in the shape of prisms, with four oblique faces truncated at their summits. Most commonly, however, it forms a shapeless mass, consisting of an immense number of small needle-shaped crystals, crossing each other irregularly. These crystals deliquiate in the air, and are soluble in half their weight of water. When exposed to the heat of ignition, they fuse; a few bubbles of oxygen gas first escape; and the nitric acid then passes undecomposed.

ART. 8.—*Nitrate of Alumine.*

This salt is but little known. It may be formed by the solution of fresh precipitated alumine, which has been well-washed with distilled water, but not dried, in diluted nitric acid, with the assistance of heat. The solution, which has always an excess of acid, after evaporation, crystallizes in thin ductile plates. The crystals are extremely soluble; and, on the application of a high temperature, abandon their acid. They are decomposed by most alkalis and earths. Pure potash, added in excess, re-dissolves the precipitate.

ART. 9.—*Nitrate of Glucine.*

The nitrate of glucine is a sweet tasted salt, which cannot be brought to crystallize. When evaporated to dryness, it rapidly absorbs moisture from the atmosphere. It is soluble in alcohol. A high temperature decomposes it, without effecting its previous fusion.

ART. 10.—*Nitrate of Zircon.*

The nitric acid dissolves, but cannot be saturated with, fresh precipitated zircon. The solution has always an excess of acid. When evaporated, it forms a yellowish transparent mass, extremely tenacious and viscid, and difficultly dried. It has a styptic astringent taste, and leaves on the tongue a thick substance, in consequence of its partial decomposition by the saliva. This dry nitrate is extremely soluble. The solution is decomposed by sulphuric acid, and by carbonate of ammonia, which throw down a precipitate soluble in an excess of the acid, or of the carbonate. Tincture of galls forms a white precipitate, which is soluble in an excess of the tincture.

ART. 11.—*Nitrate of Yttria*

May be prepared by dissolving yttria in nitric acid. The solution has a sweetish astringent taste; and, in most properties, resembles nitrate of glucine. It can scarcely be obtained in crystals; and if too great a heat be applied during evaporation, the salt becomes soft, assumes the appearance of honey, and concretes, on cooling, into a hard stony mass. Exposed to the air, it attracts moisture, and is resolved into a liquid.

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SECTION VI.

Nitrites.

THE direct combination of nitrous acid with alkalis and earths cannot be effected: for this acid, consisting of nitric acid with a redundancy of nitrous gas, is decomposed during all these combinations, the nitrous gas escapes, and the result is a compound of nitric acid with the base employed. The only mode of obtaining nitrites, is to deprive the acid, contained in the nitrates, of part of its oxygen, by exposure for a short time to the temperature of ignition. This method, it must be obvious, cannot be used with those nitrates that abandon their acid on the application of heat, or which, like nitrate of ammonia, are more completely decomposed.

In a recent memoir of Gay Lussac it is asserted, however, that the nitrites may be composed by the direct combination of alkaline and earthy bases with nitrous acid, formed in the new mode proposed by him and already described under the article *nitrous acid.*

Nitrate of potash, after fusion in a crucible, becomes a nitrite of that alkali. It has a smell, when powdered, of nitrous gas. When diluted nitric acid, or even acetic acid, is poured upon it, vapours of nitrous acid are disengaged; and hence it appears, that the affinity of this acid for its base is weakened by dis-oxygenation; for no such effect arises on adding these acids to the nitrate. The solution of the salt in water changes the syrup of violets to green. Its other properties are little known.

CHAPTER XIV.

MURIATIC ACID,—OXY-MURIATIC ACID,—AND THEIR COMBINATIONS WITH ALKALIS.

OF the base of the muriatic acid, no history can be at present given; for the analysis of this acid has not yet been satisfactorily accomplished, even with the assistance of the new powers and agents, which have been unfolded by the discoveries of Mr. Davy. From his experiments, as well as from those of the French chemists, it appears that the muriatic acid, in the purest form under which it has hitherto been obtained, still holds in combination one third its weight of water, which is even essential to its existence as a gas. For the acid cannot be detached from its combination with alkalis, earths, or metallic oxides, by any third substance, unless that third substance contains a proportion of water.

Oxy-muriatic acid is a compound of the same base with a certain proportion of oxygen. Into this compound, the water, which is essential to the constitution of ordinary muriatic acid, does not enter. It follows, therefore, that a certain base, united with water, composes muriatic acid gas; and, united with oxygen, oxymuriatic acid.

SECTION I.

Muriatic Acid.

I. THE muriatic acid, in its purest form, exists in the state of a gas, permanent over mercury only. For exhibiting its properties, therefore, a mercurial apparatus is absolutely necessary.

To obtain muriatic acid gas, let the tubulated gas bottle (plate ii. fig. 17) be about one fourth, or one third, filled with well dried muriate of soda (common salt). To this adapt the acid-holder, filled with concentrated sulphuric acid; and let the aperture of the bent pipe terminate under a jar filled with, and inverted in, quicksilver. Open the communication between the acid and the salt, by turning the cock; and immediately on the contact of these two bodies, an immense quantity of muriatic acid gas will be disengaged. The first portions, that come over, may be allowed to escape, by holding the gas bottle under a chimney; because they are contaminated by the admixture of the common air present in the bottle. The subsequent portions may be preserved for use; and the pure gas will exhibit the following qualities:

(*a*) It has a very pungent smell; and is sufficiently caustic to blister the skin, when applied to it for some time.

(*b*) When brought into contact with common air, it occasions a white cloud. This is owing to its union with the moisture always present in the atmosphere.

(*c*) It extinguishes a lighted candle. Before the flame goes out, the upper part of it assumes a greenish hue, the cause of which has not yet been explained. A white vapour also surrounds the extinguished wick, owing to the combination of water, produced by the combustion of the candle, with the muriatic acid gas.

(*d*) It is heavier than common air, in the proportion of 173 to 100. Biot and Gay Lussac, however, state its specific gravity at only 1.278. Hence 100 cubic inches (taking that quantity of common air at 31 grains) should weigh only 39.51 grains.

(*e*) It effects the liquefaction of a piece of ice, almost as rapidly as a red-hot coal.

(*f*) It is very rapidly absorbed by water. A drop or two of water, admitted to a large jar full of this gas, causes the whole of it instantly to disappear. According to Mr. Kirwan, an ounce-measure troy of water absorbs 800 cubical inches (*i. e.* 421 times its bulk, or its own weight) of muriatic acid gas; and the water, by this absorption, is increased about one third its original volume. Dr. Thomson's experiments indicate a still larger absorption, *viz.* 515 cubical inches, or 308 grains by one cubic inch, equal to 252 grains, of water, at 60° Fahrenheit; the barometer standing at 29.4. The recent experiments of Berthollet have shown that 100 grains of water absorb 12.467 of muriatic acid gas deprived of all redundant water by passing it through a tube, surrounded by a freezing mixture. By this absorption, we obtain an acid of the specific gravity 1061.4. From calculation it appears that acid of this density contains in 100 grains only 8.55 of real acid.

(*g*) Muriatic acid gas is not decomposed by being transmitted over red-hot charcoal, nor by a succession of electrical shocks. The action of electricity, however, evolves a small portion of hydrogen gas, proceeding from the water which the acid gas holds in combination. From an estimate, founded on this experiment, I have inferred, that 100 cubic inches, or 60 grains of muriatic acid gas, after being dried by muriate of lime, hold in combination 1.4 grain of water. (Philosophical Transactions, 1800.) The more recent researches of Berthollet and Davy, however, have discovered a much larger proportion. According to the former*, one fourth, and according to the latter a third, of the weight of muriatic acid gas is water.

Nature of Muriatic Acid.

From the powerful agency of potassium in depriving bodies of oxygen, it was natural to expect that if the muriatic acid be a compound of oxygen with a peculiar basis, its decomposition might possibly be effected by means of this new metal. The experiment was, therefore, made by Mr. Davy with considerable attention, and with the following result.

When potassium, he observes, is introduced into muriatic acid gas, dried by muriate of lime, it immediately becomes covered with a white crust; it heats spontaneously; and, by the assistance of a lamp, acquires, in some parts, the temperature of ignition, but does not inflame. If the potassium and the gas be in proper proportions, they both entirely disappear; a white salt is formed, and a quantity of pure hydrogen gas is evolved, which is equal to about one third the original volume of the acid gas. Eight grains of potassium effected the absorption of nearly twenty-two cubic inches of muriatic acid gas: and the quantity of hydrogen gas produced amounted to more than eight cubical inches.

It is remarkable, however, that a given weight of potassium, by its action on muriatic acid gas, evolves exactly the same quantity of hydrogen, as would arise from its agency on water. This fact may be regarded almost as a proof that the phenomena depend merely on the water, which the gas holds in combination. But the conclusion is decisively established, by comparing the real muriatic acid (which may be done by precipitation with nitrate of silver) in the muriate of potash generated in the experiment, with that contained in an equal volume of the gas itself; and if this be accurately performed, no difference, Mr. Davy observes, will be found in the weight of the results.

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* *Mémoires d'Arcueil,* ii. 56.

The action, then, of potassium on muriatic acid gas consists entirely in its decomposing the water of the gas, and not at all in the destruction of its really acid part. The decomposition of this acid, therefore, remains as much a problem as ever.

Though the principal object of Mr. Davy's experiments on muriatic acid was not accomplished, yet several curious facts occurred in the investigation, some of which it may be proper to state.

A variety of expedients were attempted to obtain muriatic acid gas perfectly free from water, but without success. The materials, which are capable of affording it when either one or both contain water, yield no gas whatsoever when both are in a state of complete dryness. Thus dry phosphoric acid or boracic acid, mixed with dry muriate of lime, and strongly heated in a forge, give no gas. Hence it would appear that water is an essential constituent of this gas. The quantity, which it contains, Mr. Davy estimates at about one third of its weight.

Though muriatic acid, however, cannot be obtained in a state of complete separation from water, yet in some of its combinations, it appears not to contain any of this fluid. When in this state it is very remarkable, that though its acidity be not neutralized, yet it loses one of the distinguishing characters of an acid, *viz.* that of reddening litmus paper. Thus in the liquid compound of muriatic and phosphorous acids, perfectly dry litmus paper sustains no change of colour; but is immediately reddened when introduced moist.

All fluid acids, also, are good conductors of electricity; but the liquid compounds of muriatic acid free from water (such as that with phosphorous acid) are complete non-conductors.

It is still farther remarkable that the muriatic acid, existing in these liquid compounds, is rendered gaseous by the addition of water, a confirmation of the principle that water is essential to the constitution of this gas.

Of all known acid substances, Mr. Davy observes (Philosophical Transactions, 1809, page 468), the dry muriatic acid is that which seems to possess the strongest and most extensive powers of combination. It unites with all acids except the carbonic; and with all oxides (including water); and with all the compound inflammable bodies except carbonaceous matter and metals. Should it ever be obtained in a pure form, it will probably be one of the most powerful agents in chemistry.

Process for preparing Liquid Muriatic Acid.

Into a tubulated retort, placed in a sand-bath, put eight parts of dried muriate of soda: and, to the tubulure, lute the bent tube (fig. 26, *a*) with fat lute. To the neck of the retort, affix a tubulated receiver (fig. 30, *b*) by means of the same lute; and to the aperture of this adapt a tube, twice bent at right angles, and furnished with Welter's contrivance for preventing absorption (fig. 31, *b*), the longer leg of which terminates beneath the surface of water contained in a two-necked bottle. From the other neck, let a second right-angled pipe proceed; and this may terminate in a similar manner, in a second bottle containing water. Let the junctures be all carefully luted; and, when they are sufficiently hardened, pour very gradually through the bent tube five parts by weight of strong sulphuric acid, making the additions at several distant intervals. On each affusion of the acid a large quantity of muriatic acid gas will be liberated, and will be absorbed by the water of the first bottle, till this has become saturated. It will then pass on to the second bottle, and be there absorbed. The water employed may amount to half the weight of the salt, and may be equally distributed between the two bottles. These it is better to surround with cold water, or, still preferably, with ice or snow; because the condensation of the gas evolves considerable heat, which prevents the water from attaining its full impregnation. When the whole of the sulphuric acid has been added, and the gas no longer issues, let a fire be lighted in the furnace, beneath the sand-bath, removing the bent tube *a*, and substituting a well-ground glass stopper. This will renew the production of gas; and the temperature must be preserved, as long as gas continues to be evolved. At this period it is necessary to keep the luting, which connects the retort and receiver, perfectly cool; otherwise it will be apt to melt. To this juncture, indeed, I prefer the application of the clay and sand lute; but this requires some address. Towards the close of the process, a dark-coloured liquid is condensed in the first receiver, consisting of a mixture of sulphuric and muriatic acids. When nothing more comes over, the operation may be suspended, and the liquid in the two bottles must be preserved in bottles with ground stoppers. It consists of liquid muriatic acid.

The liquid muriatic acid may also be obtained by diluting the sulphuric acid with the water necessary for the condensation of the gas, and adding the dilute acid, when cold, to the salt in the retort. To the retort, an adopter may be luted with the clay and sand lute; and this may terminate in a large tubulated receiver, from the aperture of which a right-angled Welter's tube proceeds, and is conveyed beneath a few ounces of water, contained in a two-necked bottle. A fire must then be lighted under the sand-bath, and continued as long as any liquid comes over. The adopter and receiver must be kept cool, by the constant application of moistened cloths.

The proportions, directed by the London College of Physicians, are four parts of dried salt, three of sulphuric acid, and three of water, all by weight. The acid and one third of the water are ordered to be mixed together in a glass retort. When they are cold, the muriate of soda is to be added; the remainder of the water poured into a receiver; and, the retort and receiver being luted together, the muriatic acid is to be distilled into the latter, by a heat gradually raised to redness. If the whole charge of acid, water, and salt, be introduced at once into the retort, the first application of heat detaches muriatic acid gas, without water enough to condense it; and a portion passes to waste. Care should be taken not to close the luting, till the whole of the common air is first expelled. The charge should not occupy more than half the bulk of the retort, and this vessel may be preserved after the operation by pouring water at 212° into it, when the materials, remaining after distillation, have cooled down to about that temperature.

The weight of the acid produced should equal, or a little exceed, that of the salt employed*.

If the muriatic acid, thus obtained, should contain sulphuric acid, which may be discovered by muriate of barytes occasioning a white precipitate, the acid is to be re-distilled from a fresh portion of muriate of soda. When prepared by Woulfe's apparatus, the product in the second bottle is always perfectly pure.

The acid formed by the latter process has the specific gravity only of about 1170; but by Woulfe's apparatus, and especially when the bottles are surrounded by ice or snow, it approaches that of 1500. The lower degree of specific gravity, however, which has been mentioned, is best adapted for keeping; for the dense acid emits a large quantity of fumes, which are extremely inconvenient and injurious to all metallic instruments.

The caput mortuum consists of sulphate of soda with some undecomposed muriate of soda. The former may be obtained, in a crystallized form, by first driving off, by a strong heat, the excess of sulphuric acid that adheres to it; and then dissolving it in hot water. The product of sulphate of soda exceeds that of the muriate employed in the proportion of about eight to five.

Liquid muriatic acid has the following properties:

1. It emits white suffocating fumes. These consist of

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* Calculating on the data already quoted from Berthollet, and assuming that muriate of soda contains 46 *per cent.* of acid, 100 parts of common salt, supposing it to be entirely decomposed by sulphuric acid, and none of the gas to be lost, should impregnate 368 parts of water to the density 1160, and afford 414 parts of liquid acid of that specific gravity.

muriatic acid gas, which becomes visible by contact with the moisture of the air.

2. When heated in a retort, or gas bottle, muriatic acid gas is disengaged, and may be collected over mercury.

3. Liquid muriatic acid is not decomposed by the contact of charcoal, essential oils, or other combustible bodies.

4. When diluted with water, no remarkable elevation of temperature is produced.

5. In a perfectly pure state it is quite colourless; but it has frequently a yellowish hue. This may proceed, either from a portion of oxygenized muriatic acid, or of muriate of iron, but, I believe, most commonly of the latter. This colour is instantly destroyed by a few drops of muriate of tin.

6. Muriatic acid combines readily with alkalis, and with most of the earths, both in their pure and carbonated states.

SECTION II.

Muriates.

ART. 1.—*Muriate of Potash.*

MURIATE of potash may be obtained by saturating muriatic acid with carbonate of potash, and evaporating the solution till the salt crystallizes. These crystals have a cubical shape, and a bitter, disagreeable taste; they dissolve in three times their weight of water, at 60°, and in a rather less proportion of boiling water. They undergo little change when exposed to the air; they decrepitate when thrown on the fire, but abandon only a small portion of their acid.

From the experiments of Berthollet it appears that this salt is composed, exclusive of water, of

66.66	potash
33.34	acid
100	

Art. 2.—*Muriate of Soda.*

Muriate of soda is that well known substance, common salt, which is become a necessary ingredient in the food of man, and is of essential utility in several of the arts.

I. Its composition may be proved, by the direct union of soda with muriatic acid.

But for purposes of experiment, the common salt may be employed, which is to be found in the shops. This may be purified, by adding to a solution of it in water a solution of carbonate of soda, as long as any milkiness ensues; filtering the solution, and evaporating it till it crystallizes.

II. Its qualities are as follow:

1. It crystallizes in regular cubes, which, when the salt is pure, are but little changed by exposure to the air. The common salt of the shops, however, acquires an increase of weight, in consequence of the absorption of moisture. The various forms under which it appears, of stoved salt, fishery salt, bay salt, &c., arise rather from differences in the size and compactness of the grain, than from any essential difference of chemical composition.

2. It requires, for solution, twice and a half its weight

of water, at 60° of Fahrenheit, and hot water takes up very little more. Hence its solution crystallizes, not like that of nitre, by cooling, but by evaporation.

3. When heated gradually it fuses, and forms, when cold, a solid compact mass.

4. If suddenly heated, as by throwing it on red-hot coals, it decrepitates. It does not, however, after being dried at the temperature of boiling water, lose by ignition more than two or three parts of water *per cent.*

5. It is not decomposed when ignited in contact with inflammable substances.

6. When mixed with powdered charcoal or sulphur, and fused in a crucible, it does not undergo any decomposition or essential change; because the muriatic acid, if it contain any oxygen (which has not yet been proved), holds that basis more strongly combined than it is attracted by combustible bodies.

7. It is decomposed by the carbonate of potash, the alkali of which combines with the muriatic acid of the salt, and the carbonic acid is transferred to the soda.—Hence we obtain muriate of potash and carbonate of soda. A process for effecting this decomposition, on a large scale, is described by Westrumb, in Crell's Journal, English translation, ii. 127.

8. It is decomposed by the sulphuric acid in the mode already described. Nitric acid also separates the muriatic acid.

It consists according to Darcet, whose estimate has the sanction of Gay Lussac, of

50.73	alkali
49.27	acid
100.	

Dr. Marcet, however (Philosophical Transactions, 1807), whose analyses are distinguished by peculiar neatness and precision, calculates its composition to be

46 acid
54 soda

100

Art. 3.—*Muriate of Ammonia.*

1. If equal measures of ammoniacal gas and muriatic acid gas be mixed together, over mercury, they are immediately and totally condensed, a white cloud is formed, and a solid substance is deposited on the sides of the vessel.—This is the muriate of ammonia. It is composed, by weight of

38.35 alkali
61.65 acid

100

For the purposes of experiment, it may be procured in the shops, under the name of sal-ammoniac.

Its properties are as follow:

(*a*) It is volatilized, without being liquefied or decomposed, and hence may be sublimed.

(*b*) It is readily soluble in water, three parts and a half of which, at 60° take up one of the salt. During its solution much caloric is absorbed. In boiling water it is still more soluble; and the solution, on cooling, shoots into regular crystals.

(*c*) It slightly attracts moisture from the air.

(*d*) On the addition of a solution of pure potash, or pure soda, the alkali is disengaged, as is evinced by the pungent smell that arises on the mixture of these two bodies, though perfectly inodorous when separate.

(*e*) It is also decomposed by barytes, strontites, lime, and magnesia.

For the purpose of separating the volatile alkali, and obtaining it in a liquid form, lime is generally employed. Two parts of powdered and sifted lime are to be mixed with one of muriate of ammonia, and put into a retort, the neck of which is to be luted (with slips of moistened bladder, bound down by string), to a tubulated receiver (fig. 31, *a*). From the other opening of the receiver, a right-angled tube, with Welter's addition (fig. 31, *b*), is to proceed, and to terminate beneath the surface of water, contained in a two-necked bottle (*c*). This bottle is to be connected, by a second right-angled tube, with another similar bottle (*d*), containing (as should the first also) a quantity of water, equal in weight to one fourth that of the muriate of ammonia. The decomposition is effected by heat; and the gas is absorbed by the water contained in the bottles, which should be surrounded by ice or snow. The water acquires a strong smell, and has its specific gravity diminished, when fully impregnated, to 0.9684. This watery solution, when heated in a gas bottle, gives up its ammonia, which passes over in an aëriform state, and may be received over quicksilver.

A solution of ammonia, in water, may also be obtained, as follows: Slake two parts of quicklime, with two of water; add to this, when cold, one part of muriate of ammonia, and six parts of water. Stir the mixture, and transfer it into a retort; lute on a receiver; and separate, by distillation, one part of liquid. The former process, however, is incomparably the best.

When a mixture of one part of powdered muriate of ammonia with two of powdered carbonate of lime (chalk), both perfectly free from moisture, is distilled together in a retort, a solid white substance condenses on the inner surface of the receiver. This is the carbonate of ammonia; and the process now described is that by which, with the substitution of proper subliming vessels, the carbonate of ammonia is prepared for sale. This operation furnishes an example of double affinity. The carbonic acid, being transferred from the lime to the ammonia, forms carbonate of ammonia; and the muriatic acid, passing to the lime, composes muriate of lime.

Art. 4.—*Muriate of Barytes*

Is best prepared, by dissolving either the artificial or native carbonate in muriatic acid much diluted; or, if neither of these can be had, the sulphuret. The iron and lead, which are occasionally dissolved, along with the barytes, may be separated by the addition of a small quantity of liquid ammonia, or by boiling and stirring the solution in contact with a little lime. When filtered and evaporated, the solution yields regular crystals, which have most commonly the shape of tables, bevelled at the edges, or of eight-sided pyramids, applied base to base. They dissolve in five parts of water, at 60°, or in a still smaller quantity of boiling water; and also in alcohol. They are not altered by exposure to the atmosphere; nor are they decomposed, except partially, by a high temperature. The sulphuric acid separates the muriatic; and the salt is also decomposed by alkaline carbonates and sulphates.

According to Mr. Aikin (Nicholson's Journal, xxii.), muriate of barytes, after ignition, contains

26.86	acid
73.14	base
100	

And the fully crystallized muriate consists of

22.93	acid
62.47	base
14.6	water
100	

Art. 5.—*Muriate of Strontites*

May be obtained by following the same process as that employed in preparing the barytic salt. The solution affords long slender hexagonal prisms, which are soluble in two parts of water, at 60°; and to almost any amount in boiling water. In a very moist atmosphere they deliquiate. They dissolve in alcohol, and give a blood-red colour to its flame.

Art. 6.—*Muriate of Lime.*

This salt may be prepared by dissolving carbonate of lime in muriatic acid, or by washing off the soluble part of the mass, which remains after the distillation of the solution of pure ammonia from muriate of ammonia and lime.

The solution crystallizes in six-sided striated prisms, terminated by very sharp pyramids. If it be evaporated to the consistence of a syrup, and exposed in a temperature of 32°, it forms a compact mass, composed of bundles of needle-shaped crystals, crossing each other confusedly.

The crystals dissolve in half their weight of cold water, and to an unlimited extent in boiling water, being, in fact, soluble in their water of crystallization.—They deliquiate rapidly in the air, and enter into fusion when heated. After being melted by a strong heat, the fused mass still contains water; for by ignition with iron filings, it yields much hydrogen gas. If fused in a crucible, and treated in the same manner as the nitrate of lime, the crystals yield a solar phosphorus, called, from its discoverer, *Homberg's phosphorus.* When mingled with snow, they produce intense cold, as has already been described.

Art. 7.—*Muriate of Magnesia.*

This is also a deliquescent and difficultly crystallized salt. It has an intensely bitter taste; is soluble in its own weight of water, or in five parts of alcohol. Unlike the preceding muriates, it is decomposed by ignition in a high temperature.

The muriates of magnesia and lime are generally contained in muriate of soda, and impart to that salt much of its deliquescent property. They impair, too, its power of preserving food. They are also ingredients of sea-water.

Art. 8.—*Muriate of Alumine*

May be formed by dissolving fresh precipitated alumine in muriatic acid; but the acid is always in excess. It is scarcely possible to obtain this salt in crystals; for, by evaporation, it assumes the state of a thick jelly. It is extremely soluble in water, and deliquescent when dry. In a high temperature it abandons its acid entirely.

Art. 9.—*Muriate of Glucine.*

This salt is little known. Like all the salts of glucine, it has a sweet taste, and crystallizes more readily than the nitrate.

Art. 10.—*Muriate of Zircon.*

Fresh precipitated zircon is readily dissolved by muriatic acid. The compound is colourless; has an astringent taste; and furnishes, by evaporation, small needle-shaped crystals, which lose their transparency in the air. It is very soluble in water and in alcohol. It is decomposed by heat, and by the saliva of the mouth. The gallic acid, poured into the solution, precipitates, if it be free from iron, a white powder. Carbonate of ammonia gives a precipitate, which is re-dissolved by an excess of the carbonate.

Art. 11.—*Muriate of Yttria.*

This compound has a striking resemblance to nitrate of yttria. Like that salt it dries with difficulty, and attracts moisture from the air. It does not crystallize, when evaporated, but forms a jelly.

SECTION III.

Oxygenized Muriatic Acid.

1. THIS acid may be formed by either of the following processes; which do not, however, afford a gas of precisely the same composition. That obtained by the third method is best adapted for exhibiting its effects on combustible substances.

Process 1. Into a stoppered retort introduce eight ounces of liquid muriatic acid, and four ounces of finely powdered manganese, and apply the heat of a lamp. A gas will be produced, which may be received over water in the usual manner. From the foregoing materials about 160 cubical inches of gas may be obtained.

Process 2. Mix eight ounces of muriate of soda (common salt) with three ounces of powdered manganese; put them into a stoppered retort, and pour on them four ounces of sulphuric acid, diluted previously with four ounces of water, and which has been suffered to cool after dilution. Or the proportions recommended by Thenard may be employed, *viz.* 1750 muriate of soda, 450 oxide of manganese, water and sulphuric acid each 800. On applying a gentle heat gas will be produced, as in Process 1. As the gas is absorbed by contact with water, though not rapidly, it should be received, when it is intended to be kept, in bottles filled with, and inverted in, water, and provided with ground stoppers. The stoppers must be introduced under water, while the bottle remains inverted.

Process 3. To an ounce-measure of liquid muriatic acid, contained in a gas bottle, add three or four drachms of a salt, which will presently be described, the hyper-oxygenized muriate of potash. If the acid be very concentrated, gas will be evolved without the application of heat; and, when it has ceased to be generated, the production will be renewed by a gentle heat, such as that obtained by immersing the gas bottle in a vessel of warm water. Receive the gas into bottles, furnished with extremely well-ground stoppers; and, when they are full, put the stopper intb its place, and invert the bottles, with their mouths downwards, in mercury.

As the gas, evolved by all these processes, is rapidly absorbed by cold water, the water of the receiving trough should be of the temperature of between 80° and 90° Fahrenheit; and the gas should remain, as short a time as possible before use, in contact with water. Mercury is not adapted for its reception, because on this metal the gas exerts a considerable action.

II. Oxygenized muriatic acid gas has the following properties:

(*a*) It has a deep yellow colour.

(*b*) It has a pungent and suffocating smell. In experiments on this gas, great care should be taken that it does not escape, in any considerable quantity, into the apartment; for its action on the lungs is extremely injurious and oppressive. It is heavier than common air (taking the statement of Gay Lussac) in the proportion of 2470 to 1000. It is composed of

77.65	muriatic acid
22.35	oxygen
100	

(*c*) By a temperature of 40°, it is reduced into a liquid form, and is condensed on the sides of the vessel. Hence some chemists have contended, that it should be classed among vapours, and not among gases.

When a receiver, filled with this gas, is surrounded by snow, or pounded ice, the gas forms on its surface a solid concretion, of a yellowish colour, resembling, in its ramifications, the ice which is deposited on the surface of windows during a frosty night. By a moderate increase of heat, such as to 50° Fahrenheit, this crust melts into a yellowish oily liquid, which, on a farther elevation of temperature, passes to the state of a gas.

(*d*) The oxygenized muriatic acid exerts powerful effects on various combustible bodies, both aëriform, liquid, and solid.

1. Let a phial, provided with a well-ground stopper, be filled with a mixture of hydrogen gas, and oxy-muriatic gas, procured by Process 3, in the proportion of three of the former to four of the latter. Put the stopper into its place, and keep the bottle, 24 hours, inverted with its mouth under water. On withdrawing the stopper, nearly the whole of the gas will have disappeared: and the remainder will be absorbed by the contact of water. In this experiment, the oxygen of the acid gas combines with the hydrogen, and forms water; and the acid returns to the state of common muriatic acid.

2. Mingle, in the detonating tube (fig. 28 or 29), three measures of hydrogen gas with four of oxygenized muriatic gas. When an electric spark is passed through the mixture, a detonation will ensue, and nearly the whole will be absorbed. The proportion of hydrogen and pure oxy-muriatic gases, required for mutual saturation, according to Mr. Cruickshank, is 8 of the former to $8\frac{1}{2}$ of the latter. If it be correct, however, as Gay Lussac has

stated, that the oxygen in this gas is equal to exactly half its volume, it ought to saturate precisely an equal bulk of hydrogen gas.

A singular fact, respecting the mutual action of oxy-muriatic acid and hydrogen gases, has been observed by Gay Lussac, and, without any knowledge of his experiments, by Mr. Dalton. A mixture of the two gases, in the proper proportion, is slowly absorbed under ordinary circumstances; but if the direct rays of the sun happen to fall on the mixture, the two gases diminish with considerable rapidity; and, if the quantity be large, they even explode. This is a singular instance of the agency of light in promoting chemical combination.

3. Both the foregoing experiments may be repeated, with the substitution of carburetted hydrogen gas from moistened charcoal, from distilled coal, or from stagnant water. A larger proportion of the oxygenized acid, however, must be used, *viz.* three or four measures to one of the combustible gas. In this case, also, the diminution is not complete; for there remains a quantity of carbonic acid gas, which may be absorbed by lime-liquor. If a proportion of oxy-muriatic gas, less than what is required for saturation, be fired with any of these gases, there is an abundant precipitation of charcoal. The direct rays of the sun's light quicken, also, the action of oxy-muriatic acid on carburetted hydrogen gas, and even occasion an explosion of the mixture.

4. The carbonic oxide is converted, by contact with the oxy-muriatic gas, into carbonic acid. Mix two measures of the carbonic oxide with two thirds of the oxygenized gas; and allow them to stand, for 24 hours, in a bottle which is entirely filled by the mixture. On withdrawing the stopper at this period under water, the water will rush

in, and will fill two thirds of the bottle. The remaining one third is carbonic acid gas, absorbable by lime-water. It is remarkable, that a mixture of these two gases cannot be set on fire by the electric spark; nor is it affected by the direct incidence of the sun's rays, like hydrogen and its compounds.

5. The olefiant gas exhibits, when mixed with oxy-muriatic gas, a very singular appearance. It is diminished in bulk, as rapidly as oxygen gas is by nitrous gas, and a thin film of oil forms on the surface of the water. The proportions required, for complete condensation, are $2\frac{1}{2}$ measures of the olefiant to 3 of the oxygenized gas. This phenomenon is not perfectly understood.

6. Oxy-muriatic gas exerts no action on nitrogen gas; but nitrous gas is condensed by it, in the same manner as by oxygen gas.

7. When mixed with sulphuretted hydrogen gas, it occasions a condensation, and a precipitation of sulphur. It condenses sulphurous acid gas into sulphuric acid.

8. Phosphorus introduced into the oxygenized acid gas, takes fire spontaneously, and burns vehemently. The results of this combustion were particularly examined by Mr. Davy. He introduced phosphorus into an exhausted receiver, and then admitted oxy-muriatic acid gas. The phosphorus took fire and burnt with a pale white flame. A white sublimate collected within the vessel, and a fluid as limpid as water trickled down its side. No muriatic acid gas was liberated; but this acid was found to exist in intimate combination with the phosphor*ic* acid, constituting the white sublimate; and, also, united with phosphor*ous* acid in the condensed liquid.

9. Sulphur is not inflamed by it; but a piece of sulphur fastened to the end of a glass rod, and confined in the

gas, is slowly oxygenized, and drops down in a liquid form.

By passing streams of this gas through flowers of sulphur, Dr. Thomson obtained a new combination of oxide of sulphur with muriatic acid, which he terms sulphuretted muriatic acid. His account of the properties of this substance may be found in the 6th volume of Nicholson's Journal. More recently it has been examined by Berthollet, jun. (*Mémoires d'Arcueil*, i.) who is of opinion that it is a triple compound of oxygen, sulphur, and muriatic acid, perfectly free both from sulphuric and sulphurous acids.

10. The charcoal of beech-wood, when finely powdered, and perfectly dry, is inflamed by this gas. So also is the pyrophorus of Homberg.

In the Philosophical Transactions (for 1809, page 466), Mr. Davy, however, has given an account of an experiment, in which charcoal, in a state of ignition, on being brought into contact with oxy-muriatic acid gas, was instantly extinguished. When charcoal was kept intensely ignited in a globe filled with oxy-muriatic acid gas, very little effect was produced. The hydrogen, evolved from the charcoal, united with a part of the oxygen of the acid, and formed aqueous vapour, which again composed common muriatic acid with the de-oxygenated portion of the acid. After this had been accomplished, no farther change ensued. This furnishes one proof among others that oxy-muriatic acid cannot be decomposed without the presence either of hydrogen or of water.

11. Almost every metal, in a state of minute division, takes fire spontaneously, and burns in this gas. The very malleable metals, such as gold, silver, &c., which can be

reduced to extremely thin leaves, are best applied to the gas in this state. Others, as iron, zinc, copper, &c., must be introduced in the state of fine filings. The most readily oxydized metals burn with the greatest brilliancy. The proportion is about 40 grains of each metal to 40 cubic inches of gas; and, into the bottom of the receiver a little sand may be poured, to prevent it from being broken.

Metallic antimony burns with a very brilliant white flame, and throws out sparks. Arsenic exhibits a fine green or blue flame, attended with sparks, and a dense white smoke; bismuth, a blueish flame; nickel, a yellowish white one; cobalt, a blueish white; zinc, a white flame and sparks; tin, a blueish white light; lead, a clear white flame; copper, a red and slowly spreading light; and iron, a bright red light. In all these experiments, the temperature of the gas should not fall short of 70°; and, to ensure their success, it should be prepared by the third process.

(*e*) The oxygenized muriatic gas destroys all vegetable colours. This may be shown by passing into it, through water, a piece of cloth, or of paper, stained with litmus, the colour of which will speedily disappear. Hence the application of this gas to the purpose of bleaching. Its efficacy, in this mode, may be seen by confining in it a pattern of unbleached calico.

(*f*) This gas is absorbed by water; slowly, if allowed to stand over it quiescent, but rapidly when agitated.

The best method of effecting the impregnation of water with this gas, is by means of a Woulfe's apparatus, the bottles of which should be surrounded by ice-cold water. The precise bulk of the gas, which water is capable of absorbing, is not ascertained. According to the proportions stated by Berthollet, 1000 grains of water, at the temperature of 43° Fahrenheit, take up 1073 grains of the gas, and acquire the specific gravity of 1003.

(*g*) The watery solution, if perfectly free from common muriatic acid, has not the usual taste of an acid, but an astringent one. Its purity from muriatic acid may be ascertained by a solution of nitrate of mercury, which is precipitated by the common, but not by the oxygenized acid.

(*h*) The watery solution acquires the colour and peculiar smell of the gas, and has a similar property of discharging vegetable colours. Hence it may be employed in bleaching.

That its action in bleaching is attended with the decomposition of the acid, and the deprivation of its oxygen, may be ascertained by examining a portion of the liquid, by means of which several patterns of unbleached calico have been successively whitened. The liquor will be found almost entirely to have lost its smell, and to precipitate nitrate of mercury abundantly.

(*i*) When the watery solution is exposed to a temperature a little above that of freezing water, the gas, which is combined with it, separates in the form of a liquid, heavier than water.

(*k*) The oxygenized acid is not decomposed by the temperature of boiling water; for it may be raised in distillation, and again condensed without change.

(*l*) When this solution is exposed to the direct rays of the sun, the oxygenized acid is decomposed; its oxygen escapes in the form of a gas, and it is reduced to the state of common muriatic acid. The oxygen gas may be collected, by exposing the solution in a gas bottle furnished with a ground stopper and bent tube, which terminates in the pneumato-chemical apparatus.

(*m*) The oxygenized muriatic acid combines with alkalis, and forms peculiar compounds.

* * *

SECTION IV.

Hyper-oxygenized Muriates.

ART. 1.—*Hyper-oxygenized Muriate of Potash.*

THE properties of this salt were discovered by Berthollet. It may be formed by passing the oxygenized muriatic acid gas, as it proceeds from the mixture of muriate of soda, sulphuric acid, and manganese (see the preceding section, Process 2) through a solution of caustic potash. This may be done by means of Woulfe's apparatus, using only one three-necked bottle in addition to the balloon. The tube, which is immersed in the alkaline solution, should be at least half an inch in diameter, to prevent its being choked up by any crystals that may form. The solution when saturated with the gas, may be gently evaporated, and the first products only of crystals are to be reserved for use; for the subsequent products consist of common muriate of potash only.—Now, since the gas, when it first came into contact with the alkaline solution, was purely oxygenized muriatic acid, it follows that a part of this acid must have been dis-oxygenized during the absorption, and have returned back to the state of common muriatic acid. Let us suppose the oxygenized acid, when first presented to the alkaline solution, to be divided into two portions; one of these gives up its excess of oxygen to the other half, returns to the state of common muriatic acid, and, combining with the alkali, forms muriate of potash.—The latter portion, therefore, is oxygenized acid, *plus* a certain quantity of oxygen; and this, uniting with another portion of alkali, forms a salt, which Mr. Chenevix has termed hyper-oxygenized muriate. Strictly speaking, simple oxygenized muriates do not exist; for, in all this class of salts, the acid contains 65 *per cent.* of oxygen; whereas, the oxygenized acid contains only 16 *per cent.* It might be expected that a stronger acid, such as the sulphuric, would expel from these salts the hyper-oxygenized acid in the form of gas; but this acid, by the temperature necessary for its liberation, is partially decomposed, and again returns nearly to the state of oxygenized acid.

The reader, who wishes for farther information on this subject, is referred to a masterly paper of Mr. Chenevix, in the Philosophical Transactions, 1801; reprinted in Nicholson's 8vo. Journal, vol. i. and in the Philosophical Magazine; and also to Mr. Hoyle's Essay, in the 5th volume of the Manchester Memoirs.

The hyper-oxygenized muriate of potash has the following qualities:

(*a*) It has the form of shining hexaedral laminæ, or rhomboidal plates.

(*b*) One part of the salt requires 17 of cold water for solution, but five parts of hot water take up two of the salt.

(*c*) It is not decomposed by exposure to the direct rays of the sun, either in a crystallized or dissolved state.

(*d*) When the hyper-oxygenized muriate is submitted to distillation in a coated glass retort, it first fuses, and, on a farther increase of temperature, yields oxygen gas of great purity. A hundred grains of the salt afford 75 cubic inches of gas, containing only about three *per cent.* of nitrogen gas.

(*e*) The hyper-oxygenized muriate of potash has no power of discharging vegetable colours; but the addition of a little of the sulphuric acid, by setting the oxygenized acid at liberty, developes this property.

(*f*) The salt is decomposed by the stronger acids, as the sulphuric and nitric acids. This will be proved by dropping a few grains of the salt into a little concentrated sulphuric acid. A strong smell will arise, and, if the quantities be sufficiently large, an explosion will ensue. The experiment should, therefore, be attempted with great caution. When this mixture is made at the bottom of a deep vessel, the vessel is filled with oxygenized muriatic gas, which inflames sulphuric ether, alcohol, or oil of turpentine, when poured into it; and also camphor, resin, tallow, elastic gum, &c. (Davy.)

Muriatic acid, as has already been stated, disengages the oxygenized acid, and the addition of a few grains of the salt to an ounce measure of the acid, imparts to it the property of discharging vegetable colours.

(*g*) This salt exerts powerful effects on inflammable bodies.

1. Rub two grains into powder in a mortar, and add one grain of sulphur. Mix them very accurately, by gentle triture, and then, having collected the mixture to one part of the mortar, press the pestle down upon it suddenly, and forcibly. A loud detonation will ensue.—Or, if the mixed ingredients be wrapped in some strong paper, and then struck with a hammer, a still louder report will be produced.

2. Mix five grains of the salt with half the quantity of powdered charcoal in a similar manner. On triturating the mixture strongly, it will inflame, especially with the addition of a grain or two of sulphur, but not with much noise.

3. Mix a small quantity of sugar with half its weight of the salt, and on the mixture pour a little strong sulphuric acid. A sudden and vehement inflammation will be produced. This experiment, as well as the following, requires caution.

4. To one grain of the powdered salt, in a mortar, add about half a grain of phosphorus. The phosphorus will detonate, on the gentlest triture, with a very loud report. The hand should be covered with a glove in making this experiment, and care should be taken that the phosphorus, in an inflamed state, does not fly into the eyes.—Phosphorus may also be inflamed under the surface of water by means of this salt. Put into a wine glass, one part of phosphorus with two of the salt; fill it nearly with water, and pour in, by means of a glass tube, reaching to the bottom, three or four parts of sulphuric acid. The phosphorus takes fire, and burns vividly under the water. This experiment requires caution, lest the inflamed phosphorus should be thrown into the eyes. (Davy.) Oil may also be thus inflamed on the surface of water, the experiment being made with the omission of the phosphorus, and the substitution of a little olive or linseed oil.

5. Hyper-oxygenized muriate may be substituted for nitre in the preparation of gunpowder, but the mixture of the ingredients requires extreme circumspection. It may be proper also to state, that this salt should not be kept mixed with sulphur in considerable quantity, such mixtures having been known to detonate spontaneously.

ART. 2.—*Hyper-oxygenized Muriate of Soda.*

This salt may be obtained, by following the process already described, with the substitution of pure soda for potash. It is exceedingly difficult, however, to obtain it pure; because it nearly agrees, in solubility, with the common muriate of soda. It is soluble in three parts of cold water, and in rather less of hot, and is slightly deliquescent. It is soluble also in alcohol; but so also, according to Mr. Chenevix, is the common muriate. It crystallizes in cubes, or in rhomboids approaching the cube in form. In the mouth it produces a sensation of cold, and a taste scarcely to be discriminated from that of muriate of soda. In other properties it agrees with the similar salt with base of potash.

ART. 3.—*Hyper-oxygenized Muriate of Ammonia.*

This salt cannot be procured by the direct union of the oxygenized acid with pure ammonia, because these two bodies mutually decompose each other; as will appear from the following experiments.

1. Fill a pint receiver with the oxygenized acid, prepared by the third process: and pour into it half a drachm of the strongest solution of ammonia that can be procured. A detonation will presently ensue.

2. Fill a four-ounce bottle with the oxygenized acid, and invert it in a cup containing four ounce-measures of the solution of pure ammonia. Presently the liquor will be absorbed, and a detonation will ensue, which will throw down the bottle, unless firmly held by the hand. In the bottle there remains a portion of nitrogen gas.

3. Pass the oxygenized acid through a solution of ammonia in a Woulfe's bottle; from one neck of which a tube proceeds, and terminates under the inverted funnel of the pneumatic trough. Bubbles of gas will be formed and may be collected in an inverted receiver. They consist of pure nitrogen gas.

In all these experiments, the redundant oxygen of the acid unites with the hydrogen of the ammonia, and forms water, while the nitrogen of the ammonia is liberated in a gaseous state. A combination of ammonia, with the hyper-oxygenized muriatic acid may be formed, by adding a solution of carbonate of ammonia to one of hyper-oxymuriate of lime. A salt is obtained, on evaporation, which is very soluble in water and in alcohol, and which is decomposed at a moderate temperature, yielding a quantity of gas and a smell of oxymuriatic acid. Its properties have not been much investigated.

The property which ammonia possesses, of decomposing oxygenized muriatic acid, renders it extremely useful in correcting the offensive vapours of that gas, which are sometimes accidentally set at liberty in places where it is prepared. And when suffocation threatens to come on, in consequence of the fumes of the oxygenized acid, the most effectual remedy is to hold a stopper, moistened with ammonia, to the mouth and nostrils.

ART. 4.—*Remaining Hyper-oxygenized Muriates.*

To effect the combination of barytes and strontites with the hyper oxygenized acid, those bases must be dissolved in hot water, which must be kept hot while the current of gas is transmitted through the solution. Lime may also be combined with the oxygenized acid, and must, to this end, be kept suspended by mechanical means. This compound derives importance from its application to the art of bleaching; for it possesses, when perfectly saturated, bleaching properties; and in this state produces whiteness in the unbleached part of goods, without destroying any delicate colours which they may contain. The salt, with base of lime, is extremely deliquescent; liquefies at a low heat, and is soluble in alcohol. It produces much cold by solution, and a sharp taste in the mouth.

For an account of the remaining salts formed with this acid, Mr. Chenevix's paper may be consulted.

SECTION V.

Nitro-muriatic Acid.

THIS acid is a compound of the nitric and muriatic acids, and may be formed most commodiously by mixing two parts of nitric acid with one of muriatic. Though the acids employed are both perfectly pale, yet the mixture becomes of a deep red colour, a brisk effervescence takes place, and pungent vapours of oxy-muriatic acid are evolved. The production of the latter acid is occasioned by the combination of the muriatic with part of the oxygen of the nitric acid, in consequence of which the latter is converted either into nitrous gas or nitrous acid. Hence the nitro-muriatic acid does not contain nitric acid, but one of an inferior degree of oxygenation. Its most distinguishing property, that of dissolving gold, will be described hereafter.

The nitro-muriatic acid does not form, with alkaline, or other bases, a distinct genus of salts, entitled to the name of nitro-muriates; for, when combined with an alkali, or an earth, the solution yields, on evaporation, a mixture of a muriate and a nitrate; and metallic bodies, dissolved in it, yield muriates only. In the latter case, the nitric acid is decomposed, oxydizes the metal, and renders it soluble in muriatic acid.

SECTION VI.

Murio-Sulphuric Acid.

MURIATIC acid gas is absorbed in considerable quantity by sulphuric acid. The compound has a brown colour, and when exposed to the air emits copious white fumes. It has no particular uses.

CHAPTER XV.

PHOSPHORUS,—PHOSPHORIC ACID,—PHOSPHOROUS ACID,—PHOSPHATES.

SECTION I.

Phosphorus.

I. PHOSPHORUS is an inflammable substance, and is known by the following external characters.

(*a*) It has generally a flesh-red colour, but, when carefully purified, may be obtained as free from colour, and as transparent, as melted white wax.

(*b*) It is so soft that it readily yields to the knife.

(*c*) It melts with a very gentle heat. To show this, it must be covered with water, to prevent it from inflaming.

(*d*) In the atmosphere it emits a white smoke, and peculiar smell; and a faint and beautiful light arises from it.

II. Phosphorus is inflamed by the application of a very gentle heat. According to Dr. Higgins, a temperature of 60° is sufficient to set it on fire, when perfectly dry. It burns with a very brilliant light, a white smoke, and a suffocating smell.

1. It may be set on fire by friction. Rub a very small bit between two pieces of brown paper; the phosphorus will inflame, and will set the paper on fire also.

2. In oxygen gas it burns with a very beautiful light; and also in nitrous oxide, and in oxygenized muriatic acid.

III. Phosphorus is volatile at 550°. Hence it may be raised by distillation; but, to prevent its taking fire on the application of heat, the retort should previously be filled with azotic or hydrogen gas, and the mouth of the retort be immersed in water.

To accomplish this, the quantity of phosphorus, which it is intended to rectify, should first be put into the retort, with a sufficient portion of water to cover it. The water must then be made hot enough to melt the phosphorus, which, on cooling, forms a compact mass, of the shape of the bottom of the retort. When cold, fill the retort, and its neck also, with water, and invert it in water. Displace the water by hydrogen gas, forced from a bladder through a bent pipe; keep the finger on the open end of the retort neck; place it in a sand bath; and immerse the mouth of it in water. Then apply heat very cautiously. A bladder should also be provided, furnished with a stop-cock and brass pipe, and filled with hydrogen gas. During the distillation, the gas, in the retort, is absorbed, and it is necessary to add more from the bladder, otherwise the water will rush into the retort, and occasion an explosion. By distillation, in this mode, phosphorus is rendered much purer. In the neck of the retort a substance is condensed of a beautiful red or carmine colour, which is a combination of carbon and phosphorus, or a *phosphuret of carbon.*

The only information, which we possess, respecting the nature of phosphorus, is derived from the electro-chemical researches of Mr. Davy. When acted upon by a battery of 500 pairs of plates in the same manner as sulphur, gas was produced in considerable quantities, and the phosphorus became of a deep red-brown colour. The gas proved to be phosphuretted hydrogen, and was equal in bulk to about four times the phosphorus employed. Hence hydrogen is proved to be one of its components.

A grain of potassium and a grain of phosphorus, being fused out of the contact of air, combined with a most vivid light and intense ignition. During the process, one tenth of a cubic inch of phosphuretted hydrogen was evolved. The compound, when made to act on water, separated less hydrogen than the potassium itself would have given; and the more the proportion of potassium to phosphorus was increased, the less was the hydrogen evolved. This fact proves that the potassium must have derived oxygen from the phosphorus, which base may be considered as another of its constituents.

What is the nature of the substance, with which the hydrogen and oxygen, contained in phosphorus, are combined, we are still however ignorant. The base of phosphorus, therefore, remains a subject for future investigation.

IV. Phosphorus may be oxygenized in various modes*.

(*a*) By mere exposure to atmospheric air. Let a stick of phosphorus be placed in a funnel, the pipe of which terminates in an empty bottle. The phosphorus will be slowly oxygenized, and, after some time, will be wholly changed into an acid, which will fall into the bottle in a liquid state.

A large quantity of acid may be obtained, if a number of sticks be thus exposed: and as they would be in danger of taking fire, if heaped together, each stick should be enclosed in a glass tube, of rather larger diameter than itself. These tubes must be disposed round a funnel, the pipe of which terminates in a bottle. The whole should be covered by a bell-shaped receiver, the air of which is to be frequently changed. The acid thus obtained is termed the *Phosphorous Acid.*

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* On the oxides of phosphorus, see Nicholson's Journal, vi. 133.

(*b*) By combustion in oxygen gas, or in atmospheric air. When burnt in this manner, every hundred parts of phosphorus, according to Lavoisier, gain an addition of 154. (See the account of this experiment in the 5th chapter of his Elements.)

(*c*) By the nitric acid. If phosphorus be cautiously added, by a little at once, to nitric acid, heated in a mattras, the nitric acid is decomposed, and its oxygen, uniting with the phosphorus, constitutes phosphoric acid.

A tubulated retort must be used for this purpose; and its neck may terminate in the apparatus already described for procuring nitric acid. By this contrivance a considerable quantity of acid will be saved.

(*d*) A similar effect is produced by oxygenized muriatic acid in a liquid state. The operation of this acid, in a gaseous form, has already been described, under the article oxy-muriatic acid; and two products, it has been stated, are obtained by this operation, *viz.* a solid white sublimate and a greenish yellow liquid.

The solid sublimate Mr. Davy considers as a compound of muriatic and phosphoric acids, both quite free from water; and hence it might be termed *murio-phosphoric acid.* When brought into contact with water, muriatic acid gas was evolved from it, and phosphoric and muriatic acids remained combined with the water. It was a non-conductor of electricity, and did not burn when heated; but sublimed at about 212°, leaving no residuum.

Accordingly as the phosphoric acid is differently prepared, its degree of oxygenation differs, and its properties are found to vary proportionably.

SECTION II.

Phosphoric Acid.

1. To prepare this acid, the process *b*, *c*, or *d*, sect. I. may be employed; but the following is the most economical method.

On 20 pounds of bone, calcined to whiteness and finely powdered, pour 20 quarts of boiling water, and add eight pounds of sulphuric acid, diluted with an equal weight of water. Let these materials be well stirred together, and be kept in mixture about 24 hours. Let the whole mass be next put into a conical bag of sufficiently porous and strong linen, in order to separate the clear liquor, and let it be washed with water till the water ceases to have much acidity to the taste. Evaporate the strained liquor in earthen vessels, placed in a sand-heat, and, when reduced to about half its bulk, let it cool. A white sediment will form in considerable quantity, which must be allowed to subside; the clear solution must be decanted, and boiled to dryness in a glass vessel. A white mass will remain, which is the dry phosphoric acid. This may be fused in a crucible, and poured out on a clean copper dish. A transparent glass is obtained, which is the phosphoric acid in a glacial state; not, however, perfectly pure, but containing sulphate and phosphate of lime.—According to Fourcroy and Vauquelin, it is, in fact, a super-phosphate of lime, containing, in 100 parts, only 30 of uncombined phosphoric acid, and 70 of neutral phosphate of lime. The glacial acid, however, may be prepared from perfectly pure phosphoric acid. It is singular that, according to the experiments of Berthier, it contains at least one fourth its weight of water, a proportion which could scarcely have been expected in so hard a substance.

To procure the phosphoric acid perfectly pure, the oxygenation of phosphorus by nitric acid, is the most eligible process (*c*, of the preceding article). The undecomposed nitric acid must be separated by distillation in a glass retort, and the dry mass, when fused, affords also glacial phosphoric acid.

II. The phosphoric acid has the following properties:

(*a*) When pure it dissolves readily in water. That obtained immediately from bones is rendered insoluble by the admixture of earthy salts. But the glacial acid, prepared with nitric acid, is readily soluble.

(*b*) It is not volatile, nor capable of being decomposed by heat only, nor does it emit any smell when heated.

(*c*) It is composed, according to the experiments of Rose, of

46.72	phosphorus
53.28	oxygen
100	

(*d*) When distilled in an earthen retort with powdered charcoal, it is decomposed; its oxygen, uniting with the carbon, forms carbonic acid, and the phosphorus rises in a separate state. This is the usual and best mode of obtaining phosphorus.

The phosphoric acid may either be employed for this purpose in the state of glass, finely powdered, and mixed with its weight of pulverized charcoal; or to the evaporated acid of bones, when acquiring a thick consistence, powdered charcoal may be added, in sufficient quantity, to give it solidity. In the latter mode, however, the materials are apt to swell, and to boil over. The mixture of acid and charcoal is then to be put into a stoneware retort, coated with Willis's lute, and the neck of which is lengthened out by a tin pipe. The open end of the pipe is to be immersed in a vessel of water. The heat is to be slowly raised, and at length made very intense. An enormous quantity of gas escapes, which takes fire on coming into contact with the atmosphere; and the phosphorus distils over in drops, which congeal in the water. As it is apt also to condense in, and to stop up, the neck of the retort and tin-pipe, it must be occasionally melted out of these, by a shovel full of hot cinders, held under them. The process is rather a difficult one; and though it is proper that the student should repeat it once, in order to complete a course of experiments, it will be found more economical to purchase, from the London preparers, the phosphorus which may be required for experiments.

Phosphorus may also be procured, by adding to urine a solution of lead in nitric acid, which precipitates a phosphate of lead. This, when well washed, dried, and distilled in a stoneware retort, yields phosphorus (see Crell's Journal, Translation, iii. 56): or a solution of phosphate of soda (which may be bought at the druggists), mixed with one of acetite of lead, in the proportion of one part of the former salt to 1¼ of the latter, yields a precipitate of phosphate of lead, from which phosphorus may be procured by distillation.

SECTION III.

Phosphates.

With alkaline and earthy bases, the phosphoric acid composes a class of salts called Phosphates, which have the following generic characters.

1. When heated with charcoal, they are not decomposed, nor is phosphorus obtained.

2. They melt, before the blow-pipe, into a hard globule, sometimes transparent, at others opake.

3. They are soluble in nitric and muriatic acids, without effervescence, and are precipitated from those acids by lime-water and pure ammonia.

4. They are decomposed, in part, by sulphuric acid, and yield a liquor which, on evaporation and distillation with charcoal, affords phosphorus.

The phosphate of soda is the only one of these salts which has any important use. It has been introduced into medicine, by Dr. Pearson, as a purgative, the purposes of which it answers, unaccompanied by any nauseous taste. The phosphate of lime, besides being found in the earth in a mineral form, constitutes a large part of the solid matter of animal bones.

As an enumeration of these salts can scarcely be interesting to the general student, and as the properties of the various salts, already described, furnish abundant discriminating characters of the different alkalis and earths, I deem it sufficient to refer, for a detailed account of them, to Dr. Thomson's Elements, or La Grange's Manual.

SECTION IV.

Phosphorous Acid—Phosphites.

The phosphorous acid (procured by exposing sticks of phosphorus to the slow action of air in the way already described, sect. 1) exhales a disagreeable fœtid odour; and yields, when heated, penetrating white vapours. When heated in a glass ball, blown at the end of a small tube, a gas issues from the orifice of the tube, which inflames on coming into contact with the atmosphere. Hence it appears to contain an excess of phosphorus. The residuum in the ball is phosphoric acid. From the experiments of Rose on the phosphoric acid, Gay Lussac infers that, conformably to his own hypothetical views, phosphorous acid must consist of

56.81	phosphorus
43.19	oxygen
100	

Phosphorous acid (it has already been stated under the article oxy-muriatic acid) combines with muriatic acid, and forms a liquor of a pale greenish yellow tint, and very limpid. When exposed to the air this liquid rapidly disappears, emitting dense white fumes, the smell of which differs a little from that of muriatic acid. It may perhaps not improperly be denominated *murio-phosphorous acid.*

This acid reddens litmus paper in its common state, but has no effect on litmus paper dipped into it, instantly after being well dried. This fact is extremely curious, inasmuch as it suggests that the property of reddening litmus belongs to acids, only when combined with water. It is a non-conductor of electricity. When mingled with water, much heat is produced, and muriatic acid gas is evolved.

The same compound may be obtained by distilling together oxy-muriate of mercury and phosphorus, or calomel and phosphorus, but in largest quantity from the first mentioned mixture.

The combinations of phosphorous acid with alkaline and earthy bases are called phosphites.

The phosphites differ considerably in their characters from phosphates.

1. They exhale a smell of phosphorus.

2. When heated, they emit a phosphorescent flame.

3. Distilled in a strong heat, they yield a little phosphorus, and are converted into phosphates.

4. They detonate, when heated with oxy-muriate of potash.

5. They are changed into phosphates by nitric, and by oxygenized muriatic acid.

SECTION V.

Binary Compounds of Phosphorus.

I. Phosphorus is susceptible of combination with sulphur, and affords a compound, the properties of which vary, according to the proportion of its ingredients. It may be obtained by melting these substances together in a tube, the mouth of which is loosely stopped by paper; or by fusing these two bodies, very cautiously, at the bottom of a Florence oil flask, nearly filled with water. The process is attended with some danger; and requires several precautions, which will be suggested by the essays of Accum and Briggs, published in the 6th and 7th volumes of Nicholson's Journal. The compound is much more fusible and combustible, than the separate components.

II. Phosphorus combines with the pure fixed alkalis, and with earths, and composes the class of phosphurets. That of lime is the most readily formed, and exhibits, extremely well, the properties of these compounds. It is prepared as follows:

Take a glass tube, about 12 inches long and one third of an inch diameter, sealed hermetically at one end. Let this tube be coated with clay, except within about half an inch of the sealed end. Put first into it a drachm or two of phosphorus, cut into small pieces, and then fill the tube with small bits of fresh burnt lime, of the size of split peas. Stop the mouth of the tube loosely with a little paper, in order to prevent the free access of air.—Next, heat to redness that part of the tube which is coated with clay, by means of a chafing-dish of red-hot charcoal; and, when the lime may be supposed to be ignited, apply heat to the part containing the phosphorus, so as to sublime it, and to bring the vapour of it into contact with the heated lime. The lime and phosphorus will unite, and will afford a compound of a reddish-brown colour.

If the carbonate of lime be substituted for pure lime, the carbonic acid is decomposed. Its carbon is set at liberty, and appears in the state of charcoal; while its oxygen unites with the phosphorus; and the phosphoric acid, thus produced, forms phosphate of lime. In this experiment, carbonic acid is decomposed by the conspiring affinities of phosphorus for oxygen, and of lime for phosphoric acid, though the former affinity only would be inadequate to produce the effect.

The phosphuret of lime has the remarkable property of decomposing water at the common temperature of the atmosphere. Drop a small piece of it into a wine-glass of water, and in a short time bubbles of phosphuretted hydrogen gas will be produced; which, rising to the surface, will take fire, and explode. If the phosphuret of lime be not perfectly fresh, it may be proper to warm the water to which it is added.

Into an ale-glass put one part of phosphuret of lime, in pieces about the size of a pea (not in powder), and add to it half a part of hyper-oxygenized muriate of potash. Fill the glass with water, and put into it a funnel, with a long pipe, or narrow glass tube, reaching to the bottom. Through this pour three or four parts of strong sulphuric acid, which will decompose the hyper-oxygenized salt; and the phosphuret also decomposing the water at the same time, flashes of fire dart from the surface of the fluid, and the bottom of the vessel is illuminated by a beautiful green light. (Davy.)

Another combination of phosphorus, the properties of which render it a fit subject of amusing experiments, is the phosphuretted hydrogen gas.

ART. 3.—*Phosphuretted Hydrogen Gas.*

1. This gas may be procured, by boiling, in a retort, a little phosphorus with a solution of pure potash. The water is decomposed; its oxygen, uniting with the phosphorus, forms phosphoric acid, which combines with the alkali, while the hydrogen dissolves another portion of phosphorus, constituting phosphuretted hydrogen gas.—This gas may also be obtained, by putting into five parts of water half a part of phosphorus, cut into very small pieces, with one of finely granulated zinc, and adding three parts of strong sulphuric acid. This affords a pretty experiment. The gas is disengaged in small bubbles, which cover the whole surface of the fluid, and take fire on reaching the air; these are succeeded by others, and a well of fire is produced. (Davy.)

In preparing this gas, the body of the retort should be filled, as nearly as possible, with the alkaline solution*; otherwise the gas, when produced, will inflame and diminish the air within the retort, and the water will ascend from the trough. This accident may be effectually prevented, by previously filling the retort with hydrogen gas.

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* The phosphorus should first be melted, under water, in the retort; which is to be emptied when the phosphorus has congealed, and then entirely filled by the alkaline solution. Of this, a sufficient portion is to be displaced by hydrogen gas, forced through a bent pipe from a bladder.

II. The properties of this gas are the following:

(*a*) It takes fire immediately on coming into contact with the air. This may be shown by letting it escape into the air, as it issues from the retort, when a very beautiful appearance will ensue. A circular dense white smoke rises in the form of a horizontal ring, which enlarges its diameter as it ascends, and forms a kind of corona.

(*b*) When mixed suddenly with oxygen gas it detonates. This experiment should be made cautiously, and in small quantity.

(*c*) The same phenomenon ensues on mixing it with oxygenized muriatic acid gas, or with nitrous oxide.

When mingled with any of these gases, it should be passed up by not more than a bubble or two at once.

(*d*) Sulphurous acid and phosphuretted hydrogen gases, when mingled together, mutually decompose each other.

(*e*) It deposits phosphorus, by standing, on the inner surface of the receiver, and loses its property of spontaneous accension.

Three grains of potassium were heated by Mr. Davy in sixteen cubic inches of phosphuretted hydrogen gas. As soon as it was fused, the retort was filled with white fumes, and a reddish substance precipitated upon the sides and upper part of it. Though the heat was continued for some minutes, yet no inflammation took place. When the retort was cool, the absorption was found to be less than a cubical inch. The potassium externally was deep brown, and internally of a dull lead colour. The potassium afterwards evolved less hydrogen from water than it would have detached before; and hence it must have been partially oxidated. Now the only source of oxygen, in this experiment, is the phosphorus, which the gas holds in solution;

and we derive, therefore, from the fact, an additional proof of the presence of oxygen in phosphorus.

Phosphorus is also soluble in oils; and, when thus dissolved, forms what has been called liquid phosphorus, which may be rubbed on the face and hands without injury. It dissolves too in ether, and a very beautiful experiment consists in pouring this phosphoric ether in small portions, and in a dark place, on the surface of hot water.

The phosphoric matches consist of phosphorus extremely dry, minutely divided, and perhaps a little oxygenized.— The simplest mode of making them is to put a little phosphorus, dried by blotting-paper, into a small phial; heat the phial, and when the phosphorus is melted, turn it round, so that the phosphorus may adhere to the sides. Cork the phial closely; and it is prepared. On putting a common sulphur-match into the bottle, and stirring it about, the phosphorus will adhere to the match, and will take fire when brought out into the air.

A correspondent in Nicholson's Journal (xi. 137) proposes to effect the oxydation of phosphorus, by exposing 100 grains to half a pint of oxy-muriatic acid gas. The vessel must afterwards be well corked up. On applying a sulphur-match to the oxygenized phosphorus, it is inflamed, as in the foregoing example.

CHAPTER XVI.

BORACIC ACID.

I. THIS acid is very rarely found native; and, for purposes of experiment, is obtained from the purified borax of commerce, by one of the following processes:

1. To a solution of borax, in boiling water, add half its weight of sulphuric acid, previously diluted with an equal quantity of water. Evaporate the solution a little; and, on cooling, shining scaly crystals will appear, which consist of boracic acid. Let them be well washed with distilled water, and dried on filtering paper.

2. Let any quantity of borax be put into a retort, with half its weight of sulphuric acid, and half its weight of water. Boracic acid may be obtained by distillation, and may be purified, by washing in water, &c., as before. By neither of these processes, however, is it obtained perfectly pure; for electrical analysis discovers in it a minute portion both of alkali and of sulphuric acid. (Davy.)

II. Boracic acid has the following qualities:

1. It has a solid form, is destitute of smell, and nearly so of taste.

2. It fuses, when heated, and loses its water of crystallization. If the heat be increased suddenly, before it has lost its water of crystallization, it sublimes; but, otherwise, it melts into a glass, which is permanent in the strongest fire.

3. It is soluble in twelve parts of cold water, and in three or four of boiling water.

4. This solution reddens vegetable blue colours, and effervesces with alkaline carbonates.

5. It is soluble in alcohol, and the solution burns with a beautiful green flame.

The boracic acid, which had resisted all other means of analysis, has at length yielded to the attempts of Mr. Davy to decompose it by the action of Voltaic electricity. When moistened with water, and exposed, between two surfaces of platina, to a battery of 500 pairs of plates, an olive-brown matter immediately began to form on the negative surface, which gradually increased in thickness, and at length became almost black. It was not changed by water, but dissolved with effervescence in warm nitrous acid. When heated to redness on platina, it burned slowly, and gave off white fumes, which had acid properties. A black mass remained which, when examined by the magnifier, appeared vitreous at the surface, and evidently contained a fixed acid.

Boracium.

As this peculiar combustible substance was a non-conductor of electricity, it was found impossible to obtain it in this way, except in very thin films. By the action of potassium, however, it was procured in larger quantities. Twelve or fourteen grains of boracic acid were heated, in a green glass tube, with the same quantity of potassium. A most intense ignition ensued; and the potassium, where it was in contact with the boracic acid, entered into vivid inflammation. A quantity of hydrogen gas appeared, equal to about twice the bulk of the acid. To collect the results; tubes of metal were employed; for the most part of brass, which appears to answer best. The residue in the tube was dissolved in water, and the insoluble part collected on a filter. Its properties are described as follows:

1. It is in the form of a powder, in colour of the darkest shades of olive. It is very friable, and not sufficiently hard to scratch glass. It is a non-conductor of electricity.

2. When it has been dried at only 100° or 120°, it gives off moisture by increase of temperature. In the atmosphere it takes fire, at a heat below that of boiling olive oil; and burns with a red light, and scintillations like charcoal.

3. It is not decomposed by heat in a platina tube, though raised to whiteness. The only change in it appears to be an increase of specific gravity.

4. In oxygen gas it burns with a most brilliant light, and is partly converted into boracic acid, and partly into a black substance, which requires a higher temperature for its inflammation, and produces a fresh quantity of boracic acid.

5. In oxymuriatic acid gas, it takes fire at common temperatures, and boracic acid is regenerated with a portion of the black matter already described.

6. It was not soluble either in nitrogen or hydrogen gases.

7. It decomposed the nitric and sulphuric acids, and boracic acid was produced.

8. It combined with alkalis, and gave pale olive coloured compounds, from which dark precipitates were separated by muriatic acid.

9. It slowly combined with melted sulphur, which acquired an olive tint; but with phosphorus scarcely any

union seemed to take place. Neither did it combine with mercury.

These qualities are sufficient to show that the combustible substance, obtained from boracic acid, and constituting its base, is different from every other known species of matter. Mr. Davy has, therefore, proposed for it the term BORACIUM. As to its nature, he is of opinion that it is probably a compound, and that one of its ingredients, which enters into alloy with potassium and with iron, is the true basis of the boracic acid. The olive coloured substance, whose properties have been already described, he believes to consist of this basis, united with a little oxygen; that when farther oxydized it forms the black matter; and that, in its full state of oxygenation, it constitutes boracic acid.

The proportion of ingredients in the boracic acid has not been accurately determined. It is stated, merely as an approximation, that it consists of one part by weight of inflammable base united with two parts of oxygen. In the black substance, Mr. Davy supposes that about three parts of the inflammable base are combined with only one of oxygen.

Boracic acid combines with alkalis and earths; but the only important combination, which it forms, is with soda. This compound is found native in India, and is brought to this country, under the name of tincal, or brute borax, which, when purified, affords the borax of the shops. In the borate of soda, the alkaline ingredient is in excess, and hence the salt converts vegetable blue colours to green. It is therefore, in strictness, a sub-borate.

Sub-borate of soda crystallizes in prisms with six irregular sides. It effloresces in the air. It fuses when ignited; loses its water of crystallization; and leaves a glass, which is transparent when cold, and which is of great use in experiments with the blow-pipe. The salt dissolves in twelve parts of cold water, or in six of boiling water. It is susceptible of combination, by fusion, with silex and with alumine; and hence is employed in making artificial gems.

For a description of the remaining borates, I refer to the 2d volume of Thomson's System of Chemistry, or the 1st volume of La Grange's Manual.

CHAPTER XVII.

FLUORIC ACID.

I. THE fluoric acid may be obtained from a substance found abundantly in Derbyshire, under the name of *fluor spar*. In converting this spar to ornamental purposes, small pieces are broken off, which may be had at a cheap rate.

The fluoric acid may be separated from this combination, *in the form of gas*, by adding, to the powdered spar, in a block tin or leaden gas bottle, half its weight of concentrated sulphuric acid. The gas, which is disengaged, may be conducted, by a bent glass tube, to the mercurial trough. The receivers employed should be previously coated with wax, by making them hot enough to melt that substance, and then allowing the wax to run in fusion over their whole inner surface. This gas, in its properties, bears a considerable resemblance to muriatic acid gas. It is somewhat of a similar smell; produces a white smoke in the atmosphere; and is rapidly absorbed by water and by ice. It has the remarkable property of corroding glass. This may be observed by passing up, into a jar filled with it, a small slip of glass, which will be considerably acted on; or by holding a piece of glass over a mixture of concentrated sulphuric acid, and powdered fluate of lime. If partially applied, by covering the glass with a cement of wax and resin, and removing the cement in part only, the gas may be employed for writing or engraving on glass; and its application to this purpose has been proposed as an important improvement in the art of engraving, by Professor Wilson of Glasgow. (Nicholson's Journal, 4to. ii. 60.)

On the presumption that the fluoric acid might agree with other acids in containing oxygen, Mr. Davy was induced to submit it to the de-oxydizing power of potassium.

When potassium was exposed to fluoric acid gas, which had been prepared in glass vessels, white fumes were immediately perceived. The metal lost its splendour, and became covered with a greyish crust. On heating the bottom of the retort, the fumes increased; and continued for some time to be emitted, but at last ceased altogether. When water was admitted, the whole of the gas did not disappear, but a small residue was left which consisted of hydrogen.

On making the experiment in a similar manner, but with a temperature nearly equal to the sublimation of the metal, the potassium inflamed and burned with a brilliant red light. After the combustion, the whole of the fluoric acid was found to be destroyed, and no gas left but a residuum of hydrogen. Ten grains and a half of potassium destroyed 14 cubic inches of fluoric acid gas, not artificially dried; and left about two and a quarter of hydrogen. But from fluoric acid, long exposed to calcined sulphate of soda, hardly one tenth its bulk of hydrogen could be thus developed.

The bottom of the retort, in these cases, was covered with a crust of various colours, in some parts chocolate, in others yellow. When a portion of it was thrown into water, it effervesced violently, and an inflammable gas was evolved, in smell resembling phosphuretted hydrogen.

A portion of the dry mass, being heated in contact with air, burned slowly, lost its colour, and became a white saline mass. In oxygen gas, it burned with an absorption of oxygen, but not with any great intensity.

By the action of this substance on water, fluoric acid and potash were generated; and some chocolate coloured particles were separated by the filter, which, when dried, and heated in oxygen gas, burned and absorbed oxygen. By this combustion, fluoric acid was also generated.

That the fluoric acid, then, is decomposed by the action of potassium, and that its base exists in the solid matter which results, is scarcely questionable; but whether in a state of purity, or, if combined, in what state, remains to be decided by farther experiments. Of the regeneration of the acid, also, by the combustion of this matter in oxygen gas, it is hardly possible to entertain a doubt. As a portion of silex appeared, always, in the fluate of potash thus produced, Mr. Davy is disposed to consider the substance in question as consisting of the siliceous and fluoric bases in a low state of oxygenation united with some potash.

The fluoric acid may be obtained *in a liquid state*, by using a leaden retort and leaden receiver. An ingenious apparatus, invented for this purpose by Mr. Knight, is described and figured in the 17th volume of the Philosophical Magazine.

The liquid acid must be preserved in leaden bottles, as it soon corrodes and penetrates glass ones. In this state of watery solution, it readily combines with alkalis, and forms soluble compounds. Its combinations with the earths are for the most part highly insoluble. The fluates have no properties that can render them interesting to the student, except the use of the alkaline ones as tests, which will be described in a subsequent part of the work.

To the liquid solution of fluoric acid, Gay Lussac and Thenard have given the name of silici-fluoric acid. It emits, they observe, dense vapours; becomes hot and even boils when suddenly mixed with water; acts strongly on glass; and has a most powerful effect on the skin, on which it speedily raises small pustules.

A piece of potassium thrown into it detonates violently. The products are hydrogen, water and fluate of potash; consequently it is a compound of water, and fluoric acid. Potash, poured into the liquid acid, gives a salt which is almost insoluble; muriate of barytes affords fluate of silex and fluate of barytes, both of which are insoluble.

Fluoboric Acid.

With the view of obtaining fluoric acid gas perfectly free from water, both Mr. Davy and Gay Lussac appear to have had recourse to the same expedient, *viz.* that of distilling perfectly dry boracic acid with fluate of lime. When these substances were exposed to a strong heat in an iron tube, in the proportion of one part of the former to two of powdered fluor spar, a gas was collected in great quantity, which exhibited singular properties, and to which Messrs. Gay Lussac and Thenard have given the name of *gas fluoborique*, or fluoboric acid gas.

This gas, according to the latter chemists, appears to contain no water, and to have so strong an affinity for it as to take it from other gases which hold water in combination. Hence, when mixed with most of those gases, on which it does not exert a chemical action, such as atmospheric air, it loses its transparency and becomes cloudy.

With ammoniacal gas it unites in two proportions. If the alkaline gas be put first into the tube, equal measures combine together, and the compound is neutral. But if we admit fluoboric gas by bubbles to the alkaline gas, we obtain a compound, with an excess of base, and consisting of one measure of fluoboric gas to two of ammonia.

Fluoboric gas is absorbed copiously by water. The saturated solution has the causticity and aspect of strong sulphuric acid; requires for ebullition a temperature considerably exceeding 212° Fahrenheit; and is condensed again in striæ which contain much gas. From analogy, Gay Lussac supposes that nitric and even sulphuric acids are, in their pure states, equally elastic with this.

The liquid acts almost as intensely as sulphuric acid on vegetable substances. It blackens paper, and affords a true ether with alcohol. It has no effect in corroding glass.

From analysis, Gay Lussac and Thenard, as well as Mr. Davy, have determined it to be a compound of boracic and fluoric acids, in proportions not yet ascertained.

APPENDIX.

DESCRIPTION OF THE PLATES.

PLATE I.

FIG. 1. (*a*) *A plain retort,* the neck of which is shown introduced a proper length into the mouth of a plain receiver *b*. The dotted lines at *c* show the receiver with the addition of a tubulure, into which either a stopper, or bent glass tube, may be occasionally fixed.

FIG. 2. *A glass alembic; a* the body, and *b* the head, which are ground so as to fit accurately, and may be separated when necessary. The head *b* is so shaped, that any liquid, which may be condensed, collects into a channel, and is carried by the pipe *c* into the receiver.

FIG. 3. *A separator,* for separating liquids of different specific gravities. It is furnished with a ground stopper at *a*, and a glass stop-cock at *b*. The vessel is filled with the liquids that are to be separated (oil and water for example), which are allowed to stand till the lighter has completely risen to the top. The stopper *a* is then removed, and the cock *b* opened, through which the heavier liquid descends; the cock being shut, as soon as the lighter one is about to flow out.

FIG. 4. A glass vessel, termed a *mattrass,* useful for effecting the solution of bodies, which require heat before

they can be dissolved, or long continued digestion, see vol. i. p. 47. The upper extremity of the long neck generally remains cool, and allows the vessel and its contents to be shaken occasionally.

Fig. 5. A glass bottle with a very thin bottom, and a projecting ring round the neck for suspending it over a lamp. These are useful for effecting *solutions* on a small scale.

Fig. 6. An apparatus, contrived by Mr. Pepys, *for ascertaining the quantity of carbonic acid discharged from any substance by the addition of an acid.* It consists of a bottle closed by a ground stopper. This stopper is perforated, and forms the lower part of a tube, which is twisted into the shape of a still-worm. In this worm, any water that escapes along with the gas, is condensed, and falls down again into the bottle. The experiment is made precisely as described, vol. i. p. 334: and the loss of weight is determined at the close of the effervescence.

Fig. 7. Mr. Leslie's *differential thermometer* described, vol. i. p. 114.

Fig. 8. (*a*) *An air thermometer,* for ascertaining the temperature of liquids. It consists of a bottle, partly filled with any coloured liquid, and partly with air, a glass tube of small bore, open at both ends, being either cemented or hermetically sealed into the bottle, so that its lower extremity may nearly touch the bottom of the bottle. The expansion of the included air, on the application of heat, drives the coloured liquid up the tube, and to an extent which may be measured by the application of a scale. The fig. *b* is another variety of the same instrument, described vol. i. p. 113.

Fig. 9. The original *air thermometer of Sanctorio;* see vol. i. p. 112.

Fig. 10. A bent funnel for introducing liquids into retorts, without soiling their necks.

Fig. 11. *An adopter.* The wider end admits the neck of

a retort; and the narrower, is passed into the mouth of a receiver.

Fig. 12. A section of an evaporating dish of Wedgwood's ware.

Fig. 13. (*a*) *A tubulated retort* luted to (*b*) *a quilled receiver*, the pipe of which enters the neck of a bottle (*c*) supported by a block of wood.

Fig. 14. Different forms of jars for precipitations, with lips for conveniently decanting the fluid from the precipitate.

Fig. 15. A tube, blown in the middle into a ball, for dropping liquids. The ball is filled by the action of the mouth applied to the upper orifice, while the lower one is immersed in the liquid. To the former the finger is then applied; and, on cautiously removing it, the liquid is expelled in drops.

Fig. 16. *A bottle for ascertaining the specific gravity of liquids.* When filled up to a mark in the neck, with distilled water of a given temperature, it should hold 1000, 2000, or any even number of grains. The quantity, which it is found to contain, of any other liquid of the same temperature, shows the specific gravity of the latter. For example, if it hold 1000 grains of water, and 1850 of sulphuric acid, the specific gravity of the latter is to that of water as 1850 to 1000.

PLATE II.

Fig. 17. An apparatus for procuring gases, without the possibility of their escaping into the room during the process, a circumstance which is of considerable importance, when the gas has an unpleasant smell or deleterious properties. Suppose that sulphuretted hydrogen gas is to be obtained from sulphuret of iron and diluted sulphuric acid. The sulphuret of iron, in coarse powder, is put into the body of the gas bottle *c*, with a proper quantity of water. The acid

holder *a* is filled with the diluted acid, the cock *b* being shut, and is then fixed into the tubulure of the gas bottle, to which it is accurately adapted by grinding. The bent tube *d* being made to terminate under a receiver filled with, and inverted in water, the perforated cock *b* is gradually opened, in consequence of which the acid descends into the gas bottle; and acts on the sulphuret of iron. If it be found necessary to renew the acid, without disturbing the apparatus, this may be done as follows. The cock *b* being shut, the stopper, which closes the acid holder, may be removed, and fresh acid be poured in, through the aperture. This may be repeated as often as is found neeessary. The acid holder may be advantageously adapted, also, to a retort for certain distillations, such as that of muriatic acid.

Fig. 18. A plain gas bottle with sigmoid tube, the end, which is received into the bottle, having a ground stopper accurately fitted to the neck. For ordinary purposes (such as obtaining hydrogen gas from diluted sulphuric acid and iron filings) this apparatus answers perfectly well, and is much less costly. It is frequently made with a tubulure and glass stopper, and is then called a tubulated gas bottle.

Fig. 19. A gas funnel, useful in transferring any gas, from a wide-mouthed vessel into a jar of narrower diameter, or into a bottle. When employed for this purpose, it is held inverted, as shown by the figure, the pipe being admitted into the aperture of the bottle or jar, which is filled with and inverted in water, and the gas being made to pass into it in bubbles.

Fig. 20. Dr. Hope's Eudiometer. The manner of using it has already been described, vol. i. p. 191.

Fig. 21. A modification of Dr. Hope's Eudiometer described, vol. i. p. 192.

Fig. 22. A gas receiver, into the neck of which is cemented a brass cap, with a female screw for receiving a stop-cock. The vessel *b* is a glass flask, which may be made to communicate with the interior of the jar *a*, by opening the cock. When the apparatus is used, it is necessary to employ two stop-cocks, and not one only, as represented by the figure. Supposing that the weight of any gas is to be ascertained, the flask *b* is exhausted, by screwing it on the transfer of an air pump; and, if great accuracy be required, it is proper to enclose a gage in the vessel. Let the flask be weighed when exhausted; then screw it upon the top of the receiver, containing the gas which is to be weighed; and open the communication, observing, by using a graduated jar, how much gas has been admitted. Suppose this to be 50 cubic inches. By weighing the flask again when full, we determine the weight of 50 cubic inches of the gas under examination. The experiment should be made when the temperature of the room is 60°, and when the barometer stands at 29.8.

FIG. 23. A plain jar for receiving gases, with a ground stopper.

FIG. 24. An Eudiometer for trying the purity of a mixture of gases containing oxygen gas, by means of nitrous gas. The process has already been described, vol. i. p. 415. The instrument should be accompanied with a phial, holding, when completely full, precisely a cubic inch.

FIG. 25. A wire stand, with a leaden foot, for the purpose of raising, above the surface of water within a jar, any substance which is to be exposed to the action of gas.

FIG. 26. An apparatus for showing that caloric exists in gases in a latent form. The application of it has been already described, vol. i. p. 173.

FIG. 27. An apparatus for drying precipitates by steam, described, vol. i. p. 47.

FIGS. 28 and 29. Tubes for exploding mixtures of hydrogen and other inflammable gases with oxygen gas, commonly termed the Eudiometer of Volta; see vol. i. p. 197.

PLATE III.

Fig. 30. The common form of a Woulfe's apparatus. In this figure the retort *a* is represented plain, but it is better to employ a tubulated one. The use of this apparatus has already been described, vol. i. p. 42.

Fig. 31. A modification of the apparatus, which has also been already described. In this figure, the mercurial trough is shown with a jar standing inverted in it, for the purpose of receiving any gas that may escape condensation by water.

Fig. 32. Mr. Pepys' improvement of Woulfe's apparatus described, vol. i. p. 32.

PLATE IV.

Figs. 33 and 34. Cuthbertson's apparatus, for exhibiting the composition of water, with the substitution of gazometers for the receivers originally employed by him. The apparatus has been described, vol. i. p. 208. Fig. 33 is an enlarged view of the conical brass piece, which is cemented into the bottom of the receiver, and through which the gases are conveyed.

Fig. 35. A gazometer of the most simple and common construction; see vol. i. p. 164.

Fig. 36. A gas holder, described, vol. i. p. 166.

Fig. 37. A galvanic trough; see vol. i. p. 226. The tube *b* shows the arrangement for decomposing water. The upper wire may be hermetically sealed into the tube, and the lower one passed through a cork, which should have a small slit cut in it, to allow the water to escape in drops as the gas is produced.

Fig. 38. The manner in which a candle may be burned in oxygen gas; see vol. i. p. 178.

FIG. 39. The combustion of iron-wire in oxygen gas.

FIG. 40. Apparatus for decomposing water over red-hot iron or charcoal; see vol. i. p. 210.

FIG. 41. An apparatus for showing the diminution effected in the volume of hydrogen and oxygen gases by their slow combustion; see vol. i. p. 199.

FIG. 42. A very simple and cheap contrivance for freezing quicksilver by muriate of lime and snow. The outer vessel of wood may be twelve and a half inches square, and seven inches deep. It should have a wooden cover, rabbeted in, and furnished with a handle. Within this is placed a tin vessel *b b*, standing on feet which are one and a half inch high, and having a projection at the top, half an inch broad, and an inch deep, on which rests a shallow tin-pan *c c*. Within the second vessel is a third *d*, made of untinned iron, and supported by feet two inches high. This vessel is four inches square, and is intended to contain the mercury. When the apparatus is used, a mixture of muriate of lime and snow is put into the outer vessel *a a*, so as completely to surround the middle vessel *b b*. Into the latter, the vessel *d*, containing the quicksilver to be frozen, previously cooled down by a freezing mixture, is put; and this is immediately surrounded by a mixture of snow and muriate of lime, previously cooled to 0° Fahrenheit, by an artificial mixture of snow and common salt. The pan *c c* is also filled with these materials, and the wooden cover is then put into its place. The vessels are now left, till the quicksilver is frozen. A more elegant, but more expensive, apparatus, by Mr. Pepys, intended for the same purpose, is figured in an early volume of the Philosophical Magazine.

FIG. 43. A wire stand, consisting of an interior circle, and three straight pieces of wire proceeding from it in the same plane. Its use is noticed, vol. i. p. 181.

FIG. 44. Mr. Davy's apparatus for the analysis of soils

described in his paper, which is copied into the third part of this work.

PLATE V.

FIG. 45. Pictet's arrangement of an apparatus for showing the radiation of caloric, unaccompanied by light; see vol. i. p. 123.

FIG. 46. An oval copper boiler, for exhibiting the most important facts respecting latent caloric. The size of its different parts (except the width, which is 4 inches) may be learned from the scale affixed to the plate, which is abundantly sufficient to enable any intelligent workman to construct the apparatus. The collar joint and stuffing box, however, it is indispensably necessary to describe, especially as the former article of apparatus is generally constructed on a bad plan.

FIG. 47 is a section upon a larger scale, of the collar joint at *b* (fig 46), made for the convenience of screwing together long or crooked metal tubes, without turning them round. *a* is a section of the end of one of the tubes, and *b* that of the other which is to be attached to it; *c* is a collar which turns loose upon the shoulder of *a*, and screws upon *b*. By screwing this collar upon *b*, the end *e e* of the tube *a* is brought to press upon the part *d d* of the tube *b*, without turning round either of those tubes. If upon *d* be laid a ring of linen cloth soaked in boiled linseed-oil, the joint, when screwed up (if tolerably well made), will be impervious to steam as well as to water or air. The projection at *d* is for preserving the ring of cloth from being displaced, and for guiding the ends of both tubes, so as to meet properly.

FIG. 48 is a section of a socket, for fixing the stem of a thermometer into a boiler or a digester, where there is much heat and pressure; *b* is a socket fixed on the outside of the boiler or digester, having a hole through it large enough to admit the bulb of the thermometer; *a* is a plug which screws into *b*, having a hole through its centre large enough to admit only the stem of the thermometer; *c c* is a loose round plate, concave on the upper side, having a hole through its centre, just sufficient also to admit the stem of the thermometer. When the instrument is to be inserted, the plug *a*, and the plate *c*, must both be taken out of the socket. The bulb is then passed through it. The plate *c* is next slipped over the stem, and dropped into its place. Some flax, soaked in linseed-oil, must next be wrapped round the stem, so as nearly to fill the socket. The plug *a* must then be screwed in, till the flax be compressed so as to make the whole sufficiently tight. The opposite surfaces of the plate *c*, and the plug *a*, are made concave, for the purpose of compressing the flax round the stem of the thermometer.

PLATE VI.

Figs. 49, 50, 51. Sections of crucibles.

Fig. 52. A muffle; see vol. i. p. 39.

Fig. 53. Stands for raising the crucible above the bars of the grate; *a* one adapted to Mr. Aikin's blast furnace; *b* one of the common form.

Fig. 54. A skittle-shaped crucible.

Fig. 55. Mr. Aikin's portable blast furnace. It is composed of three parts, all made out of the common thin black-lead melting pots, sold in London for the use of the goldsmiths. The lower piece *c* is the bottom of one of these pots, cut off so low as only to leave a cavity of about an inch, and ground smooth above and below. The outside diameter over the top is five and a half inches. The middle piece, or fire-place *a*, is a larger portion of a similar pot, with a cavity about six inches deep, and measuring seven and a half inches over the top, outside diameter, and perforated with six blast holes at the

bottom. These two pots are all that are essentially necessary to the furnace for most operations; but when it is wished to heap up fuel above the top of a crucible contained within, and especially to protect the eyes from the intolerable glare of the fire when in full heat, an upper pot *b* is added, of the same dimensions as the middle one, and with a large opening in the side, cut to allow the exit of the smoke and flame. It has also an iron stem, with a wooden handle (an old chisel answers the purpose very well) for removing it occasionally.

The bellows, which are double (*d*), are firmly fixed, by a little contrivance which will take off and on, to a heavy stool, as represented in the plate; and their handle should be lengthened, so as to make them work easier to the hand. To increase their force on particular occasions, a plate of lead may be firmly tied on the wood of the upper flap. The nozzle is received into a hole in the pot *c*, which conducts the blast into its cavity. From hence the air passes into the fire-place *a*, through six holes of the size of a large gimlet, drilled at equal distances through the bottom of the pot; and all converging in an inward direction, so that if prolonged, they would meet about the centre of the upper part of the fire. Fig. 56 shows the distribution of these holes in the bottom. The large central hole is intended to receive the stand *a*, fig. 53, which serves for supporting the crucible.

No luting is necessary in using this furnace, so that it may be set up and taken down immediately. Coak, or common cinders, taken from the fire when the coal just ceases to blaze, sifted from the dust, and broken into very small pieces, forms the best fuel for higher heats. The fire may be kindled at first by a few lighted cinders, and a small quantity of wood-charcoal.

The heat which this little furnace will afford is so intense, that its power was, at first, discovered accidentally by the fusion of a thick piece of cast iron. The utmost heat procured by it was 167° of Wedgwood's pyrometer piece, which was withdrawn from a Hessian crucible, when actually sinking down in a state of porcellanous fusion. A steady heat of 155° or 160° may be depended on if the fire be properly managed, and the bellows worked with vigour. (See Philosophical Magazine, vol. xvii., p. 166.)

By a letter from Mr. Aikin, I have learned, also, a convenient way of exhibiting, in a lecture, and performing at other times, the process of cupellation, by means of this furnace. It consists in causing a portion of the blast to be diverted from the fuel, and to pass through a crucible in which the cupel is placed. This arrangement supplies air; and the whole may be seen by a sloping tube, run through the cover of the crucible. Fig. 57 shows the furnace when used for this purpose; *a a* the furnace; *b* the perforated stopper for the central blast; *c c* a portion of earthen tube, through which the air passes, and is heated during this transit; *e* a piece of soft brick, perforated to admit the earthen tube *f*, which may be kept open for inspecting the process. No luting is required, except to join *f* to *e*.

Fig. 58. Knight's portable furnace, composed of strong iron plate lined with fire lute, the inside diameter six inches. *a* shows the grate; *b* the ash pit door; *d* the door of the fire-place when used as a sand heat; *e e* two holes opposite to each other for transmitting a tube; *g* an opening for a retort neck, when used for distilling with the naked fire.

Fig. 59. A different view of the same furnace; *a* the grate; *c* the register to the ash pit; *f* a small door, with a contrivance for supporting a muffle. The other letters correspond with the explanation of the preceding figure.

For this furnace the proper fuel, when it is used as a wind furnace, is wood-charcoal, either alone, or with the admixture of a small proportion of coak. For distillation with a sand heat, charcoal, with a little pit coal, may be employed.

PLATE VII.

Fig. 60 represents a fixed furnace, which I find very useful, because it may either be employed as a wind furnace or for distillation with a sand heat. Its total height outside is thirty-three inches, and the outside square is eighteen inches, or two bricks laid lengthwise. The thickness of the sides of the furnace is the breadth of a brick, or four and a half inches; but whenever there is room, it is better to make them nine inches in thickness. From the top of the furnace to the grate, which is moveable, and supported by two bearers, the height is thirteen inches; and at *c* is a double Rumford door; or in preference, a hole closed by a moveable earthern stopper, for introducing fuel. The ash pit should have a register door. The chimney is four inches wide by three high, and may either be furnished with a damper or not. On the top of the furnace a cast-iron ring is fixed, ten inches inside diameter, three inches broad, and half an inch thick. It is secured in its place by three iron pins, passing through three equidistant holes in the ring, and bent at the distance of nine inches at a right angle. These serve the purpose of binding the ring firmly into the brick-work. The sand pots are of different sizes; and a variety of them may be made to fit the same ring, by varying the breadth of their rims, as shown fig. 71. The bricks should be cemented together, at least for the inner half of their breadth, by loam, or by a mixture of Stourbridge clay, with two or three parts sand, and a proper quantity of water.

When this is used as a wind furnace, the opening in the side is to be closed by its stopper; or, if a Rumford door be employed, it must be defended from the fuel by a fire tile. The fuel (coak) is introduced at the top, which is occasionally covered by a fire tile. When distillation with a sand heat is performed, the sand pot rests on the iron ring, and the fuel, which may be common pit coal, is added through the opening in the side. It may be proper to state, that, in order to receive a sand pot of as large a size as possible, the upper course of brick should be bevilled within the furnace; and the width at the top may exceed a little that at the grate.

The best Stourbridge or Newcastle-on-Tyne fire-bricks are necessary in constructing this and the following furnaces.

FIG. 61 is a longitudinal section of a wind furnace, invented by Mr. Knight, with an additional chamber for applying the waste heat to useful purposes. *a* the internal cavity, which is square, for containing the fuel and the crucible. *b* the flue passing into a hot chamber *c*; an appendage particularly useful for drying luted crucibles, or bringing them to a proper temperature for the furnace; for roasting ores and various other purposes. *d* the flue connecting it with the vertical chimney *e*; which, to produce a strong heat, should never be less than thirty or forty feet high. *f f* covers, consisting of twelve-inch Welsh tiles, with handles. *g* the stoke hole, through which no more of the fire is seen than appears between the grate and the bearing bar *h*. This space is left for the double purpose of raking the fire, and occasionally taking out the bars. *k* the ash pit, which is sunk below the level of the ground, and is covered, where it projects at *l*, by an iron grating.

The best situation for this furnace, is an angle of the laboratory, the chimney being in the corner, as represented in the sketch. By this arrangement, the operator is spared the disagreeable necessity of scorching his legs, by standing opposite the stoke hole, while the backs of his legs are exposed to a current of cold air rushing to the furnace.

FIGS. 62 and 63 are different views of a furnace invented by Mr. Knight, and convertible to various purposes.

The inside of this furnace is nine inches square, and sixteen inches deep from the top to the grate. The face of the opening at *g* rises at an angle, which makes the back part five inches higher than the front. This contrivance enables us completely to cover a large retort with fuel, without obstructing the passage of the air, and also relieves partly the weight of the cover, when it requires to be moved. The walls of the furnace are at least a brick and a half in thickness, and as much more as local convenience will allow. By sinking the ash pit below the level of the ground, at *i*, the height of the furnace needs not exceed eighteen inches, which renders the management of the fuel much more easy, and subjects the face and hands less to the action of the heat. The ash pit *a*, must be at least eighteen inches deep, below the surface of the ground, and more if convenient. It must have an opening, projecting from it three or four feet, to be covered with boards, and with an iron grating next the furnace. This preserves the legs of the operator from the action of the fire.

The grate *b* is formed of separate bars, each of a triangular shape, three fourths of an inch apart, and resting on two bearers. In the front of the furnace, an iron bar is to be placed to support the brick-work, and to leave an opening, through which the bars may occasionally be drawn out, and the fire be raked and cleared of the slag. The chimney *e* is two and a half inches from the top, and four and a half wide, by two and a half high.

To fit this furnace for occasional distillation with the naked fire, an opening, *d*, fig. 62, is left on one side, which is filled up, when not wanted, by five pieces of soft fire-brick, cut to a proper shape, and secured by a clay lute. It is proper, also, to be provided with other pieces, having arched openings for transmitting the neck of a retort. One of these pieces may have a round hole for occasionally transmitting a tube, and a corresponding hole, *h*, fig 63, must then be made in the opposite side of the furnace, to be closed, when not wanted, with a stopper.

Figs. 64 and 65 represent a sand heat, for containing flat evaporating vessels; the depth from back to front two feet; the width, agreeably to the scale, six feet. At the front is a rim four inches deep, consisting of a piece of iron plate, which is fastened at each end into the wall. The floor or bottom, *e e*, is formed of cast-iron plates, which rest upon each other in corresponding rabbets. The advantage of several small plates, over one large one, is the cheapness and facility, with which they are replaced, if cracked by the heat, an accident of not infrequent occurrence. The joints are secured by a fire lute, which effectually prevents the sand from falling through. The fire place is shown by *b*; at the bottom it has a grate ten inches long, by eight wide. The flame and smoke circulate first through the flue *c*, and then through the returning flue *d*, which conveys the smoke to the chimney *g*. In constructing the flue beneath the grate, a row of bricks, set edgeways, answers the purpose, and serves also to support the inner edge of the plates.

It is advisable to cover the sand heat with a sloping roof, which may be formed of lath and plaster, and supported by side walls. The lowest part of the roof may be foremost, and about three feet above the edge of the iron plates. It is, also, necessary to have an air flue, nearly at the top of the back wall, under the dome or roof, to be closed occasionally by a door. This must open into the chimney, in which case it serves the purpose of carrying off noxious vapours.

PLATE VIII.

Figs. 66, 67, 68, are the section and plans of a reverberatory furnace for experimental purposes. In this furnace, the fuel is contained in an anterior fire-place; and the substance, to be submitted to the action of heat, is placed on the floor of another chamber, situated between the front one and the chimney. The flame of the fuel passes into the second compartment; by the form of which it is concentrated upon the substance exposed to heat, which is not confined in a separate vessel or crucible, but placed on the floor of the furnace. When reduced to a state of fusion, the melted mass is allowed to flow out through a tap-hole at *h*. The dimensions of this furnace it is scarcely possible to state, as they vary so considerably in different parts of it; but they may be ascertained by referring to the figures, and by the application of the scale. In all three figures, *a* represents the ash pit; *b* the grate composed of moveable bars; *c* the door at which the fuel is introduced; *d* a door in the side of the chamber, for the purpose of inspecting the process; *e* the floor of the furnace which descends, and is gradually contracted towards the back part; *f* another door for introducing and stirring the materials; *g* the back part of the furnace, immediately under the chimney; *h* the tap-hole; *i* the chimney.

Figs. 69 and 70, exhibit a cupelling or *enamelling furnace.* The form of this should be an oblong square; its dimensions being regulated by that of the muffle, which should go home to the back, its front edge lodging on the mouth of the furnace. On each side of the muffle, two inches and a half must be left, to let the fuel pass readily underneath, where there should also be a similar space. A stoke hole must be left on the other side, but the situation of the view will not admit its being shown. Before the muffle, is a projecting ledge or shelf, shown at *e*, which is intended to support any thing that is to be put into the muffle. Two twelve-inch tiles, worked in along with the bricks, will answer this purpose. In both figures, *a* shows the ash pit; *c* the grate; *d* the muffle; *e* the opening for introducing the muffle; *f* the chimney, and *g* the cover.

Fig. 71. Sand pots with rims of different sizes.

Figs. 72, 73. Dr Black's portable furnace, made of sheet iron lined with fire clay. Its dimensions, as they vary in almost every part, will best be learned from the scale; *a* the fire place; *b* the chimney; *c* the ash pit; *d* the door of the ash pit; *e* a register for regulating the quantity of air admitted to pass through the fuel.

Fig. 74. Mr. Chenevix's wind furnace. This is rudely sketched in Nicholson's Journal, from which the more accurate figure in plate viii. is taken. This furnace Mr. Chenevix describes as follows: "I have constructed a wind furnace, which, in some respects, is preferable to the usual form. The sides, instead of being perpendicular, are inverted; so that the hollow space is pyramidal. At the bottom the space is twelve inches square, and at the top only eight. The perpendicular height is seventeen inches, from the top to the grate. This form unites the following advantages. 1. A large surface is exposed to the air, which, having an easy entrance, rushes through the fuel with great rapidity. 2. The inclined sides act as reverberators. 3. The fuel falls of itself, and is always close to the grate."

In the figure, *a* represents the grate; *c c* are two bricks which can be let in at pleasure, to diminish the capacity: *b* is another grate which can be placed on the bricks *c c*, for occasional purposes: *d d* are bricks, which can be placed on the grate *b*, to diminish the capacity of this part of the furnace; *e* the cover. Both sets of bricks should be ground to the slope of the furnace.

In the construction of every furnace, which is intended to produce a strong heat, lime or mortar should be avoided, and the bricks should be set in loam, or Stourbridge clay, worked up with water and sand, inserting occasionally pieces of sheet iron, bent twice in opposite directions at right angles. The furnace should be allowed to remain some weeks, after setting up, before it is used; and before raising a strong heat, a gentle fire should be sometimes kindled in it, the strength of which may be gradually increased. When a strong blast is expected, it is necessary to bind the brick-work together, externally, by strong iron bars and plates, kept in their places by screws.

The chimney should be nine inches wide, and raised to as great a height as circumstances will admit

The coak of pit coal is the only fuel fitted for exciting an intense heat, and should be used in all cases, except in the reverberatory, and in distillations with the sand bath, when pit coal may be employed. The charcoal of wood is adapted principally to portable furnaces.

PLATE IX.

Fig. 75. The galvanic battery called *couronne de tasses*, described vol. i. p. 227.

Fig. 76. Apparatus for obtaining the elements of water in separate tubes; see vol. i. p. 236.

Fig. 77. The pile of Volta; see vol. i. p. 226.

Fig. 78. Section of a galvanic trough, to explain the theory of the excitation of galvanic electricity; see vol. i. p. 249.

Fig. 79. Apparatus for obtaining oxygen and hydrogen gases, from separate quantities of water not in contact with each other; see vol. i. p. 236.

Fig. 80. Two agate cups connected by moistened amianthus; see vol. i. p. 238.

Fig. 81. Two gold cones similarly connected, *ibid.*

Fig. 82. Agate cups similarly connected with an intermediate vessel *i*; see vol. i. p. 240.

Fig. 83. Apparatus for procuring potassium from potash and iron filings, described vol. i. p. 262.

Fig. 84. Apparatus for firing gases by electricity, or submitting them to electrical discharges, vol. i. p. 168.

Fig. 85. Pepys' improved gas-holder. *a* a small iron retort placed in the fire with a jointed conducting tube *b*, which is admitted into the vessel at *c*. This is shown on a larger scale in a different part of the plate. The letter *d* is placed on the body of the reservoir, and near the central pipe, which descends from the cistern *e* to nearly the bottom of the vessel. At *f* a glass tube is fixed, which shows the height of the water within the vessel. When a jar is intended to be filled with gas from the reservoir, it is placed, filled with water and inverted, in the cistern *e*. The cocks 1 and 2 being opened, the water descends through the pipe attached to the latter, and the gas rises through the cock 1. By raising the cistern *e* to a greater elevation, any degree of pressure may be obtained; and a blow-pipe may be screwed on the cock at the left side of the vessel.

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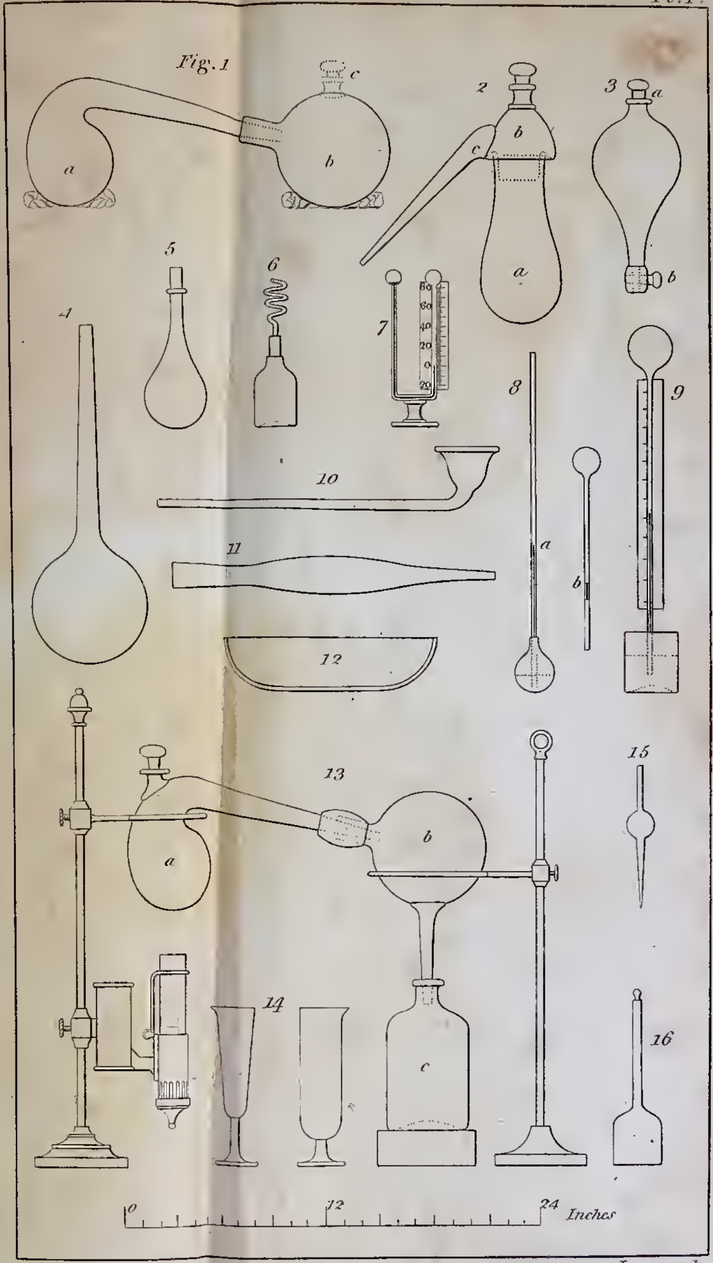

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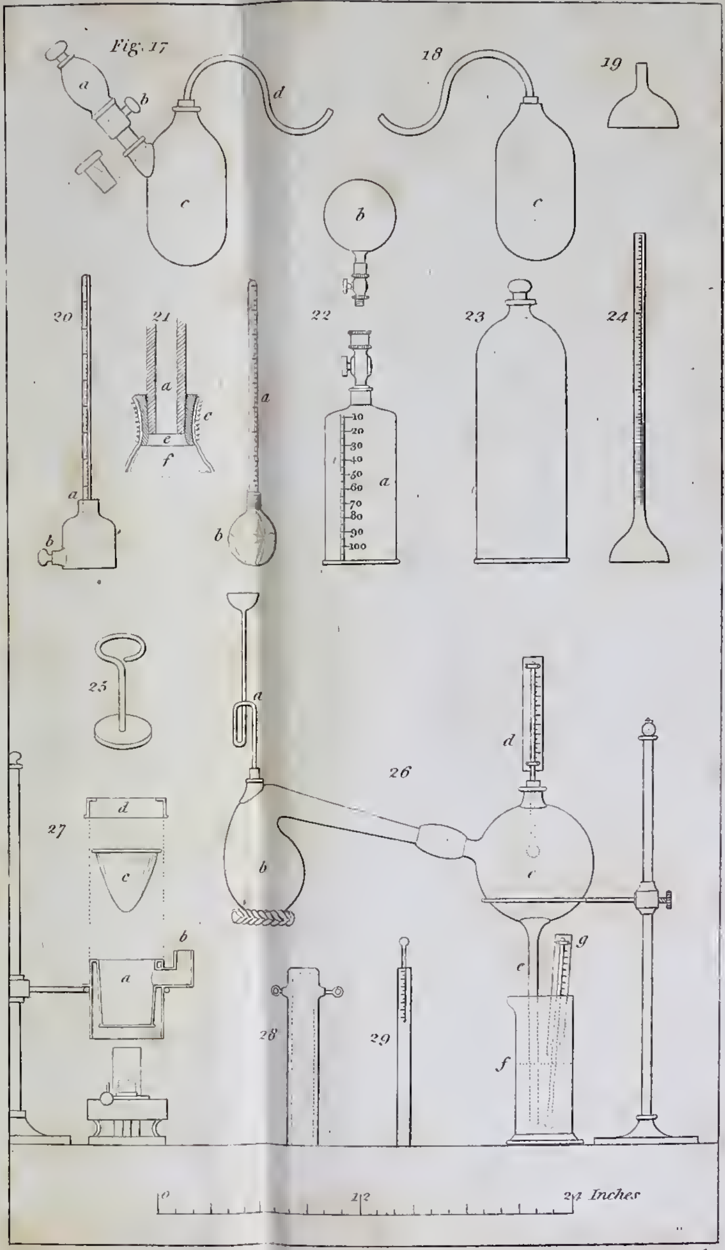

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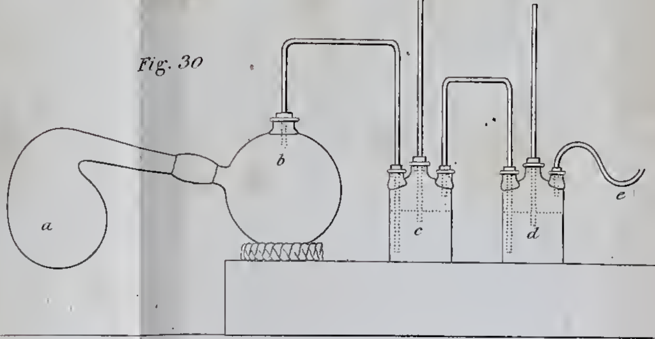
Pl. III.

Fig. 30



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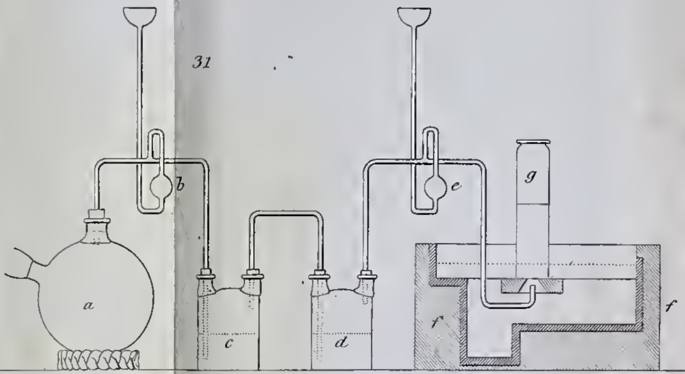

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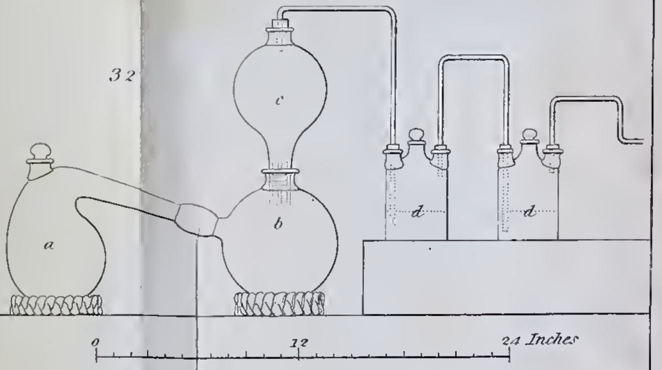

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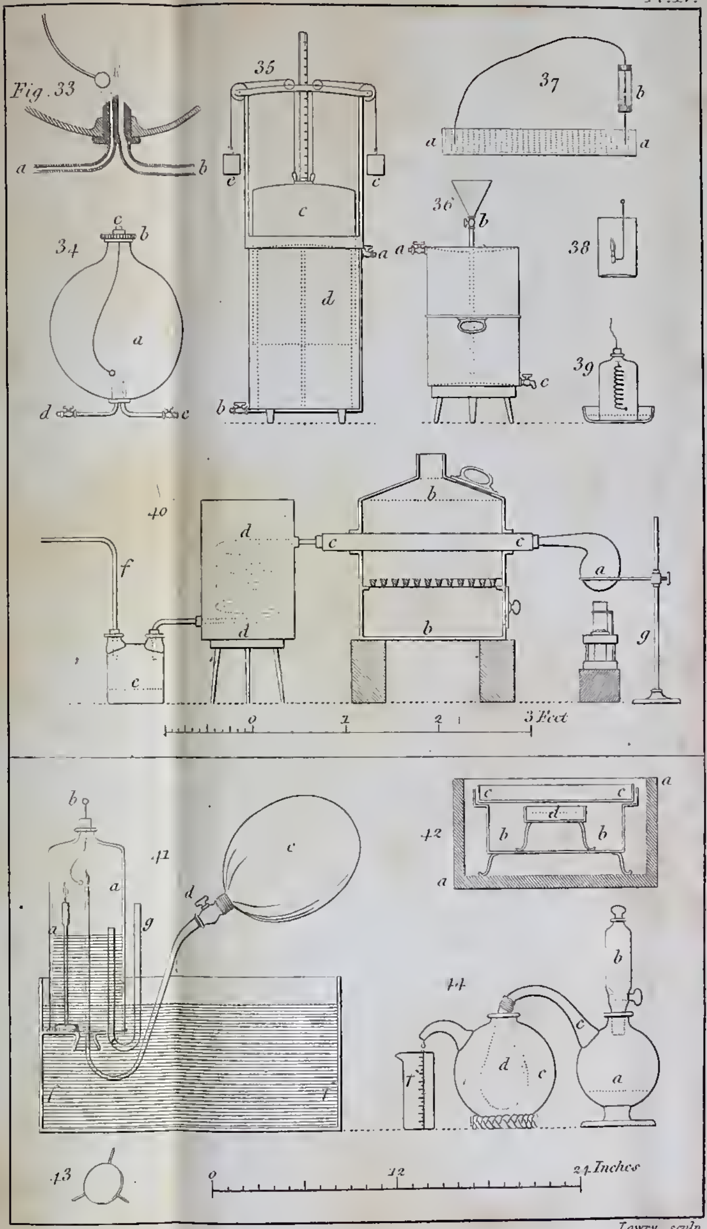

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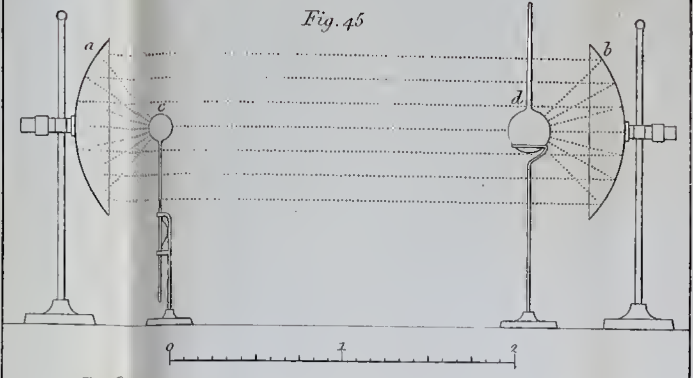

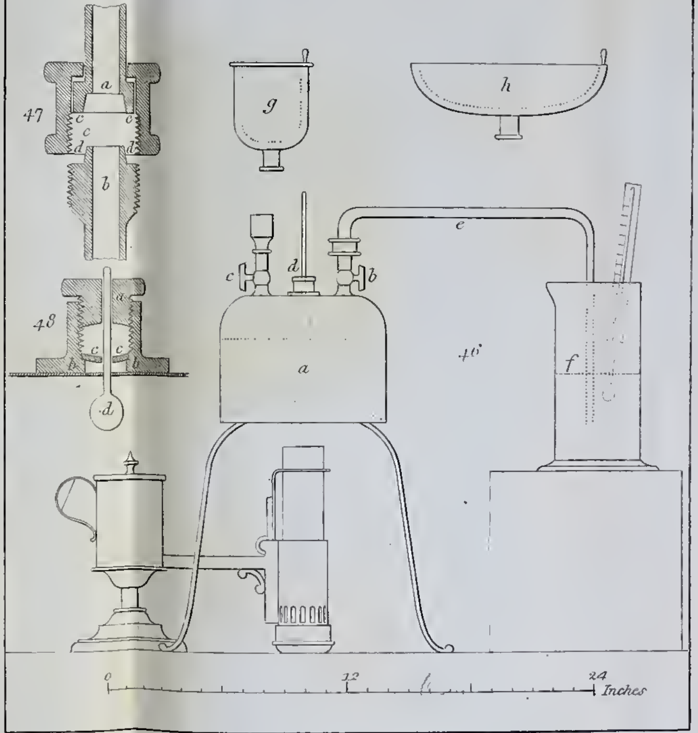

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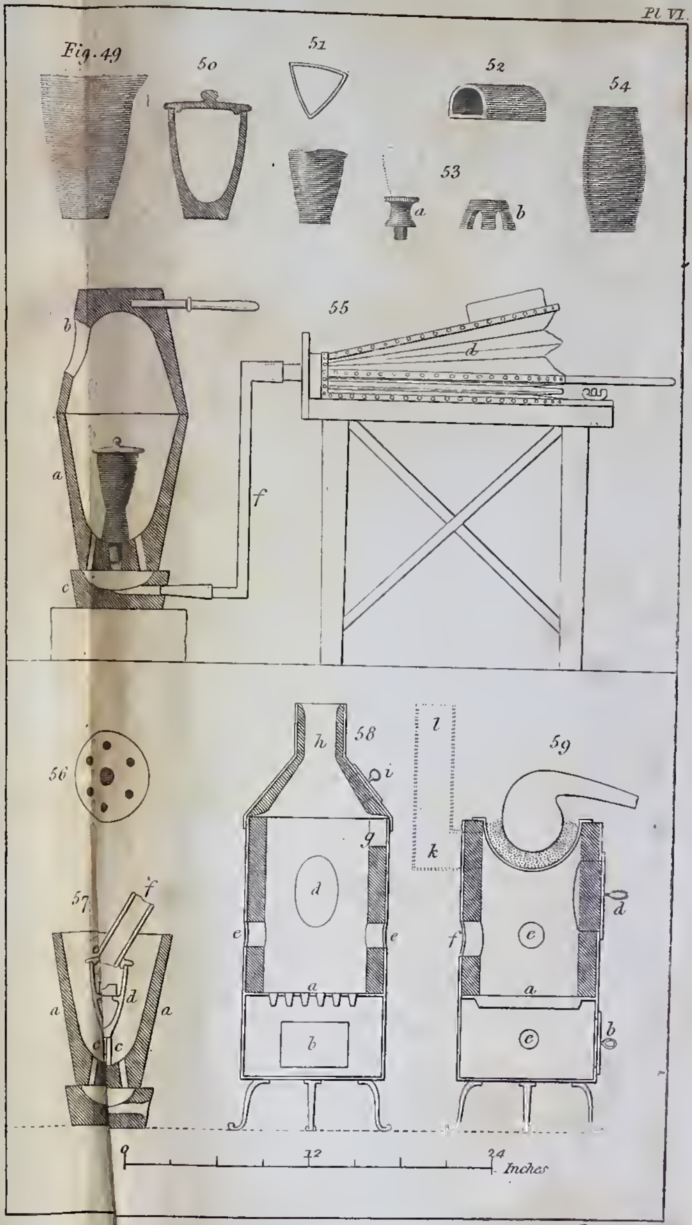

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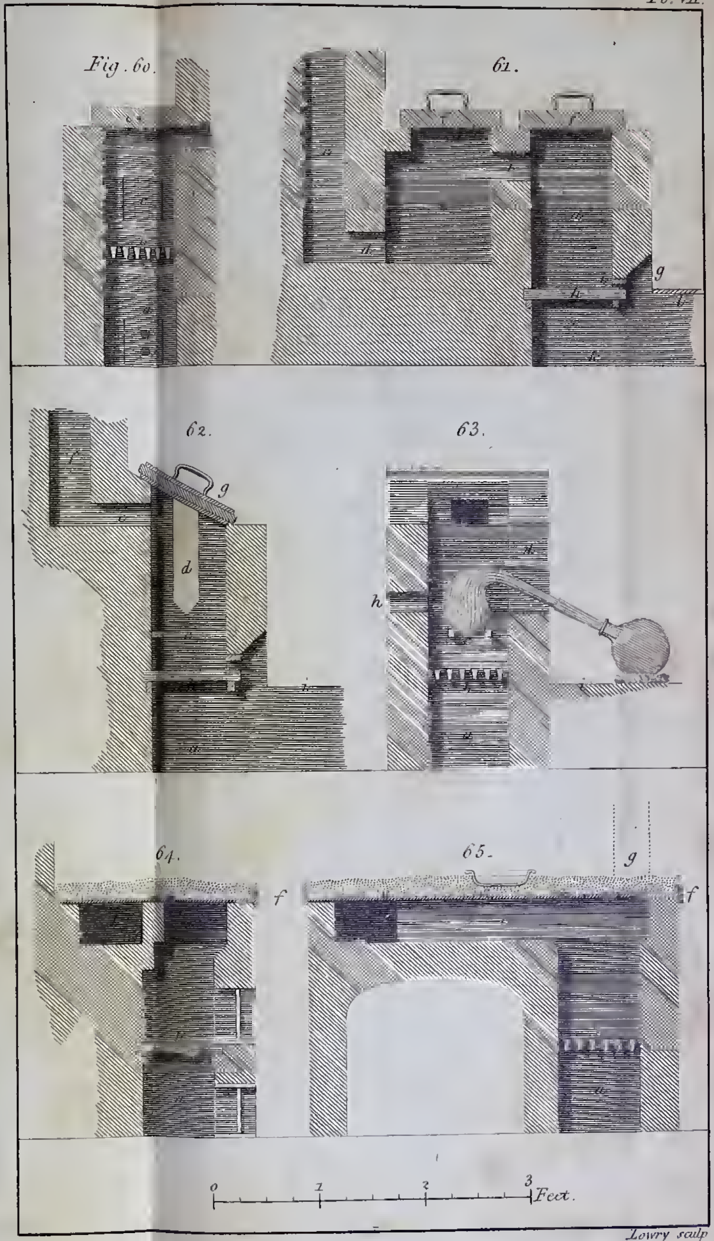

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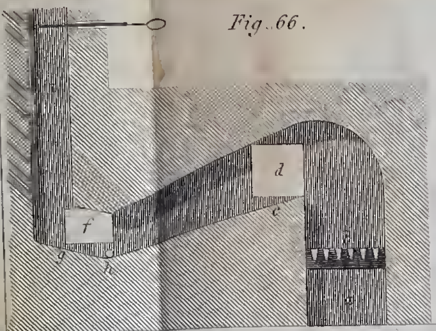

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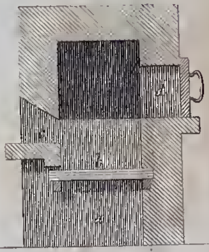

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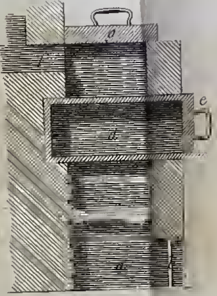

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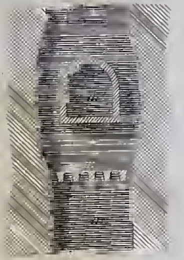

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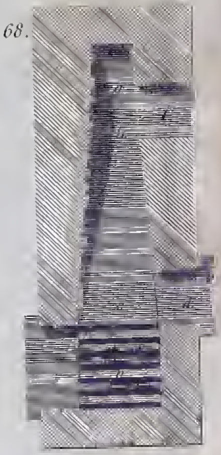

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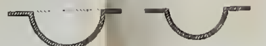

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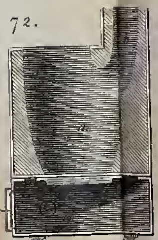

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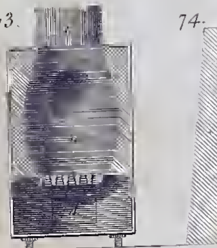

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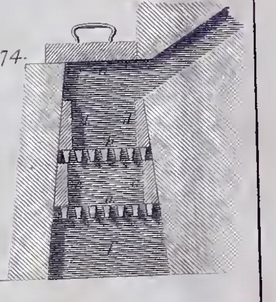

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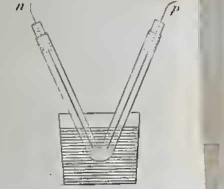

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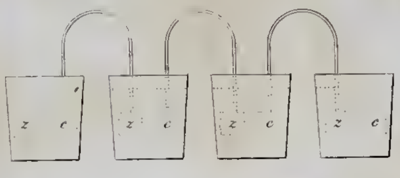

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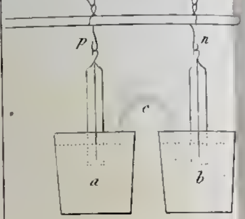

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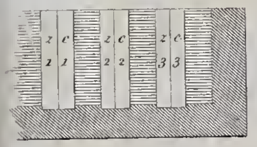

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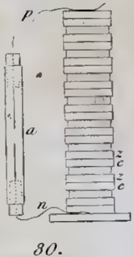

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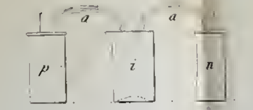

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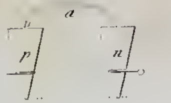

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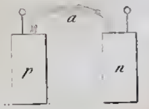

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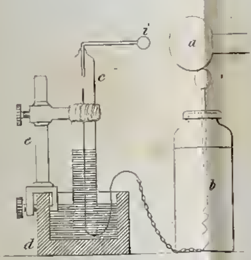

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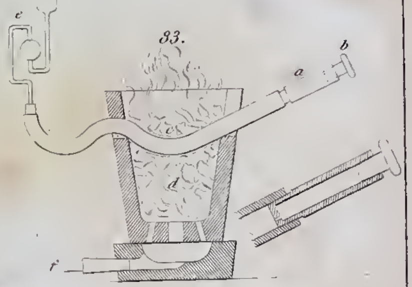

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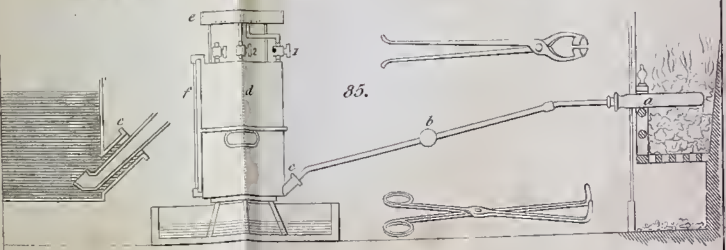

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